

## Studies on Pillaromycin A. V.<sup>1)</sup> The X-Ray Analysis of Pillaronone Monobromoacetate

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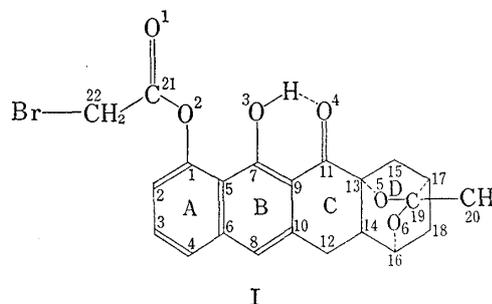
In order to confirm the chemical and stereochemical structure of pillarone, a derivative of the aglycone obtained from pillaromycin A, an X-ray analysis of its monobromoacetate has been carried out. Pillaronone monobromoacetate crystallizes in the orthorhombic system, space group  $P2_12_12_1-D_2^5$ , with four molecules of  $C_{22}H_{19}O_6Br$  in a unit cell. Application of a three-dimensional Patterson superposition method facilitated the elucidation of the structure which is given in formula (I). The final  $R$ -value was 0.133. The absolute configuration was determined using the anomalous dispersion of the bromine atom. The molecular structure in the crystal was also discussed.

Pillaromycin A,<sup>3)</sup>  $C_{28}H_{28}O_{11}$ , is an anti-tumor antibiotic compound produced by *Streptomyces flavovirens* No. 65786. The present X-ray work<sup>4)</sup> has been carried out in order to establish the chemical and stereochemical structure of pillarone,  $C_{20}H_{18}O_5$ , which is prepared by catalytic reduction of pillaromycinone. A preliminary study indicated that the monobromoacetate of pillarone was the most promising for this purpose.

### Experimental

**Preparation of the Crystal**—To a solution of 100 mg of pillarone in 2 ml of benzene was added dropwise 0.2 ml of monobromoacetic acid anhydride. After being stirred for 6 hr at 70° to complete the reaction, the mixture was poured into ice water and extracted with benzene. The benzene solution was washed with water and dried over anhydrous sodium sulfate. The solution was concentrated under a reduced pressure, petroleum ether was added and crystals separated were recrystallized several times from a mixture of ethyl acetate, ethyl ether and petroleum ether to afford yellow prisms, mp 194°. *Anal.* Calcd. for  $C_{22}H_{19}O_6Br$ : C, 57.53; H, 4.17; Br, 17.40. Found: C, 57.56; H, 4.14; Br, 17.37. IR  $\nu_{max}^{KBr}$   $cm^{-1}$ : 1772 (s), 1630 (vs). UV  $\lambda_{max}^{EtOH}$   $m\mu$  ( $\epsilon \times 10^{-3}$ ): 270 (37), 290 (6.8), 300 (6.8), 395 (6.5).

**Measurements of X-ray Diffraction**—Oscillation and Weissenberg photographs were taken about the two principal crystal axes, using  $CuK\alpha$  radiation. The specimens used had the following cross-sections at right angles to the axis of rotation;  $0.2 \times 0.3$  mm for the  $c$  axis rotation,  $0.2 \times 0.2$  mm for the  $a$  axis. Intensities of reflections were visually estimated from multiple film integrating Weissenberg photographs (8 layers about the  $c$  axis and 12 layers about the  $a$  axis, respectively) and obtained 2000 independent data (excluding zero intensity reflections). The linear absorption coefficient for  $CuK\alpha$  radiation was estimated to be  $\mu = 21.84$   $cm^{-1}$ . Thus the absorption and extinction errors were neglected. The absolute structure factors were deduced from these intensity data by the ordinary procedure, using a NEAC 2206 electronic computer.



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- 2) Location: a) *Juso, Higashiyodogawa-ku, Osaka*; b) *Uegahara, Nishinomiya*.
- 3) M. Asai, *Chem. Pharm. Bull.* (Tokyo), **18**, 1699 (1970).
- 4) Preliminary report, H. Matsuda, M. Nishikawa, K. Kamiya, M. Asai, Y. Tomiie, I. Nitta, S. Yamamura, Y. Hirata; presented at 9th Symposium on the Chemistry of Natural Products, Osaka, Oct., 1965.

## Result and Discussion

### Unit Cell and Space Group

It has been found from X-ray examination that the crystal belongs to the orthorhombic system with the unit cell parameters,  $a=12.46$ ,  $b=19.54$ ,  $c=7.85$  Å.

The space group was uniquely determined to be  $P2_12_12_1-D_2^2$ . Judging from the cell size and the space group, it is clear that the unit cell contains four molecules as the calculated density becomes a reasonable value,  $1.596$  g cm<sup>-3</sup> ( $D_0=1.591$  g cm<sup>-3</sup>).

### Determination of the Structure

As the first step of the structure analysis, a three-dimensional Patterson function  $P(uvw)$  was calculated. The coordinates of the bromine atom were determined without ambiguity from the three Harker sections, one of which at  $v=1/2$  is shown in Fig. 1. With phase angles calculated from these coordinates, a three-dimensional Fourier synthesis was calculated. In co-operation with this analysis, a minimum function diagram was obtained with four superposition of bromine vectors. In spite of quite different principles, the results obtained by these two methods were fully consistent with each other and the positions of all light atoms except hydrogen were determined undoubtedly.

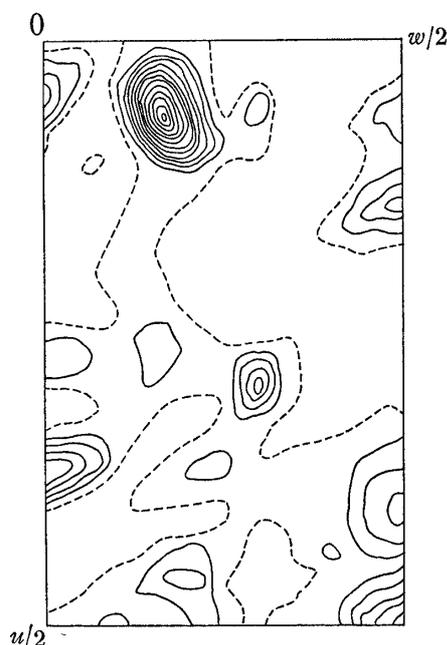


Fig. 1. Harker Section at  $v=1/2$

The refinement of the atomic positions was carried out using the least squares method, and in accordance,  $R$ -value fell to 0.133. Maps of the final electron density distribution are shown in Fig. 2a and 2b. The final  $x$ ,  $y$  and  $z$  atomic parameters and temperature factors are listed in Table I. The observed and the calculated structure factors are listed in Table II.

### Absolute Configuration

Of the 20 pairs of reflections for which the intensity differences between  $I(hkl)$  and  $I(\bar{h}\bar{k}\bar{l})$  were expected to be discernible, 14 pairs showed significant differences in the 1 $^{\text{th}}$  layer Weissenberg photographs. The results are shown in Table III. The comparison between observed and calculated intensities indicated that the assumed configuration was actually correct. The absolute configuration was then established as (I). In this paper, all figures are drawn in such a way that they represent the correct configuration.

### Molecular and Crystal Structure

The bond lengths and angles, calculated from the atomic parameters of Table I, are shown in Fig. 3 and Table IV. The mean values of bond lengths are 1.40 Å for aromatic carbon bond, 1.53 Å for C-C single bond, 1.42 Å for C-O single bond and 1.24 Å for C-O double bond. These values seem to be quite reasonable.

The rings A and B form a naphthalene skeleton and C(11), C(12), C(13), O(2), O(3) and O(4) atoms are nearly on this plane. The OH group attached to the ring B seems to form a strong six-membered ring hydrogen bond with the carbonyl group at C(11), the bond between O(3) and C(7) being somewhat contracted.

The juncture between the ring C, which is puckered at C(14), and the ring D, which is of a chair form, is in a *trans*-fusion. The atoms, O(5), O(6) and C(19) are axially bonded to

TABLE I. The Final Atomic Coordinates and Temperature Factors

Atom	$a/x$	$b/y$	$c/z$	$b$
Br	0.0266	0.8425	0.1649	4.65
O(1)	0.4733	0.5466	0.1030	4.82
O(2)	0.3614	0.0445	0.2810	3.31
O(3)	0.3253	0.9376	0.4792	3.31
O(4)	0.2314	0.1672	0.1442	4.29
O(5)	0.2292	0.3042	0.1965	3.21
O(6)	0.1911	0.4144	0.1084	3.57
C(1)	0.3982	0.9960	0.1536	3.01
C(2)	0.4367	0.0200	0.0084	4.00
C(3)	0.0372	0.0239	0.3782	4.32
C(4)	0.0541	0.0945	0.3987	3.93
C(5)	0.3808	0.9248	0.1877	2.78
C(6)	0.4038	0.8792	0.0531	3.03
C(7)	0.3439	0.8963	0.3464	3.10
C(8)	0.3837	0.8078	0.0728	3.79
C(9)	0.3234	0.8246	0.3578	2.69
C(10)	0.3463	0.7806	0.2125	3.33
C(11)	0.2179	0.2038	0.0177	3.73
C(12)	0.3176	0.7044	0.2200	3.63
C(13)	0.2447	0.2795	0.0254	2.85
C(14)	0.3275	0.6785	0.4036	3.09
C(15)	0.1352	0.7066	0.4950	3.37
C(16)	0.2950	0.6002	0.4262	3.19
C(17)	0.3636	0.3691	0.0524	3.28
C(18)	0.1780	0.5888	0.4020	3.53
C(19)	0.2720	0.3742	0.1797	3.38
C(20)	0.2969	0.3998	0.3591	4.52
C(21)	0.4323	0.0654	0.3921	3.32
C(22)	0.3769	0.1141	0.5059	4.91

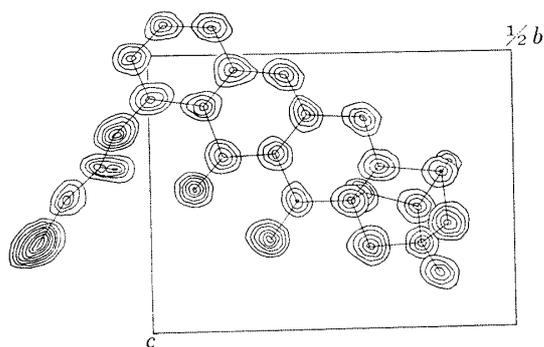


Fig. 2a. The Final Three-dimensional Electron Density Distribution shown by Means of Superimposed Contour Section Projected on (100)

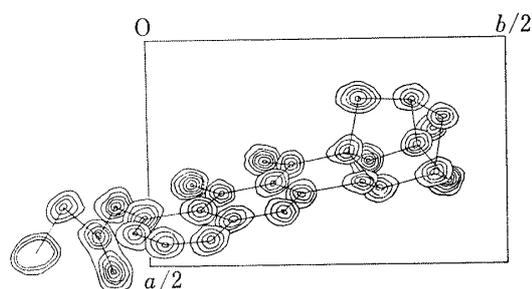


Fig. 2b. Projected on (001)

the ring D and form a ketal, leading to a cage structure of this part of the molecule. Thus considerable steric strains seem to exist at C(15) and C(18) so that the bond angles C(13)-C(15)-C(17) and C(16)-C(18)-C(17) are smaller than  $100^\circ$ .





TABLE III. Comparison of the Observed and Calculated Intensity Differences  
Used for the Establishment of Absolute Configuration

$h$	$k$	$l$	$F_c^2(h\bar{k}l)/F_c^2(hkl)$	$I_o(h\bar{k}l)/I_o(hkl)$	$h$	$k$	$l$	$F_c^2(h\bar{k}l)/F_c^2(hkl)$	$I_o(h\bar{k}l)/I_o(hkl)$
11	3	4	0.627	< 1	8	9	3	2.215	> 1
13	8	1	1.150	> 1	13	2	3	0.824	< 1
11	9	1	0.805	< 1	15	5	1	0.781	< 1
11	2	2	2.057	> 1	11	9	1	0.805	< 1
11	1	3	0.858	< 1	11	10	1	0.776	< 1
13	2	3	0.824	< 1	10	5	1	1.078	> 1
13	1	3	0.891	< 1	3	20	2	0.950	< 1

TABLE IV. Bond Angles

C(1)-O(2)-C(21)	116°	C(10)-C(12)-C(14)	109°
C(13)-O(5)-C(19)	100	O(5)-C(13)-C(11)	109
C(16)-O(6)-C(19)	102	O(5)-C(13)-C(14)	108
O(2)-C(1)-C(2)	118	O(5)-C(13)-C(15)	102
O(2)-C(1)-C(5)	124	C(11)-C(13)-C(14)	111
C(2)-C(1)-C(5)	124	C(11)-C(13)-C(15)	113
C(1)-C(2)-C(3)	120	C(14)-C(13)-C(15)	112
C(2)-C(3)-C(4)	120	C(12)-C(14)-C(13)	112
C(3)-C(4)-C(6)	121	C(12)-C(14)-C(16)	113
C(1)-C(5)-C(6)	116	C(13)-C(14)-C(16)	108
C(6)-C(5)-C(7)	117	C(13)-C(15)-C(17)	97
C(1)-C(5)-C(7)	126	O(6)-C(16)-C(14)	105
C(4)-C(6)-C(5)	118	O(6)-C(16)-C(18)	102
C(4)-C(6)-C(8)	121	C(14)-C(16)-C(18)	112
C(5)-C(6)-C(8)	120	C(15)-C(17)-C(18)	107
O(3)-C(7)-C(5)	119	C(15)-C(17)-C(19)	105
O(3)-C(7)-C(9)	121	C(18)-C(17)-C(19)	102
C(5)-C(7)-C(9)	120	C(16)-C(18)-C(17)	99
C(6)-C(8)-C(10)	124	O(5)-C(19)-O(6)	107
C(7)-C(9)-C(10)	119	O(5)-C(19)-C(17)	106
C(7)-C(9)-C(11)	120	O(5)-C(19)-C(20)	107
C(10)-C(9)-C(11)	121	O(6)-C(19)-C(17)	108
C(8)-C(10)-C(9)	119	O(6)-C(19)-C(20)	110
C(8)-C(10)-C(12)	121	C(17)-C(19)-C(20)	118
C(9)-C(10)-C(12)	120	O(1)-C(21)-O(2)	126
O(4)-C(11)-C(9)	120	O(1)-C(21)-C(22)	128
O(4)-C(11)-C(13)	121	O(2)-C(21)-C(22)	107
C(9)-C(11)-C(13)	119	Br-C(22)-C(21)	113

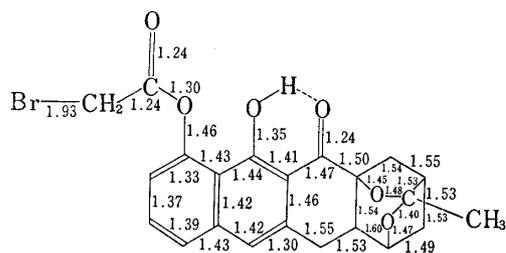


Fig. 3. Bond Distances

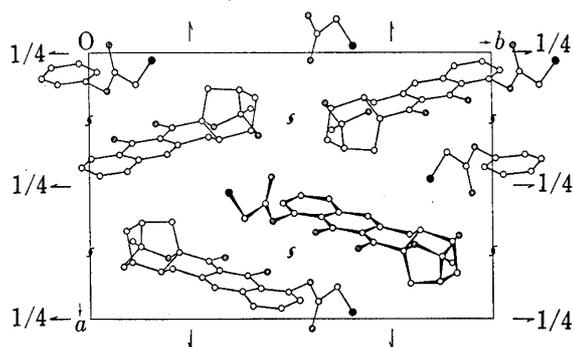


Fig. 4. Packing of the Structure Seen Down the  $c$  Axis

In Fig. 4 is shown a view of the structure along the  $c$  axis, together with the packing of the molecules in a unit cell. Since the molecule has no polar groups available for intermolecular hydrogen bonds, the association of the molecules seems to be mainly due to the van der Waals interactions.

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