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Studies on Benzodiazepinooxazoles. II.<sup>1)</sup> Crystal and Molecular Structure of  
10-Bromo-2,3,5,6,7,11b-hexahydro-2-methyl-11b-phenylbenzo[6,7]-1,4-  
diazepino[5,4-*b*]oxazol-6-one, C<sub>18</sub>H<sub>17</sub>O<sub>2</sub>N<sub>2</sub>Br·C<sub>2</sub>H<sub>5</sub>OH

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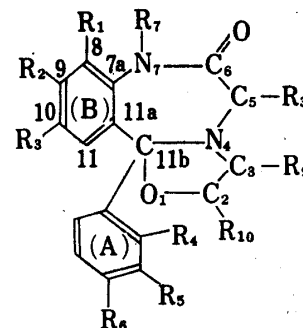
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The crystal structure of 10-bromo-2,3,5,6,7,11b-hexahydro-2-methyl-11b-phenylbenzo[6,7]-1,4-diazepino[5,4-*b*]oxazol-6-one, C<sub>18</sub>H<sub>17</sub>O<sub>2</sub>N<sub>2</sub>Br·C<sub>2</sub>H<sub>5</sub>OH, has been determined by the three dimensional Patterson method. This substance occurs as monoclinic crystals, space group P2<sub>1</sub>/c, with Z=4 in the unit cell dimensions a=16.34, b=8.88, c=13.89 Å, and β=102°. The block-diagonal least-squares refinements using 2780 reflections gave the final R=11.3%. The methyl and phenyl ring substituents are trans with respect to the plane of the oxazole ring.

### Introduction

In the preceding paper,<sup>1)</sup> the authors reported the synthesis and the pharmacology activity of benzo[6,7]-1,4-diazepino[5,4-*b*]oxazole derivatives and their analogues.

In these derivatives, compounds having substituents at 2- or 3-position can exist as two configurational isomers. The nuclear magnetic resonance (NMR) spectrum of 10-bromo-2,3,5,6,7,11b-hexahydro-2-methyl-11b-phenylbenzo[6,7]-1,4-diazepino[5,4-*b*]oxazol-6-one (R<sub>1</sub>=R<sub>2</sub>=R<sub>4</sub>=R<sub>5</sub>=R<sub>6</sub>=R<sub>7</sub>=R<sub>8</sub>=R<sub>9</sub>=H, R<sub>3</sub>=Br and R<sub>10</sub>=CH<sub>3</sub>) in deuteriochloroform showed two pairs of methyl protons at 1.25 (*J*=6.5 Hz) and 1.38 ppm (*J*=6.5 Hz) and two kinds of amino-protons at 8.32 and 8.53 ppm suggesting the existence of such isomers. After the compound was recrystallized from aqueous ethanol, the NMR spectrum of the crystals showed that the only one of the crystalline isomers had separated. Such crystals would be suitable for determining its structure and stereochemistry. The compound was labile in solution and a peak assigned to amino-proton splitted to the above values within a few minutes.



### Experimental

We have succeeded in obtaining crystals which belong to the monoclinic system by slow evaporation from 70% aqueous ethanol in the form of needles elongated along the *c*-axis. The crystals consisted of only one component isomer at the NMR level. The cell constants are determined by Weissenberg and oscillation photographs: a=16.34, b=8.88, c=13.89 Å, and β=102.0°. From the systematic absence of reflection, the space group was determined to be P2<sub>1</sub>/c. Intensity data are collected by an equi-inclination Weissenberg technique for the reflections of *hk*0 to *hk*11. After several days under the X-ray beam the crystal turned opaque so that two crystals were used to take a total of 2780 independent reflections. They

- 1) This paper forms part II of a Series entitled "Anxiolytic Sedative". For previous paper, see Ref. 1. Part I: T. Miyadera, A. Terada, M. Fukunaga, Y. Kawano, T. Kamioka, C. Tamura, H. Takagi and
- 2) Location: No. 2-58, 1-Chome, Hiromachi, Shinagawa-ku, Tokyo. R. Tachikawa, *J. Med. Chem.*, 14, 520 (1971).

were estimated by visual comparison of the calibrated standard film. Two sets of five multiple film were exposed until  $l=4$ , and one set of five for over  $l=5$ . The observed density was  $1.452 \text{ g cm}^{-3}$  measured by floatation method while the calculated value was  $1.451 \text{ g cm}^{-3}$  which was obtained by taking into account that four molecules of  $\text{C}_{18}\text{H}_{17}\text{O}_2\text{N}_2\text{Br}\cdot\text{C}_2\text{H}_5\text{OH}$  are contained in the unit cell dimensions. Lorentz and Polarization corrections were calculated but no absorption correction was made because of the small size of the crystals.

### Structure Analysis

The structure was determined by the three dimensional heavy atom method, using Sim's weighting scheme.<sup>3)</sup> The position of the bromine atom was readily obtained on the basis of the  $\text{P2}_1/\text{c}$  vector sets in the crystal. The first Fourier synthesis were calculated with phases derived from the contribution of the bromine atom alone ( $R=46.9\%$ ). The 22 peaks in the electron density map were corresponded to the proposed structure. The values of the temperature parameters in the early stages of structure factor calculation were used 3.5 for bromine and 4.0 for all light atoms. The whole atomic coordinates of the molecule were determined in the next cycle of Fourier calculation and refined by means of block diagonal least squares procedures. After 6 cycles of refinements with isotropic temperature parameters for whole atoms the R-factor reduced to 15%. At this stage the molecule of ethanol which is contained as the solvent of crystallization came out in the vacant areas of the molecular packing in the electron density map. In order to find out the positions of hydrogen atoms, the difference Fourier synthesis was computed. All hydrogen atoms were obtained around the expected positions.

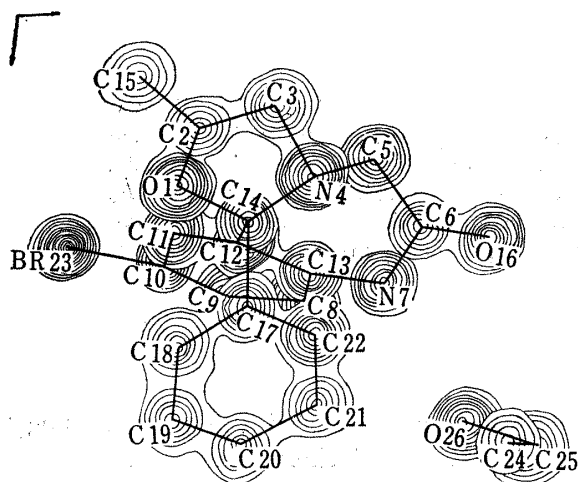


Fig. 1. The Final Electron Density Map Projected on the (010) Plane, Contours are Arbitrary in Equal Intervals except Bromine Atom

By two more cycles of the block diagonal least squares refinements with anisotropic thermal parameters for bromine the R-factor decreased to 12.4%, further refinements with anisotropic thermal parameters for all atoms except hydrogen reduced the R to 11.3%. The final Fourier map and the molecular packing diagram viewed along the b-axis are shown in Fig. 1 and 3, and the bond distances and angles are shown in Fig. 2. The least squares planes for two benzene, oxazole, and diazepine ring systems are calculated as Fig. 4 which show the deviations of atomic coordinates from these planes. The final parameters of the atomic positions and the temperature factors are listed in Table I, and their standard deviations in Table II. The final values of  $F_o-F_c$  are shown in Table III.

All computations have been carried out with the CDC 3600 system at C. Itoh Electronic Computing Service Co., Ltd. using the set of programs registered as UNICS.

### Result and Discussion

The C15 methyl and C17 phenyl groups which are attached to the oxazole ring deviate  $0.30$  and  $-1.44 \text{ \AA}$  respectively from the five membered ring least squares plane. Therefore, the stereo-chemistry of the substituents on the oxazole ring is determined as *trans* for the methyl group and the benzene ring B. The planarity calculation of the C3, N4, C5 and C14, shows that the lone pair of the N4 atom is directed towards the same side of the C17 phenyl group that is the lone pair of nitrogen and phenyl group are *cis*. The least squares plane in the seven membered ring is constructed with four atoms of C14, N4, C6 and N7. The C12, C13 and C5 atoms in this ring deviate  $+0.88$ ,  $+0.81$  and  $+0.75 \text{ \AA}$  from this least squares plane respectively. The deviations of all plus signs indicate that the ring system takes a boat form, this is a common feature found in the compounds containing unsaturated seven membered rings, *e.g.*, fulvene,<sup>4)</sup>

3) G.A. Sim, *Acta Cryst.*, **11**, 123 (1958).

4) R.E. Davio and A. Tulinsky, *J. Am. Chem. Soc.*, **88**, 4583 (1966); H. Shimanouchi, T. Hata and Y. Sasada, *Tetrahedron Letters*, **1968**, 3573.

TABLE I. Fractional Atomic Coordinates and Anisotropic Thermal Parameter. Anisotropic Temperature Factor:  $T = \exp[-(B_{11}h^2 + B_{22}k^2 + B_{33}l^2 + B_{12}hk + B_{13}hl + B_{23}kl)]$   
Each Thermal Parameter is multiplied by  $10^4$

Atom	X	Y	Z	B11	B22	B33	B12	B13	B23
O 1	.1412	.8894	.0525	19	55	15	0	0	14
C 2	.1561	1.0118	-.0110	34	66	36	3	-3	22
C 3	.2347	.9687	-.0411	26	81	21	0	2	12
N 4	.2830	.8990	.0529	20	52	8	-4	4	2
C 5	.3432	.7885	.0308	24	70	18	0	7	1
C 6	.4033	.7513	.1296	14	55	18	1	6	0
N 7	.3723	.6662	.1964	19	55	18	0	2	3
C 8	.2877	.4474	.2075	23	52	21	0	5	2
C 9	.2118	.3732	.1898	21	57	31	-2	3	-2
C10	.1422	.4504	.1427	18	65	17	-12	3	2
C11	.1442	.6033	.1152	16	58	11	-5	1	1
C12	.2202	.6792	.1326	21	44	12	-4	3	0
C13	.2926	.5986	.1776	15	57	7	-2	3	-4
C14	.2208	.8501	.1118	17	50	14	2	6	4
C15	.0811	1.0371	-.0956	30	142	42	3	0	25
O16	.4765	.7941	.1473	19	75	41	-3	4	10
C17	.2331	.9337	.2139	23	44	13	2	2	2
C18	.1675	.9422	.2659	31	74	31	-4	11	-2
C19	.1765	1.0170	.3570	37	102	36	2	15	-5
C20	.2526	1.0824	.3990	48	88	21	9	3	-7
C21	.3190	1.0764	.3487	35	60	26	4	-2	-2
C22	.3091	1.0031	.2571	24	69	18	3	1	-3
BR23	.0376	.3457	.1163	26	91	51	-29	-3	5
C24	.5253	.8476	.4091	36	108	45	-2	12	5
C25	.5509	.6844	.4192	28	99	48	4	7	-12
O26	.4806	.5906	.3836	22	69	32	3	1	0

TABLE II. Estimated Standard Deviations of the Fractional (Å) and Thermal Parameters. Each Thermal Parameter is multiplied by  $10^4$

Atom	X	Y	Z	B11	B22	B33	B12	B13	B23
O 1	.0061	.0057	.0065	2	7	4	3	2	3
O 2	.0111	.0104	.0119	4	13	8	6	4	7
O 3	.0100	.0095	.0105	3	11	7	5	3	6
N 4	.0073	.0068	.0076	2	8	5	3	2	4
O 5	.0095	.0090	.0102	3	11	6	4	3	6
O 6	.0083	.0080	.0098	3	9	6	4	3	5
N 7	.0074	.0070	.0081	2	8	5	3	2	5
O 8	.0092	.0084	.0101	3	10	6	4	3	5
O 9	.0094	.0086	.0107	3	10	7	4	3	6
O10	.0086	.0084	.0096	3	10	6	4	3	5
O11	.0084	.0081	.0092	3	9	6	4	3	5
O12	.0086	.0079	.0092	3	9	6	4	3	5
O13	.0082	.0079	.0090	3	9	5	4	3	5
O14	.0082	.0080	.0091	3	9	6	4	3	5
O15	.0119	.0125	.0132	4	17	9	7	4	9
O16	.0066	.0066	.0076	2	8	5	3	2	4
O17	.0090	.0079	.0094	3	9	6	4	3	5
O18	.0106	.0097	.0114	4	11	7	5	4	6
O19	.0119	.0111	.0125	4	14	7	5	4	8
O20	.0125	.0108	.0118	5	13	7	6	4	7
O21	.0109	.0093	.0110	4	11	7	5	4	6
O22	.0094	.0089	.0101	3	10	6	4	3	6
BR23	.0011	.0011	.0013	0	1	0	0	0	0
O24	.0125	.0117	.0134	5	15	9	6	4	8
O25	.0114	.0111	.0130	4	14	8	6	4	8
O26	.0066	.0066	.0073	2	8	5	3	2	4









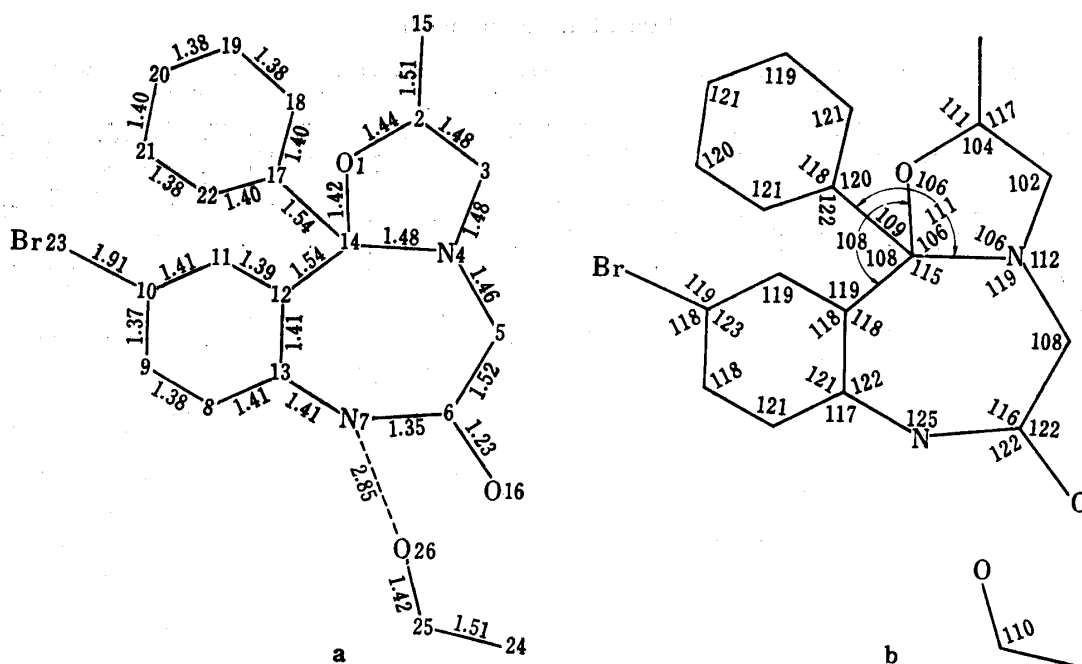


Fig. 2. (a) Bond Distances (b) Bond Angles

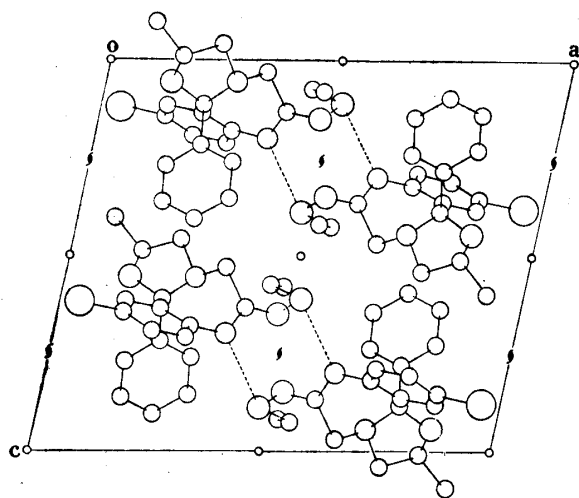


Fig. 3. The Crystal Structure as Viewed along the b Axis

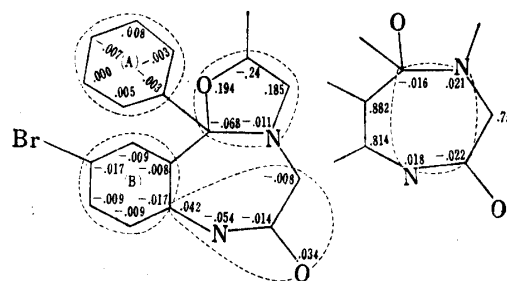


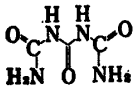
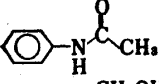
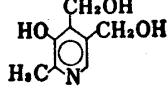
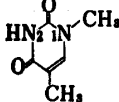
Fig. 4. Deviation in Ångstrom Units from Least-Squares Planes

saturated carbon atoms in a pyramidal form. The bond lengths of C3-N4, C5-N4 and C14-N4 are normal, and their values are 1.48, 1.46 and 1.48 Å. The angles of C3-N4-C14, C3-N4-C5 and C5-N4-C14 are 106.2, 111.5 and 118.6° respectively. On the other hand, the N7 atom is bonded to a carbonyl group and a benzene ring; the bond lengths of C6-N7 and C13-N7, 1.35 and 1.41 Å respectively are shorter than the normal C-N single bond of 1.47 Å. Shorter carbon-nitrogen bond lengths have been pointed out by Karle<sup>5)</sup> such that nitrogen atoms adjacent to C=S, C=O, and C=C bonds, tend to be in a planar configuration with short C-N bond lengths of in the range 1.34–1.40 Å. The bond angle for C6-N7-C13 is 125.3°. Several molecules containing such systems have usually larger bond angles, some examples of these compounds are shown in Table IV. The value of the mean bond lengths of the benzene rings is 1.39 Å which is normal. The deviations of the atoms from the benzene least squares planes are within 0.01 Å for benzene ring A and 0.02 Å for ring B. The intermolecular contacts within 3.7 Å are given in Table V. They are



within the range of normal van der Waals radii. Intermolecular hydrogen bonds are observed in  $N7 \cdots O26 = 2.85 \text{ \AA}$  and  $O26 \cdots O16 = 2.78 \text{ \AA}$ .

TABLE IV. Comparison of  $Csp^2-N-Csp^2$  Bond Angles

Compound	Angle	Reference
	125.3	This work
	128.4	H. Ringertz (1966) <sup>a)</sup>
	127.5	
		
	127.6	B.F. Pedersen (1967) <sup>b)</sup>
	124.5	F. Hanic (1966) <sup>c)</sup>
	1; 120.6 2; 126.3	K. Hoogsteen (1963) <sup>d)</sup>

a) H. Ringertz, *Acta Cryst.*, **20**, 932 (1966)

b) B.F. Pederson, *Acta Chem. Scand.*, **21**, 1415 (1967)

c) F. Hanic, *Acta Cryst.*, **21**, 332 (1966)

d) K. Hoogsteen, *Acta Cryst.*, **16**, 28 (1963)

TABLE V. Intermolecular Contacts within 3.7 Å

O26...O26(I)	3.47	C21...C 9(III)	3.61
N 4...C25(II)	3.67	C22...C 9(III)	3.68
C 6...O26(II)	3.58	C25...O16(IV)	3.54
O16...C25(II)	3.58	O26...C 5(IV)	3.46

I:  $1-x, 1-y, 1-z$

II:  $1-x, 1/2+y, 1/2-z$

III:  $x, 3/1+y, z$

IV:  $x, 3/2-y, 1/2+z$