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## Synthesis of 2,5-Di-t-butylpyrrole and of 2,3,5-Tri-t-butylpyrrole

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WE report two syntheses of hindered pyrrole derivatives, one of which is another1 example of direct introduction of a t-butyl group adjacent to one already present in an aromatic system. 2,5-Di-t-butylpyrrole<sup>2</sup> (DBP) (m.p. 32-33°) is prepared in 93% yield by heating 2,2,7,7-tetramethyloctane-3,6-dione3 under reflux with ammonium acetate in acetic acid.4 2,3,5-Tri-t-butylpyrrole<sup>2</sup> (TBP) (m.p. 46-47°) is formed in 72% yield (n.m.r. analysis) by alkylation of DBP with t-butyl chloride in carbon disulphide, using tin(IV) chloride as a catalyst (the yield is slightly lower with aluminium chloride). It can be separated from the di-t-butyl compound by chromatography on alumina and recrystallisation from methanol, or isolated in ca. 20% yield by direct recrystallisation of the reaction product from methanol.

The u.v. and i.r. spectra of the two compounds provide evidence for the substituted pyrrole structure: in cyclohexane, DBP has  $\lambda_{\text{max}}$  218 m $\mu$ 

 $(\epsilon 10,000)$ , shoulder at 289 m $\mu$  ( $\epsilon 20$ ), and TBP has  $\lambda_{\text{max}}$  218 m $\mu$  ( $\epsilon$  9600) shoulder at 283 m $\mu$  ( $\epsilon$  33). DBP (pure liquid) has  $v_{max}$  3500 (NH), 3000 (CH), 1380, 1360, 1260, 1200 (t-butyl groups), 1580 (ring stretching), 1045 (ring breathing), 765, and 710 cm. $^{-1}$  (out-of-plane). TBP (in nujol) has  $\nu_{max}$ 3500, 1355, 1245, 1195, 1575, 1045, 795, and 690 cm<sup>-1</sup>).

The n.m.r. spectra give the position of the t-butyl groups. DBP (in CDCl3) has a sharp peak at  $\tau$  8.75, a doublet (J 2.7 c./sec.) at  $\tau$  4.28 [J(13C-H) 167 c./sec.], and a broader signal at  $\tau 2.5$  (intensity ratio: 18:2:1). TBP (in CCl<sub>4</sub>) has three peaks at  $\tau$  8.77, 8.68, 8.60, a doublet (J 3.5 c./sec.) at  $\tau$  4.37 [ $I(^{13}C-H)$  165.5 c./sec.] and a broader signal at  $\tau = 2.5$  (Intensity ratio: 9:9:9:1:1).

Both compounds undergo spontaneous oxidation at room temperature, more rapidly in solution than in the solid state.

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- <sup>1</sup> H. Winberg and U. E. Wiersum, Chem. Comm., 1965, 1.
- <sup>2</sup> Satisfactory microanalysis has been obtained for this compound.
- <sup>3</sup> R. Ramasseul and A. Rassat, Bull. Soc. chim. France, 1963, 2214.
- <sup>4</sup> T. Ajello and S. Cusmano, Gazzetta, 1939, 69, 207.
- <sup>5</sup> In pyrrole derivatives,  $\alpha$  and  $\beta$ -protons show  $\tau$  3·35  $\pm$  0·3 and 4  $\pm$  0·3 respectively, <sup>6</sup> and J(<sup>13</sup>C–H) 183 and 169 c./sec. respectively (ref. 7).

  - <sup>6</sup> cf. Varian N.M.R. spectra catalogue.
    <sup>7</sup> K. Tori and T. Nakagawa, J. Phys. Chem., 1964, 68, 3163.