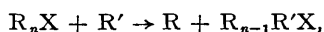


Interaction of Free Radicals and Triethylborine

J. GROTEWOLD and E. A. LISSI

(Departamento de Química Inorgánica, Analítica y Química Física, Facultad de Ciencias Exactas y Naturales, Universidad de Buenos Aires, Argentina)

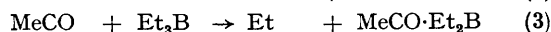
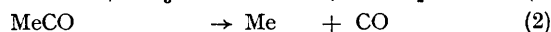
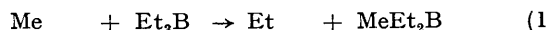
LITTLE is known about such reactions as



where R is an alkyl radical and R' is either an alkyl radical or a halogen atom. The only such reaction so far investigated¹ is that between methyl radicals and dimethylmercury. We have investigated the interaction of free radicals with alkyl-boron compounds by gas-phase photolysis of azomethane, and biacetyl in the presence of triethylborine. Triethylborine did not give volatile material when photolysed alone or when heated in the dark below 90° with the other reactants. The volatile products were:

- (a) With acetone; ethane, ethylene, propane, and butane.
- (b) With azomethane; ethane, propane, and small amounts of butane.
- (c) With biacetyl; (i) at low temperature, ethane, butane, and small amounts of uncondensables; (ii) at 50°, the same products and small amounts of propane.

These products are compatible with the following reactions:



We could not find any reference of compounds containing acetyl-boron bonds in the literature, but they could be postulated, at least as intermediates, because the B-C bond would be stabilized by the double bond located in the α position to the boron atom.

The average value for the disproportionation ratio found from runs with biacetyl at room temperature was 0.12 (from the runs with acetone a slightly higher value was obtained). This agreement with the accepted value not only shows that the postulated mechanism is in the main correct, but that the system is a potentially useful secondary source of radicals.

¹ R. E. Rebert and P. Ausloos, *J. Amer. Chem. Soc.*, 1963, **85**, 3086.

If we take the value for k_2 given by E. O'Neal and S. W. Benson,² k_3 is $= 1.4 \times 10^9$ c.c. sec.⁻¹ mole⁻¹, which would give a very low energy of activation (lower than 3 kcal, if 10^{11} c.c. sec.⁻¹ mole⁻¹ is taken for the A factor). The attack with a methyl radical (reaction 1) seems to be 100 times slower at room temperature.

Reaction 5' is faster than similar hydrogen abstraction from alkyl groups in other compounds.³ At room temperature $k_4 = 5 \times 10^6$ c.c. sec.⁻¹ mole⁻¹. This points to a particularly active hydrogen atom (probably that α to the boron atom).

(Received, December 3rd, 1964.)

² E. O'Neal and S. W. Benson, *J. Chem. Phys.*, 1962, **36**, 2196.

³ A. F. Trotman-Dickenson, "Progress in Reaction Kinetics," Pergamon Press, Vol. 1, 1960, p. 107.