## Isolation of the Intermediate in a Heterocyclic Ring Expansion Reaction

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When methyl-lithium is added to indole in methylene chloride solution the indole undergoes ring expansion to quinoline.1 Quinoline reacts rapidly with methyl-lithium in ether solution to give 1,2-dihydro-2-methylquinoline. Yet, when indole (1 mole) is allowed to react with excess of methyl-lithium in ether solution in the presence of a limited quantity of methylene chloride (1 mole), for a prolonged period, a mixture of quinoline and 1,2-dihydro-2-methylquinoline is formed. The presence of quinoline in the final product indicates the existence of an intermediate, that does not react readily with methyl-lithium, and is converted into quinoline in the final stage of the reaction when the excess of methyl-lithium is decomposed.

were found to be 1:1.8 respectively, by using a radiotracer technique. This confirms that the reactive species is an electrophile.

3-Methylindole was treated with methyl-lithium and carbon-14-labelled methylene chloride. Acetyl chloride was added, and the excess finally decomposed with water. From the reaction mixture a radioactive oil, that was insoluble in acid, was isolated. On hydrolysis with alcoholic potassium hydroxide the oil was shown by dilution analysis to yield 4-methylquinoline. Purification of the neutral compound by column chromatography yielded an amber oil which showed well-defined absorption peaks at 860 and 1020 cm.-1, providing strong evidence for a cyclopropane structure (II).

Labelled methylene chloride and methylene

It has been postulated<sup>1,2</sup> that the ring expansion proceeds via attack of chlorocarbene on the 2,3double bond of indole with the formation of a cyclopropane intermediate (I). The relative, competitive reaction rates of indole and 3-methylindole with methyl-lithium and methylene bromide

bromide both gave radioactive neutral compounds when treated with indole, methyl-lithium, and acetyl chloride. On hydrolysis both compounds were converted into radioactive quinoline.

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<sup>&</sup>lt;sup>1</sup> G. L. Closs and G. M. Schwartz, J. Amer. Chem. Soc., 1961, 26, 2609.

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