The Synthesis of (\pm) -Prostaglandins E_2 , $F_{2\alpha}$, and $F_{2\beta}$

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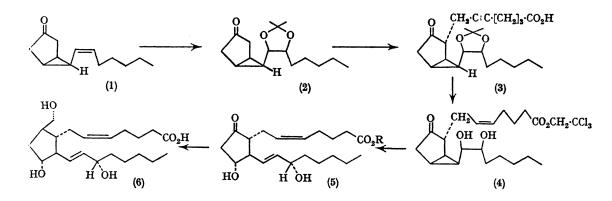
PROSTAGLANDIN E₂ (PGE₂) (5; R = H) and F₂ α (PGF₂ α) (6) are recognized as the most commonly occurring natural prostaglandins,¹ and show both qualitative and quantitative differences in biological activity² from their 5,6-dihydroanalogue, PGE₁ and PGF_{1 α}. Total syntheses of (±)-PGE₁ and (±)-PGF_{1 α} have recently been reported.³ This report details a synthesis of (±)-PGE₂, and PGF_{2 α}, and PGF_{2 β} based on 6-endo-substituted bicyclo[3,1,0]hexane intermediates described in an accompanying communication.^{3f}

6-endo-(1'-Heptenyl)bicyclo[3,1,0]hexane-3-one^{3f} (1) was hydroxylated with osmium tetroxide and the resulting vicglycol was converted into its acetonide (2). This was alkylated with 1-bromo-7-tetrahydropyranyloxyhept-2-yne[†] by means of potassium t-butoxide in tetrahydrofuran. Selective hydrolysis of the tetrahydropyranyl ether with oxalic acid in methanol and oxidation of the resulting primary alcohol by Jones' reagent⁴ gave the carboxylic acid (3). The triple bond was hydrogenated to a *cis*-double bond using palladium on barium sulphate catalyst in pyridine. Removal of the acetonide grouping in aqueous tetrahydrofuran containing hydrochloric acid, followed by esterification of the carboxy-group with trichloroethanol gave compound (4) as a mixture of two isomeric racemates. This was converted into the bis-mesylates which were solvolysed in aqueous acetone to produce the trichloroethyl esters of (\pm)-prostaglandin E₂ (5; R = CH₂·CCl₃) and its 15epimer, each in about 15% yield, easily separable by silica gel chromatography.[‡] Treatment of the more polar of these with zinc in acetic acid removed the trichloroethyl ester group,⁵ producing (\pm) -PGE₂ (5; R = H), which gave n.m.r. (in CDCl₃), i.r. (in CH₂Cl₂), and mass spectra identical with those of natural PGE₂. Treatment with base gave a material

 \dagger This was prepared by C-alkylation of the dilithium salt of prop-2-yn-1-ol with 1-bromo-4-tetrahydropyranyloxybutane, mesylation of the resulting propargylic alcohol and displacement of mesylate by lithium bromide. This sequence was previously carried out by G. Just and E. S. Ferdinandi, private communication to J. E. Pike.

The above intermediates were all noncrystalline, and were characterized by n.m.r., i.r., and mass spectra, and by g.l.c. and t.l.c.

showing u.v. absorption at 278 nm. (ϵ 26,000) and t.l.c. mobility identical with that of authentic PGB₂. The mobility and colour reactions (with vanillin-phosphoric acid spray) of synthetic and natural PGE₂ were identical on both silica gel and silver nitrate-impregnated silica gel thin-layer plates (AIX system⁶), as was the g.l.c. analysis of the ratio. These were separated on acid-washed silica gel and the more polar (±)-PGF_{2 β} fractions easily crystallized, m.p. 90-92° after recrystallization from ethyl acetate These products were identical with those obtained from natural PGE, by t.l.c. on several systems, including silver nitrateimpregnated plates, by mass spectra, and by g.l.c. analysis



derived methoxime, trimethylsilyl derivatives.§ These g.l.c. peaks from both natural and synthetic PGE₂ were collected and showed the same mass-spectral fragmentation patterns and molecular-ion peaks (597). The synthetic (\pm) -PGE₂ showed at least 50% of the biological activity of natural PGE₂ in its effects on blood pressure in rats and on contraction of smooth muscle.¶

of their trimethylsilyl derivatives. The (\pm) -PGF_{2a} had at least 50% the activity of natural $\mathrm{PGF}_{2\alpha}$ in its affect on blood pressure and smooth muscle.¶

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Reduction of (\pm) -PGE₂ with sodium borohydride gave (\pm) -PGF₂₀ (6) and its 9-epimer, (\pm) -PGF₂₆, in about a 45:55

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 \P The biological assays were performed by Dr. J. R. Weeks and associates of the Experimental Biology Unit of the Upjohn Company.

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⁶ M. Hamberg and B. Samuelsson, J. Biol. Chem., 241, 257, (1966).