

The High-temperature Hydration of Norbornadiene¹

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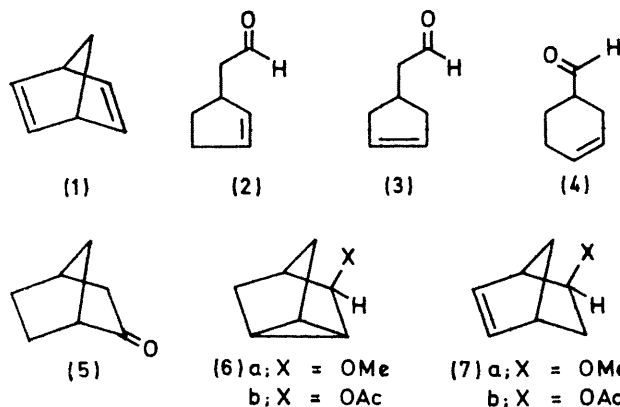
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Summary The high-temperature hydration of norbornadiene at 250° in 10% HOAc-H₂O initially yields norbornenols, nortricyclanol, and norbornanediols which are gradually converted into cyclopent-2- and -3-ene-acetaldehydes, cyclohex-3-enecarboxaldehyde, and norcamphor.

THE acid-catalysed additions of methanol and acetic acid² to norbornadiene (1) at 25→230° yield a mixture of (6a), (7a) and (6b), (7b), respectively. Norbornadiene, in acetic acid-BF₃, yields a mixture of norbornyl diacetates.³

We have found that the hydration of (1) in 10% HOAc-H₂O at 250° *initially* yields norbornenols, nortricyclanol, and norbornanediols, which are gradually converted into (2), (3), (4), and (5). The results of a preliminary study are shown in the Table.

That the reaction is acid-catalysed is shown by entries i and v. Entries i and iv clearly show that (1) rapidly gives norbornenols, nortricyclanol, and norbornanediols, which are converted into (2), (3), (4), and (5) *via* rearrangement⁴ and cleavage processes.⁵ The preparation of (4), (5)



and especially (2) and (3)⁶ in one step from (1) and aqueous acetic acid is remarkable. When nortricyclanol is heated in 10% HOAc-H₂O at 250° (2), (3), (4), and (5)⁷ are formed.

Norcamphor is most likely formed (a) *via* enols produced by 1,2- and 1,3-eliminations of a proton or H₂O and (b) from

appropriate hydroxy-substituted cations by a 2,3- or 2,6-hydrogen shift.⁴ Cleavage of a 1,6- or a 1,7-bond in suitable hydroxy-cations would yield the aldehydes.⁵

some other mechanism which involves HOAc-AcO⁻. That is, the hydrogens are exchanged prior to formation of the final products by a protonation-deprotonation pathway[‡]

Data for the hydration of norbornadiene at 250°

Entry ^a	Time (h)	Rel. % yield ^b					% Yield ^d		D/molecule ^e		
		(2) + (3)	(4)	(5)	Diols	Other ^c	(2) + (3) + (4)	(5)	(2) + (3)	(4)	(5)
i	0.5	4	3	8	55	30					
ii	1.5	15	18	40	20	7					
iii	3	17	19	45	15	4					
iv	9.5	14	17	54	14	1	21	33	2.5	2.6	2.4
v ^f	10	norbornadiene + polymer									

^a Typically, a solution of (1) in 10% (v/v) HOAc-H₂O or 10% (v/v) DOAc-D₂O was de-gassed and heated in a sealed tube.

^b The relative percent yields were determined by g.l.c. on $\frac{1}{8}$ in. \times 10 ft. 10% SE-30 and 10% Carbowax on Chromosorb W columns.

^c The norbornenols and nortricyclanol are included here.

^d The yields were determined by preparative g.l.c. Each is corrected for collection losses.

^e The enolizable deuterium in (2), (3), and (4) was washed out by three injections through a $\frac{3}{8}$ in. \times 10 ft. 5% KOH + 20% Carbowax column maintained at 170°. (5) was treated three times with a KOH-MeOH-H₂O solution.

^f The reaction was run in H₂O only.

When (1) is heated in deuteriated medium (entry iv), (2) and (3) combined, (4), and (5) contain 2.5, 2.6, and 2.4 nonenolizable deuterium atoms, each a composite of up to ²H₆-species. Control experiments have shown that (a) >95% of the deuterium must be incorporated into (2), (3), and (4), prior to cleavage and (b) deuterium is not incorporated into (5)[†] via acid-catalysed homoenolization⁸ or

which includes as steps 1,2-Wagner-Meerwein shifts, 2,6- and 2,3-hydrogen shifts and 1,2- and 1,3-eliminations of a proton.

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[†] Camphenilone also incorporates no deuterium on treatment with 10% DOAc-D₂O or 10% CF₃CO₂D-D₂O at 250° for 120 h.

[‡] Interestingly, when norbornene is heated in 10% DOAc-D₂O at 250° for 15 h, the resulting *exo*-2-, *endo*-2-norborneol mixture (80:20) contains 8.3 deuterium atoms per molecule, a composite of up to ²H₁₁ species. Experiments with other substrates are presently under way to determine whether or not these conditions are useful generally to prepare perdeuterio-materials.

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² S. J. Cristol, W. K. Seifert, D. W. Johnson and J. B. Jurale, *J. Amer. Chem. Soc.*, 1962, **84**, 3918; S. J. Cristol, T. C. Morrill, and R. A. Sanchez, *J. Org. Chem.*, 1966, **31**, 2726; 2733.

³ H. Krieger, *Suomen. Kem.*, 1962, **35B**, 127.

⁴ B. M. Benjamin and C. L. Collins, *J. Amer. Chem. Soc.*, 1966, **88**, 1556; J. A. Berson, J. H. Hammons, A. W. McRowe, R. G. Bergman, A. Remanick, and D. Houston, *ibid.*, 1967, **89**, 2590; A. W. Bushell and P. Wilder, *ibid.*, p. 5721.

⁵ S. J. Cristol and R. K. Bly, *J. Amer. Chem. Soc.*, 1960, **82**, 6155; D. E. Gwynn and L. Skillern, *Chem. Comm.*, 1968, 490; V. K. Jones and L. B. Jones, *Tetrahedron Letters*, 1970, 3171.

⁶ For synthesis of cyclopent-3-ene-ethanols, see E. L. Allred, J. Sonnenberg, and S. Winstein, *J. Amer. Chem. Soc.*, 1959, **81**, 5833; *J. Org. Chem.*, 1960, **25**, 26; R. G. Lawton, *J. Amer. Chem. Soc.*, 1961, **83**, 2399.

⁷ N. H. Werstiuk and T. Kadai, results to be published.

⁸ A. Nickon and J. L. Lambert, *J. Amer. Chem. Soc.*, 1966, **88**, 1905.