

Selenium Dithizonate

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It is unfortunate that Dr. Starý (preceding communication) does not refer to the evidence presented in our papers¹ that a yellow organic extract (which we showed to be due to the formation of a disulphide) can also be obtained in the complete *absence* of selenium and that apparently the same compound is produced by the notoriously easy oxidation of dithizone by Se^{IV} which is thereby reduced to Se⁰.

However, an element of ambiguity remains. It is not inconceivable that the yellow product of the interaction

of dithizone in an organic solvent with Se^{IV} in hydrochloric acid could contain both the disulphide and Se(HDz)₄ according to conditions and that Starý has succeeded in securing those in which no oxidation takes place.

All known primary metal dithizonates can readily be obtained as analytically pure solids. When it has been isolated in a pure state, the properties and structure of solid selenium(IV) dithizonate should reveal a new aspect of the chemistry of this element.

¹ R. S. Ramakrishna and H. M. N. H. Irving, *Chem. Comm.*, 1969, 1356; *Analyt. Chim. Acta*, 1970, 49, 9.