Syntheses of Naphtho[2,3-c]pyran-5,10-quinones using Ceric Ammonium Nitrate

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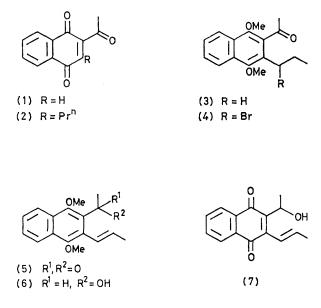
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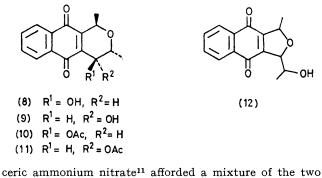
Summary Two 2-(alk-1-enyl)-3-(1-hydroxyalkyl)-1,4-dimethoxynaphthalenes are oxidised with silver(II) oxide to the corresponding naphthoquinones, while oxidation with cerium(IV) ammonium nitrate affords derivatives of the title compounds.

VARIOUS natural quinones containing the naphtho[2,3-c]pyran ring system have been identified. These include the eleutherins,^{1,2} ventilagone,³ the nanaomycins^{4,5} and related compounds,⁶ and the aphid pigments.⁷ We report here a convenient synthesis of the ring system.



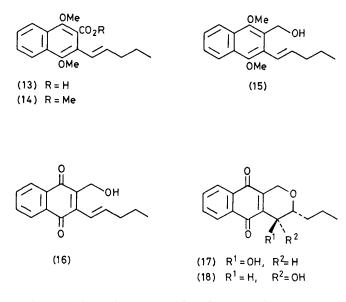
2-Acetyl-1,4-naphthoquinone $(1)^8$ was propylated with butanoic acid in the presence of silver nitrate and potassium peroxodisulphonate⁹ to afford the quinone (2)† (55%). Reductive methylation gave the naphthalene (3)† (100%), which was brominated with N-bromosuccinimide to give the benzylic bromide (4). This was converted into the *trans*-olefin (5)† [61% from (3)] with 1,5-diazabicyclo-[4.3.0]non-5-ene, and the ketone (5) was in turn reduced to the corresponding alcohol (6)† (96%) using sodium borohydride.

Oxidative demethylation¹⁰ by silver(II) oxide of the naphthalene dimethyl ether (6) provided as the only product the quinonoid alcohol (7) \dagger (59%), which did not undergo spontaneous cyclisation to a demethoxy-eleutherin. Alternatively, oxidation of the same dimethyl ether (6) with



ceric ammonium nitrate¹¹ afforded a mixture of the two naphthopyranquinones[‡] (8)[†] (20%) and (9)[†] (59%), 7,9dideoxy-analogues of quinones A and A'.¹² None of the quinone (7) was observed in the latter oxidation. Assignments (8) and (9) were distinguished from the alternative (12) on the basis of a comparison of the ¹H n.m.r. spectra of

† All new compounds gave satisfactory elemental analyses and their spectroscopic data were in accord with the assigned structures.‡ Compounds depicted as single enantiomers represent racemates.



the former pair and those of their respective acetates (10) and (11), † while individual assignments for compounds (8)and (9) were made from the distinctive coupling constants published for quinones A and A'.13

A second example supported the generality of the oxidative cyclisation. Thus the methyl ester (14)[†] of the trans-unsaturated acid $(13)^{14}$ was reduced to the alcohol (15)[†] (89%) with lithium aluminium hydride. Silver(II) oxide oxidation of (15) afforded only the quinonoid alcohol (16) \dagger (62%), whereas ceric ammonium nitrate oxidation of the alcohol (15) gave a mixture of the naphthopyranguinones^{\ddagger} (17)^{\dagger} (22%) and (18)^{\dagger} (32%). In this latter reaction some of the alcohol (16) (29%) was also observed.

The acid (13) did not yield readily identifiable products with either oxidant.

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- ¹ H. Schmid, A. Ebnöther, and Th. M. Meijer, *Helv. Chim. Acta.*, 1950, **33**, 1751. ² W. Eisenhuth and H. Schmid, *Helv. Chim. Acta*, 1958, **41**, 2021.

- ² W. Eisenhuth and H. Schmid, Helv. Chim. Acta, 1958, 41, 2021.
 ³ R. G. Cooke, A. Liu, C. L. Raston, and A. H. White, Aust. J. Chem., 1980, 33, 303.
 ⁴ S. Omaru, H. Tanaka, Y. Okada, and H. Marumo, J. Chem. Soc., Chem. Commun., 1976, 320.
 ⁵ T. Li and R. H. Ellison, J. Am. Chem. Soc., 1978, 100, 6263.
 ⁶ E.g. G. A. Ellestad, M. P. Kunstmann, H. A. Whaley, and E. L. Patterson, J. Am. Chem. Soc., 1968, 90, 1325.
 ⁷ D. W. Cameron and A. R. Todd, in 'Oxidative Coupling of Phenols,' eds. W. I. Taylor and A. R. Battersby, Arnold, London, 1967.
 ⁸ G. Read and V. M. Ruiz, J. Chem. Soc., Perkin Trans. 1, 1973, 235.
 ⁹ N. Jacobsen and K. Torsell, Leibigs Ann. Chem., Soc., 1972, 94, 227.
 ¹¹ P. Jacob, III, P. S. Callery, A. T. Shulgin, and N. Castagnoli, Jr., J. Org. Chem., 1976, 41, 3627.
 ¹² D. W. Cameron, R. I. T. Cromartie, D. G. I. Kingston, and A. R. Todd, J. Chem. Soc., 1964, 51.
 ¹³ D. W. Cameron, D. G. I. Kingston, N. Sheppard, and A. R. Todd, J. Chem. Soc., 1964, 98.
 ¹⁴ R. G. F. Giles, M. K. Reuben, and G. H. P. Roos, S. Afr. J. Chem., 1979, 32, 127.