## Green light-emitting diodes from poly(2-dimethyloctylsilyl-1,4-phenylenevinylene)

## Do-Hoon Hwang,<sup>*a*</sup> Sung Tae Kim,<sup>*b,c*</sup> Hong-Ku Shim,<sup>*d*</sup> Andrew B. Holmes,<sup>\**a,c*</sup> Stephen C. Moratti<sup>*c*</sup> and Richard H. Friend<sup>*b*</sup>

<sup>a</sup> University Chemical Laboratory, Department of Chemistry, Lensfield Road, Cambridge, UK CB2 1EW

<sup>b</sup> Cavendish Laboratory, Department of Physics, Madingley Road, Cambridge, UK CB3 0HE

e Melville Laboratory for Polymer Synthesis, Department of Chemistry, Pembroke Street, Cambridge, UK CB2 3RA

<sup>d</sup> Department of Chemistry, Korea Advanced Institute of Science and Technology, Taejon 305-701, Korea

A novel silyl-substituted solvent processable poly(1,4-phenylenevinylene) (PPV) derivative, poly(2-dimethyloctylsilyl-1,4-phenylenevinylene) (DMOS-PPV) is synthesized by the dehydrohalogenation route from 2-dimethyloctylsilyl-1,4bis(bromomethyl)benzene, and the light-emitting properties of the polymer are studied; single layer electroluminescent devices (ITO/polymer/Ca or Al) exhibit an emission maximum at 520 nm with internal quantum efficiency in the range 0.2–0.3%.

Light-emitting polymers have been extensively investigated in recent years since the Cambridge group first reported a green light-emitting diode (LED) using poly(1,4-phenylenevinylene) (PPV) as an emitting layer.<sup>1-4</sup> Organic polymer LEDs have many advantages for the development of a large-area visible light-emitting display, because of the good processability, low operation voltage, fast response time and colour tunability over the full visible range by control of the HOMO-LUMO bandgap of the emissive layer. PPV has been most widely used as the emissive layer for the light-emitting diodes, and has been prepared through a thermal elimination process from a water-5 or organic-soluble precursor polymer.6-8 Several organic solvent soluble PPV derivatives have been developed in order to improve processability.9-11 Recently, Zhang et al. reported the improved quantum efficiency in green polymer light-emitting diodes with a silyl-substituted soluble PPV derivative, poly(2cholestanoxy-5-thexylsilyl-1,4-phenylenevinylene) (CS-PPV).<sup>11</sup> They reported<sup>12</sup> that CS-PPV showed high quantum efficiency with an air-stable aluminium electrode by adding an 2-(4-biphenyl)electron transporting molecular dopant, 5-(4-tert-butylphenyl)-1,3,4-oxadiazole (PBD).<sup>13,14</sup>

The effects of silicon substitution on the luminescence properties are of interest in the field of polymer LEDs, and here we report the synthesis of a new silyl-substituted soluble PPV derivative, poly(2-dimethyloctylsilyl-1,4-phenylenevinylene) (DMOS-PPV). Single layer electroluminescent (EL) devices have been fabricated using this polymer as the emissive layer. The synthetic route is outlined in Scheme 1. Silylation of the Grignard reagent derived from 1 afforded the silyl derivative 2 which after radical bromination gave the dibromo-compound 3.<sup>†</sup> Dehydrohalogenation condensation polymerization afforded DMOS-PPV 4.<sup>15,16</sup>

DMOS-PPV **4** is completely soluble in common organic solvents such as chloroform, tetrahydrofuran and toluene without evidence of gel formation.

Fig. 1 shows the UV–VIS, photoluminescence (PL) and EL spectra of the DMOS-PPV film. DMOS-PPV 4 shows a slightly narrower absorption band compared with PPV. The absorption maximum and edge of the DMOS-PPV are at *ca*. 414 and 500 nm, respectively, at room temperature. These positions are blue-shifted compared with those of the unsubstituted PPV (420 and 530 nm, respectively), presumably owing to the steric effect of the bulky dimethyloctylsilyl group.

DMOS-PPV **4** shows an emission maximum at *ca*. 520 nm which corresponds to the green region. The absolute photoluminescence quantum efficiency for a solid film of DMOS-PPV was 60%. By comparison, the reported PL efficiencies of PPV and MEH-PPV are 27 and 15%, respectively.<sup>17</sup>

Fig. 2 shows the current density–electric field characteristics measured for a typical ITO/DMOS-PPV/Al device with film thickness of 700 Å. The forward current density increases with increasing forward bias field and the curve shows typical diode



Fig. 1 UV–VIS (crosses), PL (open circles) of DMOS-PPV film and EL (solid circles) of the ITO/DMOS-PPV/Al device

characteristics. The voltage dependence of emission intensity from the device shows that light emission becomes observable at a bias of about 15 V at a current density of 0.93 mA cm<sup>-2</sup>. The devices showed reproducible internal quantum efficiencies of 0.2% (*ca.* 0.05% external efficiency).‡

Fig. 3 shows the current density–field characteristics of an ITO/DMOS/Ca device (700 Å thickness). The threshold voltage of the device was about 11 V at a current density of 1.8 mA cm<sup>-2</sup>. The measured maximum internal quantum efficiency of the diode was 0.3% (*ca.* 0.1% external efficiency). These values compare favourably with the efficiency of the single layer green polymer LED reported by Son *et al.*<sup>8</sup>

Recently we have fabricated multilayer EL devices with DMOS-PPV and various charge-transporting materials, such as 2-(4-biphenylyl)-5-(4-tert-butylphenyl)-1,3,4-oxadiazole (PBD) or poly(aromatic oxadiazole)s. These devices showed highly improved quantum efficiencies using an aluminium cathode.

The DMOS-PPV film has good processability and the high quantum efficiency may make it a good candidate for application in polymer LEDs.



Fig. 2 Current density-electric field and light intensity-voltage (inset) characteristics of ITO/DMOS-PPV/Al device



Fig. 3 Current density-electric field and light intensity-voltage (inset) characteristics of ITO/DMOS-PPV/Ca device

We thank the Korea Science and Engineering Foundation and the British Council (D. H. H.), the European Commission (Brite Euram Project, BRE2-CT93-0592 'PolyLED'), the Engineering and Physical Sciences Research Council (UK) (Swansea mass spectrometry service, Daresbury database<sup>19</sup>), and LG Electronics (S. T. K.) for financial support.

## Footnotes

† Selected spectroscopic data. **2**, <sup>1</sup>H NMR (CDCl<sub>3</sub>, 200 MHz) δ 7.28 (1 H, s), 7.09 (2 H, s), 2.43 (3 H, s), 2.34 (3 H, s), 1.41–1.14 (12 H, m), 0.97–0.75 (5 H, m), 0.33 (6 H, s); **3**, <sup>1</sup>H NMR (CDCl<sub>3</sub>, 200 MHz) δ 7.48 (1 H, s), 7.43 (2 H, s), 4.61 (2 H, s), 4.48 (2 H, s), 1.43–1.19 (12 H, m), 0.99–0.81 (5 H, m), 0.42 (6 H, s); DMOS-PPV **4**, GPC (polystyrene standard) showed a  $M_w$  of 1.1 × 10<sup>6</sup> and polydispersity index of 7.2. (Found: C, 77.5; H, 10.05. Calc. C, 79.34; H, 10.36%). FTIR (NaCl)  $\nu_{max}/cm^{-1}$  2955, 2922, 2853, 1729, 1468, 1377, 1251, 1137, 1066, 961, 836.

<sup>‡</sup> The internal efficiency is a factor of  $2n^2$  larger than the external efficiency where *n* is the refractive index of the emissive layer.<sup>18</sup>

## References

- J. H. Burroughes, D. D. C. Bradley, A. R. Brown, R. N. Marks, K. Mackay, R. H. Friend, P. L. Burn and A. B. Holmes, *Nature*, 1990, 347, 539.
- 2 N. C. Greenham, S. C. Moratti, D. D. C. Bradley, R. H. Friend and A. B. Holmes, *Nature*, 1993, **365**, 628.
- 3 G. Gustafsson, Y. Cao, G. M. Treacy, F. Klavetter, N. Colaneri and A. J. Heeger, *Nature*, 1992, **357** 477.
- 4 P. L. Burn, A. B. Holmes, A. Kraft, D. D. C. Bradley, A. R. Brown, P. H. Esind and P. W. Commun. N. 4 1002, 256, 47
- R. H. Friend and R. W. Gymer, *Nature*, 1992, **356**, 47.
- 5 R. A. Wessling, J. Polym. Sci., Polym. Symp., 1985, 72, 55.
- 6 F. Louwet, D. Vanderzande, J. Gelan and J. Mullens, *Macromolecules*, 1995, **28**, 1330.
- 7 P. L. Burn, D. D. C. Bradley, R. H. Friend, D. A. Halliday, A. B. Holmes, R. W. Jackson and A. Kraft, *J. Chem. Soc., Perkin Trans. 1*, 1992, 3225.
- 8 S. Son, A. Dodabalapur, A. J. Lovinger and M. E. Galvin, *Science*, 1995, **269**, 376.
- 9 F. Wudl, P. M. Allemand, G. Srdanov, Z. Ni and D. McBranch, in Materials for Nonlinear Optics: Chemical Perspectives, ed. S. R. Marder, J. E. Sohn and G. D. Stucky, ACS Symp. Ser., 1991, 45, 683.
- 10 F. Wudl, S. Höger, C. Zhang, K. Pakbaz and A. J. Heeger, *Polym. Prepr.*, 1993, 34, 197.
- 11 S. Höger, J. J. McNamara, S. Schricker and F. Wudl, *Chem. Mater.*, 1994, 6, 171.
- 12 C. Zhang, S. Höger, K. Pakbaz, F. Wudl and A. J. Heeger, J. Electron. Mater., 1994, 23, 453.
- 13 P. L. Burn, A. B. Holmes, A. Kraft, A. R. Brown, D. D. C. Bradley and R. H. Friend, *Mater. Res. Soc. Symp. Proc.*, 1992, **247**, 447.
- 14 A. R. Brown, D. D. C. Bradley, J. H. Burroughes, R. H. Friend, N. C. Greenham, P. L. Burn, A. B. Holmes and A. Kraft, *Appl. Phys. Lett.*, 1992, **61**, 279.
- 15 H. G. Gilch and W. L. Wheelwright, J. Polym. Sci. A-1, 1966, 4, 1337.
- 16 W. S. Swatos and B. Gordon, Polym. Prepr., 1990, 31, 505.
- 17 N. C. Greenham, I. D. W. Samuel, G. R. Hayes, R. T. Philips, Y. A. R. R. Kessener, S. C. Moratti, A. B. Holmes and R. H. Friend, *Chem. Phys. Lett.*, 1995, **241**, 89.
- 18 N. C. Greenham, R. H. Friend and D. D. C. Bradley, *Adv. Mater.*, 1994, 6, 491.
- 19 D. A. Fletcher, R. F. McMeeking and D. Parkin, J. Chem. Inf. Comput. Sci., 1996, 36, 746.

Received, 24th June 1996; Com. 6/04364J