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A COMBINED 1,2-OXYGEN AND 1,4-CARBONYL GROUP TRANSPOSITION IN TETRALONES VIA HETEROCYCLIC INTERMEDIATES^{+,1)}

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> The conversions of the oxanorbornadiene $\underline{1}$ to the hydroxy-tetralones $\underline{2}$ and $\underline{3}$ are described. The sequence $\underline{1} - - \frac{1}{2}$ includes a combined 1,2-oxygen and 1,4-carbonyl group transposition.

Treatment of the easily available oxanorbornadiene \underline{l}^{1} with acid (conc. hydrochloric acid in methanol, 20 hr under reflux) yields the tetralone $\underline{2}$ (56 %, mp 134-136°C from ether) without migration of functional groups. The synthesis of the rearranged hydroxy-tetralone $\underline{3}$ starts with the photolysis of $\underline{1}$ to the corre-

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sponding oxaquadricyclane and subsequent thermal isomerisation to the oxepine $\frac{4}{2}$ (over all yield 55 %)¹⁾. Acid catalysed isomerisation of $\frac{4}{2}$ with trifluoroacetic acid³⁾ leads in 40 % yield to the 4-hydroxy-8-keto-derivative $\frac{3}{2}$ (trifluoroacetic acid: tetrachloromethane 1:10, 2 hr at 25°C, subsequent hydrolysis with



sodium hydrogencarbonate solution; mp 203-205°C from ethyl acetate).

2 and 3 differ in the positions of the phenolic hydroxyl group (C-3 in 2, C-4 in 3; 1,2-transposition) and of the carbonyl group (C-5 in 2, C-8 in 3; 1,4-transposition) as well. The assignment of structure for 2 and 3 is consistent with spectroscopic data (Table 1)⁴⁾. The ir and ¹H-nmr spectra of 2 indicate hydrogen bonding between the phenolic hydroxyl group and one carbomethoxy group. Moreover the uv spectrum of the corresponding alcohol (2, CHOH instead of CO on C-5) obtained from 2 and sodium borohydride (methanol, 5 hr at 0°C; mp 75-76°C from ether) is highly similar to that of 3-hydroxy-6-methyl-dimethylphthalate³⁾. On the other hand the same spectra indicate the absence of intramolecular hydrogen bonding in 3. Thereby other possible isomers are excluded⁵⁾. 3 is probably formed from $\frac{4}{2}$ by the mechanistic pathway established for other $oxepines^{3,6)}$. Table 1. Spectroscopic data for the compounds 2 - 4 2: ir (KBr): 1670(CO), 1695(CO₂CH₃), 1745(CO₂CH₃), 3100 cm⁻¹(OH). - $uv(C_{2}H_{5}OH): \lambda_{max}(log\epsilon) = 220$ (4.36), 257 (4.05), 341 nm (3.66). - ¹H-nmr(CDCl₃): δ = 1.8-2.3 (m,2H,CH₂), 2.5-2.95 (m,4H, 2CH₂), 3.90 (s, 3H, CO₂CH₃), 3.92 (s, 3H, CO₂CH₃), 7.68 (s, 1H, 4-H), 10.6 (s,1H,0-H). <u>]</u>: ir (KBr): 1665(CO), 1725-1745(CO₂CH₃), 3200 cm⁻¹(OH). -

uv($C_{2}H_{5}OH$): λ_{max} (log ϵ)= 234 (4.27), 322 nm (3.20). - ¹H-nmr (CDCl₃): δ = 1.6-2.35 (m,2H,CH₂), 2.5-2.7(t,2H,CH₂), 2.7-3.1(t, 2H,CH₂), 3.82 (s,3H,CO₂CH₃), 3.97(s,3H,CO₂CH₃), 6.85 (broad,1H, OH), 7.5(s,1H,3-H). $\frac{2}{2} (CHOH instead of CO on C-5): ir(KBr): 1680(CO_2CH_3), 1720(CO_2-CH_3), 3110(OH), 3460 cm⁻¹(OH). - uv(C_2H_5OH): <math>\lambda_{max}(\log \varepsilon) = 215 (4.41), 250(3.92), 320 nm(3.70). - ^{1}H-nmr(CDCl_3): \delta = 1.5-2.25 (m,5H,2CH_2 + 1 OH), 2.55(t,2H,CH_2), 3.87 (s,3H,CO_2CH_3), 3.90(s, 3H,CO_2CH_3), 4.65(t,1H,5-H), 7.15(s,1H,4-H), 10.62(s,1H,OH).$ $<math>\frac{1}{2}^{11}: ^{13}C-nmr(CDCl_3): \delta = 18.7(t,J_{CH}=130 Hz,8-C), 29.8(t,J_{CH} = 129 Hz,9-C), 32.9(t,J_{CH} = 129 Hz,7-C), 52.2(q,J_{CH} = 148 Hz, CH_3OOC), 64.4(t,J_{CH} = 150 Hz,dioxolane ring-C), 108.2(m,6-C), 116.4(dd,J_{CH} = 165 Hz,J_{CCH} = 13 Hz,3-C), 120.9(s,5a-C), 130.4 (d,J_{CCCH} = 7 Hz,5-C), 140.9(d,J_{CCH} = 7 Hz,4-C), 145.6 (dd,J_{CH} = 209 Hz,J_{CCH} = 6 Hz,2-C), 163.3(m,9a-C), 165.4(d,J_{CCCCH} = 4 Hz, CO_2CH_3 on C-5), 167.8 (d,J_{CCCH}=4 Hz,CO_2CH_3 on 4-C). ACKNOWLEDGEMENT: The financial support of the Deutsche Forschungs-gemeinschaft is gratefully acknowledged. REFERENCES$

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