Registry No. $EuCl_3$, 10025-76-0.

References and Notes

(1) C. E. Myers, *Inorg. Chem.*, **14**, 199 (1975).
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(3) C. E. Habermann and A. H. Daane, *J. Chem. Phys.*, **41**, 2818 (1964).
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(5) (a) R. H. Schumm, D. D. Wagman, S. Bailey, W. H. Evans, and V. B. Parker, *Natl. Bur. Stand. (U.S.), Tech. Note*, **No. 270-7** (1973); (b) D. A. Johnson, *J. Chem. Soc. A*, 2578 (1969); (c) L. R. Morss and H. O. Haug, *J. Chem. Thermodyn.*, **5**, 513 (1973).
(6) R. C. Feber, "Heats of Dissociation of Gaseous Halides", Los Alamos Scientific Laboratory Report No. LA-3164, USAEC No. TID-4500, 40th ed, Los Alamos Scientific Laboratory, Los Alamos, N.M., 1965.

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Additions and Corrections

1973, Volume 12

George F. McKnight and G. P. Haight, Jr.*: Reactions of Octacyanomolybdate(VI). I. Kinetics of Oxidation by Chromium(VI).

Pages 1619 and 1620. An error of sign in the exponent of the rate constant has been made in (1) the abstract, page 1619, $k = (8.3 \pm 2.0) \times 10^4$, (2) the last line before the **Discussion**, page 1620, $k = (8.8 \pm 1.5) \times 10^4$, and (3) the third column heading in Table III which should read $10^{-4}k''/[H^+]^2$. Parts of the discussion based on the reported low value of k are thus irrelevant since $Mo(CN)_8^{4-}$ is actually faster than $Fe(CN)_6^{4-}$, but the conclusions concerning inner sphere electron transfer processes are not affected.—G. Haight

1974, Volume 13

Daniel D. Poulin and Ronald G. Cavell*: Phosphoranes. I. Tris(trifluoromethyl)bis(dimethylamino)phosphorane, $(CH_3)_3P[N(CH_3)_2]_2$, and Related Chlorodimethylaminotrifluoromethylphosphoranes.

Page 2332. The first sentence after the bold face heading in column 1 should read: "The nmr results obtained herein and elsewhere[26] consistently suggest that CF_3 groups preferentially occupy equatorial positions in halogenophosphoranes even when the CF_3 group has a *higher* electronegativity than the halogen."—Ronald G. Cavell

Nick Serpone* and Ken A. Hersh: Kinetic Analysis of the Configurational Rearrangements in and the Stereochemistry of Some Organotin(IV) β-Ketoenolate Complexes.

Page 2905. In the second column, the sentence beginning on line 22 should read: "To describe this exchange process,.....six independent first-order rate constants will...".—N. Serpone

1975, Volume 14

Philip S. Bryan and James C. Dabrowiak*: Synthesis and Characterization of Manganese Complexes Containing a Synthetic Macrocyclic Ligand.

Page 297. In Table II, the column headings of the first and second columns should be interchanged, and the headings of the third and fourth columns should be interchanged.—James C. Dabrowiak

Fred W. Moore and Kenneth W. Hicks*: Mechanism of the Permanganate Ion Oxidation of Vanadium(IV).

Page 413. The middle initial of the first author should read "W".—Kenneth W. Hicks

D. L. Reger: Cyanide, Isocyanide, and Nitrile Derivatives of Cyclopentadienyliron. Interaction of Chiral Metal Complexes with an Optically Active Shift Reagent.

Page 662. Twenty lines were omitted from the top of the first column on page 662. Beginning with the last line on page 661, the text should read as follows:

2. Potassium Cyanide. A mixture of $[(\eta^5\text{-}C_5H_5)Fe(CO)(CNEt)(PPh_3)]PF_6$ (0.7 g, 1.15 mmol) and KCN (0.15 g, 2.3 mmol) was heated at 70° for 3 hr in ethanol (40 ml). As no reaction took place under these conditions, the solution was then refluxed for 2 hr. The infrared spectrum of this solution indicated that starting reagents were still mainly present with small amounts of $(\eta^5\text{-}C_5H_5)Fe(CO)(CN)(PPh_3)$ also present. More KCN (0.1 g, 1.5 mmol) was added and the solution was refluxed for 6 hr more. Additional $(\eta^5\text{-}C_5H_5)Fe(CO)(CN)(PPh_3)$ had formed, but the solution still contained mainly starting reagents.

3. Sodium Borohydride. A mixture of $[(\eta^5\text{-}C_5H_5)Fe(CO)(CNEt)(PPh_3)]PF_6$ (1.0 g, 1.6 mmol) and $NaBH_4$ (0.3 g, 7.9 mmol) was stirred for 2 hr in THF (40 ml). The solvent was evaporated, and the resulting residual oil was dissolved in benzene (8 ml) and chromatographed on alumina (2 × 8 cm) with hexane–diethyl ether (2:1 v/v) as eluent. A single yellow band was eluted that was concentrated and cooled to 0° to yield $(\eta^5\text{-}C_5H_5)Fe(CO)(PPh_3)H$ (0.32 g, 48%), identified by comparison to an authentic sample.[12,13]

$(\eta^5$-Cyclopentadienyl)carbonyl(cyanomethyl)(triphenylphosphine)iron(II) (2). A solution of *ca.* 20 mmol of $LiCH_2CN$ in THF (500 ml) at −78° was prepared by the method of Crouse and See-

Page 663. The first 18 lines on this page duplicate the last 18 lines on page 662 and should be deleted.—D. L. Reger

Itamar Bodek, Geoffrey Davies,* and John H. Ferguson: Studies of Ammineaquocobalt(III) Chemistry. Kinetics and Mechanisms of Formation and Dissociation of Monochlorodiammineaquocobalt(III) Complexes and of the Reduction of *cis*-Diammineaquocobalt(III) Species by Br^- in Acid Perchlorate Solution.

Page 1709. In column 1, the first sentence of the paragraph on High-Temperature Synthesis should begin: "An ice-cold solution of $CoCl_2 \cdot 6H_2O$ (20 g, 0.084 mol) and H_2O_2 (5 ml, 10 *M*, 0.05 mol) in 50 ml of water.....". Later in the same sentence the values in parentheses after $KHCO_3$ should be 40 g, 0.4 mol.—Itamar Bodek