Additions and Corrections

1975, Volume 14

Rudi van Eldik and Gordon M. Harris*: Substitution Reactions of Oxalato Complex Ions. XI. Kinetics of the Anation of Aquopentaamminecobalt(III) Ion by Oxalate in Aqueous Acidic Solution.

Page 12. Equation 4 is wrong and should read:

$$k_{obsd} = \frac{k_0[H_2C_2O_4] + k_1Q_1[HC_2O_4^-] + k_2Q_2[C_2O_4^{2-}]}{1 + Q_1[HC_2O_4^-] + Q_2[C_2O_4^{2-}]}$$

Fortunately, since we did our experiments in three different acidity regions we essentially could utilize three different equations separately. These were as follows: (a) at high pH

$$k_{obsd} = \frac{k_2Q_2[C_2O_4^{2-}]}{1 + Q_2[C_2O_4^{2-}]}$$

(b) at pH 2.5

$$k_{obsd} = \frac{k_1Q_1[HC_2O_4^-]}{1 + Q_1[HC_2O_4^-]}$$

(c) at low pH

$$k_{obsd} = \frac{k_0[H_2C_2O_4] + k_1Q_1[HC_2O_4^-]}{1 + Q_1[HC_2O_4^-]}$$

The results of our analysis of the data therefore are completely correct in the high pH range and at pH 2.5 and are only subject to small errors at low pH (pH ≤1), since the concentration of $[HC_2O_4^-]$ is rather small and Q_1 is only 1.8. Our conclusions are not influenced by the error so the arguments we have given are still completely sound.—G. M. Harris

Ira B. Goldberg,* Karl O. Christe, and Richard D. Wilson: Electron Spin Resonance Study of O_2^+ Salts. Influence of Crystal Field Effects and Relaxation.

Page 155. Equation 6c should read:

$$g_y = g_e\left(\frac{\Delta^2}{\lambda^2 + \Delta^2}\right)^{1/2} + \frac{\lambda}{E}\left[\left(\frac{\Delta^2}{\lambda^2 + \Delta^2}\right)^{1/2} - \left(\frac{\lambda^2}{\lambda^2 + \Delta^2}\right)^{1/2} - 1\right] \quad (6c)$$

Ira B. Goldberg

G. Joy, A. P. Gaughan Jr., I. Wharf, D. F. Shriver,* and J. P. Dougherty: Single-Crystal Raman Evidence for and X-Ray Analysis of the Distorted Square-Pyramidal Pentachlorothallate and Pentachloroindate Complexes in $[(C_2H_5)_4N]_2[TlCl_5]$ and $[(C_2H_5)_4N]_2[InCl_5]$.

Page 1799. In Table IV, $4(C_4)$ in column one should correlate with $4(C_4)$ in the center column instead of $2/m(C_{2h})$.

Page 1800. The caption to Figure 3 should give the standard errors in the bond angles as ±0.5° and ±0.6° in place of ±5° and ±6° respectively.—A. P. Gaughan

Billy J. McKerley, Gary C. Faber, and Gerard R. Dobson*: Octahedral Metal Carbonyls. XXXV. Mechanism of the Reaction of 2,2'-Dipyridyltetracarbonylchromium(0) with Alkyl Phosphites.

Page 2276. The script beginning with line 11 from the bottom, first column, should read: "The greater rigidity of the *o*-phen's presumably would lead to greater $k_{-2}/k_3[L]$ ratios (eq 5), rendering ligand-dependent rate behavior more likely to be observed. Moreover..."—G. R. Dobson