ternary arsenide, whereas a difference of ~ 0.04 Å might be expected solely on the basis of the covalent radii.[17] This leaves one with the thought that introduction of charge into such isoelectronic systems, viz., for Si–Se $\rightarrow$ Si–As$^-$ K$^+$, lengthens the Si–As bond appreciably. The same may occur when the higher field Ba^{2+} is present in Ba$_3$Si$_2$As$_4$ (above) although the two types of arsenic atoms present there complicate the situation.

Previously, we were able to discern a concerted (but hypothetical) process of ring opening and bond formation whereby the structure of SiAs could be reductively converted to that of the anion sheet found in KSi$_3$As$_3$. This is not as direct for K$_2$SiAs$_2$

since the structure of the possible precursor SiAs$_2$[18] (GeAs$_2$ type[19]) includes As–As bonds in the five-membered rings that have been condensed into sheets.

(17) Pauling, L. *The Nature of the Chemical Bond*, 3rd ed.; Cornell University Press: Ithaca, NY, 1960; p 225.

(18) Wadsten, T. *Acta Chem. Scand.* **1967**, *21*, 593.
(19) Bryden, J. H. *Acta Crystallogr.* **1962**, *15*, 167.

Additions and Corrections

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Karen Libson, Mary Noon Doyle, Rudy W. Thomas, Theodore Nelesnik, Mary Woods, James C. Sullivan, R. C. Elder, and Edward Deutsch*: Structural and Kinetic Investigations of a Tc(III)/Tc(II) Redox Couple. X-ray Crystal Structures of *trans*-[Tc^{II}(DPPE)$_2$Cl$_2$] and *trans*-[Tc^{III}(DPPE)$_2$Cl$_2$]NO$_3$·HNO$_3$, Where DPPE = 1,2-Bis-(diphenylphosphino)ethane.

Page 3615. In Table I the y coordinates of atoms C(34) and C(35) were incorrectly listed as positive numbers. The proper values are -0.0646 (4) for C(34) and -0.0255 (3) for C(35).—R. C. Elder