

isomers to an identical new configuration remains possible but seems less likely than the red \rightarrow green transformation proposed above because of the need to postulate the existence of a third isomer for the cluster. This interpretation also fits with chemical intuition, since the green isomer, which has a staggered geometry, is expected to be the more stable under high-pressure conditions. However a structure intermediate between the eclipsed and staggered forms might be stable at high pressure, but we have no specific evidence for this.

These results differ in an interesting fashion from those reported earlier for the gold clusters $[\text{Au}_9(\text{PR}_3)_8]^{3+}$,¹⁰ where the application of pressure induced rearrangement of the entire cluster skeleton. In this example, the cluster skeletal geometries in the two modifications differ only at an interior link, and applied pressure appears to be inducing a conformational transformation around this bond.

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Tin(II) Poly(1-pyrazolyl)borates

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Ever since their first preparation,¹ poly(1-pyrazolyl)borates, $M[(\text{pz})_{4-n}\text{BR}_n]$ (M = alkali metal, Hpz = pyrazole or C-substituted derivative thereof, R = noncoordinating substituent, n = 0-2), have found extensive use as chelating ligands in transition-metal chemistry.² However, studies of the poly(1-pyrazolyl)borate chemistry of main-group elements have been rather limited. Only most recently have some characterized poly(1-pyrazolyl)borates of main-group (other than alkali and alkaline earth) metals become available. The data include four (somewhat contradictory) reports on tin(IV)³⁻⁶ and one on tin(II) derivatives;⁷ silicon or germanium poly(1-pyrazolyl)borates have not yet been reported.

In an extension of recent studies on the interaction of boron compounds with poly(1-pyrazolyl)borates, the synthesis and characterization of other main-group element derivatives of the latter is currently being investigated. The present report describes a study of a series of Sn(II) poly(1-pyrazolyl)borates.

Experimental Section

Elemental analyses were performed by the Schwarzkopf Microanalytical Laboratory, Woodside, NY. Melting points (uncorrected) were determined on a Mel-Temp block.

NMR spectra were recorded for solutions in CDCl_3 (unless otherwise noted) on a Varian VXR-400 (^1H , ^{119}Sn , variable-temperature, high-resolution) or GEMINI-200 (^1H , ^{13}C) instrument. Chemical shift data are given in ppm with positive values indicating downfield from the reference (internal $(\text{CH}_3)_4\text{Si}$ for ^1H and ^{13}C NMR, external $(\text{C}_2\text{H}_5)_2\text{O}-\text{BF}_3$ for ^{119}Sn NMR, external $(\text{CH}_3)_4\text{Sn}$ for ^{119}Sn NMR); s = singlet,

d = doublet, t = triplet, q = quartet, m = unresolved multiplet, and an asterisk denotes a broad signal. Coupling constants J are given in hertz. ^{13}C NMR spectra were recorded in the proton-decoupled mode.

All nonreferenced reagents were obtained from Aldrich Chemical Co., Milwaukee, WI, and used as received.

$\text{Na}[(\text{C}_6\text{H}_5)_3\text{B}(\text{pz})_2]$ (Hpz = pyrazole) was prepared by the previously described reaction of $\text{Na}[\text{B}(\text{C}_6\text{H}_5)_4]$ with Hpz .⁸ After extensive purification, ultimately by subliming off traces of adhering Hpz , the pure compound has a mp 284-287 °C. NMR data: $\delta(^1\text{H})$ 7.52 (1 H, d, J = 2.2), 7.41 (1 H, d, J = 1.7), 7.2 (3 H, m), 6.95 (2 H, m), 6.20 (1 H, unsym t = two overlapping d); $\delta(^{11}\text{B})$ 1.3 (s, $h_{1/2}$ = 325 Hz).

$[\text{B}(\text{pz})_4]_2\text{Sn}$ (1). A mixture of 0.94 g (5 mmol) of SnCl_2 , 3.15 g (10 mmol) of $\text{K}[\text{B}(\text{pz})_4]$,¹ and 50 mL of methylene chloride was stirred at room temperature for 5 h. The mixture was filtered, and solvent was evaporated from the clear filtrate under reduced pressure to leave 3.2 g (94%) of crude product. It was recrystallized from toluene to give a material of mp 208-210 °C. Anal. Calcd for $\text{C}_{24}\text{H}_{24}\text{B}_2\text{N}_{16}\text{Sn}$ (M_r = 676.5): C, 42.57; H, 3.58; B, 3.20; N, 33.11; Sn, 17.54. Found: C, 43.15; H, 3.50; B, 3.20; N, 33.37; Sn, 17.55.

NMR data: $\delta(^1\text{H})$ 7.63 (1 H, unresolved d), 7.26 (1 H, d, J = 2.4), 6.32 (1 H, unsym t = two overlapping d, J = ca. 1.9); $\delta(^{11}\text{B})$ 0.4 (s, $h_{1/2}$ = 16 Hz); $\delta(^{13}\text{C})$ 142.2, 135.9, 106.5; $\delta(^{119}\text{Sn})$ -744 (s, $h_{1/2}$ = 50 Hz); (solution in CD_2Cl_2) $\delta(^1\text{H})$ 7.62 (1 H, d, J = 1.3), 7.27 (1 H, d, J = 2.1), 6.32 (1 H, unsym t = two overlapping d, J = ca. 2.1); (at -85 °C) $\delta(^1\text{H})$ 7.93 (1 H, 7.59 (3 H), 7.40 (3 H), 6.81 (1 H), 6.38 (4 H).

$[\text{B}(\text{pz})_4]_2\text{Sn}$ (2, Hpz' = 3-methylpyrazole) was obtained from the reaction of 0.33 g (1.74 mmol) of SnCl_2 with 1.30 g (3.48 mmol) of $\text{K}[\text{B}(\text{pz}')_4]$ ⁹ (30 mL of methylene chloride, 4 h stirring). The crude material was obtained in quantitative yield; it was washed with cyclohexane to give a product of mp 176-180 °C. Anal. Calcd for $\text{C}_{32}\text{H}_{40}\text{B}_2\text{N}_{16}\text{Sn}$ (M_r = 788.6): C, 48.69; H, 5.11; B, 2.75; N, 28.40; Sn, 15.05. Found: C, 47.15; H, 4.70; B, 3.50; N, 26.99; Sn, 14.84.

NMR data: $\delta(^1\text{H})$, 7.4* (s) + 7.2* (s) (1 H total, ratio 1:3), 6.05 (1 H, d, J = 1.7), 2.3* (s) + 2.1* (s) (3 H total, ratio 1:3); $\delta(^{11}\text{B})$ -0.8 (s, $h_{1/2}$ = 30 Hz); $\delta(^{13}\text{C})$ 151.8, 136.9, 106.5, 13.1; $\delta(^{119}\text{Sn})$ -808 (s, $h_{1/2}$ = 50 Hz); (at 50 °C) $\delta(^1\text{H})$ 7.2* (1 H, s), 6.05 (1 H, d, J = 1.7), 2.1* (3 H, s).

$[\text{B}(\text{pz}')_4]\text{SnCl}$ (8) was prepared from 0.50 g (2.64 mmol) of SnCl_2 and 0.99 g (2.64 mmol) of $\text{K}[\text{B}(\text{pz}')_4]$ ⁹ (30 mL of methylene chloride, 1 h stirring) in essentially quantitative yield. The crude material was washed with cyclohexane to give a product of mp 165-169 °C dec. Anal. Calcd for $\text{C}_{16}\text{H}_{20}\text{BClN}_8\text{Sn}$ (M_r = 489.1): C, 39.25; H, 4.12; B, 2.21; Cl, 7.26; N, 22.90; Sn, 24.26. Found: C, 39.27; H, 4.11; B, 2.34; Cl, 7.65; N, 22.40; Sn, 24.02.

NMR data: $\delta(^1\text{H})$ 7.51* (1 H, s), 6.08* (1 H, s), 2.51 (3 H, s); $\delta(^{11}\text{B})$ -1.3 (s, $h_{1/2}$ = 35 Hz); $\delta(^{13}\text{C})$ 151.7, 136.6, 106.6, 13.7; $\delta(^{119}\text{Sn})$ -533 (s, $h_{1/2}$ = 230 Hz); (at -70 °C, solution in CD_2Cl_2): $\delta(^1\text{H})$ 7.80 (d, J = 2.1) + 7.41* (s) (1 H total, ratio 1:3), 6.32 (d, J = 2.0) + 6.04 (d, J = 2.0) (1 H total, ratio 1:3), 2.45 (s) + 2.34 (s) (3 H total, ratio 3:1).

$[\text{HB}(\text{pz})_3]_2\text{Sn}$ (3) was prepared from 0.53 g (2.8 mmol) of SnCl_2 and 1.40 g (5.6 mmol) of $\text{K}[\text{HB}(\text{pz})_3]$ ¹ (30 mL of methylene chloride, 6 h stirring) to give 1.3 g (86%) of crude product. Traces of pyrazole were removed by sublimation under vacuum to leave a material of mp 162-165 °C. Anal. Calcd for $\text{C}_{18}\text{H}_{20}\text{B}_2\text{N}_{12}\text{Sn}$ (M_r = 544.4): C, 39.67; H, 3.70; B, 3.98; N, 30.86; Sn, 21.79. Found: C, 39.39; H, 3.51; B, 3.70; N, 30.50; Sn, 21.63.

NMR data: $\delta(^1\text{H})$ 7.69 (1 H, d, J = 2.2, of d, J = 0.6), 7.35 (1 H, d, J = 1.5), 6.17 (1 H, unsym t = two overlapping d, J = ca. 2.1); $\delta(^{11}\text{B})$ -2.8 (d, J = 95); $\delta(^{13}\text{C})$ 140.3, 135.7, 104.9; $\delta(^{119}\text{Sn})$ -877 (s, $h_{1/2}$ = 750 Hz; at -50 °C, $h_{1/2}$ = 270 Hz).

$[\text{HB}(\text{pz}^*)_3]_2\text{Sn}$ (4, Hpz^* = 3,5-dimethylpyrazole).⁷ NMR data: $\delta(^1\text{H})$ 5.72 (1 H, s), 2.26 (3 H, s), 1.80 (3 H, s); $\delta(^{11}\text{B})$ -8.3 (broad s, $h_{1/2}$ = 320 Hz); $\delta(^{13}\text{C})$ 148.9, 144.9, 106.0, 12.6, 12.3; $\delta(^{119}\text{Sn})$ -933 (s, $h_{1/2}$ = 350 Hz; at -50 °C, $h_{1/2}$ = 240 Hz). Lit.:⁷ $\delta(^1\text{H})$ 5.79 (3 H, s), 4.75* (1 H, s), 2.33 (9 H, s), 1.85 (9 H, s); $\delta(^{119}\text{Sn})$ (solution in CD_2Cl_2) -935 (s, $h_{1/2}$ = 300 Hz).

$[\text{HB}(\text{pz}^*)_3]\text{SnCl}$ (9) decomposes near 250 °C. NMR data: $\delta(^1\text{H})$ 5.82 (1 H, s), 2.50 (3 H, s), 2.37 (3 H, s); $\delta(^{11}\text{B})$ -9.6 (d, J = 92); $\delta(^{13}\text{C})$ 149.8, 145.4, 106.3, 13.8, 12.5; $\delta(^{119}\text{Sn})$ -579 ($h_{1/2}$ = 200 Hz). Lit.:⁷ $\delta(^1\text{H})$ (solution in C_6D_6) 5.40 (1 H, s), 2.36 (3 H, s), 1.96 (3 H, s); $\delta(^{119}\text{Sn})$ (solution in CD_2Cl_2) -1460 (s, $h_{1/2}$ > 300 Hz).

$[\text{H}_2\text{B}(\text{pz})_2]_2\text{Sn}$ (5) was prepared from 0.75 g (3.95 mmol) of SnCl_2 and 1.47 g (7.9 mmol) of $\text{K}[\text{H}_2\text{B}(\text{pz})_2]$ ¹ (30 mL of methylene chloride, 2 h stirring) to give 1.23 g (75%) of material, sintering at 110 °C and decomposing at 116-124 °C. The material could not be purified by recrystallization but deteriorated in all relevant attempts; it decomposes

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readily at elevated temperatures. Anal. Calcd for $C_{12}H_{16}B_2N_8Sn$ ($M_r = 412.4$): C, 34.92; H, 3.91; B, 5.24; N, 27.16; Sn, 28.77. Found: C, 32.89; H, 3.64; B, 4.40; N, 25.15; Sn, 30.29.

NMR data: $\delta(^1H)$ 7.64 (1 H, d, $J = 1.7$), 7.36* (1 H, s), 6.21* (1 H, unresolved), ca. 4* (1 H); $\delta(^{11}B)$ -8.7 (s, $h_{1/2} = 280$ Hz); $\delta(^{13}C)$ 139.4, 136.5, 105.3; $\delta(^{119}Sn)$ -654 (s, $h_{1/2} = 450$ Hz; at $-50^\circ C$, $h_{1/2} = 150$ Hz).

$[H_2B(pz)_2]_2SnCl$ (10) was obtained from 0.68 g (3.6 mmol) of $SnCl_2$ and 0.67 g (3.6 mmol) of $K[H_2B(pz)_2]^{11}$ (20 mL methylene chloride, 2 h stirring); 0.8 g (74%), mp 142–144 °C dec. Anal. Calcd for $C_6H_8B_2ClN_4Sn$ ($M_r = 301.1$): C, 23.92; H, 2.68; B, 3.59; Cl, 11.79; N, 18.60; Sn, 39.42. Found: C, 24.09; H, 2.50; B, 3.21; Cl, 12.30; N, 18.04; Sn, 39.17.

NMR data: $\delta(^1H)$ 7.8* (1 H, s), 7.65 (1 H, d, $J = 1.5$), 6.35* (1 H, unsym t), 3.85* (1 H); $\delta(^{11}B)$ -8.9 (t, $J = 98$); $\delta(^{13}C)$ 138.7, 137.3, 106.2; $\delta(^{119}Sn)$ -305 (s, $h_{1/2} = 1250$ Hz); (at $-50^\circ C$) $\delta(^{119}Sn)$ -262 (small and broad), -336 (s, $h_{1/2} = 330$ Hz).

$[H_2B(pz')_2]_2Sn$ (7) was prepared from 1.08 g (5.7 mmol) of $SnCl_2$ and 2.44 g (11.4 mmol) of $K[H_2B(pz')_2]^{9,10}$ (30 mL of methylene chloride, 3 h stirring) to give 2.2 g (82%) of crude product. This was redissolved in a minimal amount of methylene chloride, and the clear solution was cooled to $-40^\circ C$; the precipitated product decomposed at 122–124 °C. Anal. Calcd for $C_{16}H_{24}B_2N_8Sn$ ($M_r = 468.5$): C, 40.98; H, 5.16; B, 4.61; N, 23.91; Sn, 25.34. Found: C, 41.15; H, 5.14; B, 4.51; N, 23.71; Sn, 25.43.

NMR data: $\delta(^1H)$ 7.53 (1 H, d, $J = 1.8$), 5.98 (1 H, d, $J = 1.8$), 3.9* (1 H, ill-resolved q), 1.63 (3 H, s); $\delta(^{11}B)$ -9.8 (unresolved t?, $h_{1/2} = 350$ Hz); H-decoupled, $h_{1/2} = 190$ Hz; $\delta(^{13}C)$ 149.9, 136.8, 105.1, 11.7; $\delta(^{119}Sn)$ -685 (s, $h_{1/2} = 360$ Hz).

$[H_2B(pz')_2]_2SnCl$ (12) was prepared from 1.13 g (5.96 mmol) of $SnCl_2$ and 1.28 g (5.96 mmol) of $K[H_2B(pz')_2]^{9,10}$ (30 mL of methylene chloride, 3.5 h stirring). The product, mp 144–146 °C dec, precipitated on concentration of the solution and was obtained in 66% yield (1.3 g). Anal. Calcd for $C_8H_{12}BClN_4Sn$ ($M_r = 329.1$): C, 29.17; H, 3.68; B, 3.28; Cl, 10.79; N, 17.02; Sn, 36.06. Found: C, 29.65; H, 3.72; B, 3.31; Cl, 10.77; N, 16.80; Sn, 35.82.

NMR data: $\delta(^1H)$ 7.49 (1 H, d, $J = 1.9$), 6.08 (1 H, d, $J = 1.5$), 3.9* (1 H, very broad), 2.52 (3 H, s); $\delta(^{11}B)$ -9.9 (t, $J = 98$); $\delta(^{13}C)$ 151.0, 137.6, 106.2, 13.2; $\delta(^{119}Sn)$ -271 (s, $h_{1/2} = 365$ Hz).

$[(C_6H_5)_2B(pz)_2]_2Sn$ (6) was prepared from 2.0 g (6.1 mmol) of $Na[(C_6H_5)_2B(pz)_2]^{18}$ and 0.59 g (3.1 mmol) of $SnCl_2$ (30 mL of methylene chloride, 4 h stirring) to give 2.0 g (89%) of crude product. It was recrystallized from cyclohexane to give a material of mp 276–280 °C dec. Anal. Calcd for $C_{36}H_{32}B_2N_8Sn$ ($M_r = 717.0$): C, 60.30; H, 4.50; B, 3.02; N, 15.63; Sn, 16.55. Found: C, 60.46; H, 4.44; B, 3.30; N, 15.92; Sn, 16.08.

NMR data: $\delta(^1H)$ 7.58 (1 H, d, $J = 1.5$), 7.40 (1 H, d, $J = 1.2$), 7.25* (3 H, unresolved), 6.95* (2 H, very broad s), 6.19 (1 H, unsym t = two overlapping d, $J = ca. 1.9$); $\delta(^{11}B)$ 1.3 (s, $h_{1/2} = 500$ Hz); $\delta(^{13}C)$ 147*, 140.1, 137.7, 134.4, 127.7, 127.4, 104.2; $\delta(^{119}Sn)$ -728 (s, $h_{1/2} = 75$ Hz).

$[(C_6H_5)_2]_2SnCl$ (11) was obtained from 1.0 g (5.3 mmol) of $SnCl_2$ and 1.78 g (5.27 mmol) of $K[(C_6H_5)_2B(pz)_2]^{11}$ (30 mL of methylene chloride, 2 h stirring). The crude material (2.2 g, 92%) was purified by recrystallization from toluene to give a product of mp 234–238 °C dec. Anal. Calcd for $C_{18}H_{16}BClN_4Sn$ ($M_r = 435.1$): C, 47.67; H, 3.56; B, 2.38; Cl, 7.83; N, 12.36; Sn, 26.20. Found: C, 47.73; H, 3.41; B, 2.60; Cl, 7.70; N, 12.29; Sn, 26.22.

NMR data: $\delta(^1H)$ 7.94 (1 H, d, $J = 2.1$), 7.68 (1 H, d, $J = 2.4$), 7.33* (4 H, unresolved), 6.60* (1 H, unresolved), 6.40 (1 H, unsym t = two overlapping d, $J = 2.3$); $\delta(^{11}B)$ 1.5 (s, $h_{1/2} = 225$ Hz); $\delta(^{13}C)$ 139.9, 139.0, 135.7*, 135.6*, 128.3*, 105.4; $\delta(^{119}Sn)$ -353 (s, $h_{1/2} = 160$ Hz).

Results and Discussion

Preparation of Tin(II) Poly(1-pyrazolyl)borates and General Remarks.

Several Sn(II) poly(1-pyrazolyl)borates of the types L_2Sn and $LSnCl$ have been prepared by reaction of $SnCl_2$ with an alkali-metal poly(1-pyrazolyl)borate. The compounds include the species L_2Sn , where $L = [B(pz)_4]^-$ (1), $[B(pz')_4]^-$ (2), $[HB(pz)_3]^-$ (3), $[R_2B(pz)_2]^-$ ($R = H$ (5), C_6H_5 (6)), or $[H_2B(pz)_2]^-$ (7), and $LSnCl$ where $L = [B(pz)_4]^-$ (8), $[H_2B(pz)_2]^-$ (10), $[(C_6H_5)_2B(pz)_2]^-$ (11), or $[H_2B(pz')_2]^-$ (12) ($Hpz =$ pyrazole, $Hpz' =$ 3-methylpyrazole). All of these species were obtained by stirring a mixture of stoichiometric amounts of an alkali-metal poly(1-pyrazolyl)borate and $SnCl_2$ in methylene chloride at room temperature. Species of the type $LSnCl$ formed with particular

ease, whereas prolonged reaction times were frequently required for the preparation of L_2Sn species. The two compounds L_2Sn (4) and $LSnCl$ (9) with $L = [HB(pz^*)_3]^-$ ($Hpz^* =$ 3,5-dimethylpyrazole) have previously been described and were prepared by this same route.⁷

The complexes derived from $L = [H_2B(pz)_2]^-$ and those containing the 3-methylpyrazole moiety were extremely temperature sensitive and could not be recrystallized without apparent decomposition. Furthermore, many of the complexes, in particular those of the type $LSnCl$, underwent slow autogeneous changes, even when stored in the solid state, much more readily in solution. For example, the complex $[B(pz)_4]SnCl$ was also prepared and the 1H NMR (and mass spectroscopic) data of a freshly prepared sample substantiated the formulation. However, even after a short period of time the spectrum had deteriorated and new signals emerged; no reliable ^{119}Sn NMR data could be obtained. Recent studies in our laboratory suggest that the noted deterioration of some of the complexes on standing is likely due to hydrolysis. It has been observed that alkali-metal poly(1-pyrazolyl)borates are hygroscopic, a feature which has not been noted earlier. Thus, small amounts of water can be introduced into a system with the reagents. This may not necessarily interfere with the initial formation of the complexes but, obviously, can cause complications on storage. This feature may also explain the reported slow formation of pyrazole on standing of solutions of $[HB(pz)_3]Sn(CH_3)_3$ in methylene chloride.⁴

The X-ray crystal structures of $[HB(pz^*)_3]_2Sn$ (4) and $[HB(pz^*)_3]SnCl$ (9) have been determined earlier. In the solid state 9 was found to have approximately trigonal-bipyramidal geometry about tin (three N atoms, the Cl atom, and the lone pair), whereas the tin in 4 has an approximately octahedral environment (five N atoms and the lone pair). Hence, one of the pz^* groups of 4 was found to be different from the other five, which was not observed in the (room temperature) 1H NMR spectrum (solution in $CDCl_3$).⁷ However, it has now been observed (by 1H NMR spectroscopy) that in solution even at low temperature the equivalence of the six pz^* groups of 4 is maintained. This observation suggests that the solid-state structure and the solution structure are not identical, since coordination of the Sn(II) to the six N atoms of equivalent pz^* groups give it an effective coordination sphere of seven, when the lone pair is included.

The arrangement of the following discussion is based on the assumption that in solution the Sn(II) will achieve the maximum coordination of a discrete molecule.

Tin(II) Poly(1-pyrazolyl)borates with Formal Coordination Number 7.

The room-temperature 1H (and ^{13}C) NMR spectrum of $[B(pz)_4]_2Sn$ (1) exhibited only one set of signals for the pz groups. However, on cooling of 1 to $-40^\circ C$, considerable line broadening of the 1H NMR signals at 7.63 and 7.26 ppm occurred, whereas the signal at 6.32 ppm was essentially unaffected. New additional broad signals at about 6.8 and 7.9 ppm appeared near $-70^\circ C$, and at $-85^\circ C$ two distinct sets of pyrazolyl signals with $\delta(^1H)$ 7.93/6.81/6.38 (1 pz) and 7.59/7.40/6.38 (3 pz), respectively, were observed, indicating the presence of two different types of pz groups in 1:3 ratio. Hence, the species is fluxional in solution at room temperature, but the low-temperature 1H NMR data suggest that only six of the eight pyrazolyl groups in 1 are coordinated to tin. Under these circumstances the coordination sphere of tin comprises six N atoms and the lone pair. As expected, the ^{11}B NMR spectrum of 1 featured one sharp signal at 0.4 ppm, and $\delta(^{119}Sn)$ was observed at -744.

The 1H NMR spectrum of $[B(pz')_4]_2Sn$ (2) exhibited the anticipated two sets of pz' signals in 1:3 ratio even at room temperature; they merged into a single set near $50^\circ C$. With the exception of that of H(4), the room-temperature signals were extremely broad. Again, only one ^{11}B NMR and one ^{119}Sn NMR signal were observed and both were quite comparable to those of 1.

The room-temperature 1H NMR spectrum of $[HB(pz)_3]_2Sn$ (3) exhibited signals for only one type of pz group, and no additional signals emerged on lowering the temperature to $-95^\circ C$. This observation correlates well with the data obtained for 1, since

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Table I. ^{119}Sn chemical Shift Data for Tin(II) Poly(1-pyrazolyl)borates (Hpz = Pyrazole, Hpz' = 3-Methylpyrazole, Hpz* = 3,5-Dimethylpyrazole; Solvent CDCl_3)

no.	compd	$\delta(^{119}\text{Sn})$	$h_{1/2}$, Hz	presumed coordn sphere of Sn
1	$[\text{B}(\text{pz})_4]_2\text{Sn}$	-744	50	7
2	$[\text{B}(\text{pz}')_4]_2\text{Sn}$	-808	50	7
3	$[\text{HB}(\text{pz})_3]_2\text{Sn}$	-877	750 ^a	7
4	$[\text{HB}(\text{pz}^*)_3]_2\text{Sn}^b$	-933	350 ^a	7
5	$[\text{H}_2\text{B}(\text{pz})_2]_2\text{Sn}$	-654	450 ^a	5
6	$[(\text{C}_6\text{H}_5)_2\text{B}(\text{pz})_2]_2\text{Sn}$	-728	75	5
7	$[\text{H}_2\text{B}(\text{pz}')_2]_2\text{Sn}$	-685	360	5
8	$[\text{B}(\text{pz}')_4]\text{SnCl}$	-533	230	5
9	$[\text{HB}(\text{pz}^*)_3]\text{SnCl}^c$	-579	200	5
10	$[\text{H}_2\text{B}(\text{pz})_2]\text{SnCl}$	-305	1250 ^d	4
11	$[(\text{C}_6\text{H}_5)_2\text{B}(\text{pz})_2]\text{SnCl}$	-353	160	4
12	$[\text{H}_2\text{B}(\text{pz}')_2]\text{SnCl}$	-271	365	4

^a $h_{1/2}$ narrows at lower temperature. ^b Lit.:⁷ $\delta(^{119}\text{Sn})$ -935 ($h_{1/2}$ = 300 Hz) in CH_2Cl_2 . ^c Lit.:⁷ $\delta(^{119}\text{Sn})$ -1460 ($h_{1/2}$ = >300 Hz) in CH_2Cl_2 . This datum seems to be in error. ^d At low temperatures the signal splits into two with δ -262 and -336, respectively.

the only difference between the compounds is the replacement of the terminal pz group (i.e., the one which is not bonding to the Sn(II) in the static structure) by H. It suggests that in solution all three pz groups of the $\text{HB}(\text{pz})_3$ ligands in **3** are bonding with the tin, even at low temperatures. Also, as noted above, no new ^1H NMR signals emerged in the ^1H NMR spectrum of **4** at low temperatures. Hence, unless even at -95 °C rapid (on the NMR time scale) fluxionality persists, the solution structure of **4** is different from that in the solid state.

Both **3** and **4** exhibited only one ^{11}B NMR signal, and the ^{119}Sn chemical shift of **3** (δ -877) is quite comparable to that reported for **4** (δ -935).⁷ Furthermore, the ^{119}Sn NMR data of the species **1-4** are all very similar and fall in the range of about -750 to -950 ppm (see Table I). This suggests that in solution the coordinative environments of the Sn(II) in these compounds are identical. Since the NMR data indicate the binding of a total of six N atoms of the ligands to the tin, the cited chemical shift range seems to reflect 7-fold coordination (when including the lone pair) for Sn(II) in these poly(1-pyrazolyl)borates.

Tin(II) Complexes with Formal Coordination Number 5. A maximum coordination sphere of 5 may be expected for complexes of the type $[\text{R}_2\text{B}(\text{pz})_2]_2\text{Sn}$ (R = noncoordinating substituents), which results from the binding of four N atoms of the ligands to the Sn(II) and the lone pair of the latter.

The compound $[\text{H}_2\text{B}(\text{pz})_2]_2\text{Sn}$ (**5**) was only obtained as crude (although, on the basis of ^1H NMR data, reasonably pure) material, which could not be purified by recrystallization or sublimation. Rather, all purification attempts resulted in deterioration of the sample and the formation of increasing amounts of free pyrazole was observed. The compound $[\text{H}_2\text{B}(\text{pz}')_2]_2\text{Sn}$ (**7**) could be somewhat purified by sacrificing yield.

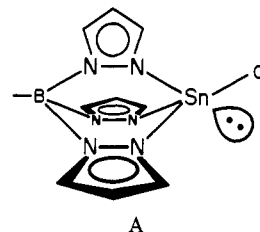
Furthermore, the reaction of SnCl_2 with 2 molar equiv of $\text{K}[(\text{C}_2\text{H}_5)_2\text{B}(\text{pz})_2]$ gave no pure identifiable product and there were indications (NMR spectra) that extensive ethyl group migration had occurred. In addition, the formation of substantial amounts of free pyrazole was observed, although all work was done under anhydrous conditions. On the other hand, when $\text{Na}[(\text{C}_6\text{H}_5)_2\text{B}(\text{pz})_2]$ was employed in an analogous reaction, the desired $[(\text{C}_6\text{H}_5)_2\text{B}(\text{pz})_2]_2\text{Sn}$ (**6**) was obtained in good yield.

The ^{119}Sn chemical shifts of **5-7** were quite comparable with each other ranging from about -650 to -730 ppm. The broadness of the signal of **5** was somewhat disturbing; however, the signal sharpened considerably on lowering of the temperature. This feature may be due to conformational changes in these complexes with two puckered BN_4Sn rings, i.e., containing the skeleton $\text{B}(\mu\text{-pz})_2\text{Sn}(\mu\text{-pz})_2\text{B}$.

In poly(1-pyrazolyl)borate complexes of Sn(II) of the type LSnCl , a maximum of three pyrazolyl groups can bind to the

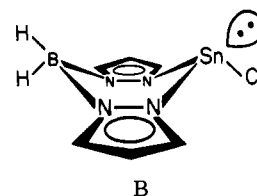
central metal atom if they are available. The room-temperature ^1H NMR spectrum of $[\text{B}(\text{pz}')_4]\text{SnCl}$ (**8**) exhibited only one set of signals for the pz' groups. However, a second set began to emerge at -50 °C, and at -70 °C two sets of pz' signals in 1:3 ratio were clearly seen. This documents the expected binding of three N atoms of the ligand to the tin. The ^{11}B and the ^{119}Sn NMR spectrum of **8** exhibited only one signal each. On the other hand, the ^{119}Sn chemical shift, δ -533, differed significantly from that reported⁷ for $[\text{HB}(\text{pz}^*)_3]\text{SnCl}$ (**9**) with δ -1460. However, this latter datum is apparently in error: When the spectrum of freshly prepared **9** was recorded in CDCl_3 during the course of the present work, $\delta(^{119}\text{Sn})$ was observed at -579.

It is apparent from the NMR data that the coordination of Sn in compounds **8** and **9** in solution is the same and most likely to be 5-fold: three N atoms of the ligands, one Cl atom, and the lone pair, as is illustrated in structure A. The ^{119}Sn chemical



shifts of the compounds fall in the range between those of compounds of types **1-4** (with a presumably higher coordination of the Sn) and those of **5-7** (with different ligands surrounding the Sn). Apparently, the ^{119}Sn chemical shift range of about -530 to -730 ppm seems to reflect 5-fold coordination about the Sn(II) in these poly(1-pyrazolyl)borate species.

Species with Formal Coordination Number 4. Finally, the species $[\text{H}_2\text{B}(\text{pz})_2]\text{SnCl}$ (**10**), $[(\text{C}_6\text{H}_5)_2\text{B}(\text{pz})_2]\text{SnCl}$ (**11**), and $[\text{H}_2\text{B}(\text{pz}')_2]\text{SnCl}$ (**12**) were examined. In all of these the tin of discrete species should be in a four-coordinate environment, i.e., bonded with two N atoms from the poly(1-pyrazolyl)borate moiety and the Cl atom, in addition to the lone pair. The ^{119}Sn chemical shifts of these three species fall in a fairly narrow range from about -270 to -350 ppm. This is not surprising considering the cited environment of the Sn(II). On the other hand, the signal of **10** seemed to be exceedingly broad as compared to those of **11** and **12**. However, when the temperature was lowered, the ^{119}Sn and NMR signal of **10** began to split, and at -40 °C two separate signals at -262 and -336 ppm, respectively, were observed. It is reasonable to assume that both signals indicate the same coordination sphere for the Sn(II). This can be explained by assuming that the puckered BN_4Sn ring of **10** (structure B) readily



inverts at room temperature thus causing the broadness of the signal, whereas at low temperature two structures of different geometries are frozen out. Such an inversion may be impaired by the methyl groups of the ligand in **12** and by a relatively large phenyl group in **11**.

Conclusion

On the basis of the preceding data, ^{119}Sn NMR spectroscopy seems to be a useful probe for the determination of the coordination sphere of Sn(II) in poly(1-pyrazolyl)borate complexes in solution. However, it should be emphasized that the solution structures are not necessarily identical with those in the solid state, as is seen in the case of compound **4**.

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