

New Insights into the Coordination Chemistry and Molecular Structure of Copper(II) Histidine Complexes in Aqueous Solutions

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Aqueous solutions of Cu²⁺/histidine (his) (1:2) have been analyzed in parallel with infrared, Raman, ultraviolet/visible/near-infrared, electron spin resonance, and X-ray absorption spectroscopy in the pH range from 0 to 10. Comprehensive interpretation of the data has been used to extract complementary structural information in order to determine the relative abundance of the different complexes. The formation of six different, partly coexisting species is proposed. Structural proposals from literature have been unambiguously confirmed, refined, or, in several cases, corrected. At highly acidic conditions, Cu²⁺ and his are present as free ions, but around pH = 2, coordination starts via the deprotonated carboxylic acid group. This results in the intermediate species Cu²⁺[H₃his⁺(O_c)] and Cu²⁺[H₃his⁺(O_c)₂]. The coordination via O_c is attended with a drop in the pK_a value of the other receptor groups resulting in a concomitant conversion to the bidentates Cu²⁺[H₂his⁰(O_c,N_{am})] and Cu²⁺[H₂his⁰(O_c,N_{am})₂], with the latter being dominant at pH = 3.5. Coordination of the imidazole ring begins around pH = 3 and leads to the formation of the mixed ligand complexes Cu²⁺[H₂his⁰(O_c,N_{am})] [His⁻(O_c,N_{am},N_{im})] and Cu²⁺[His⁻(N_{am},N_{im})] [His⁻(O_c,N_{am},N_{im})] around pH = 5. It is demonstrated that coordination of the imidazole ring occurs predominantly via the N⁷ atom. At pH > 7, the double-tridentate ligand complex Cu²⁺[His⁻(O_c,N_{am},N_{im})₂] is the major species with the N atoms in the equatorial plane and the O atoms in the axial position. This complex decomposes at pH > 10 into a copper oxide/hydroxide precipitate. The overall results provide a consistent picture of the mechanism that drives the coordination and complex formation of the Cu²⁺/his system.

Introduction

Cu is one of the most abundant transition metals present in living systems, and Cu²⁺ ions are well-known to play a key role in the working mechanism of several important metalloenzymes, such as galactose oxidase and superoxide dismutase. A driving force in the activity of these enzymes is the folding of a peptide, which consists of amino acid side chains, around the Cu²⁺ cation. Histidine (his) moieties are known to be of major importance in this process because this type of N-coordinating ligand forms fairly stable complexes with the Cu cation. To unravel the formation processes of these types of enzymes and to find the exact

coordination geometry of the Cu/his moieties, a lot of research has been conducted in this area,^{1–17} much of which has been summarized in two recent reviews.^{18,19}

A wide variety of analytical tools have been applied to obtain this type of information directly from the enzyme systems, but the analyses are often hampered by interference

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of the protein matrix. As an alternative, complexes of Cu^{2+} with amino acid ligands are used as model systems, and this approach is even more interesting because complexes of Cu^{2+} with his can be used to mimic the working mechanism of Cu enzymes. Another important motivation to study Cu^{2+} /his species is their expected role in the development of man-made enzymes via functionalization of inorganic hosts, such as zeolites and clay minerals, with these types of complexes.^{20–24}

Elucidation of the coordination geometry of the Cu^{2+} /his complexes, however, is not easy, because this amino acid ligand has four potential sites for (de)protonation and thus for metal binding, i.e., the carboxylate group, the primary amino group, and two different N atoms of the imidazole ring. In principle, each of these sites can be involved in the coordination to Cu^{2+} , which may result in mono-, bi-, or tridentate binding forms. Besides, the state of protonation of the different sites, the manner of coordination, and the coexistence of several different complexes are all pH-dependent. Furthermore, the Cu/his ratio, the Cu^{2+} concentration, and the ionic strength of the solution are generally known to affect the structures of the complexes. To elucidate the molecular structures of Cu^{2+} /his complexes and to obtain insight into the forces that drive anchoring of the metal ion, many studies have been carried out since the late 1960s. In these studies, a wide variety of techniques have been used to obtain the desired information, ranging from spectroscopic techniques, like circular dichroism,^{10,25,26} ultraviolet/visible/near-infrared (UV/vis/NIR),^{10,25–29} infrared (IR),^{3,14,26–33}

Raman,^{3–5,10,13,15–17,33–35} ^1H NMR,^{10,32,36–39} ^{13}C NMR,³⁶ electron spin resonance (ESR),^{10,27,37,40–48} and extended X-ray absorption fine structure (EXAFS),^{6,49} to potentiometric^{10,26,28,50–54} and calorimetric⁵⁵ experiments.

Several Cu^{2+} /his complexes consisting of one or two central Cu atoms have been proposed in these studies, with one or two his ligands in a mono-, bi-, or tridentate coordination and neutrally, negatively, or positively charged. However, despite all of these efforts, controversy on the mode of his binding to Cu^{2+} and the geometrical structure of the different complexes still exists. As an example, a point of dispute is that Cu^{2+} prefers a square-planar or highly distorted octahedral configuration, which would exclude coordination via more than two binding sites of the same his molecule.^{31,54} Furthermore, Itabashi and Itoh³⁴ reported the existence of four different species and stated that Cu^{2+} is bound predominantly as a glycine (gly)-like chelate via the deprotonated NH_2 group and the carboxylate group. In contrast, Kruck and Sarkar²⁹ suggested eight different complexes, of which several involve coordination of the imidazole N atoms.

A major cause for the conflicting views is that conclusions are often based on only one or two characterization techniques, occasionally combined with theoretical calculations. Moreover, the corresponding experimental conditions are usually different because they have been optimized for a specific characterization technique. To mention a few

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examples, the weak scattering efficiency in Raman spectroscopy requires high Cu concentrations, whereas the opposite situation is desired in UV/vis/NIR because of the intense blue color of the complexes dissolved in water. On the contrary, IR spectroscopic studies on aqueous systems are largely hampered by the interference of water absorptions. For that reason, the preferred method for this technique is analysis after lyophilization, crystallization, or the use of D₂O. Finally, (pulsed) ESR and electron–nuclear double resonance studies have been carried out on frozen solutions at temperatures even down to 4 K.^{20,40,41,43} Hence, it is very difficult to compare data and to draw objective and integrated conclusions.

In our opinion, a more successful strategy to gain more insight into the coordination chemistry of Cu²⁺/his complexes is to combine several complementary analytical techniques in one integrated approach, which allows analysis at more or less the same sample conditions. To our knowledge, such a multitechnique approach has not been reported for Cu²⁺/his complexes so far. Five techniques that, within certain limits, meet the above-mentioned requirements are IR, Raman, UV/vis/NIR, ESR, and X-ray absorption spectroscopy (XAS). As demonstrated in a recent paper,⁵⁶ the integrated use of IR and Raman has proven to be an ideal tool to distinguish all possible states of protonation of his, while UV/vis/NIR is the method of choice to determine the identity and the amount of atoms that are coordinating to the central Cu²⁺ ion from the Cu d–d transition band. In addition, ESR can be used to discriminate between coordination of O and N in the equatorial plane of the complexes. Finally, EXAFS can supply information on the distance between the central Cu cation and the O and/or N atoms in the first coordination sphere and thus on changes in the local structure around the Cu²⁺ ion. Moreover, an important advantage of these characterization techniques is that they can be applied not only at physiological conditions but also at more extreme pH values. For that reason, we decided to carry out a systematic study, using this multispectroscopic approach, on Cu²⁺/his complexes in aqueous solutions, which covers the widest possible pH range. As a result, we propose a more consistent picture for the mechanism that drives the complex formation and of the molecular structure of this very important Cu²⁺/his system.

Experimental Section

1. Chemicals and Solutions. All chemicals (p.a. grade) were obtained from Acros Organics (Geel, Belgium). For each experiment, solutions of L-histidine, Cu(NO₃)₂·3H₂O, and CuCl₂ (all p.a. grade) were freshly made by dissolving each of the components in demineralized water at various concentrations. All solutions were prepared in a Cu²⁺/his ratio of 1:2 and acidified by the addition of HCl or made basic by the addition of NaOH in order to prevent interfering buffer absorption in the IR or Raman spectra. To minimize the reducing effect of the X-ray beam on the samples, Cu(NO₃)₂·3H₂O was used as the precursor salt instead of CuCl₂ and HNO₃ was used instead of HCl to adjust the pH for the EXAFS

analyses.⁵⁷ Concentrations were 0.001 M (Cu²⁺) for ESR, 0.01 M for EXAFS and simultaneous UV/vis/NIR, and 0.25 M for IR, Raman, and UV/vis/NIR analyses. IR, Raman, and UV/vis/NIR measurements were carried out simultaneously with the same 0.25 M solutions. Additional UV/vis/NIR analyses of the 0.01 M solutions were performed to detect possible concentration effects and, hence, to facilitate a comparison of the IR and Raman data with the results of the more diluted samples used for EXAFS and ESR. All experiments were done at ambient temperature (298 K), and a pH meter (model PHM 210 from Radiometer Analytical, USA, Loveland, CO) was used to determine the pH values. To establish the stability of the Cu²⁺/his complexes and the reversibility of the complexation process, titration experiments were carried out from pH = 0 up to pH = 12 and backward and from pH = 0 up to pH = 10 and backward. Both experiments were monitored by IR, Raman, and UV/vis/NIR. Solutions of Cu with gly and histamine (him) were studied as structural analogues of the Cu²⁺/his complex.

2. Characterization. IR measurements during titrations were performed at room temperature on an Autochem ReactIR-4000 system (Mettler Toledo) equipped with a Dicomp attenuated total reflectance probe and a nitrogen-cooled mercury/cadmium/telluride detector. The sample compartment was flushed with dry air to reduce interference from water and carbon dioxide. For each spectrum, 256 scans were accumulated at a data-point resolution of 2 cm⁻¹ (optical resolution of 4 cm⁻¹). Spectral interpretation was carried out after subtraction of the spectrum of water at the corresponding pH as the background. Multivariate curve resolution (MCR) was carried out on the IR spectra, using model calculations in Matlab. Details on the MCR procedure can be found in the Supporting Information.

Raman measurements were carried out on a Kaiser RXN spectrometer equipped with a 70-mW, 532-nm diode laser for excitation, a holographic grating for dispersion, and a Peltier-cooled Andor CCD camera for detection. Scanning was performed via a 10-mm objective probe head that was mounted just above the surface of the solution. The solution was stirred and kept at room temperature during the measurements to prevent local heating, i.e., degradation of the sample. The detector pixel resolution was about 2 cm⁻¹, and 15 scans were accumulated for one spectrum at an exposure time of 30 s/scan. Subtraction of the water background has not been applied for the Raman spectra.

UV/vis/NIR spectra of the aqueous solutions were measured from 200 to 1000 nm at a spectral resolution of 0.15 nm with a Cary 50 spectrometer (Varian) equipped with a probe head that was connected to the spectrometer via optical fibers. The effective optical path length in the probe head was 2 mm for the 0.25 M samples, whereas the path length was 10 mm for the 0.010 M samples.

X-band ESR measurements were carried out on a Bruker ESP300 (Utrecht) and a Bruker ESP300E spectrometer (Leuven), both operating near 9.5 GHz. ESR spectra were recorded at room temperature, using 80- μ L quartz capillaries as sample cells. Spectra were taken at pH values of 2.0, 2.9, 3.4, 4.4, and 7.3. An in-house-developed computer program was used for simulation of the Cu²⁺/his spectra, taking into account the natural abundances of the ⁶³Cu (69.2%) and ⁶⁵Cu (30.8%) isotopes. The EPR spectra are motionally narrowed and can be described in the Redfield limit as a superposition of Lorentzians because, at room temperature, the Cu complexes are rapidly tumbling in the solution. For that reason, the spectrum

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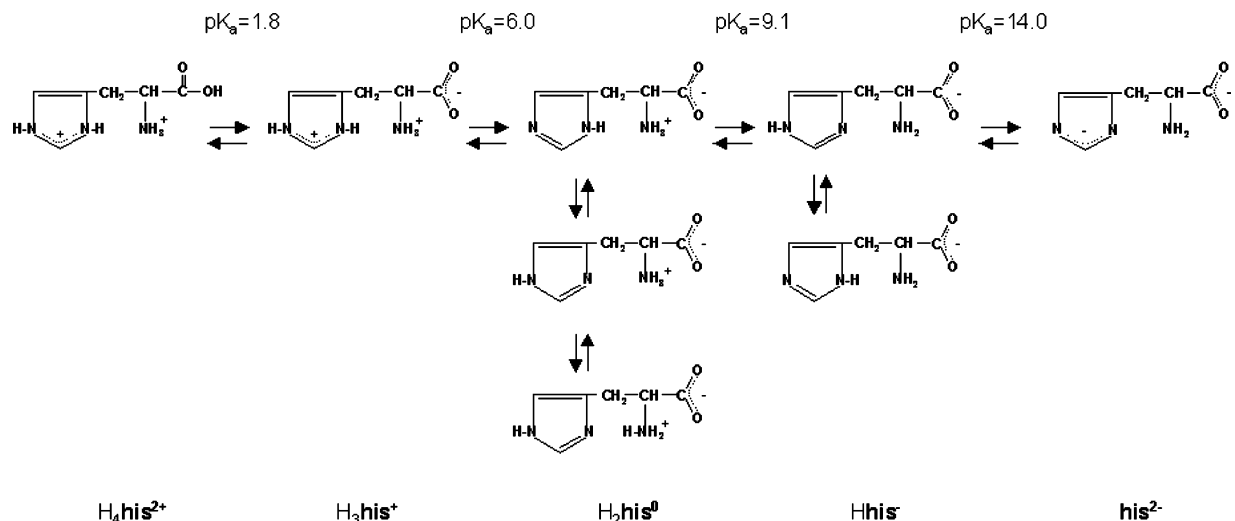


Figure 1. Different states of protonation and corresponding pK_a values of his.

of each species has been simulated as the weighted superposition of eight Lorentzians, i.e., four from each Cu isotope ($I = 3/2$). The different line widths of the Lorentzians as a result of the difference in the hyperfine quantum number were taken into consideration by fitting the relaxation parameters α , β , and γ , according to the formula $T_2^{-1} = \alpha + \beta m + \gamma m^2$. In this way, the g and A values, together with the relative abundance of the different Cu^{2+}/his complexes at various pHs, were determined. To compensate for the difference in the nuclear moments of ^{63}Cu and ^{65}Cu , the A values were calculated with $A^{63} = A_0$ and $A^{65} = 1.07A_0$.

XAS data were collected on Beamline BM26A at the ESRF (Grenoble, France), which operates at 6 GeV with a typical beam current of 200 mA. The station was equipped with an Si(111) double-crystal monochromator and ion chambers for measuring incident and transmitted beam intensities. Measurements were performed at the Cu K edge (8979 eV) in transmission mode on samples at pH = 2.0, 2.9, 3.4, 4.4, 6.0, and 8.0 at room temperature in a stainless steel cell (diameter of 2.5 mm) with Kapton windows. Higher harmonic contributions were filtered by the optics of the setup. A Cu foil was applied to calibrate the energy of the monochromator. Several scans were averaged to improve the signal-to-noise ratio of the data. The XAFS data were processed using the suite of programs available at Daresbury Laboratory, namely, *EXCALIB* (for converting the raw data to energy versus the absorption coefficient) and *EXBROOK* (for *pre*- and *post*-edge background removal) to obtain the normalized X-ray absorption near-edge structure (XANES) part of the spectra and *EXCURV98*⁵⁸ to carry out EXAFS data analysis. EXAFS refinements were carried out on k^1 -weighted spectra plotted over a k range of 3–13.3 \AA^{-1} , considering only single-scattering paths. An amplitude reduction factor (S_0^2) value of 0.9 obtained from the fitting of the Cu metal foil was also used in the analyses. Because the Cu–O and Cu–N bond distances and backscattering amplitudes are about similar, it is difficult to distinguish between their respective contributions in EXAFS. This is a particular problem for the Cu^{2+}/his system, in which the two different atoms are coordinating in the same coordination shell. Because the fitting of two shells with the respective atoms leads to unphysical results, the refinements were performed considering only a single shell of either Cu–O or Cu–N contributions, with the best fit to the data being determined by the EXAFS R factor (goodness of fit).

Results

As described in a previous paper,⁵⁶ his in an aqueous solution can be present in five different pH-dependent states of protonation, i.e., H_4his^{2+} , H_3his^+ , H_2his^0 , $Hhis^-$, and his^{2-} (Figure 1). The molecule has four potential sites for deprotonation and thus for metal coordination, but metal binding via the deprotonated imidazole ring (his^{2-}) is generally rejected regarding the extremely high pK_a value of this element ($pK_a = 14$). Furthermore, the ionic state H_4his^{2+} lacks a site available for coordination and this leaves, in principle, H_3his^+ , H_2his^0 , and $Hhis^-$ as possible ligating structures. On the other hand, the presence of two different N atoms in the imidazole ring also allows H_2his^0 and $Hhis^-$ to be present in two pH-dependent tautomeric forms. The commonly accepted notation to distinguish both N atoms is N^π for the one closest to the CH_2 group and N^τ for the other one. Unfortunately, a uniform nomenclature to describe the different Cu^{2+}/his complexes is absent in the literature. To prevent further confusion, we define the state of protonation of his as described in Figure 1, while for the different metal binding sites, the following notation is used: O_c for the carboxylate group, N_{am} for the amino group, and N_{im} for the imidazole N atoms, which can be denoted in more detail, if appropriate, as N^π for the N closest to the CH_2 group and N^τ for the other one.

Furthermore, it should be noted that at certain pH values two, three, or even more different complexes could coexist, which implies that spectra will not always correspond to single species. Obviously, this hampers a straightforward interpretation and unambiguous structural elucidation. Spectral subtraction, deconvolution, and difference spectroscopy have been proposed in the literature to partially circumvent this problem, but, in general, these methods were found to be not very successful because of the lack of reliable standards and the large number of factors that may affect band positions and intensities. For those reasons, these methods have not been applied here. Instead, to obtain as much information as possible, titration experiments have been carried out under well-controlled experimental conditions

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and, when appropriate, with increments of about 0.1 pH units between two subsequent measurements. As described in the Experimental Section, two different titration experiments were performed to check the reversibility of the complex formation. In the first experiment, the pH of the $\text{Cu}^{2+}/\text{his}$ solution was raised from pH = 0 up to pH = 12 and back, but we observed precipitation and a change of color of the solution above pH = 10. The same observations were reported by Itabashi and Itoh,³⁴ and in accordance with these authors, we identified the precipitate as copper oxide/hydroxide. However, we also noticed that the complex formation process was irreversible when passing pH = 12 and the same result was obtained when titration was started at pH = 12. Furthermore, precipitation turned out to proceed faster when the solution was not freshly prepared. These phenomena were not observed when a pH value of 10 was not exceeded, and for that reason, spectral data at pHs higher than 10 have been rejected. Implicitly, this implies that we also reject the existence of a binuclear $\text{Cu}^{2+}/\text{his}$ complex at pH > 10.5, as proposed in the literature.²⁹

1. Vibrational Spectroscopy. The large number of IR and Raman spectra, recorded during the titration experiments, allowed representation of the data in 3-D plots (see Figure S1 in the Supporting Information). For IR, the resulting picture illustrated the overall changes in the spectra as a function of the pH, but determination of small spectral changes required closer examination of the individual spectra. A similar 3-D representation of the Raman spectra turned out to be less informative because the band intensities drop dramatically at higher pH as a result of strong darkening of the solution. As an alternative, a number of representative IR and Raman spectra at pH values of 1, 2, 4, and 8 are presented in Figures 2 and 3 (for a more detailed overview of the spectra, we refer to the Supporting Information). Assignments (as indicated in Figures 2 and 3) are proposed after spectral interpretation in the Discussion section, taking into account IR and Raman selection rules and assignments from our recent work.⁵⁶ In addition, spectral details, which are relevant to the determination of coordination sites and the geometrical structure of complexes, are presented.

2. UV/Vis/NIR Spectroscopy. The UV/vis/NIR absorption spectra of aqueous 0.01 M $\text{Cu}^{2+}/\text{his}$ solutions as a function of the pH are presented in Figure 4.

According to the literature, the band with λ_{max} between 600 and 800 nm is assigned to the d–d transition of Cu^{2+} .^{59,60} The intensity of this band increases with the pH, and the position shifts to shorter wavelength up to pH = 5, after which it shifts back to longer wavelength for pH > 5. The absorption band around 285 nm increases very strongly with the pH and shifts to longer wavelength. To determine the origin of the intensity increase and the shift in the position of the d–d transition band, the spectra of the $\text{Cu}^{2+}/\text{his}$ system were compared with the UV/vis/NIR data of solutions containing Cu^{2+} complexes with him and gly, recorded over

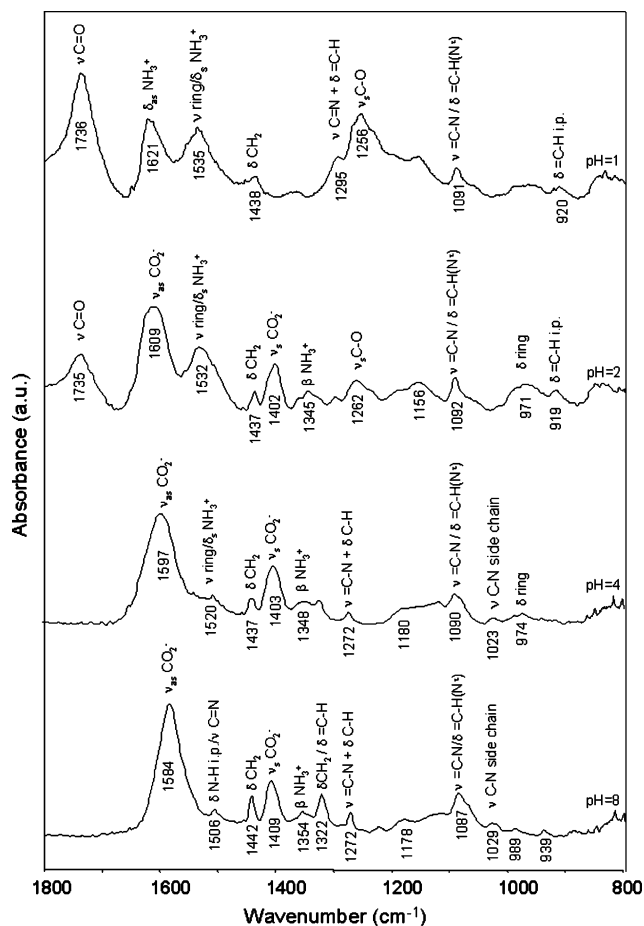


Figure 2. IR spectra of $\text{Cu}^{2+}/\text{his}$ solutions at pH = 1 (top), pH = 2, pH = 4, and pH = 8 (bottom). The assignments of the vibrations are indicated in the figure.

the same pH range. These ligands have been chosen because they have only two coordinating sites, which are highly comparable to the ones of his, i.e., N_{am} and N_{im} for him and N_{am} and O_c for gly. Besides, the coordination geometry of these Cu^{2+} complexes is well-known and relatively simple. The peak position of the d–d transition band turned out to be the same for all compounds at pH = 1 but differ at higher pH values. Furthermore, all samples exhibit a considerable shift to shorter wavelength, but the band of $\text{Cu}^{2+}/\text{his}$ passes a minimum at pH = 5, whereas the positions of the $\text{Cu}^{2+}/\text{gly}$ and $\text{Cu}^{2+}/\text{him}$ solutions stabilize at higher pH values (see Figure S4 in the Supporting Information). A deviating behavior of $\text{Cu}^{2+}/\text{his}$ is also observed for the intensity of the d–d transition band as it increases up to pH = 10, whereas for $\text{Cu}^{2+}/\text{him}$ and $\text{Cu}^{2+}/\text{gly}$, it remains constant at pH > 7. Finally, it should be noted that the data obtained from the 0.01 M $\text{Cu}^{2+}/\text{his}$ solutions showed no significant differences compared to the 0.25 M samples.

3. ESR. As was expected for a Cu^{2+} ($I = 3/2$) complex in solution at room temperature, the first-derivative ESR spectra recorded at pH values of 4.4 and 7.3 (Figure 5) show four hyperfine lines. In contrast, at least five hyperfine lines are distinguished at pH = 2.0, 2.9, and 3.4, which indicates that more than one $\text{Cu}^{2+}/\text{his}$ complex must be present in the corresponding solutions. To determine the ESR parameters of the complexes and their relative contribution, spectrum

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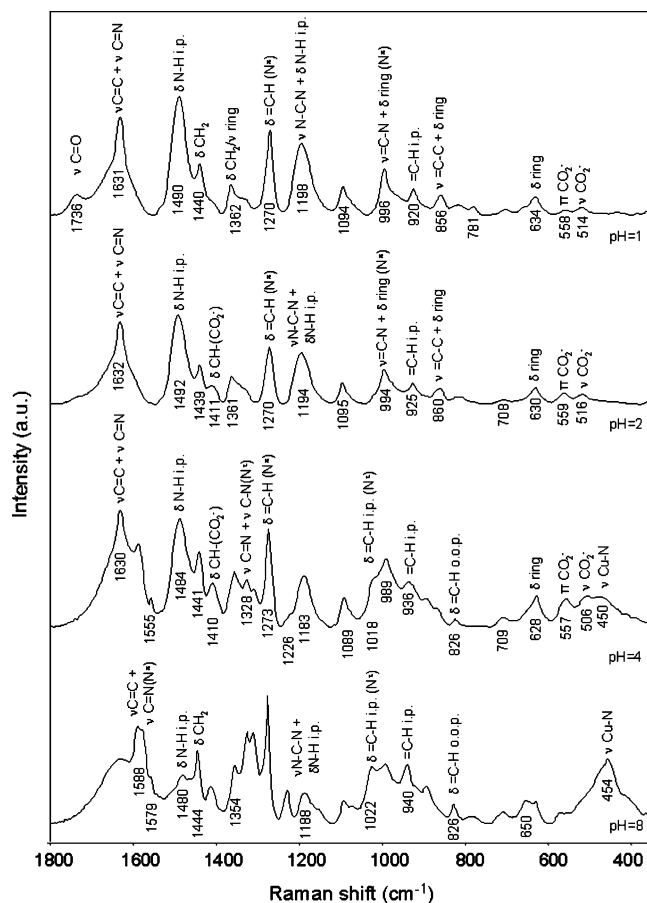


Figure 3. Raman spectra of $\text{Cu}^{2+}/\text{his}$ solutions at pH = 1 (top), pH = 2, pH = 4, and pH = 8 (bottom). The assignments of the vibrations are indicated in the figure.

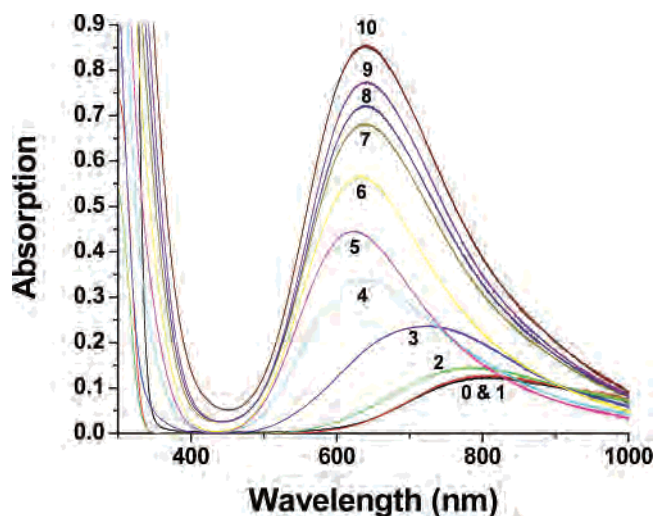


Figure 4. UV/vis/NIR spectra of the 0.01 M $\text{Cu}^{2+}/\text{his}$ (1:2) solutions in the pH range of 0–10 (pH values are indicated).

simulations were carried out. The natural abundance of the Cu isotopes and the different widths of the Cu hyperfine lines were taken into account upon fitting the relaxation parameters. The results of the calculations obtained for all samples are summarized in Table 1. The results revealed the presence of five different $\text{Cu}^{2+}/\text{his}$ complexes in the pH range of 2.0–7.3 in a relatively large abundance, each of

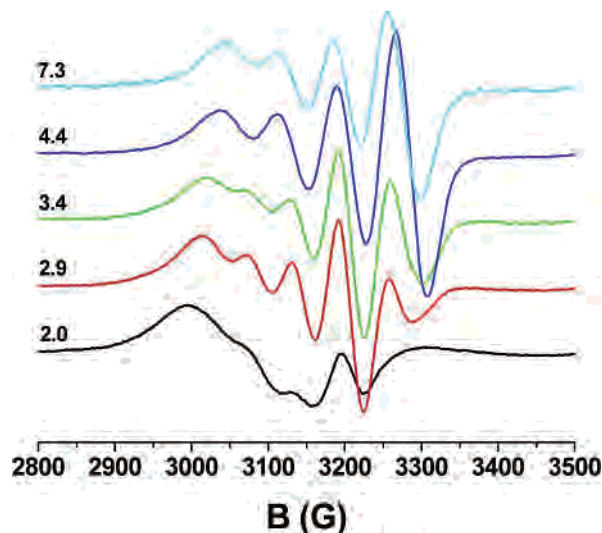


Figure 5. First-derivative ESR spectra of 0.001 M $\text{Cu}^{2+}/\text{his}$ solutions at pH values of 2.0, 2.9, 3.4, 4.4, and 7.3.

Table 1. Optimized ESR Parameters Obtained from Simulation of the $\text{Cu}^{2+}/\text{his}$ ESR Spectra at Different pH Values^a

pH	species	abundance (%)	g_0 value	A_0 value (G)	α (G)	β (G)	γ (G)
2.0	$\text{Cu}^{2+}(\text{H}_2\text{O})_6$	85	2.193	34	49	-1.6	0.9
	c	15	2.154	58	31	-8.9	1.3
2.9	c	75	2.154	58	33	-6.9	2.8
	e	25	2.126	74	28	-10	4.2
3.4	c	55	2.154	58	31	-8.9	2.3
	e	45	2.126	74	29	-10	2.9
4.4	f	90	2.119	72	30	-10	3.5
7.3	h	90	2.120	64	31	-10	2.6

^a Details on the simulation procedure are given in the Experimental Section. The species indicated as c, e, f, and h refer to the structures as depicted in Figure 7. Estimated overall accuracies: abundance, 10%; g_0 value, 0.002; A_0 value, 2 G; α , 2 G; β , 1 G; γ , 0.5 G.

them characterized by their own g_0 and A_0 values. A rough estimate of the relative contribution of these species is also included in Table 1. Occasionally, different values have been obtained for the same species at different pHs. This reflects the inaccuracy of the simulation and is probably due to differences in the physical properties of the solution (pH, viscosity, and dielectric constant).

4. EXAFS Analysis. The XANES of the XAS spectra of four $\text{Cu}^{2+}/\text{his}$ samples, the background-subtracted EXAFS data (k^1 -weighted) at various pHs, and the associated Fourier transforms (including both the real part and the transform envelope) are presented in the Supporting Information (Figures S6–S8). The results of the fitting are presented in the Supporting Information (Table S1). From the analyses, it can be concluded that the bond distance sequence as a function of the pH is as follows: pH = 2.0 \approx pH = 2.9 \approx pH = 3.4 < pH = 4.4 < pH = 6.0 < pH = 7.3 \approx pH = 8.0. Most notably, the best fit for the first shell for the samples at pH = 2.0, 2.9, and 3.4 is obtained with four O atoms at a distance of 1.95 Å. At pH = 4.4, this distance has increased to 1.96 Å. For pH = 6.0, 7.3, and 8.0, a distance of 1.98–1.99 Å was found, using N atoms in the fitting procedure.

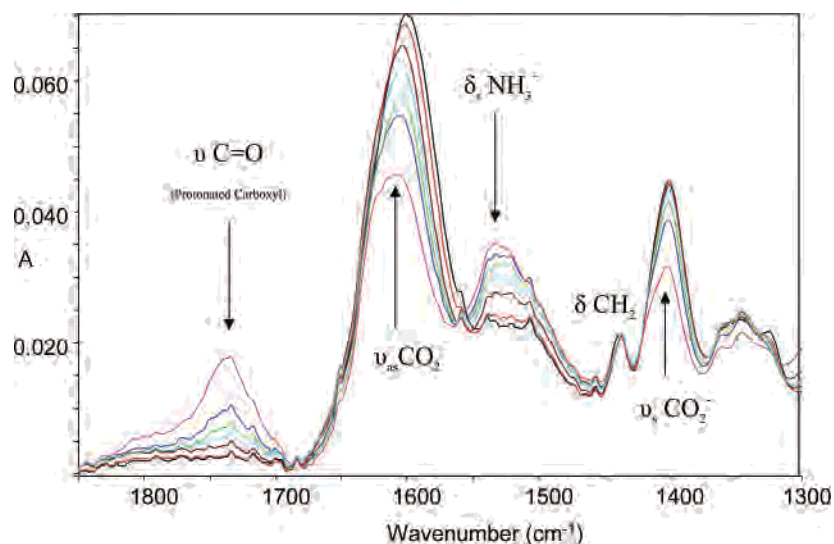


Figure 6. IR spectra of aqueous solutions of 0.25 M $\text{Cu}^{2+}/\text{his}$ at $\text{pH} = 1.95\text{--}3.55$. The arrows indicate the increasing (\uparrow) and decreasing (\downarrow) intensities at increasing pH.

Discussion

A large amount of spectral data, covering a variety of spectral changes as a function of pH, have been obtained from the different experiments. To systematically relate the data to (changes in) the geometrical structures of $\text{Cu}^{2+}/\text{his}$ complexes, the results will be discussed in subsections that represent the most important pH ranges.

pH Range of 0–2. The IR and Raman spectra of the $\text{Cu}^{2+}/\text{his}$ solutions recorded up to $\text{pH} = 1.8$ appeared to be virtually identical with the ones of the his solutions at the same conditions. The presence of the C=O and C–O(–H) stretching bands in the IR spectra around bands at 1736 and 1257 cm^{-1} , respectively, confirms that the carboxyl group is largely protonated, in accordance with the pK_a of 1.8. This implies that his is exclusively present as a free, noncoordinated ligand in the double-protonated ionic form, $\text{H}_2\text{his}^{2+}$. This is confirmed by the UV/vis/NIR data. The Cu solutions with his, him, and gly at $\text{pH} = 1$ exhibit the same d–d transition band that is characteristic for Cu^{2+} surrounded by O atoms, i.e., as a hexa-aqua complex. The same conclusion can be drawn from the ESR data. It is known that the effect of the nuclear spin of Cu on the unpaired electron will be larger when the density of the d electron at the position of the Cu nucleus increases. When only O atoms are coordinating, the density of the unpaired paramagnetic electron at Cu^{2+} is at its minimum, which is accompanied by relatively high g and low A values compared to coordination by N atoms. The simulation of the ESR spectrum of the $\text{Cu}^{2+}/\text{his}$ solution at $\text{pH} = 2$ indicates that about 85% of Cu^{2+} is present as a complex with relatively lowest A and highest g values (Table 1), and from the literature, it is known that these values are characteristic of the Cu hexa-aqua complex.⁶³ The same is concluded from the results of EXAFS. In accordance with

the literature, a perfect fit is obtained with four O atoms from water in the first shell of the equatorial plane at a distance of 1.95 Å.⁶⁴

pH Range of 2–4. The first differences compared to the IR and Raman spectra of free his are observed at about $\text{pH} = 2$. The changes in IR are well illustrated by the spectra recorded during titration from $\text{pH} = 1.95$ to 3.55, as shown in Figure 6. The intensity of the CH_2 scissoring band at 1440 cm^{-1} remains the same, but $\nu(\text{C}=\text{O})$ at 1736 cm^{-1} disappears. Concomitantly, two bands increase in intensity, i.e., the symmetric and antisymmetric stretching bands of the CO_2^- group around 1402 and 1610 cm^{-1} , respectively. These changes reflect the deprotonation of the carboxyl group with increasing pH and, as such, this behavior is identical with what we observed for free his.⁵⁶ However, the position of $\nu_{\text{as}}(\text{CO}_2^-)$ is about 10–15 cm^{-1} lower than that in the spectra of free his at the same pH value. Although a part of this shift can be attributed to deprotonation of the amino group, as will be pointed out later, the large downward shift must be the result of coordination of one of the carboxylate O atoms to the Cu^{2+} ion.^{65,66} The same conclusion has been reported by Kruck and Sarkar²⁹ from the intensity decrease of the 1736- cm^{-1} band at $\text{pH} = 2.5$ upon the addition of Cu^{2+} , but it also illustrates that the conclusion of a $\text{Cu}^{2+}/\text{his}$ complex with a protonated carboxyl group as proposed by Meyer and Bauman⁵³ is wrong. In accordance with the literature,⁶⁵ the effect of metal coordination on $\nu_{\text{s}}(\text{CO}_2^-)$ around 1402 cm^{-1} is much smaller. The Raman shift at around 560 cm^{-1} , which rapidly increases in intensity at $\text{pH} > 2$, has been proposed as additional evidence by Itabashi and Itoh.³⁴ Indeed, a weak band at this position is also present in the spectrum at $\text{pH} = 1$, but this one probably originates from another vibration. These observations point to the

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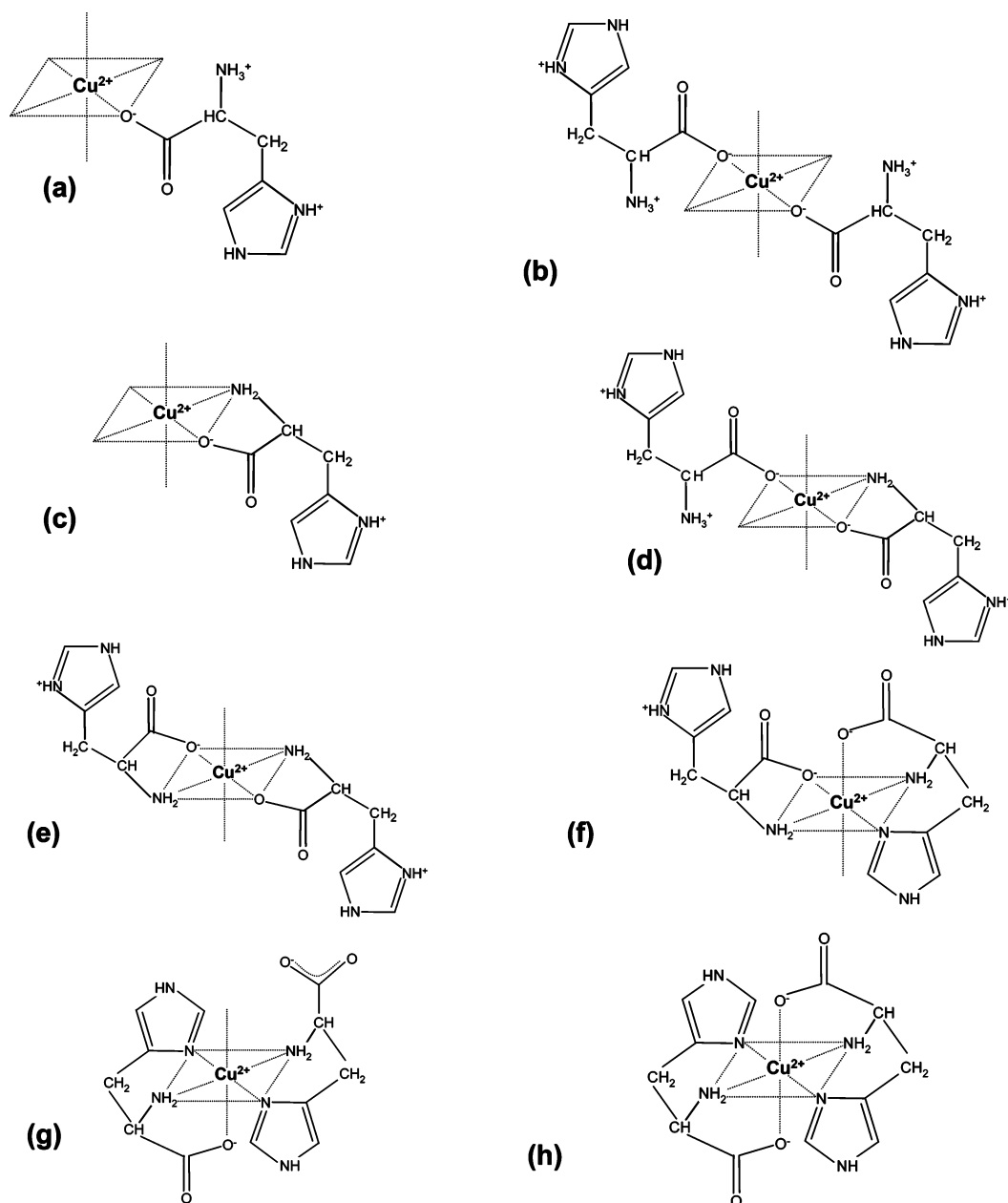


Figure 7. Proposed molecular structures of Cu^{2+} /his complexes at different pH values in the pH range of 0–10. Coordination sites not occupied by his are occupied by water.

formation of complexes of the type $\text{Cu}^{2+}[\text{H}_3\text{his}^+(\text{O}_c)]$ and $\text{Cu}^{2+}[\text{H}_3\text{his}^+(\text{O}_c)]_2$, depicted as structures a and b in Figure 7, but closer examination of the IR spectra in Figure 6 shows that the intensity of the IR band at 1535 cm^{-1} decreases rapidly too upon an increase in the pH. In the spectra of free his,⁵⁶ this absorption originates from the symmetric bending vibration of the protonated amino group [$\delta_s(\text{NH}_3^+)$], which normally loses its proton in accordance with the pK_a at $\text{pH} > 9$. It follows that the accelerated deprotonation of the amino group must be the result of coordination of N_{am} to Cu^{2+} . The same can be concluded from the absorption band at around 1610 cm^{-1} . This band consists of a peak from $\delta_{\text{as}}(\text{NH}_3^+)$ at about 1620 cm^{-1} that partly overlaps the absorption of $\nu_{\text{as}}(\text{CO}_2^-)$ at around 1600 cm^{-1} . Upon an increase in the pH, the contribution of $\delta_{\text{as}}(\text{NH}_3^+)$ decreases, and as a result, $\nu_{\text{as}}(\text{CO}_2^-)$ becomes the dominating component

of the observed band. In principle, one might expect the appearance of an NH_2 scissoring vibration in the same region, but as for the ionic states Hhis^- and his^{2-} of free his, this absorption is weak and probably coincides with that of $\nu_{\text{as}}(\text{CO}_2^-)$. Either way, the observed spectral changes can only be explained by assuming coordination of N_{am} to Cu^{2+} , which eliminates the dominating presence of a complex with (O_c , N_{im}) coordination at this pH, as has been proposed in several literature reports.^{32,38,54,67} It also follows that the previously mentioned complexes a and b must convert practically instantaneously to a bidentate complex, with O_c and N_{am} as the coordinating anchors, such as in complexes c–e (Figure 7). Clear vibrational spectroscopic differences between these complexes cannot be expected, and hence it

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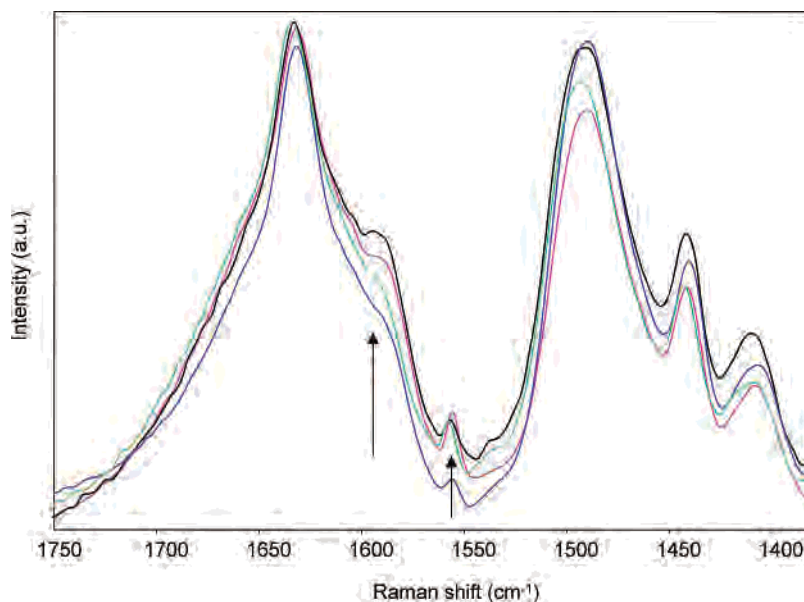


Figure 8. Raman spectra of 0.25 M aqueous solutions of $\text{Cu}^{2+}/\text{his}$ at pH = 2.78, 2.97, 3.14, and 3.47. The arrows indicate the appearance of bands at 1588 and 1555 cm^{-1} .

is not possible to draw conclusions on their relative abundance. Most likely, in accordance with Itabashi and Itoh,³⁴ the gly-like form *c* will be the major species at pH = 3. However, because all bands related to $\text{C}=\text{O}$ (1736 cm^{-1}) and NH_3^+ (1535 cm^{-1}) have disappeared at pH = 4, it is evident that free his is practically absent, and as a result, the complex $\text{Cu}^{2+}[\text{H}_2\text{his}^0(\text{O}_c, \text{N}_{\text{am}})]_2$ (structure *e*) must be the predominant species at this pH. Most likely, the formation of this species proceeds via the intermediate form *d*, but the lifetime of this complex will be negligibly short.

The shift of equilibrium from a complex with one his ligand to a structure with two his molecules is confirmed by the UV/vis/NIR and EXAFS data. EXAFS analysis reveals that there is no change in the distance of the first coordination sphere (Table S1 of the Supporting Information) and, thus, the observed shift of the UV/vis/NIR d–d transition band to shorter wavelength must be the result of an increasing amount of coordinating N atoms in the equatorial plane as in the bidentate ligand complex *e*. Besides, the position of the d–d transition band at pH = 4 is the same as that for $\text{Cu}^{2+}/\text{gly}$ in the pH range of 6–10, in which the ligands are known to be coordinating via two N_{am} and two O_c atoms in the equatorial plane. Hence, it is logical to assume an $\text{N}_{\text{am}}\text{O}_c\text{N}_{\text{am}}\text{O}_c$ -like coordination for $\text{Cu}^{2+}/\text{his}$ at pH = 4. This conclusion is in agreement with the majority of the papers in the literature.^{10,29,30,34} The ESR results support these interpretations. The Cu hexa-aqua complex is about 85% at pH = 2 with g_0 and A_0 values that are characteristic of an O coordination sphere around Cu^{2+} . The remaining 15% is of a complex(es) that becomes dominating (75%) at pH = 2.9. The corresponding g_0 and A_0 parameters point to the coordination of one N atom, which is in line with the structures *c* and *d* (Figure 7). However, as stated before, we assume that the amino group coordinates to Cu^{2+} immediately after coordination of the carboxyl group. This makes it very unlikely that structure *d* is the dominating

species at pH = 2.9 and, therefore, we conclude that the mono-his complex *c* is the major component.

The ESR parameters of the minor (25%) species at pH = 2.9 are characteristic of complexes with two N atoms coordinating to Cu^{2+} .⁴⁸ In view of the preceding arguments, it is obvious that this must be the bis-his complex *e*, in which the two ligands are coordinating gly-like. Upon an increase in the pH to 3.4, the relative amount of the two complexes changes to a 55:45 (%) mixture. Although we assume a considerable error in the calculated β values (Table 1), the decrease from -1.6 at pH = 2.0 to -10 at pH = 3.4 is in agreement with this assignment because the decrease reflects a reduced tumbling velocity of the complex as a result of the increasing mass of the complexes upon converting from the hexa-aqua structure to the mono-his complexes and finally to the bis-his complexes.

Further information on the coordination geometry of the complexes is extracted from the Raman spectra (Figure 8). At pH = 2.8, two bands grow in at around 1555 and 1588 cm^{-1} . The latter also shows up in the Raman spectra of free his but at much higher pH values. It has been assigned to coupled imidazole $\text{C}=\text{C}/\text{C}=\text{N}$ modes and is the result of deprotonation of the ring.⁵⁶ For that reason, we assign the appearance of the bands at 1588 cm^{-1} in the spectra of $\text{Cu}^{2+}/\text{his}$ as the result of coordination of (one of) the N_{im} atoms, which is in agreement with literature reports.^{15,34} The involvement of the imidazole ring in the coordination can also be concluded from the prominent imidazole $=\text{NH}$ deformation bands in the Raman spectra at 1490 and 1198 cm^{-1} , which start to reduce in intensity at about pH = 3.5 instead of around pH = 6 in free his.

The coordination to the imidazole ring and the primary amino group at a pH significantly lower than the corresponding $\text{p}K_a$ values seems strange, particularly because repulsive forces will hamper the attraction between Cu^{2+} and the positively charged NH_3^+ and imidazole ring. This implies

that a pK_a lowering mechanism must be the driving force. One explanation is that, upon coordination to O_c , the Cu^{2+} ion has an inductive effect on the ligand that promotes deprotonation and, hence, coordination of N_{am} and N_{im} . Next, the fact that coordination to N_{am} is favored over N_{im} despite its higher pK_a value can be attributed to either an energetically more stable geometry or a reduced inductive effect on N_{im} . The latter could be the result of the longer distance to the carboxylate group in combination with the resonance properties of the imidazole ring that will act as an electronic buffer. Another explanation has been reported very recently by Zhong et al.⁶⁷ The authors also observed considerable pK_a lowering effects induced by Cu^{2+} complexes and assigned these to an electrostatic interaction between the central Cu^{2+} ion and the anionic π system of the ligands. Because inductive effects should have resulted in larger effects on the vibrational frequencies, we believe that the mechanism proposed by these authors is more plausible.

pH Range of 4–5. Upon an increase of the pH to 5, the UV/vis/NIR d–d transition band shifts to 615 nm, which is similar to the position of the Cu^{2+}/him complex at $pH > 7$, where the coordination is known to be $N_{am}N_{im}N_{am}N_{im}$ in the equatorial plane. The shift for Cu^{2+}/his can only be explained by a further increase in the number coordinating N atoms because the ligand field splitting for N atoms is larger than that for O atoms.^{59,60} This conclusion is endorsed by the results of EXAFS analysis, which show that the distance of the first coordination sphere around the Cu^{2+} ion remains unaffected. For that reason, we conclude that the predominant $N_{am}O_cN_{am}O_c$ coordination at $pH = 3.5$ starts to change into an $N_{am}O_cN_{am}N_{im}$ coordination around $pH = 4.5$ and subsequently into an $N_{am}N_{im}N_{am}N_{im}$ coordination at $pH > 5$. This is in line with most of the literature^{29,34,37,38,68} but proves that the assignment of Crawford and Dalton,⁴⁶ who stated that $N_{am}O_cN_{am}O_c$ coordination remains the major structure up to $pH = 8$, is incorrect. According to the UV/vis/NIR data, $N_{am}N_{im}N_{am}N_{im}$ coordination appears to be almost complete already at $pH = 5$ because the position of the band almost matches that of Cu^{2+}/him at $pH > 6$. The growing involvement of N atoms in the coordination also follows from a comparison of the ESR data recorded at $pH = 4.4$ and 3.4 . The g_0 and A_0 parameters determined at $pH = 4.4$ clearly point to three N atoms in the coordination sphere of Cu^{2+} . The corresponding complex(es) was calculated to cover 90% of the species present.

The IR and Raman data are in accordance with this conclusion. First, it is evident that N_{am} remains coordinated to Cu^{2+} because the IR peaks related to the NH_3^+ group remain absent in the spectra. Second, the Raman band at 1588 cm^{-1} rapidly gains intensity and, at $pH = 5$, it becomes even stronger than the one at 1632 cm^{-1} , which illustrates the increasing number of coordinating N_{im} atoms. The band at around 1588 cm^{-1} has also been reported in Raman studies on Cu^{2+} complexes with histidyl side-chain peptides^{3,16} and him ³³ and has been proposed to represent N^π – Cu^{2+} ligation (N^π -protonated). Additional indications that N^π is the favored

coordination site are extracted from the weak Raman band at 990 cm^{-1} and the IR absorptions at 1272 and 1087 cm^{-1} . These bands are commonly used to distinguish the N^π -protonated tautomer from the N^τ -protonated tautomer in the free ionic states H_2his^0 and $Hhis^-$. In addition, the appearance of a broad Cu – N stretching band at 454 cm^{-1} also points to coordination to N_{im} .^{3,33} Furthermore, the IR spectra show the appearance of a weak band at 1506 cm^{-1} at around $pH = 4$, which is absent in the spectra of free his . It has been proposed as a marker for coordination to N^τ in the spectrum of Cu^{2+}/him ,³³ but the corresponding vibrational mode is not clear.

Consequently, it seems logical to conclude that the increasing coordination of N_{im} (i.e., N^τ) will be accompanied by a proportional release of carboxylate groups from the equatorial plane to the free form. However, the vibrational frequency of the IR-active $\nu_{as}(CO_2^-)$ decreases to 1584 cm^{-1} , while $\nu_s(CO_2^-)$ increases from 1403 to 1408 cm^{-1} . Indeed, the direction and magnitude of the shifts are similar to what we measured in the spectra of free his recorded in the same pH range, but the position of $\nu_{as}(CO_2^-)$ remains about 15 cm^{-1} lower. This implies that a substantial amount of the carboxyl groups must remain coordinated to Cu^{2+} , either still in the equatorial plane or converted into the axial position. Therefore, we propose that the equatorial gly-like bidentate complex e starts to convert to the complex with tridentate ligands $Cu^{2+}[Hhis^-(O_c, N_{am}, N_{im})]_2$ (Figure 7, structure h) with an equatorial him -like coordination for N_{am} and N_{im} and the carboxyl O_c atoms in the axial position. According to the ESR and UV/vis/NIR data, this process proceeds via the mixed bidentate/tridentate ligand complexes f and g as intermediates.

pH Range of 5–7. The data obtained at higher pH values confirm these conclusions but also reveal that a complex with $N_{im}N_{im}N_{im}N_{im}$ coordination around $pH = 6.5$, as proposed by Basosi et al.,⁴² is unlikely. The increase of the pH from 5 to 7 is attended not only with an increase of the intensity of the UV/vis/NIR d–d transition band but also with a shift back to longer wavelength. This is different from that of the Cu^{2+}/gly and Cu^{2+}/him solutions, where the band remains unaffected. A plausible explanation for this phenomenon is that two N_{am} atoms and two N_{im} atoms start to coordinate in the equatorial plane on proceeding pH, while the O_c atoms convert from the equatorial plane to the axial position. This results in a decrease of the ligand field splitting because this parameter is also related to the distance of the orbitals of the d metal toward the atoms of the ligand. Besides, an increase in the axial O_c coordination causes a distortion of the tetragonally elongated symmetry, which implies that the transition probability becomes larger and explains why the intensity of the d–d transition band is enhanced. Finally, the observed red shift of the band confirms the distortion of the symmetry because the axial coordination of the carboxyl group(s) will cause an increase in the distance between the central Cu ion and the chelating atoms in the equatorial plane. This also follows from the EXAFS results, which reveal a small but significant increase in the distance of Cu^{2+} to the first coordination sphere from 1.96 to 1.99 \AA upon an increase in the pH from 4.4 to 7.3 .

These conclusions confirm the conversion from the bidentate ligand complex e to the tridentate ligand form h, most likely via the mixed bidentate/tridentate ligand complexes f and g as intermediate forms. The coordination of N_{am} and N_{im} also follows from the vibrational spectra. For N_{am} , all bands related to the uncoordinated form remain absent, while the increasing coordination of N_{im} can be extracted from the drop in the intensity of the Raman-active $=NH$ deformation bands at 1492 and 1190 cm^{-1} . Similar to free his, this intensity decrease is the result of the reduced symmetry of the imidazole ring by deprotonation of one of the N_{im} atoms. Furthermore, a shoulder at around 1578 cm^{-1} arises on the Raman $C=C/C=N$ stretching band at 1588 cm^{-1} . In accordance with the pK_a value of the ring, two bands at around 1580 cm^{-1} are also observed in free his above $pH = 6$, and it therefore seems obvious to assign this shoulder to the formation of a complex with N^τ coordination. However, according to the literature,^{3,15,16,33} the vibration that is correlated to N^τ ligation (N^τ -protonated) should be found at the high-frequency side instead of the low-frequency side of the 1588- cm^{-1} band, i.e., at around 1603 cm^{-1} . Besides, model calculations revealed that coordination to N^τ is not attractive for sterical reasons. Consequently, we conclude that the coexistence of complexes with coordination to both N^τ and N^π , or a complex with both tautomeric forms coordinated to the same Cu^{2+} , as suggested by Torreggiani et al.,³³ is unlikely. Another argument to reject two types of coordination of the N_{im} atoms is the lack of doublets at around 1100 cm^{-1} in IR and at 1275 and 1000 cm^{-1} in Raman. Referring to the literature,^{9,31,33,36,68–70} such double bands mark the presence of both tautomers, but only one band is observed in these regions. The presence of one tautomeric form is further evidenced by the small bandwidth of the Raman-active ring-breathing vibration at 1273 cm^{-1} . This peak is considerably sharper than what we observed in the Raman spectra of free his, which points to a single rigid structure of the ring, i.e., only one type of N_{im} coordination. To explain the presence of the band at around 1578 cm^{-1} , we propose two possibilities: (1) the metal coordination of the ring results in decoupling of the combined $C=C/C=N$ vibration, which gives rise to separate $C=C$ and $C=N$ stretching modes and (2) two N^π -ligated complexes are present that only differ in their symmetrical orientation toward the Cu ion. The origin of the weak Raman band that comes up at 1555 cm^{-1} is not clear, but its appearance matches the explanation of decoupling because in this region only $C=C$ and/or $C=N$ stretching modes are present. Finally, the fact that only one Raman band arises at 1228 cm^{-1} can also be attributed to the deprotonation of only one of the N_{im} atoms, similarly in free his.⁵⁶

pH Range of 7–10. The spectral changes in ESR, IR, Raman, UV/vis/NIR, and EXAFS in the pH range of 7–10 are relatively small compared to the data collected at $pH =$

6, which confirms that structure h (Figure 7) has become the predominant complex at neutral and basic conditions. In accordance with this, simulation calculations on the ESR spectrum at $pH = 7.3$ clearly point to the dominating presence of a single Cu complex, and the high A_0 and low g_0 values as well as the fitted β value of -10 G are fully in line with the bis-his structure of this complex. The consistent β value also indicates that the formation of a much larger binuclear Cu^{2+} /his complex, as proposed in the literature,²⁹ is not very likely, at least at this pH.

In line with the conversion to complex h, the UV/vis/NIR absorption band recorded at $pH = 7$ is slightly red-shifted compared to the position at $pH = 6$ but stabilizes at higher pH, indicating that this process is practically completed at $pH > 7$. The same can be concluded from the increasing intensity, which is consistent with the results from EXAFS and ESR. Indeed, a further intensity enhancement is observed at $pH > 7$, but this can be attributed to the formation of copper hydroxide/oxide species, as pointed out in the Experimental Section. It also explains why Perrin and Sharma⁵⁰ and Sigel et al.³⁸ erroneously assumed the presence of $Cu^{2+}(OH)$ /his complexes at $pH > 7$.

Although the changes in the vibrational spectra are small, important information on the coordination of the CO_2^- group can be extracted from the $\nu_{as}(CO_2^-)$ band at 1584 cm^{-1} in IR. In contrast to free his, where we observed a continuous red shift to 1564 cm^{-1} upon an increase in the pH to 10, the band remains fixed at 1584 cm^{-1} . In our opinion, this confirms that the carboxyl group remains coordinated to Cu^{2+} . Furthermore, the bandwidth is significantly smaller compared to, e.g., $pH = 4$ and 5, which can be considered as a change from a mixture of equatorially and axially coordinated CO_2^- at low pH to a single fixed axial position at $pH > 7$, thus proving that reorientation into the axial position is completed. The simultaneous disappearance of the Raman band at 560 cm^{-1} , which has been assigned to a Cu–O vibration in the equatorial plane,³⁴ can be seen as additional evidence for the correctness of this assignment. Finally, the enhanced intensity of the Raman doublet at around 1580 cm^{-1} on the cost of the $C=C/C=N$ band at 1632 cm^{-1} indicates that all of his has become coordinated to Cu^{2+} through N^π , which is confirmed by the increased intensity of the Cu–N stretching band at 454 cm^{-1} . It follows that the double-tridentate ligand complex h is indeed the major species at high pH, whereas bidentate geometries are probably no longer present.

Complex Distribution. The large number of IR spectra, extracted from the titration experiments, allowed MCR calculations to obtain a quantitative distribution plot of the complexes as a function of the pH. The relative abundances of the different complexes as derived from ESR, UV/vis, and EXAFS have been used as additional input to enhance the reliability and accuracy of the plot. The result of the MCR calculations, of which details are given in the Supporting Information, is shown in Figure 9. The distribution plot slightly differs from the one proposed by Sarkar and co-workers,^{18,19} but as demonstrated, the large number of complementary spectroscopic data produced in our study not

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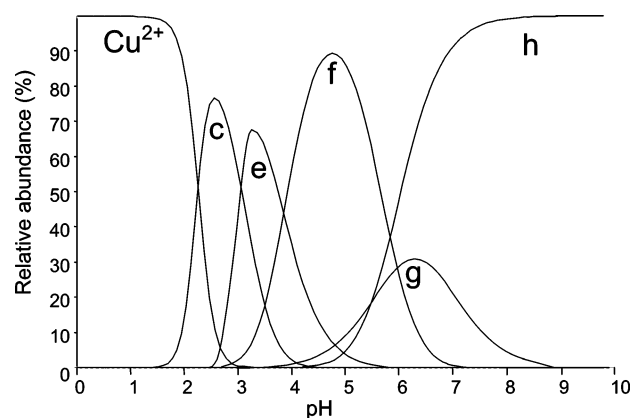


Figure 9. Proposed distribution of Cu^{2+} /his complexes as a function of the pH of the solution. Complexes a and e–h correspond to the structures depicted in Figure 7. Complexes a, b, and d are supposed to be absent or present below the limit of detection.

only can be used to determine the coordination and ionic states of the his ligands but also admits further refinement in distinguishing the different geometrical orientations within each state of protonation in a quantitative way. Six complexes evidently appear to be present, i.e., $\text{Cu}(\text{H}_2\text{O})_6^{2+}$ and complexes c and e–h, while three complexes, i.e., a, b, and d, have not been detected. However, this is in line with what was expected.

Conclusions

The combined application of IR, Raman, UV/vis/NIR, ESR, and XAS proves to be crucial to obtaining consistent and complementary information on the geometrical structures and relative abundances of Cu^{2+} /his (1:2) complexes in an aqueous solution as a function of the pH. Although intrinsic instrumental limitations hampered spectroscopic analysis at identical concentration levels, thus limiting the accuracy of the overall conclusions, the experimental conditions have shown to be a fair compromise. The results from UV/vis/NIR and ESR fully agree on the nature of the atoms in the first coordination shell of the Cu^{2+} ion, while EXAFS adds knowledge on the average distance between the coordinating atoms and the central Cu^{2+} ion. IR and Raman produce detailed complementary information on the state of protonation of the his ligands and, hence, on the sites involved in the coordination to Cu^{2+} . As a result, structural proposals from the literature have been unambiguously confirmed, clarified, or corrected where necessary.

It was found that the complex formation process is reversible as long as the pH does not exceed a value of 10. At higher pH, decomposition and precipitation of Cu^{2+} /his complexes occurs. At $\text{pH} < 2$, both Cu^{2+} and his are present

as free ions. The formation of complexes starts at around the pK_a value of the carboxyl group, i.e., at $\text{pH} = 1.8$, by anchoring an O atom in the equatorial plane (complexes a and b). Concomitantly, coordination via the amino group begins, which results in a bidentate gly-like chelation with mainly one ligand (complex c). At around $\text{pH} = 3$, this complex is the major constituent. Coordination of a second his molecule rapidly follows, and at $\text{pH} = 4$, the bidentate ligand complex e is the dominant species. The Raman data show that coordination to the N^π atom of the imidazole ring starts at around $\text{pH} = 2.8$ and, at around $\text{pH} = 5$, complex f with equatorial him-like coordination via the N_{am} and N^π atoms becomes the main component. This reorientation of the his ligand is not attended with the release of the carboxyl group to the nonbonded state, but according to IR, the coordination of the O atom converts from the equatorial plane to the axial position. This conversion continues via complex g to the double-tridentate ligand complex h with four N atoms in the equatorial plane and two carboxyl O atoms in the axial position as the dominant geometrical structure at $\text{pH} > 7$. This structure remains stable up to $\text{pH} = 10$ and decomposes at higher solution pH.

Finally, it is concluded that the pK_a value is not the primary driving force in the coordination process. After deprotonation and anchoring to the carboxyl group in accordance with its pK_a value (1.8), coordination to N_{am} and N_{im} atoms occurs at around $\text{pH} = 3$ and 5, respectively. This is in a different order and at significantly lower pH than could be expected on the basis of the pK_a values of these N groups. The effect is assigned to an electrostatic interaction between the central Cu^{2+} ion and the anionic π system of the his ligands.

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Supporting Information Available: IR, ESR, and Raman spectra of Cu^{2+} /his at various pH values; ESR, XAFS, and MCR analyses; and Fourier transforms. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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