

Dihydrogen to Dihydride Isomerization Mechanism in $[(C_5Me_5)FeH_2(Ph_2PCH_2CH_2PPh_2)]^+$ through the Experimental and Theoretical Analysis of Kinetic Isotope Effects

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The isomerization of complex $[Cp^*Fe(dppe)(\eta^2-H_2)]^+$, generated in situ by low-temperature protonation of $Cp^*Fe(dppe)H$ with either HBF_4 or CF_3COOH , to the dihydride tautomer *trans*- $[Cp^*Fe(dppe)(H)_2]^+$ is irreversible and follows first-order kinetics in the -10 to $+15$ °C range with $\Delta H^\ddagger = 21.6 \pm 0.8$ kcal mol⁻¹ and $\Delta S^\ddagger = 5 \pm 3$ eu. The isomerization rate constant is essentially independent of the nature and quantity of a strong acid. Density functional theory (DFT) calculations on various models, including the complete system at both the quantum mechanics/molecular mechanics (QM/MM) and full QM levels, probe the relative importance of steric and electronic effects for the relative stability of the nonclassical and classical isomers and identify two likely isomerization mechanisms: a "direct" pathway involving simultaneous H–H bond breaking and cis–trans isomerization and a "via Cp" pathway involving agostic C_5Me_5H intermediates. Both pathways are characterized by activation energies in close correspondence with the experimental value (21.3 and 22.2 kcal mol⁻¹, respectively). Further kinetic studies were carried out for the $Cp^*Fe(dppe)H + CF_3COOD$ and $Cp^*Fe(dppe)D + CF_3COOD$ systems at 273 K. The $[Cp^*Fe(dppe)(\eta^2-HD)]^+$ complex establishes a very rapid isotope redistribution equilibrium with the η^2-H_2 and η^2-D_2 analogues. The equilibrium constant value ($K = 3.3 \pm 0.3$) indicates a significant equilibrium isotope effect. Simulation of the rate data provides access to the individual isomerization rate constants k_{HH} , k_{HD} , and k_{DD} for the three isotopomers, yielding kinetic isotope effects: $k_{HH}/k_{HD} = 1.24 \pm 0.01$ and $k_{HD}/k_{DD} = 1.58 \pm 0.01$ (and, consequently, $k_{HH}/k_{DD} = 1.96 \pm 0.02$). The analysis of the DFT-calculated frequencies, using the $[Cp^*Fe(dppe)(\eta^2-H_2)]^+$ model system, for the $[Cp^*Fe(dppe)(\eta^2-XY)]^+$ isotopomers as well as transition states for the "direct" (TS_{dir}) and "via Cp" (TS_{rot}) pathways ($X = H, D$) allowed the computation of the expected isotope effects. A comparison with the experiment strongly suggests that the mechanism occurs via the "direct" pathway for the present system, although the small difference in the calculated energy barriers suggests that the "via Cp" pathway may be preferred in other cases.

Introduction

It is nowadays well established that most 16-electron L_nM fragments form a dihydrogen complex $L_nM(\eta^2-H_2)$ in the

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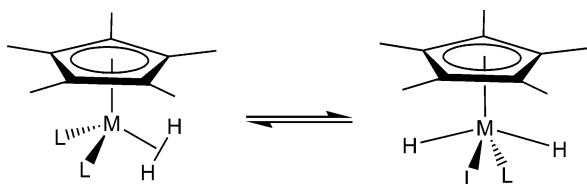
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initial stages of their interaction with dihydrogen.¹ Subsequently, depending on the properties of the metal and ligands, an oxidative addition process may take place, leading to a dihydride, or the dihydrogen complex may remain as the stable final product. An equilibrium between the two isomeric forms may also be observed. Dihydrogen complexes are also sometimes obtained by protonating hydride precursors under kinetically controlled conditions, preceding in some cases the rearrangement to the thermodynamic classical dihydride product.^{2–7} For the group 8 half-sandwich systems such as

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Scheme 1



$[\text{CpRuLL}'\text{H}_2]^+$ (Cp = cyclopentadienyl; L = tertiary phosphine)^{3,4} and iron analogues,^{6,7} this interconversion implies an important stereochemical change. In these systems, two different mutual arrangements of two ligands are possible: cisoid and transoid. Because there are no reports of the coexistence of dihydrogen and *cis*-dihydride isomers, the interconversion must entail passing from the three-legged-piano-stool dihydrogen complex to a transoid square-based four-legged-piano-stool dihydride (Scheme 1).⁸ Thus, the cleavage of the H–H bond is accompanied by a *cis*–*trans* rearrangement, which involves the migration of one H atom from one side of the molecule to the opposite one. The isomerization of the dihydrogen complex $[\text{Cp}^*\text{Fe}(\text{dppe})(\eta^2\text{-H}_2)]^+$ is particularly interesting because the chelate nature of the diphosphine ligand may impose an extra obstacle to the H migration to yield the observed *trans* structure for the dihydride product.

The isomerization kinetics have been studied by NMR spectroscopy for the cationic ruthenium dihydrogen complexes.^{3,9} Reported activation enthalpies range from 16 to 21 kcal mol⁻¹, with small and negative activation entropies. The highest values have been obtained for bidentate phosphines with sterically demanding ligands.⁹ The small ΔS^\ddagger supports an intramolecular mechanism in which the dihydrogen ligand has lost rotational freedom in the transition state.³ A dissociative mechanism has been proposed when bulky and good electron-donor ligands are present⁹ but was discarded for a simpler system on the basis of kinetic evidence (no retardation effect was observed in the presence of excess ligand).³ The possibility of an intermolecular deprotonation followed by reprotonation of the neutral hydride intermediate was ruled out for the ruthenium system on the basis of the absence of an isotope scrambling for a selectively generated (HD) system.^{3,9} Finally, an alternative intramolecular mechanism proceeding by transfer of an H atom from the dihydrogen ligand to the Cp ligand, thereby

affording an intermediate cyclopentadiene complex, followed by ring rotation and transfer back to the metal center, has also been contemplated but discarded on the basis of indirect bond energy arguments.³ A theoretical study of this isomerization process has not been tackled yet, and no reports on the transition states or a comparison of energy barriers for different mechanisms can be found.

We have recently reported experimental investigations, including a stopped-flow kinetic study, of the hydrogen bonding and proton transfer to $\text{Cp}^*\text{Fe}(\text{dppe})\text{H}$ with a variety of proton donors (HA) of different acid strength.^{10,11} This study has highlighted the presence of a hydrogen-bonded intermediate involving the hydride site (and not the metal site), $\text{Cp}^*\text{Fe}(\text{dppe})\text{H}\cdots\text{HA}$, the reversibility of the proton transfer step (see Scheme 2), and the need of a second proton donor molecule to trigger the proton transfer process. An accompanying theoretical investigation has mostly focused on the first step, i.e., proton transfer to the hydride ligand.¹¹ Concerning the subsequent isomerization process, an intramolecular mechanism is suggested by an identical rate when using proton donors of different strength at the same temperature. The alternative reversible deprotonation, followed by a slower protonation at the metal site, is inconsistent with this experimental observation.¹⁰ Our attempts to carry out an Eyring analysis of this isomerization step were complicated by the reversibility of the preceding proton-transfer step, which resulted in strong coupling between two key rate constants in the mathematical fitting procedure, with the first rate constant leading from the dihydrogen intermediate to the final dihydride product (k_2 , irreversible step) and the second one leading from the same intermediate back to the starting monohydride complex (k_{-1} ; see the Supporting Information of ref 11).

In this report, we present a new experimental study that has led to the determination of the accurate rate constant and activation parameters for the isomerization process of the $[\text{Cp}^*\text{Fe}(\text{dppe})\text{H}_2]^+$ system when the counterion is either BF_4^- or CF_3COO^- . These investigations include the determination of the kinetic isotope effects (KIEs) $k_{\text{HH}}/k_{\text{HD}}$ and $k_{\text{HD}}/k_{\text{DD}}$, obtained from the comparative isomerization rates of complexes $[\text{Cp}^*\text{Fe}(\text{dppe})(\text{H}_2)]^+$ (k_{HH}), $[\text{Cp}^*\text{Fe}(\text{dppe})(\text{HD})]^+$ (k_{HD}), and $[\text{Cp}^*\text{Fe}(\text{dppe})(\text{D}_2)]^+$ (k_{DD}). To the best of our knowledge, the experimental determination of the isotope effects for the transformation of a nonclassical $[(\text{H}_2)/(\text{HD})/(\text{D}_2)]$ to a classical $[(\text{H})_2/(\text{H})(\text{D})/(\text{D})_2]$ dihydride complex is unprecedented. This is accompanied by a theoretical investigation whose purpose is 2-fold. First, it seeks to determine the subtle factors tuning the relative stability between the nonclassical dihydrogen and the classical dihydride isomers of the $[(\text{C}_5\text{R}_5)\text{Fe}(\text{R}'_2\text{PCH}_2\text{CH}_2\text{PR}'_2)\text{H}_2]^+$ complexes (R = H, Me; R' = H, Ph, ^tPr), separating electronic and steric contributions. Second, it explores the intimate details of the possible intramolecular isomerization mechanisms. In particular, discrimination between two pathways showing very

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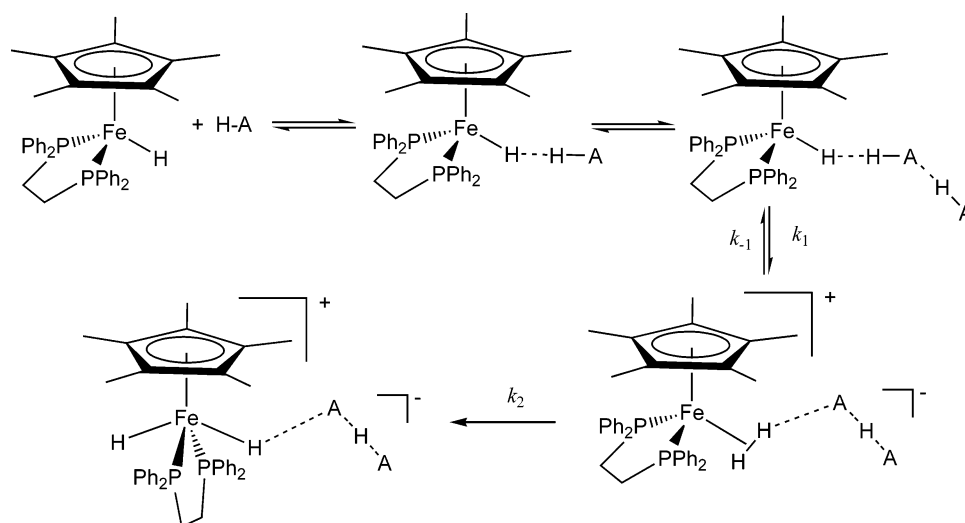
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Scheme 2



similar computed activation barriers (both close to the experimental value) will be possible through the analysis of the KIEs. The computational work involves pure quantum mechanics (QM) and mixed QM/molecular mechanics (MM) methods.

Experimental Part

All manipulations were carried out under an argon atmosphere by standard Schlenk techniques. Complexes Cp*Fe(dppe)H and Cp*Fe(dppe)D were synthesized according to the literature.⁷ NMR characterization of the deuteride complex in C₆D₆: ²H NMR δ -16.9 (t, $J_{\text{PD}} = 10.4$ Hz); ³¹P{¹H} NMR δ 111.5 (t, $J_{\text{PD}} = 10.3$ Hz). In addition, the ¹H NMR spectrum shows contamination by the hydride complex Cp*Fe(dppe)H (25%, by integration of the hydride resonance against the Cp* resonance). The acids HBF₄·OEt₂, CF₃COOH, and CF₃COOD were purchased from Aldrich and used as received.

The NMR studies were carried out in standard 5-mm NMR tubes containing solutions of the complexes in CD₂Cl₂. The ¹H and ³¹P NMR data were collected with a Bruker AMX 400 spectrometer operating at 400.13 and 161.98 MHz, respectively. The low-temperature kinetic experiments with ³¹P{¹H,²H} NMR monitoring were carried out on a Bruker AV500 spectrometer equipped with a 5-mm triple-resonance inverse probe with dedicated ³¹P channel operating at 500.33 MHz for ¹H NMR, 202.54 MHz for ³¹P NMR, and 76.80 MHz for ²H NMR. For the purpose of quantitative concentration measurements for the kinetic runs, the ³¹P NMR spectra were recorded with inverse-gated ¹H and ²H decoupling. For ³¹P and ¹H NMR, a typical pulse width corresponding to a 30° flip angle and 30-s relaxation delays was used to obtain reliable integration data. In the case of overlapping resonances, the quantitative evaluation was realized by peak deconvolution (see the Supporting Information). All chemical shifts for ¹H and ²H NMR are relative to tetramethylsilane using a residual peak of the solvent as a secondary standard. ³¹P NMR chemical shifts were referenced to an external 85% H₃PO₄ sample. The temperature was regulated with a Bruker TV-3000 unit. Temperature calibration was determined using a methanol chemical shift thermometer. The temperature accuracy and stability was ± 1 K. All mixings between the acid and hydride complexes were performed at -80 °C.

Computational Details

QM calculations were performed with the *Gaussian 98* package¹² at the density functional theory (DFT) B3PW91 level.^{13–15} Furthermore, coupled-cluster^{16,17} single-point calculations on the DFT-optimized structures, using both single and double substitutions^{18–21} and including triple excitations noniteratively,²² were performed to obtain more reliable energies. Core electrons of the Fe and P atoms were described using the effective-core pseudopotentials of Hay–Wadt,^{23,24} and valence electrons were described with the standard LANL2DZ basis set associated with the ECP.¹² In the case of the P atoms, a set of d-type functions was added.²⁵ C and H atoms nonbonded to the metal were described with a 6-31G basis set.²⁶ The two H atoms bonded to the Fe atom were described with a 6-31G(d,p) basis set of functions.²⁷

Mixed QM/MM calculations were performed with the IMOMM method²⁸ with a program built from modified versions of two standard programs: *Gaussian 98*¹² for the QM part and mm3(92)²⁹ for the MM calculations. The IMOMM approach will be used to point out the effect of

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Table 1. Results of the Isomerization Kinetics of $[\text{Cp}^*\text{Fe}(\text{dppe})(\eta^2\text{-X}_2)]^+$ to $[\text{Cp}^*\text{Fe}(\text{dppe})(\text{X})_2]^+$ ($\text{X} = \text{H}, \text{D}$)^a

run	complex	acid	[HA/Fe]	T (°C)	$10^4 k_{\text{HH}}$ (s ⁻¹) ^a	$10^4 k_{\text{HD}}$ (s ⁻¹) ^a	$10^4 k_{\text{DD}}$ (s ⁻¹) ^a
1	Cp*Fe(dppe)H	HBF ₄	1	-10	0.70 ± 0.09		
2	Cp*Fe(dppe)H	HBF ₄	1	0	3.21 ± 0.07		
3	Cp*Fe(dppe)H	HBF ₄	1	5	5.92 ± 0.15		
4	Cp*Fe(dppe)H	HBF ₄	1	15	28.3 ± 0.7		
5	Cp*Fe(dppe)H	HBF ₄	3	0	3.53 ± 0.22		
6	Cp*Fe(dppe)H	CF ₃ COOH	1	0	3.10 ± 0.12		
7	Cp*Fe(dppe)D ^a	CF ₃ COOD	1	0			1.73 ± 0.12
8	Cp*Fe(dppe)H	CF ₃ COOD	1	0	3.07 ± 0.02	2.48 ± 0.01	1.56 ± 0.01

^a The standard deviations reported for runs 1–6 were calculated keeping in mind an estimated 10% error in the NMR integration. For run 7, see the Supporting Information. ^b The starting material was contaminated by ca. 25% of CpFe(dppe)H.

the substituents located both on the Cp group and on the phosphine ligand. The substituents are methyl, phenyl, and isopropyl groups treated in the MM part through the MM3 force field. The same basis set as that in the pure quantum computations was used at the same level of theory for the QM part. In the IMOMM computations, all of the geometrical parameters were optimized except for the distances between the atoms linking the QM and MM parts. The KIEs were calculated as described in the Results and Discussion section using harmonic vibrational frequencies obtained from frequency calculations. The frequencies of each isotopically substituted system were obtained at the fixed geometry previously optimized for the corresponding unsubstituted isotopomer. These calculations were carried out with *Gaussian 03*.³⁰

Solvent effects were taken into account by means of polarized continuum model calculations^{31,32} using standard options.³⁰ Individual solvation cavities were added on the H atoms directly bonded to the Fe atom. The free energies of solvation were computed in dichloromethane ($\epsilon = 8.93$) at the geometries optimized in the gas phase.

Results and Discussion

1. Isomerization Kinetics of $[\text{Cp}^*\text{Fe}(\text{dppe})(\eta^2\text{-X}_2)]^+$ ($\text{X} = \text{H}, \text{D}$). As outlined in the Introduction, the accurate determination of the isomerization rate constant for the process shown in Scheme 2 at different temperatures was thwarted by the reversibility of the proton-transfer step when a mildly acidic proton donor (i.e., $(\text{CF}_3)_n\text{CH}_3-n\text{OH}$ with $n = 1-3$) was used. In an attempt to circumvent this problem, we have carried out stopped-flow studies using stronger acids (CF_3COOH and HBF_4), to force the proton-transfer step to quantitative conversion and then measure the subsequent isomerization under clean pseudo-first-order conditions. Unfortunately, these experimental conditions introduce another complication. The final dihydride product is unstable in dichloromethane in the presence of an excess of strong acids, leading to the oxidized chloride complex $[\text{Cp}^*\text{Fe}(\text{dppe})\text{Cl}]^+$, as described previously.¹⁰ This phenomenon is not observed when using the $(\text{CF}_3)_n\text{CH}_3-n\text{OH}$ proton donors, even in large excess amounts, whereas it is observed when

using CF_3COOH or HBF_4 in greater than a 2-fold excess. Thus, the system is not amenable to a clean kinetic study in dichloromethane. A change of solvent has been envisaged, but no ideal system has so far been found: tetrahydrofuran reduces considerably the thermodynamic drive to proton transfer (because of the free acid stabilization by hydrogen bonding with the solvent molecules) and polymerizes under strongly acidic conditions; toluene cannot keep the produced salts in solution at sufficient concentrations.

A convenient solution to this problem was found by recurring to the NMR technique. The advantage in this case is that the relative amounts of dihydrogen and dihydride complexes can be measured directly and independently, even in the presence of any amount of the $[\text{Cp}^*\text{Fe}(\text{dppe})\text{Cl}]^+$ decomposition product (a paramagnetic complex). Therefore, the rate at which the intermediate dihydrogen complex decays is directly accessible with accuracy. In addition, working at low temperatures (thereby slowing down the isomerization process to a suitable time scale for NMR monitoring) reduces the impact of the acid-catalyzed decomposition in dichloromethane. The sum of the integrated intensities for the dihydrogen and dihydride complexes, monitored against that of an internal standard, provides a direct gauge of the system stability.

After generation of the intermediate dihydrogen complex quantitatively in situ by low-temperature (-80 °C) addition of the appropriate amount of a strong acid (see Table 1 for details), monitoring was carried out at four different temperatures (-10 , 0 , $+5$, and $+15$ °C, runs 1–4, respectively). Both the decay of the dihydrogen complex resonance (examples are shown in Figure 1, top) and the growth of the final dihydride complex resonance were monitored by ¹H NMR spectroscopy. The sum of the integral, after normalization relative to the solvent peak that was used as an internal standard, was relatively constant throughout the reaction, showing limited decomposition ($<10\%$ when using 1 equiv of a strong acid; a maximum 17% decomposition during 153 min was found for run 5, where a greater excess of acid was used). However, because the acid-catalyzed decomposition occurs only for the classical product and because the isomerization is irreversible,¹⁰ the decay rate of the nonclassical reagent is an accurate measure of the isomerization rate.

The nonclassical resonance decay gives an excellent fit to the first-order rate law at each temperature (examples are shown in Figure 1, above), yielding the rate constants k_{HH}

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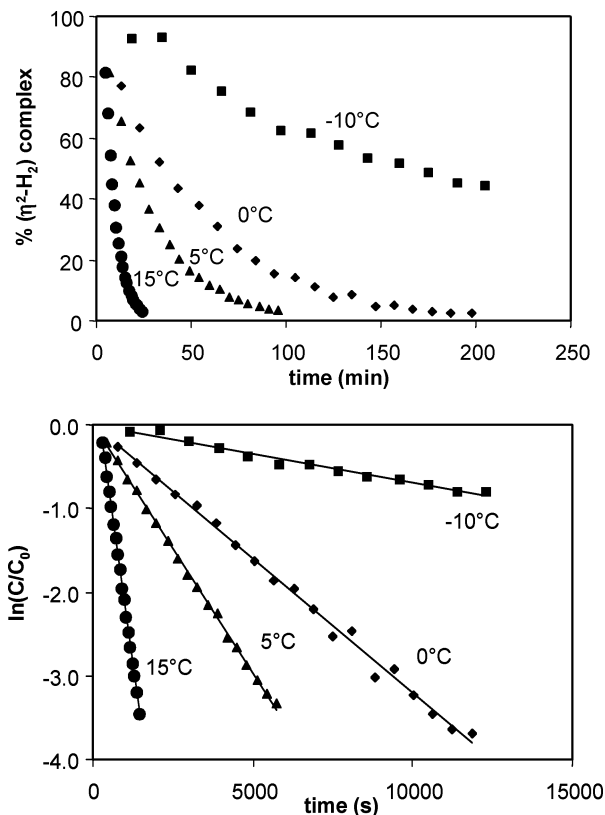


Figure 1. ^1H NMR monitoring of the $[\text{Cp}^*\text{Fe}(\text{dppe})(\eta^2\text{-H}_2)]^+$ resonance decay (BF_4^- salt in CD_2Cl_2).

reported in Table 1. The three different kinetic runs carried out at 0 °C (runs 2, 5, and 6) show that the isomerization rate does not significantly depend on the acid concentration (compare runs 2 and 5) or on its nature (compare runs 2 and 6). The latter observation confirms our previous findings of identical isomerization rates using different $(\text{CF}_3)_n\text{CH}_3-n\text{OH}$ proton donors.¹⁰ A comparison of runs 6 and 7, on the other hand, shows a significant KIE. Monitoring of run 7 could not be carried out conveniently by ^1H NMR because the Cp^* and dppe peaks of the reactant and product are not significantly spread apart. The quantitative determination was based on the integration of the inverse-gated doubly decoupled $^{31}\text{P}\{^1\text{H},^2\text{H}\}$ NMR resonance (see the Experimental Part). It should be remarked that the ca. 25% of the $\text{Cp}^*\text{FeH}(\text{dppe})$ isotopomer in the deuteride starting material may slightly affect the resulting rate constant value, given that the ^1H -containing nonclassical impurity, $[\text{Cp}^*\text{Fe}(\eta^2\text{-HD})(\text{dppe})]^+$, isomerizes faster (vide infra). Therefore, the value of k_{DD} reported in Table 1 should be considered as an upper estimate of the isomerization rate constant for complex $[\text{Cp}^*\text{Fe}(\text{dppe})(\eta^2\text{-D}_2)]^+$ (a more accurate value for k_{DD} is available from an alternative approach, vide infra). The Eyring analysis of the rate constants k_{HH} of runs 1–4 (Figure 2) yields the activation parameters $\Delta H^\ddagger = 21.6 \pm 0.8$ kcal mol $^{-1}$ and $\Delta S^\ddagger = 5 \pm 3$ eu. These values are similar to those previously reported for related ruthenium derivatives.^{3,9} The enthalpy value resulting from this analysis is useful for a comparison with the computed energies for the transition states of different potential rearrangement pathways (vide infra).

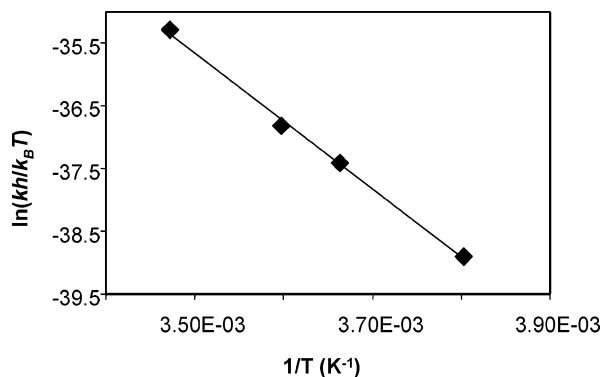


Figure 2. Eyring analysis of the isomerization rate constant k_{HH} .

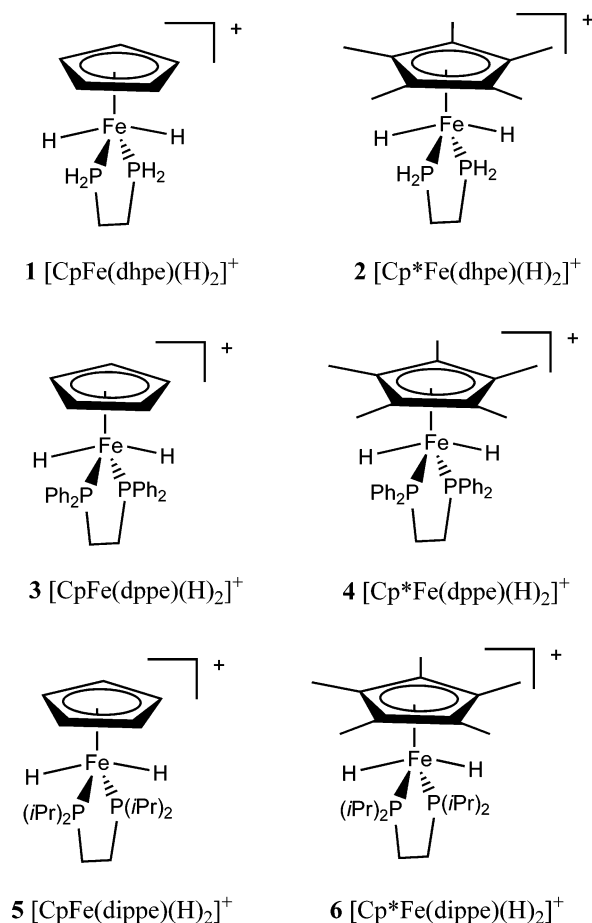
2. Role of the Electronic and Steric Effects in the Relative Stabilities of the Dihydrogen and Dihydride Complexes. Before analysis of the possible rearrangement pathways in detail from the computational point of view, it is important to gain a full appreciation of the model effect on the relative energies. As a general trend in computational chemistry, models are used to represent the real molecular systems. This is mainly done in order to deal with affordable computation time and system resources, but this procedure also allows one to introduce the complexity of the real systems in a gradual way, analyzing at each step the changes suffered by the system. This approach will be used to separate electronic and steric contributions to the relative stabilities of the dihydrogen and dihydride forms. Usually, the system is simplified by replacing substituents with H atoms. If we apply this general procedure to the $[(\text{C}_5\text{R}_5)\text{Fe}(\text{R}'_2\text{PCH}_2\text{CH}_2\text{-PR}'_2)\text{H}_2]^+$ complexes, we can define four models, as depicted in Chart 1, together with the real complexes. The smallest model is the $[\text{CpFe}(\text{dhpe})\text{H}_2]^+$ complex **1** ($\text{dhpe} = \text{PH}_2\text{CH}_2\text{-CH}_2\text{PH}_2$), where both the methyl groups of the Cp^* ligand and the phenyl or isopropyl substituents of the phosphine ligand are replaced by H atoms. In the second model $[\text{Cp}^*\text{Fe}(\text{dhpe})\text{H}_2]^+$ (**2**), only the phosphine substituents are replaced by H atoms, keeping the whole Cp^* . Two additional models can be generated by introducing the real phosphine substituents while keeping the model Cp ligand: $[\text{CpFe}(\text{dppe})\text{H}_2]^+$ (**3**) for $\text{R}' = \text{Ph}$ and $[\text{CpFe}(\text{dippe})\text{H}_2]^+$ (**5**) for $\text{R}' = ^i\text{Pr}$. The four models can be related to the two experimentally reported complexes $[\text{Cp}^*\text{Fe}(\text{dppe})\text{H}_2]^+$ (**4**)^{6,33} and $[\text{Cp}^*\text{Fe}(\text{dippe})\text{H}_2]^+$ (**6**).³⁴ The dihydride form of all of the species considered in the calculations and the numbering scheme are presented in Chart 1. The cisoid (dihydrogen) and transoid (dihydride) isomers will be labeled **c** and **t**, respectively.

The optimized structures of the isomers for the model complex **1** are depicted in Figure 3. The structural features of the dihydrogen and dihydride isomers for the systems **1–6** are analyzed in detail in the Supporting Information. The relative energies of the dihydrides **1t–6t** with respect to the corresponding dihydrogen isomers **1c–6c** are given in Table 2. Our computed values for **1** and **4** (+6.6 and -3.9 kcal

(33) Hamon, P.; Toupet, L.; Hamon, J.-R.; Lapinte, C. *Organometallics* **1992**, *11*, 1429–1431.

(34) Jiménez-Tenorio, M.; Puerta, M. C.; Valerga, P. *Organometallics* **1994**, *13*, 3330–3337.

Chart 1



mol⁻¹, respectively) compare well with those previously reported using a different functional (+4.3 and -4.4 kcal mol⁻¹, respectively).⁶ To check the accuracy of our methodology, we have recalculated the energy of the dihydrogen/dihydride couple of the smallest system, **1**, with the highly correlated CCSD(T) method. There is a reasonable agreement between both methods because the dihydride lies 9.4 kcal mol⁻¹ above the dihydrogen complex at the CCSD(T) level. Moreover, the experimental trend on the stabilities is well reproduced by our calculations. The dihydrogen complex was found to be more stable than the dihydride only for the Cp/dppe set of ligands.^{6,10,33–35} We have also checked whether solvation affects the dihydrogen/dihydride relative stabilities by calculating the energy difference between the dihydrogen and dihydride forms of **2** in dichloromethane. The gas-phase energy difference (4.9 kcal mol⁻¹ in favor of the dihydrogen complex) is only slightly decreased to 4.5 kcal mol⁻¹ in dichloromethane. Therefore, solvation plays a very minor role in this equilibrium. On the contrary, the equilibrium is strongly influenced by the Cp and phosphine substituents. The introduction of the methyl substituents in the Cp ring favors the dihydride, reversing the stabilities. The preference for the dihydride isomer is considerably enhanced by the presence of the isopropyl substituents in the phosphine. Because the classical dihydride is the thermodynamic product

(35) Jia, G.; Ng, W. S.; Yao, J.; Lau, C.-P.; Chen, Y. *Organometallics* **1996**, *15*, 5039–5045.

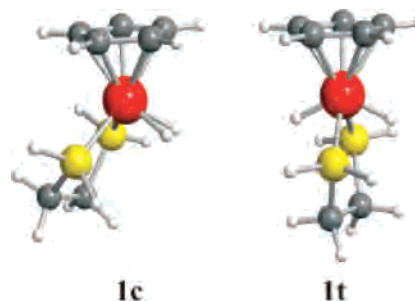


Figure 3. Optimized geometries of the dihydrogen (left) and dihydride (right) isomers of **1**.

Table 2. Relative QM Energies (in kcal mol⁻¹; QM/MM Values in Parentheses) of the Dihydride Complexes, with Respect to the Corresponding Dihydrogen Isomer

complex	ΔE
[CpFe(dhpe)H ₂] ⁺ , 1	6.6
[Cp*Fe(dhpe)H ₂] ⁺ , 2	4.9 (5.9) ^a
[CpFe(dppe)H ₂] ⁺ , 3	1.4 (5.7) ^b
[Cp*Fe(dppe)H ₂] ⁺ , 4	-3.9 (0.1) ^b
[CpFe(dippe)H ₂] ⁺ , 5	-3.6 (0.9) ^c
[Cp*Fe(dippe)H ₂] ⁺ , 6	-8.6 (-6.2) ^c

^a The methyl groups of the Cp* ligand are in the MM part. ^b The phenyl substituents of the dppe ligand are in the MM part. ^c The isopropyl substituents of the dippe ligand are in the MM part.

of the protonation of the Cp*Fe(dppe)H complex,^{6,10,11} we expect that it must be more stable than the related non-classical dihydrogen complex. However, the nonclassical dihydrogen complex is the most stable isomer for all of the models except when the computation is carried out on the real complex [Cp*Fe(dppe)(H)₂]⁺. It is clear that in our case model systems cannot account for the experimental trends on the relative stabilities of the two isomers.

The substituents tune the relative stabilities of the dihydrogen and dihydride forms, and thus the ability of the metal fragment to break the H–H bond, by a combination of electronic and steric effects. We will apply a simple analysis already used on other systems to separate and quantify both contributions on the relative energies of the two isomers.^{36–38} To this aim, we have carried out additional QM/MM computations. The approach relies on two basic assumptions: (1) the full QM calculation on the whole system describes correctly the experimental behavior and thus includes both electronic and steric contributions; (2) the QM/MM calculation describes only the steric effects of the groups included in the MM part, leaving out their electronic effects. From the small model **1**, the substituents have been added in order to build up the real complexes, but in these QM/MM calculations, they have been introduced at the MM level. By this scheme, only the steric contributions of the substituent are responsible for the changes in the relative stabilities of the two isomers.

The relative energy between the two isomers computed with model **1** is the reference. No substituents are present,

(36) Barea, G.; Maseras, F.; Jean, Y.; Lledós, A. *Inorg. Chem.* **1996**, *35*, 6401–6405.

(37) Bustelo, E.; Carbo, J.; Lledós, A.; Mereiter, K.; Puerta, M.; Valerga, P. *J. Am. Chem. Soc.* **2003**, *125*, 3311–3321.

(38) Maresca, O.; Maseras, F.; Lledós, A. *New J. Chem.* **2004**, *28*, 625–630.

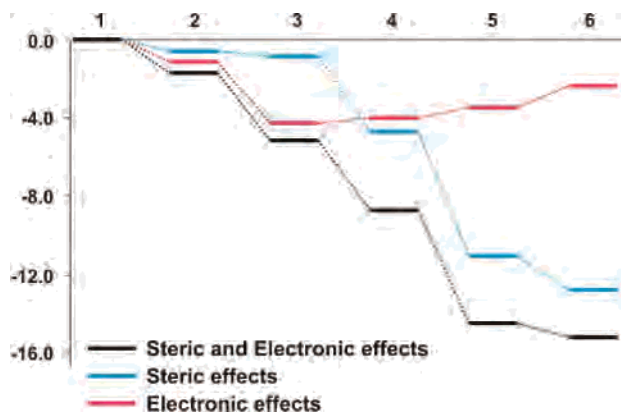


Figure 4. Substituent steric and electronic effects on the stabilization of the classical dihydride isomer for the complexes drawn in Chart 1.

neither on the C_5 ring nor on the phosphine ligand, and the dihydrogen is $6.6 \text{ kcal mol}^{-1}$ more stable than the dihydride. This value is reduced to $4.9 \text{ kcal mol}^{-1}$ when performing a full QM calculation of the $[\text{Cp}^*\text{Fe}(\text{dhpe})\text{H}_2]$ system **2**, whereas a QM/MM calculation of the same system with the methyl substituents of the Cp^* considered only at the MM level gives the dihydrogen $5.9 \text{ kcal mol}^{-1}$ more stable than the dihydride. The total contribution of the methyls is thus $1.7 \text{ kcal mol}^{-1}$, of which $0.7 \text{ kcal mol}^{-1}$ is attributed to the steric effect. Afterward, the electronic effects of the methyl groups ($1.0 \text{ kcal mol}^{-1}$) can be estimated by subtracting ($1.7 - 0.7 \text{ kcal mol}^{-1}$) the steric effect from the total contribution. Of course, the electronic effect could also be evaluated directly as the difference ($5.9 - 4.9 \text{ kcal mol}^{-1}$) between the relative QM and QM/MM energies reported in Table 2, with an identical result. The same scheme can be applied to the phenyl and isopropyl substituents. The results are summarized in Figure 4.

From **1** to **2**, both the electronic and steric effects are quite small and are of the same order of magnitude. However, the effect of the phenyl groups (from **1** to **3**) is more pronounced. The dihydride isomer is stabilized about $6.0 \text{ kcal mol}^{-1}$, but this energy is not enough to reverse the thermodynamic preference for the dihydrogen. The extra stabilization is mainly due to electronic effects, with the steric effects not being relevant. When we compare the relative energies of the real systems containing the bulky phosphines and the Cp^* ligand, a stronger influence of the Cp^* methyl groups can be appreciated: 5.3 and $5.0 \text{ kcal mol}^{-1}$ for the dppe and dippe phosphines, respectively. In both cases, the substitution of Cp by Cp^* favors the dihydride complex. Assuming that the electronic effect of the methyls will be similar in the three cases, stabilizing the dihydride by about $1.0 \text{ kcal mol}^{-1}$, it is evident that steric factors of the ring substituents work in favor of the dihydride in the real complexes. Phenyl and isopropyl substituents have a major influence in the thermodynamic preference for the dihydride isomer. Changing the H model substituents with phenyls in the Cp^* -containing complexes (**2** vs **4**) favors the dihydride by $8.8 \text{ kcal mol}^{-1}$, with a similar contribution of electronic and steric effects. The influence of the isopropyls (**2** vs **6**) in favoring the dihydride is even higher ($13.5 \text{ kcal mol}^{-1}$), although in this case, the weight of the steric effects is the main cause for

the extra stabilization of the dihydride isomer. The branched isopropyl groups have a more important steric hindrance than the planar phenyl groups, rendering the dihydride the favored isomer, even with the Cp ligand (**1** vs **5**).

At a first glance, it could seem surprising that the coordination number increase associated with the dihydrogen \rightarrow dihydride interconversion could be favored on steric grounds. However, the destabilizing steric interactions are not developed between the piano-stool legs but rather between the phosphine substituents and the C_5 ring. They are signaled by the notable increase in the $[\text{C}_5 \text{ ring}(\text{centroid})-\text{Fe}-[\text{P}_2(\text{centroid})]$ (α) angle on going from the dihydrogen (α of about 150°) to the dihydride structures (α of about 180°); see the Supporting Information. In a thorough study of the η^2 -dihydrogen complexes of ruthenium, it was already pointed out that the interaction of phosphine substituents with the Cp or Cp^* ring may be significant in determining the position of the dihydrogen \rightarrow dihydride equilibrium.³ Our study fully confirms this hypothesis and quantifies the magnitude of such an interaction.

The electronic effects also stabilize the dihydride isomer. The largest effect is found with the phenyl groups. The role of the phenyl substituents is particularly important in the preferential stabilization of the isomer with the highest oxidation state. The omission of the phenyl substituents in the model complexes is the main reason for the poor results they give for the dihydrogen/dihydride relative energies.

3. Computational Study of the Dihydrogen \rightarrow Dihydride Isomerization. We now turn to the analysis of the isomerization mechanism. As mentioned in the Introduction, several mechanisms have been suggested for the isomerization of the related dihydrogen ruthenium derivatives:^{3,9,39,40} (1) a direct intramolecular rearrangement through a trigonal-bipyramidal transition state; (2) a dissociative process with the partial or total phosphine ligand decooordination; (3) a relay of the migrating hydride via the Cp ligand; (4) a deprotonation of the cis isomer followed by protonation of the metal at the trans position. The first three possibilities are intramolecular processes, while the latest involves a proton-donor molecule. Because we have accumulated evidence, including the present study, that the isomerization rate of **4** is independent of the nature of the proton donor,^{6,10,11} only intramolecular mechanisms have been explored. In addition, the new experimental results shown in this paper set an experimental value to calibrate the highest enthalpy point along the computed reaction coordinate.

To keep the computational demands affordable, the study of the reaction mechanisms was performed at a full QM level only for the **1** and **2** model complexes and at the QM/MM level with the IMOMM method for the real **4** complex (with phenyl groups at the MM level). Additional single-point energy computations at the full QM level on the real **4** complex were performed on the geometry of the intermediates and transition states located along the QM/MM potential energy surface (QM/IMOMM).

(39) Cayuela, E.; Jalon, F. A.; Manzano, B. R.; Espino, G.; Weissensteiner, W.; Mereiter, K. *J. Am. Chem. Soc.* **2004**, *126*, 7049–7062.

(40) Ryan, O. B.; Tilset, M. *J. Am. Chem. Soc.* **1991**, *113*, 9554–9561.

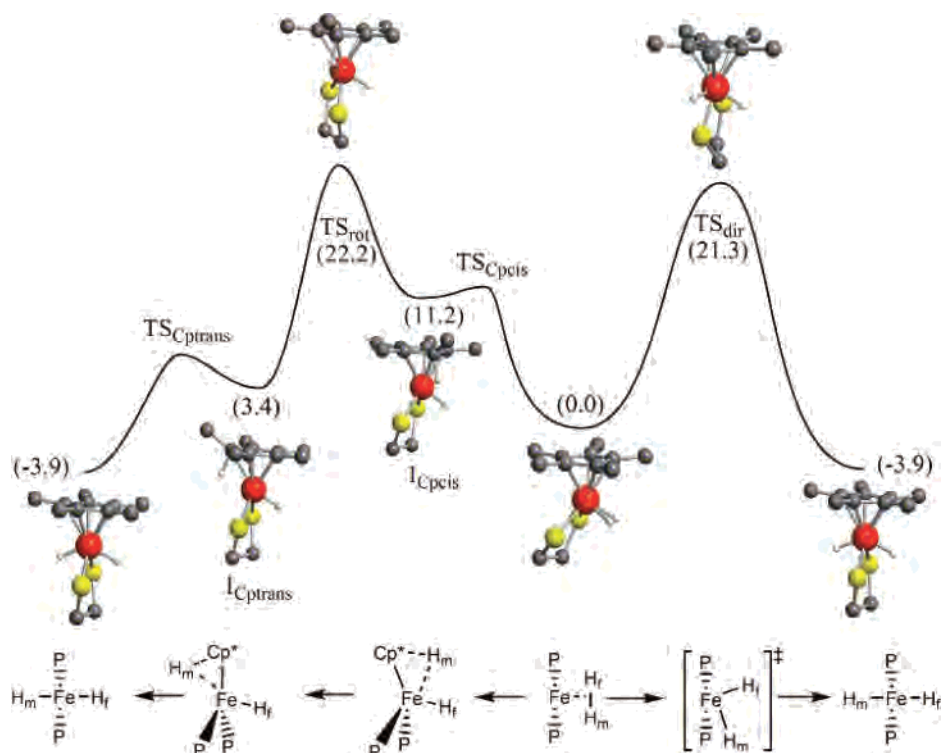


Figure 5. Direct (right) and “via Cp” (left) pathways for the *cis*-dihydrogen \rightarrow *trans*-dihydride isomerization. The optimized geometries are shown for the **2** system (the Cp* and dhpe H atoms have been omitted for the sake of clarity), whereas the energies (in kcal mol⁻¹) correspond to the QM calculations on the QM/MM optimized **4** system.

Direct Intramolecular Rearrangement. As discussed above, in these piano-stool compounds, the *cis*-dihydrogen \rightarrow *trans*-dihydride isomerization implies two processes: oxidative addition of dihydrogen and structural rearrangement of the piano-stool legs. These two processes may proceed in one or two steps. The simple cleavage of the H–H bond in the dihydrogen structure leads to a *cis*-dihydride. Although there are no reports of *cis*-dihydride [(C₅R₅)FeH₂(PP)]⁺ complexes, either isolated or spectroscopically detected, we considered also the possibility of a transient *cis*-dihydride intermediate. Starting from the dihydrogen complex and taking as a reaction coordinate the H–H distance, we calculated the energy curve for the H–H breaking. The energy was continuously increasing, and no stationary point corresponding to a *cis*-dihydride intermediate could be located. Thus, the reaction appears to be taking place in a single step.

Then, we explored the potential energy surface for the intramolecular rearrangement taking the H–Fe–H angle as the reaction coordinate. While the two hydride positions are symmetry-equivalent in the starting dihydrogen complex, they become inequivalent in the transition state. They will be identified as the “fixed” (H_f) and “migrating” (H_m) positions, respectively. The H_f–Fe–H_m angle must undergo a large change during the process: from 32° in the dihydrogen complex to 131° in the dihydride complex. Starting from the maximum of this monodimensional energy profile, we were able to locate the transition state for the direct intramolecular rearrangement (TS_{dir}). This presents a single imaginary frequency (298i cm⁻¹), with the opening of the H_f–Fe–H_m angle as the main component of its associated

Table 3. Main Geometrical Parameters (Distances in Angstroms and Angles in Degrees) of the Transition States for the Two Reported Dihydrogen \rightarrow Dihydride Interconversion Mechanisms of **2**^a

	2c	TS _{dir}	TS _{rot}	2t
Fe–H	1.576	1.482	1.483	1.482
	1.580	1.490	1.644	1.482
H _f –H _m	0.869	2.282	2.274	2.698
Fe–P	2.220	2.149	2.165	2.166
	2.208	2.203	2.183	2.167
Fe–C(Cp) _{range}	2.121	2.115	2.104	2.105
	2.159	2.202	2.238	2.135
α ^b	144.8	158.9	162.3	179.3
P ₁ –Fe–P ₂	85.2	84.0	85.1	91.5
H _f –Fe–H _m	32.0	100.3		131.0
Cp _c *–Fe–H _f –H _m ^c	36.9	100.7		179.7
Cp _c *–Fe–P ₁ –P ₂ ^c	115.8	151.8		178.8
Cp _c *–Fe–P ₁ ^c	126.0	153.6		134.0
Cp _c *–Fe–P ₂ ^c	128.0	119.0		134.0

^a Values for the initial dihydrogen and the final dihydride complexes are also included. ^b α = angle [C₅ ring(centroid)]–Fe–[P₂(centroid)]. ^c Cp_c* = Cp* centroid.

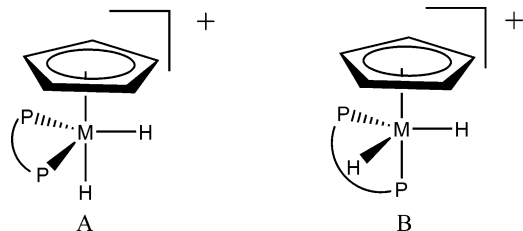
eigenvector. The mechanism of the process is illustrated in Figure 5 (right-hand side), together with the structure of the TS_{dir} for the **2** system. The main geometric parameters of the TS_{dir} for the Cp*–dhpe system are given in Table 3, while the full list of calculated frequencies as given as the Supporting Information. Similar structural values, not reported for the sake of brevity, are obtained for the Cp–dhpe and Cp*–dppe systems.

The transition state has a dihydride nature. The H–H bond appears totally broken, as reflected by the long H–H separation, whereas the two Fe–H bonds are already formed. Consequently, the oxidative addition process is already completed at the transition state level; the structure can be

Table 4. Relative Energies (ΔE , kcal mol⁻¹) of the Species Involved in the Two Mechanisms

	dihydrogen	TS _{Cp_{cis}}	I _{Cp_{cis}}	TS _{Rot}	TS _{dir}	I _{Cp_{trans}}	TS _{Cp_{trans}}	dihydride
[CpFe(dhpe)H ₂] ⁺	0.0				30.1			6.6
[Cp*Fe(dhpe)H ₂] ⁺	0.0	15.3	14.9	30.0	28.9	11.5	12.5	4.9
[Cp*Fe(dppe)H ₂] ^{++a}	0.0		12.1	23.9	24.5	5.9		0.1
[Cp*Fe(dppe)H ₂] ^{++b}	0.0		11.2	22.2	21.3	3.4		-3.9

^a QM/MM calculation with the phenyl groups in the MM part. ^b Full QM calculation at the QM/MM-optimized geometries.

Chart 2

described as an iron(IV) dihydride. Nevertheless, the α angle is closer to its initial value in the dihydrogen complex than to the 180° value in the final dihydride. The phosphine ligand is also experiencing a significant repositioning. The dihedral angle formed by the C₅ ring (centroid), the Fe atom, and the two P atoms takes a midway value (about 150°) between the 115° in the dihydrogen and the 180° in the *trans*-dihydride. The movement of the phosphine ligand can also be appreciated from the two Cp_c*-Fe-P angles. These are close to each other in both the starting dihydrogen and final dihydride complexes, and the values do not change much on going from one to the other (126° and 128° for the dihydrogen complex and 134° and 134° for the dihydride complex). However, one angle is opened to 150°, whereas the other one is closed to 120°, in the transition state. This illustrates the swing of the entire diphosphine ligand toward one side of the molecule, which is necessary to allow passage of one hydride ligand from the side, to move from the front to the back of the metal center.

Given that both the transition state and the final *trans*-dihydride are four-legged-piano-stool complexes, they can be analyzed in relation to the mechanisms invoked for the CpMo(CO)₂LR *cis*-*trans* interconversion.^{3,41} Within this model, the two trigonal-bipyramidal-like transition states presented in Chart 2 can be envisaged. The geometric parameters of the transition states are in closer agreement with a structure of type **B**, describing a movement in which one of the phosphine arms has moved down to allow the migration of the hydride ligand toward its final position. The generation of a four-legged-piano-stool product with a *transoid* disposition of the E and H ligands resulting from the oxidative addition of a E-H bond by a (C₅R₅)ML₂ fragment is a relatively common result, not well explained yet. A transition state closely related to TS_{dir} has been reported for the *cis*-*trans* isomerization of a hydridoaryl complex of rhenium.⁴² It could also be at work for the oxidative addition of an alkyne C-H bond to yield a *trans*-alkynyl hydride.³⁷

The transition state TS_{dir} lies 30.1 and 28.9 kcal mol⁻¹ above the dihydrogen for **1** and **2**, respectively (Table 4). In these model systems, the effect of the Cp methyl substituents is to decrease slightly the potential energy barrier. We have also computed the enthalpy and free energy of activation, namely, $\Delta H^\ddagger = 27.1$ kcal mol⁻¹ and $\Delta G^\ddagger = 30.7$ kcal mol⁻¹ for **2** at $T = 298.15$ K. The small, almost zero, activation entropy ($\Delta S^\ddagger = -3.9$ eu) is not inconsistent with the small value observed in this work for the iron complex (*vide supra*) and with the experimentally reported $\Delta S^\ddagger (-3.3 \pm 0.7$ eu) in related ruthenium complexes.³ As was already found for the dihydrogen-dihydride equilibrium, the energy barrier of the rearrangement is only slightly affected by solvent effects. The computed barrier in dichloromethane for the **2** system is 28.6 kcal mol⁻¹, very close to the gas-phase value (28.9 kcal mol⁻¹).

As reported for the relative stabilities of the dihydrogen and dihydride complexes, the effects of the substituents (electronic and steric) are also major on the transition-state energy. The stabilization found for the Cp* transition state is similar to those found for the minima, in agreement with the dihydride nature of the transition state. The same trend could be inferred for the phenyl effects, and this is indeed the case. The potential energy barrier for the real system **4**, with the phenyl groups described at the QM/MM level, is computed as 24.5 kcal mol⁻¹, and full QM energy computations at the fixed QM/MM minima and transition state further lower the energy barrier to 21.3 kcal mol⁻¹. Thus, the effect of the phenyl substituents is to considerably decrease the isomerization barrier. The calculated barrier is very close to the experimental value of 21.6 ± 0.8 kcal mol⁻¹ (see above) and is also close to the barrier of the related isomerization of the [Cp*Ru(dippe)(H₂)]⁺ complex.⁹

Phosphine Dissociation Mechanism. We have also considered a dissociative isomerization mechanism, in which one of the P atoms in the [Cp*Fe(dhpe)(H₂)]⁺ complex dissociates from the metal, yielding an unsaturated intermediate where the diphosphine acts as a monodentate ligand. The oxidative addition would subsequently take place in this intermediate. However, the nonclassical dihydrogen complex with one phosphine arm fully decoordinates is 30.0 kcal mol⁻¹ above the saturated dihydrogen complex. This high Fe-P binding energy value led us to discard this mechanism for the systems under study.

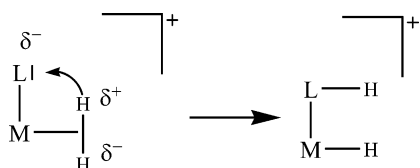
Isomerization through the Cp Ring. Some examples of hydrogen migration involving the metal center and the Cp ring can be found in the literature for iron compounds. Examples are provided by [η^4 -C₅H₅(*exo*-D)]Fe(CO)₃,⁴³

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Scheme 3



$\text{Fe}(\text{CO})_3(\eta^4\text{-C}_6\text{H}_7\text{Ph})$,⁴⁴ and $(\eta^5\text{-C}_5\text{H}_5)\text{FeH}(\text{triphos})$ (triphos = $\text{Ph}_2\text{PCH}_2\text{CH}_2\text{PPh-CH}_2\text{CH}_2\text{PPh}_2$).⁴⁵ These isomerizations take place under more drastic conditions relative to the dihydrogen \rightarrow dihydride interconversion examined here. In addition, this mechanism was briefly contemplated for the analogous ruthenium system by Chinn and Heinekey³ but considered unlikely on the basis of bond strength considerations. During the thorough exploration of the potential energy surface for the dihydrogen \rightarrow dihydride rearrangement in the $[\text{Cp}^*\text{Fe}(\text{dhpe})(\text{H}_2)]^+$ system, we discovered the existence of low-lying intermediates with a protonated C_5 ring. This finding prompted us to fully explore the “via Cp” mechanism.

This mechanism appears as a multistep pathway in which the H_m atom migrates from the cis to the trans position via a successive series of H jumps from and to the metal, together with a rotation of the protonated C_5 ring. The mechanism of the process is illustrated in Figure 5 (left-hand side), along with the structure of all intermediates for the **2** system. The most relevant geometrical parameters of the TS_{rot} (one imaginary frequency with $87i\text{ cm}^{-1}$) for the $\text{Cp}^*\text{-dhpe}$ system are given in Table 3, while the full list of calculated frequencies are given as Supporting Information. The two intermediates $\text{I}_{\text{Cp}^*\text{cis}}$ and $\text{I}_{\text{Cp}^*\text{trans}}$ are formally unsaturated, 16-electron dieneiron(II) complexes, $[(\eta^4\text{-C}_5\text{Me}_5\text{H})\text{Fe}(\text{dppe})\text{H}]^+$. However, they are stabilized by an agostic interaction established by the newly created $\sigma\text{ C-H}$ bond and the Fe center, which keeps the C_5 ring essentially planar. The C-H and Fe-H distances (1.248 and 1.661 Å in $\text{I}_{\text{Cp}^*\text{cis}}$; 1.216 and 1.757 Å in $\text{I}_{\text{Cp}^*\text{trans}}$) confirm the presence of this interaction. The H-H bond is completely broken in $\text{I}_{\text{Cp}^*\text{cis}}$ (H-H = 2.005 Å). In the rotation transition state TS_{rot} , the proton is midway between the two positions. The H-H distance is almost constant along the rotation. The TS_{rot} also presents a $\sigma\text{ C-H}$ agostic interaction with the Fe center (C-H = 1.179 Å; Fe-H = 1.483 Å). The transition states $\text{TS}_{\text{Cp}^*\text{cis}}$ and $\text{TS}_{\text{Cp}^*\text{trans}}$ have geometries very near those of the corresponding intermediates, $\text{I}_{\text{Cp}^*\text{cis}}$ and $\text{I}_{\text{Cp}^*\text{trans}}$, respectively. The optimized geometries of all stationary points are available as Supporting Information.

This first step can be viewed as an intramolecular heterolytic splitting of the H-H bond of the dihydrogen ligand. In cationic complexes, the $\eta^2\text{-H}_2$ ligand can be acidic enough to protonate a basic cis ligand; see Scheme 3,⁴⁶ whereas the basic properties of a Cp^* ligand are well-known. We have searched the related ring-protonated intermediate

for the simpler Cp model complex, but this search was unsuccessful. All of our attempts to obtain such species reverted to the dihydrogen complex. Evidently, the simpler Cp ligand is not basic enough for this system to be protonated by the dihydrogen. We must point out that, for those CpFe complexes where the hydrogen migration from the metal center to the Cp ring has been experimentally observed, this process occurs under more drastic conditions. In terms of the currently examined mechanism, replacing Cp by Cp^* should facilitate the migration, making it possible under milder conditions.

The relative energies of the intermediates and transition states for the “via Cp” mechanism for the **2** and **4** systems are given in Table 4. In the model system, the $\text{I}_{\text{Cp}^*\text{cis}}$ intermediate lies 14.9 kcal mol⁻¹ above the dihydrogen and the transition state that connects both is 0.4 kcal mol⁻¹ above the intermediate. The *trans*- Cp^* -protonated complex is 3.4 kcal mol⁻¹ more stable than the *cis*-protonated derivative. The transition state for the rotation of the protonated ring is found to be 30.0 kcal mol⁻¹ above the initial dihydrogen complex in the gas phase and 30.8 kcal mol⁻¹ in dichloromethane. The energy barrier for the Cp^* rotation is 15.1 kcal mol⁻¹. The presence of the strong agostic bond, which is preserved along the rotation, is the main reason for this high value for a Cp^* rotational process. Therefore, the rate-determining step of the “via Cp” mechanism is the rotation of the protonated Cp^* . The computed activation enthalpy and free energy are $\Delta H^\ddagger = 29.2\text{ kcal mol}^{-1}$ and $\Delta G^\ddagger = 30.7\text{ kcal mol}^{-1}$. This intramolecular mechanism also implies a small negative activation entropy ($\Delta S^\ddagger = -5.0\text{ eu}$). In the gas phase, the energy barriers of the “direct” and “via Cp” mechanisms for the dihydrogen \rightarrow *trans*-dihydride interconversion in the model **2** system are very similar (about 30 kcal mol⁻¹). Dichloromethane slightly favors the direct mechanism, increasing the difference between both barriers from 1.1 to 2.2 kcal mol⁻¹. Next, we examine how the presence of the phenyl substituents in the real system affects this result.

For the **4** complex, we have not been able to find the transition states $\text{TS}_{\text{Cp}^*\text{cis}}$ and $\text{TS}_{\text{Cp}^*\text{trans}}$ because of the flatness of the potential energy surface around the intermediates. The main effect of the phenyl groups is to stabilize all of the species involved in the mechanism with respect to the dihydrogen complex (see Table 4). Given that the phenyl groups are described at the MM level, the explanation for this behavior should be in the steric effects. On the basis of the results presented above on the relative stabilities of the dihydrogen and dihydride forms, the main reason for this behavior could be attributed to a destabilization of the dihydrogen complex by the steric interaction between phenyl and methyl groups. Very close values of the energy barriers for the “direct” and “via Cp” mechanisms are found also for the real system at the QM/MM level: 24.5 kcal mol⁻¹ (“direct”) and 23.9 kcal mol⁻¹ (“via Cp”).

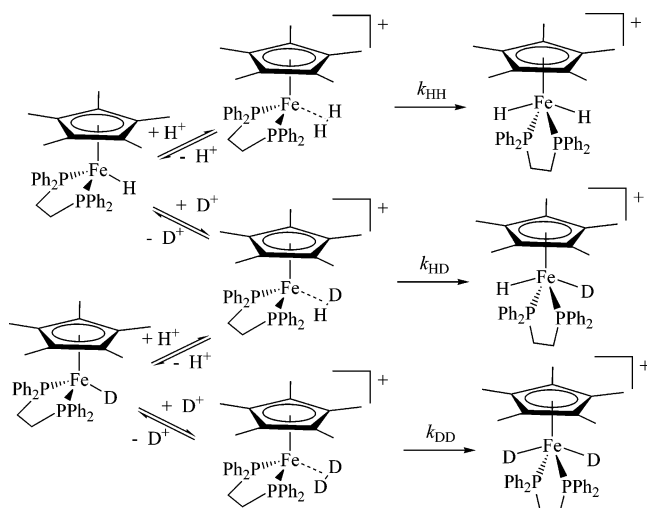
To introduce the electronic effects of the phenyl substituents, we have performed full QM computations at the fixed QM/MM structures of the optimized intermediates and transition state of the **4** system. The relative energies of all

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Scheme 4



of the intermediates and TS_{rot} slightly decrease relative to the dihydrogen (Table 4). At this level, the energy barrier for the “via Cp” pathway remains again close to that computed for the “direct” pathway and to the experimentally determined values. Thus, the possibility of a dihydrogen \rightarrow dihydride interconversion via hydrogen migration to the Cp* ring and involving an heterolytic H–H bond breaking appears as a feasible and competing pathway, together with the expected direct oxidative addition mechanism in these cationic iron half-sandwich complexes. Therefore, the computational work carried out so far does not allow a clear-cut distinction of the two alternative mechanisms. For this reason, we have decided to carry out a more detailed analysis of the KIE.

4. Isomerization Kinetics of $[Cp^*Fe(dppe)(\eta^2-HD)]^+$. The isomerization kinetics starting from the hydride complex $Cp^*Fe(dppe)H$ and the deuterated trifluoroacetic acid, CF_3COOD , in a 1:1 ratio provided additional information on the KIE. We carried out this experiment hoping to be able to measure the isomerization rate constant of the pure η^2-HD complex (k_{HD}) and compare this with the results of the rate constants given in Table 1 for the η^2-H_2 (k_{HH}) and η^2-D_2 (k_{DD}) complexes. As it turns out, the outcome of this experiment was more complex and interesting than expected because the nonclassical species obtained by low-temperature protonation establishes a rapid equilibrium with the mixture of the other two isotopomers. This means that, although the proton transfer to the hydride site is effectively quantitative when using a strong acid such as CF_3COOD , the process is still subject to rapid reversibility, generating both CF_3COOH and $Cp^*Fe(dppe)D$, which further lead to the other two nonclassical complexes as shown in Scheme 4. Therefore, subsequent isomerization yields all three classical products.

However, useful kinetics information could still be obtained from this experiment because (i) the equilibration between the three nonclassical isotopomers according to the equilibrium shown in eq 1 is much more rapid than the isomerization rate, (ii) the isomerization is irreversible, and (iii) the classical dihydride products of the isomerization

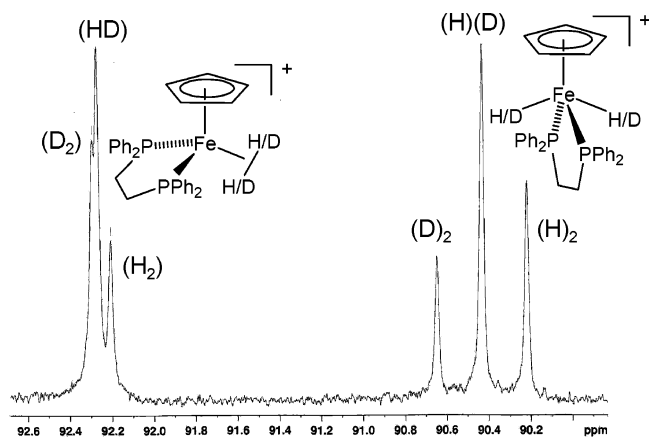
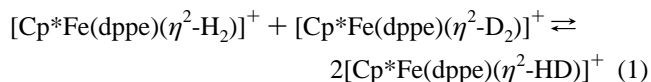


Figure 6. Example of a ^{31}P NMR spectrum obtained during the kinetic run of the isomerization experiment starting from $Cp^*FeH(dppe) + CF_3COOD$ (1:1) at 0 °C in CD_2Cl_2 .

process do not scramble the hydride ligands. The latter statement was verified by an independent experiment, where a sample of $[Cp^*Fe(dppe)(H_2)]^+$ was generated and then treated with an equivalent amount of CF_3COOD . No deuterium incorporation occurred over 2 h at 273 K (the isomerization process is extensive over this time scale, as can be seen in Figure 1).



The information on the individual concentrations of the three nonclassical and the three classical species from a kinetic run carried out at 0 °C, starting with approximately equimolar amounts of $Cp^*FeH(dppe)$ and CF_3COOD , was obtained by integration of the inverse-gated doubly decoupled $^{31}P\{^1H, ^2H\}$ NMR resonances. An example of a recorded spectrum is shown in Figure 6. The double decoupling from proton and deuterium is essential to generating sufficiently sharp resonances and allowing a satisfactory deconvolution of the individual components, especially for the nonclassical isotopomers. The inverse-gated procedure, together with long relaxation times and the use of a small pulse angle (see details in the Experimental Section), allowed one to remove or at least minimize the distortion of the integrated intensities caused by the nuclear Overhauser effect and by an incomplete magnetization recovery. Visual inspection of Figure 6 shows that the classical structure experiences a much greater isotope shift than the nonclassical one. This is quite expected because the isotope effect in the former is transmitted from the hydride ligands to the detected ^{31}P nucleus through stronger Fe–H covalent bonds.

The evolution of the signal intensities is shown in Figure 7. The presence of a KIE is already quite evident from a qualitative analysis of the experimental data. In fact, the concentration of the η^2-H_2 complex is always below that of the η^2-D_2 complex, whereas the classical dihydride product always has a greater concentration than the dideuteride product. Fitting the calculated decays of the three nonclassical species and the calculated growth of the three classical species to the experimental data gave the solid lines indicated

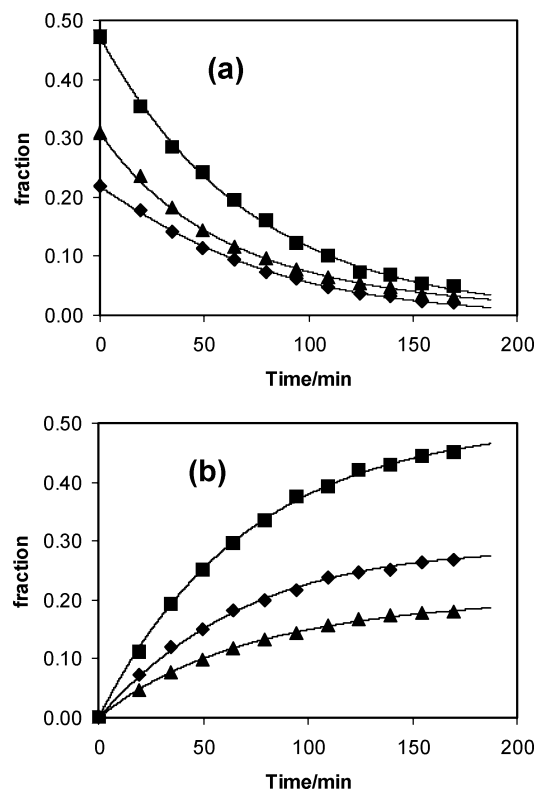


Figure 7. Evolution of the concentrations of (a) nonclassical and (b) classical complexes during the isomerization process following the reaction of $\text{Cp}^*\text{Fe}(\text{dppe})\text{H}$ with CF_3COOD at 0°C in CD_2Cl_2 : diamonds, dihydride; squares, monohydride monodeuteride; triangles, dideuteride. The solid lines are the best fit on the basis of the kinetic model, as outlined in the text and in the Supporting Information.

in Figure 7 for the optimized rate constant values reported in Table 1 (run 8). The details of the data analysis, which required the numerical integration of the coupled differential equations and a global least-squares fitting procedure, are available as Supporting Information. Note that the optimized value for k_{HH} is identical within experimental error to that independently (and more directly) obtained from the CF_3COOH protonation (Table 1, run 6). In addition, the optimized value of k_{DD} is slightly lower than that obtained directly from the CF_3COOD protonation of the deuteride sample (Table 1, run 7), which must be considered an upper estimate, as discussed above.

The data analysis (see the Supporting Information) used the assumptions that the mixture of the three nonclassical isotopomers is instantaneously equilibrated by virtue of eq 1, with an equilibrium constant taken as the average calculated from the experimental data ($K = 3.3 \pm 0.3$). Note that this value is significantly different from that expected for a statistical distribution without an isotope effect ($K = 4$). When the fitting procedure was repeated by manually altering this K value to 4 or to other values, the obtained residual was greater, confirming the significance of this deviation. Such observations prove that, besides the KIE affecting the dihydrogen–dihydride isomerization, there is an additional equilibrium isotope effect (EIE) on the isotope redistribution equilibrium among the three nonclassical species. The reason for this EIE will be analyzed in the next section.

From the results in Table 1 (run 8), we can derive $k_{\text{HH}}/k_{\text{HD}} = 1.24 \pm 0.01$ and $k_{\text{HD}}/k_{\text{DD}} = 1.58 \pm 0.01$. Isotope effects have been previously reported for the oxidative addition of dihydrogen to transition metals but only for reactions between a metal complex and free dihydrogen, to the best of our knowledge. Furthermore, they only relate to the addition of dihydrogen versus dideuterium. Examples ($k_{\text{H}}/k_{\text{D}}$ in parentheses) are the additions to $\text{Fe}(\text{CO})_4$ (1.1 ± 0.1 at 24°C),⁴⁷ to $\text{Ir}(\text{PPh}_3)_2(\text{CO})\text{Cl}$ (1.09 at 25°C),^{48,49} to $[\text{Rh}(\text{PPh}_3)_2\text{Cl}]$ (1.5 at 23°C),⁵⁰ to $[\text{Cr}(\text{CO})_5(\text{C}_6\text{H}_{12})]$ (1.9 at 23°C),⁵¹ and to $[\text{W}(\text{PMe}_3)_3\text{I}_2]$ (1.2 ± 0.2 at 60°C).⁵² While the oxidative addition process may take place in a single step, a local energy minimum for an intermediate dihydrogen complex may also occur. In the latter case, when the slow step is the intramolecular rearrangement of the dihydrogen ligand, the observed isotope effect results from the combination of an EIE for the dihydrogen coordination step and a KIE for the rearrangement. The dihydrogen coordination equilibrium constant has been measured for a number of systems and is always characterized by an inverse EIE. Examples (with the $K_{\text{H}}/K_{\text{D}}$ in parentheses) are $[\text{Ir}(\text{PCy}_3)_2\text{HCl}_2]$ (0.50 at -13°C),⁵³ $[\text{Ir}(\text{Pr}^i\text{Bu}_2\text{Me})_2\text{H}_2\text{Cl}]$ (0.37 at -13°C),⁵⁴ $[\text{Os}(\text{P}^i\text{Pr}_3)_2(\text{CO})(\text{Cl})\text{H}]$ (0.35 at 85°C),⁵⁵ $[\text{Cr}(\text{CO})_3(\text{PCy}_3)_2]$ (0.65 ± 0.15 at 22°C),⁵⁶ and $[\text{W}(\text{PMe}_3)_4\text{I}_2]$ (0.63 ± 0.05 at 60°C).⁵² Thus, a greater normal KIE for the isomerization process would be necessary to yield an observed normal KIE for the oxidative addition. Our determined $k_{\text{HH}}/k_{\text{DD}}$ value is 1.96 ± 0.02 at 273 K , i.e., comparable to the highest reported value for an oxidative addition process, namely, 1.9 for $[\text{Cr}(\text{CO})_5(\text{C}_6\text{H}_{12})]$.⁵¹ The present report appears to be the first direct investigation of the KIE for the rearrangement process of a dihydrogen intermediate to the final dihydride product and the first complete analysis encompassing the H_2 , HD , and D_2 isotopomers. Note, however, that the rearrangement does not consist of a simple H–H bond breaking completed by the formation of the two M–H bonds with little change in the rest of the coordination sphere because it occurs in the majority of the other cases. A cis to trans rearrangement also takes place. Thus, the current results should not be extrapolated in general.

We can conclude that the isomerization process from $[\text{Cp}^*\text{Fe}(\text{dppe})(\eta^2\text{-H}_2)]^+$ to **4** has a small isotope effect and

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that this is not symmetric relative to the two hydride ligands. This is not surprising because the two hydride ligands are inequivalent in the rearrangement process (one migrates to the new position, and the second one remains on the starting site). Both KIEs and the EIE will now be analyzed from a theoretical point of view, in terms of the two mechanistic pathways that were found energetically compatible with the measured activation enthalpy.

5. Theoretical Analysis of the EIE. On the basis of the energy-optimized structure of the nonclassical model complex $[\text{Cp}^*\text{Fe}(\text{dhpe})(\eta^2\text{-H}_2)]^+$, i.e., system **2**, frequency calculations were run for the HD and D₂ analogues. For the sake of clarity, the three isotopomers have been named **2-H₂** (nondeuterated), **2-HD** (monodeuterated), and **2-D₂** (dideuterated). The EIE can be obtained directly from the calculated ΔG of the isotope redistribution equilibrium (see thermochemical data in the Supporting Information). This is obtained from a statistical mechanics treatment that uses unscaled normal-mode frequencies. The resulting value of the equilibrium constant for eq 1 is 3.81. The slight disagreement with the experiment is expected, given the various approximations used by the calculation.

From the knowledge of the individual frequencies, the equilibrium constant may also be calculated through the classical Bigeleisen–Mayer treatment, where the constant is given by the expression in eq 2. For the definition of the

$$K = 4(\text{MMI})(\text{EXC})(\text{ZPE}) \quad (2)$$

MMI (mass moment of inertia), EXC (excitations of vibrational energy levels), and ZPE (zero-point vibrational energy) terms, see the Supporting Information and the literature.^{57,58} The factor of 4 takes into account the doubled concentration of the HD species in the equilibrium expression. The advantage of the Bigeleisen–Mayer treatment is to pinpoint the isotope-sensitive modes that are mostly responsible for the observed effect. Related analyses have been previously reported for the secondary isotope effects in the coordination of C₂H₄ vs C₂D₄ to the Os₂(CO)₈ moiety⁵⁹ and for the primary isotope effects in the H₂/D₂, CH₄/CH₃D, and CH₄/CD₄ oxidative additions to *trans*-Ir(PR₃)₂(CO)X,⁶⁰ in the H₂/D₂ coordination to W(CO)₃(PCy₃)₂,⁵⁶ and in the H₂/D₂/T₂ coordination and oxidative addition to W(CO)₅.⁶¹ The isotopic redistribution between H₂, HD, and D₂ complexes has never been analyzed to the best of our knowledge.

This type of analysis was carried out for our system on the basis of the selected, most isotope-sensitive frequencies, reported in Table 5 (no scaling factors have been used). Only these normal modes are sensitive to the isotope nature and contribute significantly to the EIE. A full list of vibrational frequencies is provided in the Supporting Information section. The calculated K value is 3.30 at 273 K, in agreement with

Table 5. Frequency Analysis of the EIE for the Isotope Redistribution Equilibrium in Complex $[\text{Cp}^*\text{Fe}(\text{dhpe})(\eta^2\text{-XY})]^+$ (**2-XY**) (X, Y = H, D) at 273 K

ν/cm^{-1}			assignment	MMI	EXC	ZPE	K
2-H₂	2-HD	2-D₂					
2762.7	2423.5	1956.3	$\nu(\text{H-H})$	1.09	0.95	0.71	2.94
1810.3	1561.0	1283.9	$\nu_{\text{as}}(\text{Fe-H})$	1.05	0.98	0.93	3.84
1096.9	903.1	785.6	$\nu_{\text{sym}}(\text{Fe-H})$	0.95	1.03	1.22	4.78
680.4	651.5	512.9	$\delta(\text{Fe-H}_2)_{\text{out-of-plane}}$	1.22	1.00	0.75	3.65
585.5	533.2	534.6	$\delta(\text{Fe-H}_2)_{\text{in-plane}}$	0.91	1.00	1.15	4.19
468.1	370.6	359.4	$\tau(\text{Fe-H}_2)$	0.82	1.00	1.26	4.10
			total	0.97	0.96	0.88	3.30

Table 6. Calculated Activation Free Energies in kcal mol⁻¹ [and Relative $\Delta\Delta G^\ddagger$ Values with Respect to $\Delta G^\ddagger(\text{HH})$ in Parentheses] for the “Direct” and “via Cp” Mechanisms at 273 K (Model **2** System)

mechanism	$\Delta G^\ddagger(\text{HH})$	$\Delta G^\ddagger(\text{H}_m\text{D}_f)$	$\Delta G^\ddagger(\text{D}_m\text{H}_f)$	$\Delta G^\ddagger(\text{DD})$
direct	29.26 (0.00)	29.41 (0.15)	29.54 (0.28)	29.70 (0.44)
via Cp	31.27 (0.00)	31.41 (0.14)	31.15 (-0.12)	31.31 (0.04)

the experimental value. As Table 5 shows, the EIE is dominated by the ZPE term, and the most significant contribution to the reduction of K from the statistical value of 4 is given by the $\nu(\text{H-H})$ vibration. However, the contribution of the other isotope-sensitive modes cannot be neglected.

6. KIE on the Dihydrogen–Dihydride Isomerization.

Analogous frequency calculations were run on the isotopomers of the two transition states corresponding to the two mechanisms found compatible with the experimental activation barrier. We must also consider that the symmetry of the dihydrogen ligand is broken on going from the initial nonclassical system to the transition state. While the two hydride positions are symmetry-equivalent in the starting dihydrogen complex, they become inequivalent in the transition state. They will be identified as the “fixed” (H_f/D_f) and “migrating” (H_m/D_m) positions, respectively (see Figure 5). This implies that two pathways for each mechanism must be considered in the case of the monodeuterated species, depending on whether the migrating atom is H or D.

As for the EIE analysis, the KIE can be derived directly from the calculated free energies, in this case the activation free energies. These are reported in Table 6, and the corresponding ZPE and thermal enthalpy values are included in the Supporting Information. Clearly, the migration of hydrogen is favored over the migration of deuterium in the mixed **2-HD** species for the “direct” mechanism, whereas the reverse is true for the “via Cp” mechanism. The resulting calculated isotope effects are given in Table 7. The $\Delta\Delta G^\ddagger$ values for the two different transition states of the **2-HD** systems (H_mD_f and D_mH_f) are small in each mechanism (−0.117 kcal mol⁻¹ for TS_{dir} and 0.264 kcal mol⁻¹ for TS_{rot}), allowing the calculation of $k(\text{H}_m\text{D}_f)/k(\text{D}_m\text{H}_f) = 1.241$ for the “direct” and 0.615 for the “via Cp” mechanism. Thus, a considerable fraction of the HD sample will isomerize by the less favorable pathway (H_mD_f/D_mH_f = 55.4:44.6 for the “direct” and 38.1:61.9 for the “via Cp” mechanism). The observed KIE must take this into account. Thus, the calculated effective rate of isomerization for the **2-HD**

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Table 7. Frequency Analysis of the KIE for the Irreversible Isomerization of $[\text{Cp}^*\text{Fe}(\text{dhpe})(\eta^2\text{-AB})]^+$ to $[\text{Cp}^*\text{Fe}(\text{dhpe})(\text{A})(\text{B})]^+$ (A, B = H, D) at 273 K

X/Y ^a	VP*	EXC*	ZPE*	k_X/k_Y^b	k_X/k_Y^c
(a) "Direct" Mechanism					
HH/H _m D _f	0.938	1.050	1.313	1.304	1.299
HH/D _m H _f	0.753	1.126	1.380	1.619	1.536
H _m D _f /DD	0.772	1.093	1.521	1.693	1.680
D _m H _f /DD	0.961	1.019	1.447	1.364	1.421
(b) "Via Cp" Mechanism					
HH/H _m D _f	1.018	1.027	1.150	1.282	1.215
HH/H _m D _f	0.997	1.100	0.742	0.788	0.816
H _m D _f /DD	1.018	1.068	0.824	0.826	0.897
H _m D _f /DD	1.038	0.998	1.277	1.344	1.336

^a The m and f labels refer to the "moving" and "fixed" atoms involved in the rearrangement. ^b From the ΔG^\ddagger analysis (see the text). ^c From eq 3 and the selected frequencies in Table 8.

species will be $k_{\text{HD}} = 0.554k(\text{H}_m\text{D}_f) + 0.446k(\text{D}_m\text{H}_f)$ for the "direct" pathway and $k_{\text{HD}} = 0.381k(\text{H}_m\text{D}_f) + 0.619k(\text{D}_m\text{H}_f)$ for the "via Cp" pathway. This leads to the calculation of the following values for the KIE:

(a) direct	(b) via Cp
$k_{\text{HH}}/k_{\text{HD}} = 1.428$	$k_{\text{HH}}/k_{\text{HD}} = 0.924$
$k_{\text{HD}}/k_{\text{DD}} = 1.547$	$k_{\text{HD}}/k_{\text{DD}} = 1.147$

The computed values are in much better agreement with the experimental ones (1.24 ± 0.01 and 1.58 ± 0.01 , respectively) for the "direct" mechanism. This result must, therefore, be considered in strong support of the occurrence of the "direct" mechanism and against the "via Cp" mechanism.

As for the EIE analysis, the KIEs can also be calculated by an extension of the Bigeleisen–Mayer method,^{58,62} with the relevant relationship becoming that of eq 3. The

$$\text{KIE} = k_X/k_Y = (\text{VP}^*)(\text{EXC}^*)(\text{ZPE}^*) \quad (3)$$

expressions of the VP*, EXC*, and ZPE* terms are based on the starting materials and transition states rather than on the starting materials and products. They are again available in the Supporting Information and in the literature.

Table 8 contains only those frequencies that are sensitive to the isotope nature and contribute significantly to the KIE. The full list of frequencies is available in the Supporting Information. The frequency analysis in this case is not as straightforward as that for the EIE, especially for the "via Cp" mechanism, because the Fe–X/Y and X–Y stretching and bending motions are heavily mixed with other molecular motions in several normal modes. Only those normal modes having a major component from the Fe–X/Y bonds have been explicitly assigned in Table 8. The calculation of the various terms of eq 3 leads to the values collected in Table 7. It can be seen that the ZPE* term contributes the most but the VP* term also gives an important contribution, especially for the "direct" mechanism when comparing different moving atoms (H_m/D_m). This is mostly due to the ν_X^*/ν_Y^* term contained in VP* (ratio of the imaginary frequencies; see the Supporting Information) because large changes at the level of the Fe–H_m and H_f–H_m bonds occur

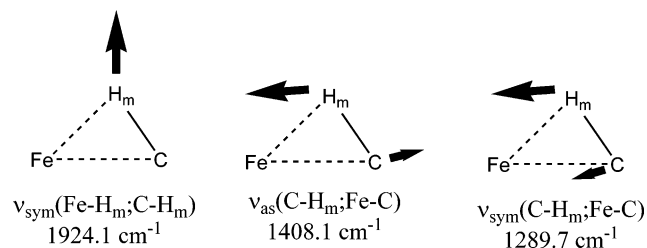
Table 8. Isotope-Sensitive Frequencies (cm⁻¹) for the Transition States of the "Direct" (TS_{dir}) and "via Cp" (TS_{rot}) Mechanisms

(a) "Direct" Mechanism				
TS _{dir} -H ₂	TS _{dir} -H _m D _f	TS _{dir} -D _m H _f	TS _{dir} -D ₂	assignment
-298.3	-296.8	-227.2	-226.6	ν^*
507.5	601.2	507.1	598.8	
536.5	523.6	536.3	522.6	
587.0	593.8	586.8	593.7	
591.7	568.5	589.0	567.3	
654.5	466.1	651.2	465.1	$\delta(\text{Fe}-\text{H}_f)_{\text{in-plane}}$
667.5	545.4	662.6	544.3	
677.0	668.4	675.6	664.7	
778.7	671.4	783.1	668.9	
1025.3	1024.9	719.3	721.3	$\delta(\text{Fe}-\text{H}_m)_{\text{out-of-plane}}$
1921.0	1369.7	1921.3	1369.0	$\nu(\text{Fe}-\text{H}_f)$
2027.7	2027.5	1448.9	1449.4	$\nu(\text{Fe}-\text{H}_m)$
(b) "Via Cp" Mechanism				
TS _{rot} -H ₂	TS _{rot} -H _m D _f	TS _{rot} -D _m H _f	TS _{rot} -D ₂	assignment
-86.8	-85.9	-86.6	-85.7	ν^*
364.8	352.1	364.3	351.8	
394.7	394.3	382.1	381.7	
440.1	411.6	439.3	411.5	
469.1	448.5	467.6	447.8	
531.2	530.6	519.6	516.8	
537.0	544.0	535.4	543.2	
548.6	542.0	546.1	539.6	
637.3	529.4	637.1	529.1	$\delta(\text{Fe}-\text{H}_f)$
686.8	686.7	681.4	681.4	
768.6	756.2	768.3	755.9	
772.9	776.7	772.8	776.6	
801.6	802.3	834.4	834.7	
819.1	603.3	813.6	602.6	
823.7	822.4	821.3	817.5	
846.6	834.0	847.3	832.6	
929.9	929.9	699.9	700.0	$\delta(\text{Fe}-\text{H}_m-\text{C})$
983.2	982.9	1121.7	1121.7	
1059.8	1059.7	998.1	997.7	
1083.3	1083.2	1075.1	1075.0	
1289.7	1289.7	906.9	906.6	$\nu_{\text{sym}}(\text{Fe}-\text{H}_m;\text{Fe}-\text{C})$
1408.1	1408.3	1353.3	1353.3	$\nu_{\text{as}}(\text{Fe}-\text{H}_m;\text{Fe}-\text{C})$
1924.1	1926.1	1386.2	1381.3	$\nu_{\text{sym}}(\text{Fe}-\text{H}_m;\text{C}-\text{H}_m)$
1952.3	1390.2	1950.4	1396.4	$\nu(\text{Fe}-\text{H}_f)$

around TS_{dir}. On the other hand, the "via Cp" mechanism involves a transition state TS_{rot} connecting two agostic intermediates and having itself a significant agostic interaction. The C–H_m and Fe–H_m bonds undergo little changes around TS_{rot}, resulting in essentially isotope-independent imaginary frequencies.

The corresponding KIE values are again listed in Table 7. We can grossly attribute this difference to the variation of the chemical environment of the migrating atom, H_m/D_m, on going from the starting dihydrogen complex and the transition state. For the "direct" mechanism, the strong H–H and two weak Fe–H bonds (in-phase and out-of-phase stretching vibrations at 1096.9 and 1810.3 cm⁻¹ for the **2-H**₂ species, respectively) are transformed into two stronger Fe–H bonds, yielding stretching vibrations at 1921.0 (for H_f) and 2027.7 (for H_m) cm⁻¹. However, the loss of the strong H–H bond in the starting material ($\nu_{\text{HH}} = 2762.7$ cm⁻¹) dominates, yielding a normal KIE. The fact that KIE(HH/HD) is greater when the migrating atom is D (1.536, primary isotope effect) and lower when it is H (1.299, secondary isotope effect) is not solely related to the bond stretching vibrations because $\nu(\text{Fe}-\text{X}_m) > \nu(\text{Fe}-\text{X}_f)$ for both X = H and D. Thus, the bending modes play a dominant role. The

Chart 3



same argument rationalizes why KIE(HD/DD) is greater when the migrating atom is H. For the “via Cp” mechanism, on the other hand, the two hydride ligands yield a relatively strong Fe–H_f bond (its stretching vibration at 1952.3 cm⁻¹ in TS_{rot} is similar to that calculated in TS_{dir}) and an agostic Fe···H_m–C moiety. The strength of the C–H bond is reduced from that of a regular C–H bond, but the combination of the C–H_m and Fe···H_m vibrations, which are mixed with an Fe–C component, yields three relatively high-frequency normal modes, as shown in Chart 3. The overall effect is a strengthening of the bonding environment for the H_m atom and a weakening for the H_f atom, leading to an inverse primary KIE (migrating D) and a normal secondary KIE (migrating H).

Conclusions

We have reported here a detailed investigation of the irreversible isomerization process of the dihydrogen intermediate [Cp*Fe(dppe)(η²-H₂)]⁺, obtained by low-temperature protonation of Cp*Fe(dppe)H, to the dihydride complex *trans*-[Cp*Fe(dppe)(H)₂]⁺. The computational tool illustrates the need to include the steric and electronic effects of the Cp and diphosphine substituents to quantitatively reproduce the relative stabilities of nonclassical and classical isomers and the activation barrier of the isomerization process. However, the calculations also indicate that two different pathways are likely candidates for the isomerization mechanism: a “direct” pathway (ΔE[‡] = 21.3 kcal mol⁻¹) and a “via Cp” pathway (ΔE[‡] = 22.2 kcal mol⁻¹), never suggested previously, which involves Cp*H intermediates (cf. 21.6 ± 0.8 kcal mol⁻¹ from the experiment). Sorting between these pathways was possible through the investigation of isotope effects. Generation of the [Cp*Fe(dppe)(η²-HD)]⁺ complex

by protonation of Cp*Fe(dppe)H with CF₃COOD results in a rapid isotope redistribution equilibrium, yielding complexes [Cp*Fe(dppe)(η²-H₂)]⁺ and [Cp*Fe(dppe)(η²-D₂)]⁺; each of the three isotopomers isomerizes to the corresponding classical product with its own rate constant (*k*_{HH}, *k*_{HD}, and *k*_{DD}). The resulting KIEs at 273 K are *k*_{HH}/*k*_{HD} = 1.24 ± 0.01 and *k*_{HD}/*k*_{DD} = 1.58 ± 0.01 (and, consequently, *k*_{HH}/*k*_{DD} = 1.96 ± 0.02). This type of isotope effect analysis, for a rearrangement process of a dihydrogen complex to the isomeric dihydride and encompassing the three H₂, HD, and D₂ isotopomers, is unprecedented to the best of our knowledge. The isotope redistribution EIE and the KIEs were also derived from the calculated normal-mode frequencies of the dihydrogen isotopomers and the transition states. The computed values are in excellent agreement with the experimental values, but only for the “direct” isomerization mechanism, providing strong support in favor of this rearrangement pathway and against the “via Cp” pathway for this system. However, given the closeness of the calculated activation barriers, it seems reasonable to think that subtle modifications in the metallic system could favor a mechanism of the “via Cp” type.

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Supporting Information Available: Details of the kinetic analysis for the Cp*Fe(dppe)H + CF₃COOD reaction, optimized geometries (Cartesian coordinates) for the calculated species, and calculated normal-mode frequencies used for the EIE and KIE analyses. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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