Additions and Corrections

1966, Volume 5

Peter Moore, Fred Basolo, and Ralph G. Pearson: Mechanism of the Acid Hydrolysis of the Iodopentaaquochromium(III) Ion. Evidence for a *trans* Effect of Iodide in a Chromium(III) Complex.

Page 227. The sentence beginning in column 1, line 5, should read: "This is the rate at which one coordinated water molecule exchanges." Delete the following sentence. The next sentence after should read: "The rate constant for release of one water molecule in $Cr(H_2O)_6^{3+}$ is 2.65×10^{-6} sec^{-1} at 25°." In the fourth paragraph, the last two lines should read "rate constant for the exchange of one water molecule of *ca.* 2.9×10^{-4} sec^{-1}."—Ralph G. Pearson

Theodore C. F. Munday and John D. Corbett: An Electromotive Force Study of Lower Oxidation States of Lead, Cadmium, and Tin in Molten $NaAlCl_4$.

Page 1264. The data given in ref 15 for $Pb(AlCl_4)_2$ are incorrect.

Page 1268. The correct powder pattern of $Pb(AlCl_4)_2$ agrees with that given in ref 22 save for evidence of excess $AlCl_3$ in the latter.—J. D. Corbett

1967, Volume 6

E. König and K. Madeja: $^5T_2$–$^1A_1$ Equilibria in Some Iron(II)–Bis(1,10-phenanthroline) Complexes.

Page 51. In Figure 2, the last terms in the energy expressions should read, for the three different groups of Zeeman levels, from the top to the bottom, respectively, $+9\beta^2H^2/4\lambda$, $-5\beta^2H^2/12\lambda$, $-2\beta^2H^2/3\lambda$. In eq 4, the brace in the denominator has been placed incorrectly. The denominator should read $x\{1 + 7 \exp[-(\Delta E/\lambda - 2)x] +$ etc.—E. König

Stephen P. Tanner, Fred Basolo, and Ralph G. Pearson: Anchimeric Assistance in the Reactions of Dichloro(ethylenediamine-N,N′-diacetic acid)platinum(II) and Its Conjugate Base.

Page 1091. Change the second display equation to

$$k_{\text{obsd}} = \frac{k_1[\text{H}^+]^2}{K_\text{A} + [\text{H}^+]^2} + \frac{k_2K_\text{A}}{K_\text{A} + [\text{H}^+]^2}$$

R. Graham Hughes and Clifford S. Garner: The Use of Chromium(IV) Diperoxo Amines in the Synthesis of Chromium(III) Amine Complexes. III. Some Bromo, Phosphato, and Sulfato Ethylenediamine Complexes.

Page 1521. In Table I, the absorption maxima and minima for two of the complexes listed were inadvertently taken from spectra of other substances. The correct values are:

Complex	$[HClO_4]$, F	λ, mμ	a_M, M^{-1} cm^{-1} [a]
Green-blue $Cr(en)(OH_2)_2Br_2^+$ [b–d]	0.3	363 (min)	10.5
		412 (max)	25.4
		440 (min)	24.2
		470 (max)	24.7
		490 (min)	24.3
		543 (max)	31.4
		565 (min)	30.8
		630 (max)	48.0
Magenta $Cr(en)(OH_2)_3Br^{2+}$ [b,c]	1	410 (max)	30.3
		460 (min)	17.8
		475 (plat)	19.6
		537 (max)	36.4
		~600 (sh)	~28

This correction is in addition to those noted earlier in *Inorg. Chem.*, 6, 2273 (1967).—C. S. Garner

W. A. Spofford, III, P. D. Carfagna, and E. L. Amma: The Crystal Structure of *trans*-Bis(phenylethynyl)bis(triethylphosphine)nickel(II), $[P(C_2H_5)_3]_2Ni(C{\equiv}CC_6H_5)_2$.

Page 1553. The *a* lattice dimension should be 14.75 Å as opposed to 15.31 Å. This introduces a systematic error into the calculated structure factors, atomic parameters, bond distances, and rms displacements. However, the changes in the calculated structure factors and atomic parameters are negligible. The changes in interatomic distances and angles are generally small, and only Ni–P and Ni–C distances are changed more than 2 standard deviations. The revised entire Table III is given below. The rms displacements of Table IV do not change significantly (within 1 standard deviation) except for Ni and P, which are revised below.

TABLE III

INTERATOMIC DISTANCES (Å), ANGLES (DEGREES), AND ERRORS[a]

Bonding Interatomic Distances and Angles

Ni–P	2.179 (3)	P–Ni–C_1	87.6 (4)
Ni–C_1	1.847 (11)	Ni–P–C_9	111.5 (5)
C_1–C_2	1.195 (14)	Ni–P–C_{11}	113.9 (5)
C_2–C_3	1.434 (15)	Ni–P–C_{13}	116.5 (5)
C_3–C_4	1.413 (17)	P–C_9–C_{10}	111.2 (11)
C_3–C_5	1.399 (16)	P–C_{11}–C_{12}	112.1 (10)
C_4–C_6	1.400 (17)	P–C_{13}–C_{14}	116.7 (11)
C_5–C_7	1.390 (16)	Ni–C_1–C_2	177.3 (11)
C_6–C_8	1.391 (20)	C_1–C_2–C_3	176.9 (13)
C_7–C_8	1.407 (21)	C_2–C_3–C_4	119.8 (12)
P–C_9	1.854 (14)	C_2–C_3–C_5	118.9 (12)
C_9–C_{10}	1.549 (24)	C_3–C_5–C_7	120.2 (13)
P–C_{11}	1.830 (15)	C_3–C_4–C_6	118.9 (13)
C_{11}–C_{12}	1.525 (21)	C_4–C_3–C_5	121.3 (13)
P–C_{13}	1.835 (12)	C_5–C_7–C_8	117.8 (14)
C_{13}–C_{14}	1.522 (20)	C_4–C_6–C_8	118.6 (14)
		C_6–C_8–C_7	123.2 (13)
		C_9–P–C_{11}	102.4 (7)
		C_9–P–C_{13}	106.0 (7)
		C_{11}–P–C_{13}	105.4 (7)

Nonbonding Intramolecular Distances

Ni–C_2	3.041 (12)	C_1–C_{13}'	3.084 (18)
Ni–C_3	4.472 (10)	C_2–C_9	3.856 (19)
P–C_1	2.797 (12)	C_2–C_{11}	3.774 (18)
P–C_1'	2.915 (12)	C_1–C_3	2.629 (15)
P–C_2	3.700 (12)	C_1–C_4	3.569 (17)
P–C_2'	3.783 (13)	C_1–C_5	3.504 (16)
P–C_{10}	2.814 (20)	C_2–C_{13}'	3.375 (18)
P–C_{12}	2.789 (18)	C_3–C_9	3.859 (19)
P–C_{14}	2.863 (15)	C_3–C_{11}	4.769 (18)
C_1–C_9	3.317 (18)	C_4–C_{11}	4.788 (19)
C_1–C_{11}	3.270 (18)	C_3–C_{13}'	4.117 (19)
C_1–C_{12}	3.907 (21)	C_5–C_{13}'	4.094 (19)

Nonbonding Intermolecular Distances

C_4–C_{14}	3.721 (24)	All others	>3.900

[a] Primes refer to atoms related by the center of symmetry of the Ni site. The dihedral angle between normals to planes defined by Ni–P–C_1 and the benzene ring is 43.6 ± 0.5°.

TABLE IV

RMS ATOMIC DISPLACEMENTS IN THE DIRECTION OF THE PRINCIPAL AXES OF THE THERMAL ELLIPSOIDS (Å)

$\sigma' = \sigma \times 10^3$

Atom	R_1	σ'	R_2	σ'	R_3	σ'
Ni	0.199	3	0.266	4	0.276	3
P	0.226	4	0.271	4	0.288	4