

CONTRIBUTION FROM THE DEPARTMENT OF CHEMISTRY,
NORTHWESTERN UNIVERSITY, EVANSTON, ILLINOIS 60201**Five-Coordinate Nickel(III). The Crystal and Molecular Structure of $\text{NiBr}_3(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2 \cdot 0.5\text{NiBr}_2(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2 \cdot \text{C}_6\text{H}_6$**

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The crystal and molecular structure of the 1:0.5:1 adduct of tribromobis(phenyldimethylphosphine)nickel(III), dibromobis(phenyldimethylphosphine)nickel(II), and benzene has been determined from three-dimensional X-ray data collected by counter methods. The structure has been refined by least-squares techniques to a final R factor on F of 0.054. The material crystallizes in space group $C_2^1\text{-P}\bar{1}$ of the triclinic system, with two molecules of $\text{NiBr}_3(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2$, one molecule of $\text{NiBr}_2(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2$, and two molecules of benzene in a cell of dimensions $a = 9.021$ (5), $b = 17.951$ (10), $c = 11.181$ (6) Å; $\alpha = 98^\circ 52$ (1)', $\beta = 94^\circ 29$ (1)', $\gamma = 90^\circ 44$ (1)'. The observed and calculated densities are 1.67 ± 0.02 and 1.68 g/cm³, respectively. The four- and five-coordinate molecules are well separated. The *trans*-planar complex $\text{NiBr}_2(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2$ is crystallographically required to possess a center of inversion; the Ni-P and Ni-Br bond lengths are 2.251 (3) and 2.297 (2) Å, respectively. The five-coordinate Ni(III) complex $\text{NiBr}_3(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2$ possesses a trigonal-bipyramidal geometry with axial phosphine groups. There is a slight distortion of the molecule in the plane of the Ni and three Br atoms, which could arise from Jahn-Teller effects. The Ni-P bond distances are 2.263 (3) and 2.273 (3) Å, and the Ni-Br bond distances are 2.349 (2), 2.375 (2), and 2.339 (2) Å.

Introduction

The preparation of a nickel(III) complex, $\text{NiBr}_3(\text{P}(\text{C}_6\text{H}_5)_3)_2$, was first reported by Jensen¹ in 1936. From dipole moment data, Jensen and coworkers² inferred that complexes of this type possess a trigonal-bipyramidal rather than a square-pyramidal geometry, as they had earlier suggested.³ Several other complexes containing trivalent nickel with a variety of donor atoms have been suggested,⁴⁻⁶ largely on the basis of observed magnetic moments consistent with low-spin Ni(III). There has been much discussion as to whether the complexes $\text{Ni}(\text{mnt})_2^-$ and $\text{Ni}(\text{diars})_2\text{Cl}_2^+$ (mnt^{2-} and diars are $(\text{CN})_2\text{C}_2\text{S}_2^{2-}$ and $o\text{-C}_6\text{H}_4(\text{As}(\text{CH}_3)_2)_2$, respectively) can properly be considered as complexes of trivalent nickel.^{5,7-10} The recent X-ray structure determination¹¹ of $\text{Ni}(\text{diars})_2\text{Cl}_2^+$ showed that this cation possesses a nearly octahedral geometry; electron spin resonance studies indicate¹¹ that the complex could be more correctly considered as a Ni(II) complex, with the unpaired electron spending a significant amount of time on the As atoms.

Tribromobis(phenyldimethylphosphine)nickel(III), $\text{NiBr}_3(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2$, has been prepared by Alyea

and Meek.¹² The determination of the structure of $\text{NiBr}_3(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2$ was undertaken to substantiate its formulation as a monomeric complex of trivalent nickel. A preliminary account of this work has appeared.¹³ Furthermore, structures of five-coordinate transition metal complexes are being extensively investigated in this laboratory; we have found that complexes of the type MX_3Y_2 frequently exhibit distortions from an idealized geometry. Two Ni(II) complexes, $\text{Ni}(\text{CN})_2(\text{P}(\text{C}_6\text{H}_5)(\text{OC}_2\text{H}_5)_2)_3$ and $\text{Ni}(\text{CN})_2(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_3$, have been shown to be distorted from a trigonal-bipyramidal toward a tetragonal-pyramidal geometry.^{14,15} The determination of the structure of a five-coordinate Ni(III) complex should provide a useful comparison with the Ni(II) structures.

Collection and Reduction of the Intensity Data

Crystals of what proved to be $\text{NiBr}_3(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2 \cdot 0.5\text{NiBr}_2(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2 \cdot \text{C}_6\text{H}_6$ were kindly supplied by Professor D. W. Meek and Dr. E. C. Alyea. The crystals were obtained by recrystallization of pure $\text{NiBr}_3(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2$ from a benzene-hexane solution. However, unit cell and density determinations followed by magnetic and spectral evidence¹³ on the recrystallized sample led us to believe that the unit cell contained one diamagnetic molecule of $\text{NiBr}_2(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2$ to every two paramagnetic molecules of $\text{NiBr}_3(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2$; this formulation was confirmed by the X-ray structural determination, which also indicated the presence of two molecules of benzene in the unit cell. An extremely dark violet, diamond-shaped crystal of approximate dimensions $0.1 \times 0.4 \times 0.4$ mm in the [010], [001], and [101] directions was sealed in a 0.5-mm

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quartz capillary and mounted on a eucentric goniometer head with [101] nearly along the ϕ axis. Weissenberg and precession photographs did not show any special symmetry or systematic absences, suggesting space groups $C_1^1-P\bar{1}$ or $C_1^1-P\bar{1}$. A triclinic unit cell was assigned, and a Delaunay reduction of the cell parameters as determined from the films failed to indicate any hidden symmetry. The cell parameters, at 22° as determined from diffractometer measurements (*vide infra*) using Cu $K\alpha$ radiation (λ 1.54056 Å), are $a = 9.021$ (5), $b = 17.951$ (10), $c = 11.181$ (6) Å; $\alpha = 98^\circ 52$ (1)', $\beta = 94^\circ 29$ (1)', $\gamma = 90^\circ 44$ (1)'. The observed density of 1.67 ± 0.02 g/cm³, obtained by flotation in aqueous ZnCl₂ solutions, agrees well with the calculated value of 1.68 g/cm³ for two molecules of NiBr₃(P(C₆H₅)(CH₃)₂)₂, one molecule of NiBr₂(P(C₆H₅)(CH₃)₂)₂, and two molecules of benzene in the unit cell.

Sixteen reflections were carefully centered using a Picker four-circle automatic diffractometer, and the cell parameters and orientation angles were determined from a least-squares refinement of the setting angles of these reflections using procedures previously described.¹⁶ Cu $K\alpha$ radiation was used for the data collection, and the diffracted beams were filtered through a 1.0-mil Ni foil. The intensities were collected by the θ - 2θ scan technique at a takeoff angle of 1.8° . The peaks were scanned from 0.5° on the low-angle side of the $K\alpha_1$ peak to 0.7° on the high-angle side at a scan rate of 1.0° /min. Stationary-crystal, stationary-counter background counts of 10 sec were taken at each end of the scan. The scintillation counter was placed 29 cm from the crystal and had an aperture of 4.0×4.0 mm.

A unique data set was initially collected to $2\theta = 70^\circ$, and a second shell to $2\theta = 88^\circ$ was subsequently collected. Four reflections were monitored as standards after every 250 reflections were processed, and each of these remained constant within counting statistics throughout the data collection process. The data were corrected for background and standard deviations were assigned to the intensities as previously described,¹⁶ using a value of 0.04 for p . A later statistical analysis (*vide infra*) of the weighting scheme as a function of $|F_o|$ indicated that this value was too large; the strong reflections had been underweighted. The data were also corrected for Lorentz and polarization effects.

In all, 2735 reflections were observed; only 97 of these were less than their standard deviations. The 2514 reflections for which $F_o^2 \geq 3\sigma(F_o^2)$ were used in the solution and refinement of the structure. The crystal was measured by means of a micrometer eyepiece for application of a numerical absorption correction, and the six faces of the forms {010}, {001}, and {101} were identified by optical and X-ray means.

Solution and Refinement of the Structure

The positions of the two Ni, four Br, and three P atoms were determined by direct methods using Sayre's

equation¹⁷ in the form of a modified version of Long's computer program,¹⁸ assuming the centric space group $P\bar{1}$. One cycle of least-squares refinement,¹⁹ assigning isotropic temperature factors to all 9 atoms, resulted in discrepancy factors $R_1 = \sum ||F_o| - |F_c|| / \sum |F_o|$ and $R_2 = (\sum w(|F_o| - |F_c|)^2 / \sum w F_o^2)^{1/2}$ of 0.346 and 0.448, respectively, where the weights w were taken as $4F_o^2 / \sigma^2(F_o^2)$. In this and succeeding refinements the function minimized was $\sum w(|F_o| - |F_c|)^2$. The atomic scattering factors for Ni, Br, P, and C were from the usual tabulation²⁰ and those of Stewart, *et al.*,²¹ were used for H. The anomalous scattering factors of Cromer²² for Ni, Br, and P were used for the real and imaginary scattering contributions to F_c .²³ A difference Fourier map gave the positions of the methyl and phenyl group carbon atoms. These were included in two cycles of least-squares refinement in which the phenyl rings were treated as rigid groups²⁴ of known geometry and dimension (C-C = 1.392 Å) with a single group temperature factor. The resultant discrepancy factors were $R_1 = 0.210$ and $R_2 = 0.304$. A difference Fourier map at this point indicated the presence of a molecule of solvent benzene in the asymmetric unit; when this was included in a cycle of least-squares refinement, the discrepancy factors were lowered to $R_1 = 0.175$ and $R_2 = 0.261$.

The observed intensities were next corrected for the effects of absorption, using a calculated absorption coefficient of 82.75 cm⁻¹; transmission coefficients ranged from 0.40 to 0.09. One cycle of least-squares refinement led to a significant improvement in the overall agreement, with $R_1 = 0.138$ and $R_2 = 0.240$. This calculation was followed by two cycles of refinement in which anisotropic thermal parameters were assigned to all nongroup atoms and individual isotropic thermal parameters were assigned to group carbon atoms (except for the solvent benzene molecule); the resultant discrepancy factors were 0.063 and 0.116. A difference Fourier map was computed using only those reflections for which $\lambda^{-1} \sin \theta < 0.3$ Å⁻¹, and all phenyl and methyl group hydrogen atoms were located. The methyl hydrogen atom positions were idealized to conform to the known geometry (C-H = 1.09 Å, \angle H-C-H = $109^\circ 28'$) in best agreement with the observed positions and were added as fixed contributions to the calculated structure factors, assuming $B = 5.0$ Å². The idealized positions of the phenyl ring H atoms (C-H = 1.0 Å, $B = B$ of C atom to which H atom is bonded) were also added. Two

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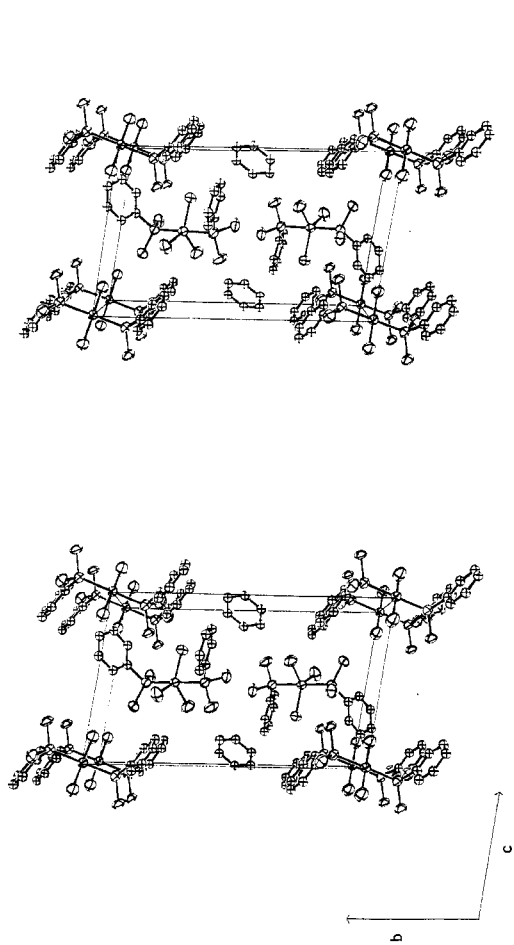


Figure 1.—A stereoscopic view of the unit cell of $\text{NiBr}_3(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2 \cdot 0.5\text{NiBr}_2(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2 \cdot \text{C}_6\text{H}_6$.

TABLE I
FINAL ATOMIC AND GROUP PARAMETERS FOR $\text{NiBr}_3(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2 \cdot 0.5\text{NiBr}_2(\text{P}(\text{C}_6\text{H}_5)(\text{CH}_3)_2)_2 \cdot \text{C}_6\text{H}_6$

Atom ^a	x	y	z	β_{11}	β_{22}	β_{33}	β_{12}	β_{13}	β_{23}	B_1 or B_2 , Å ²	B_3 , Å ²	B_4 , Å ²	B_5 , Å ²	B_6 , Å ²
Ni ₁	0	0	0	0.0106 (4)	0.0028 (1)	0.0070 (2)	0.0004 (1)	0.0002 (2)	0.0002 (1)	0.0070 (2)	0.0004 (1)	0.0002 (2)	0.0001 (1)	0.0001 (1)
Ni ₂	-0.30091 (19)	0.26330 (9)	0.50779 (14)	0.0135 (3)	0.0029 (1)	0.0062 (2)	0.0000 (1)	0.0016 (2)	0.0002 (1)	0.0029 (1)	0.0000 (1)	0.0016 (2)	0.0002 (1)	0.0002 (1)
Br ₁	-0.18289 (16)	0.00772 (7)	-0.15214 (13)	0.0222 (3)	0.0043 (1)	0.0130 (2)	0.0017 (1)	-0.0067 (2)	-0.0003 (1)	0.0130 (2)	0.0017 (1)	-0.0067 (2)	-0.0003 (1)	-0.0003 (1)
Br ₂	-0.39635 (13)	0.26668 (7)	0.30718 (10)	0.0146 (2)	0.0042 (1)	0.0071 (1)	-0.0001 (1)	0.0006 (1)	0.0010 (1)	0.0071 (1)	-0.0001 (1)	0.0006 (1)	0.0010 (1)	0.0010 (1)
Br ₃	-0.46253 (17)	0.32408 (8)	0.64783 (12)	0.0281 (3)	0.0046 (1)	0.0117 (2)	-0.0002 (1)	0.0109 (2)	-0.0010 (1)	0.0117 (2)	-0.0002 (1)	0.0109 (2)	-0.0010 (1)	-0.0010 (1)
Br ₄	-0.08448 (16)	0.21372 (8)	0.59026 (14)	0.0185 (3)	0.0049 (1)	0.0167 (2)	-0.0013 (1)	-0.0067 (2)	0.0031 (1)	0.0049 (1)	-0.0013 (1)	-0.0067 (2)	0.0031 (1)	0.0031 (1)
P ₁	-0.02529 (31)	0.12269 (15)	0.07386 (25)	0.0128 (5)	0.0029 (1)	0.0080 (3)	0.0003 (2)	0.0017 (3)	-0.0002 (1)	0.0128 (5)	0.0003 (2)	0.0017 (3)	-0.0002 (1)	-0.0002 (1)
P ₂	-0.42617 (31)	0.15318 (15)	0.50691 (24)	0.0127 (5)	0.0029 (1)	0.0079 (3)	-0.0001 (2)	0.0022 (3)	0.0006 (1)	0.0127 (5)	-0.0001 (2)	0.0022 (3)	0.0006 (1)	0.0006 (1)
P ₃	-0.17949 (34)	0.37531 (16)	0.51002 (26)	0.0165 (5)	0.0030 (1)	0.0086 (3)	-0.0006 (2)	0.0025 (3)	0.0004 (2)	0.0165 (5)	-0.0006 (2)	0.0025 (3)	0.0004 (2)	0.0004 (2)
C ₁	-0.2001 (12)	0.1671 (6)	0.0366 (10)	0.0140 (20)	0.0038 (5)	0.0125 (14)	0.0019 (8)	0.0018 (13)	0.0002 (7)	0.0140 (20)	0.0019 (8)	0.0018 (13)	-0.0002 (7)	-0.0002 (7)
C ₂	0.0000 (12)	0.1513 (6)	0.2383 (9)	0.0195 (20)	0.0039 (5)	0.0067 (11)	-0.0008 (8)	0.0057 (12)	0.0013 (6)	0.0195 (20)	-0.0008 (8)	0.0057 (12)	-0.0013 (6)	-0.0013 (6)
C ₃	-0.4190 (13)	0.1237 (6)	0.6540 (10)	0.0185 (21)	0.0040 (5)	0.0095 (12)	-0.0010 (8)	0.0042 (13)	0.0015 (6)	0.0185 (21)	-0.0010 (8)	0.0042 (13)	0.0015 (6)	0.0015 (6)
C ₄	-0.6234 (13)	0.1566 (6)	0.4617 (12)	0.0147 (21)	0.0039 (5)	0.0152 (16)	-0.0005 (8)	0.0018 (14)	0.0013 (7)	0.0147 (21)	-0.0005 (8)	0.0018 (14)	0.0013 (7)	0.0013 (7)
C ₅	-0.2998 (14)	0.4477 (6)	0.4656 (11)	0.0225 (24)	0.0025 (4)	0.0153 (16)	0.0008 (8)	0.0073 (16)	0.0007 (7)	0.0225 (24)	0.0008 (8)	0.0073 (16)	0.0007 (7)	0.0007 (7)
C ₆	-0.0961 (15)	0.4149 (7)	0.6602 (11)	0.0246 (26)	0.0053 (6)	0.0096 (14)	-0.0004 (10)	0.0018 (15)	-0.0003 (8)	0.0246 (26)	-0.0004 (10)	0.0018 (15)	-0.0003 (8)	-0.0003 (8)
Group ^b	x_0	y_0	z_0	ϵ	η	B_1 or B_2 , Å ²	B_3 , Å ²	B_4 , Å ²	B_5 , Å ²	B_6 , Å ²	B_7 , Å ²	B_8 , Å ²	B_9 , Å ²	B_{10} , Å ²
Ring 1	0.2268 (6)	0.2115 (3)	-0.0399 (4)	2.456 (4)	2.495 (6)	3.5 (2)	4.8 (3)	4.8 (3)	5.9 (3)	5.8 (3)	5.8 (3)	5.8 (3)	5.0 (3)	5.0 (3)
Ring 2	-0.2955 (6)	0.0206 (3)	0.3242 (5)	-2.381 (5)	1.948 (7)	4.1 (2)	5.6 (3)	5.6 (3)	8.2 (4)	7.6 (4)	7.6 (4)	7.1 (3)	5.4 (3)	5.4 (3)
Ring 3	0.0817 (6)	0.3636 (3)	0.3327 (5)	-2.364 (4)	1.258 (7)	3.9 (2)	4.8 (2)	4.8 (2)	6.2 (3)	6.4 (3)	6.4 (3)	6.8 (3)	5.3 (3)	5.3 (3)
Benzene	-0.3096 (8)	0.4716 (4)	0.0715 (6)	2.827 (6)	0.847 (7)	8.7 (2)								

^a x , y , and z are fractional coordinates. The form of the anisotropic thermal ellipsoid is $\exp[-(\beta_{11}h^2 + \beta_{22}k^2 + \beta_{33}l^2 + 2\beta_{12}hk + 2\beta_{13}hl + 2\beta_{23}kl)]$. The estimated standard deviations of the least significant digits are given in parentheses. ^b x_0 , y_0 , and z_0 are the fractional coordinates of the ring centers. The angles δ , ϵ , and η (in radians) which bring about alignment (except for translation) of an internal coordinate system within the ring with a fixed external coordinate system are those previously described.²⁴ B_i is the isotropic thermal parameter of atom i in a given ring. The rings are numbered so that C₁ is attached to P; C₄ is *para* to C₁.

TABLE II
 DERIVED PARAMETERS FOR GROUP CARBON ATOMS^a

Atom	<i>x</i>	<i>y</i>	<i>z</i>
Ring 1			
C ₁	0.1163 (7)	0.1752 (5)	0.0115 (6)
C ₂	0.2630 (8)	0.1766 (4)	0.0609 (5)
C ₃	0.3734 (8)	0.2129 (5)	0.0095 (7)
C ₄	0.3373 (8)	0.2479 (6)	-0.0914 (7)
C ₅	0.1906 (9)	0.2465 (4)	-0.1407 (5)
C ₆	0.0801 (6)	0.2102 (5)	-0.0893 (6)
Ring 2			
C ₁	-0.3562 (12)	0.0765 (5)	0.4051 (6)
C ₂	-0.4116 (10)	0.0631 (6)	0.2838 (7)
C ₃	-0.3508 (10)	0.0072 (5)	0.2029 (5)
C ₄	-0.2347 (13)	-0.0353 (6)	0.2433 (7)
C ₅	-0.1794 (10)	-0.0219 (5)	0.3646 (8)
C ₆	-0.2401 (8)	0.0340 (4)	0.4454 (5)
Ring 3			
C ₁	-0.0321 (7)	0.3700 (7)	0.4106 (6)
C ₂	-0.0622 (7)	0.3782 (6)	0.2894 (6)
C ₃	0.0516 (9)	0.3717 (4)	0.2114 (5)
C ₄	0.1955 (8)	0.3572 (7)	0.2548 (7)
C ₅	0.2256 (7)	0.3490 (6)	0.3760 (7)
C ₆	0.1118 (8)	0.3555 (4)	0.4539 (5)
Benzene			
C ₁	-0.4183 (11)	0.4783 (7)	0.1541 (9)
C ₂	-0.3231 (12)	0.5392 (5)	0.1492 (9)
C ₃	-0.2144 (12)	0.5325 (5)	0.0667 (10)
C ₄	-0.2009 (10)	0.4649 (7)	-0.0110 (9)
C ₅	-0.2961 (12)	0.4040 (5)	-0.0061 (9)
C ₆	-0.4047 (12)	0.4107 (5)	0.0764 (10)

^a Estimated standard deviations are derived from those of the group parameters. Intra-ring distances are fixed (C-C = 1.392 Å).

cycles of least-squares refinement yielded discrepancy factors of $R_1 = 0.055$ and $R_2 = 0.093$.

As analysis of the weighting scheme applied showed no systematic variation of the function $\Sigma w(|F_o| - |F_c|)^2$ with $\lambda^{-1} \sin \theta$, but indicated that the reflections with a large value of $|F_o|$ were being underweighted. The weights were modified empirically to give a constant value of $\Sigma w(|F_o| - |F_c|)^2$ over all ranges of $|F_o|$, and one cycle of least-squares refinement resulted in final discrepancy factors of $R_1 = 0.054$ and $R_2 = 0.070$ and an error in an observation of unit weight of 0.99. A difference Fourier map showed no peak greater than $1.1 \text{ e}^-/\text{Å}^3$, approximately 30% of the height of a carbon atom.

Table I presents the positional, thermal, and group parameters, along with the corresponding standard deviations in these parameters as estimated from the inverse matrix. Table II gives the positional parameters of the group carbon atoms which may be derived from the data in Table I. The final values of $10|F_o|$ and $10|F_c|$ (in electrons) are presented in Table III. For the 221 reflections for which $F_o^2 < 3\sigma(F_o^2)$, none had $|F_o^2 - F_c^2| > 3\sigma(F_o^2)$. These reflections accordingly are omitted from Table III.

The orientations of the thermal ellipsoids can be seen in the figures. Only the C and Br atoms exhibit significant anisotropy, with the root-mean-square amplitudes of vibration ranging from 0.20 to 0.38 Å for C atoms and from 0.18 to 0.39 Å for Br atoms.

Description of the Structure

The structure consists of well-separated four- and five-coordinate molecules, with a four-coordinate *trans*-NiBr₂(P(C₆H₅)(CH₃)₂)₂ molecule at the origin of the unit cell and two five-coordinate NiBr₃(P(C₆H₅)(CH₃)₂)₂ molecules occupying general positions; the two benzene molecules are located in a rather large hole in the structure. A stereoscopic view of the unit cell is given in Figure 1; it can be seen that layers of five-coordinate molecules alternate with layers of four-coordinate molecules and benzene molecules stacked perpendicular to the *c* axis. The closest approaches of hydrogen atoms between adjacent molecules are 2.29 and 2.40 Å. Interatomic distances and angles and their standard deviations, as computed from the final parameters and the correlation matrix, are presented in Table IV.

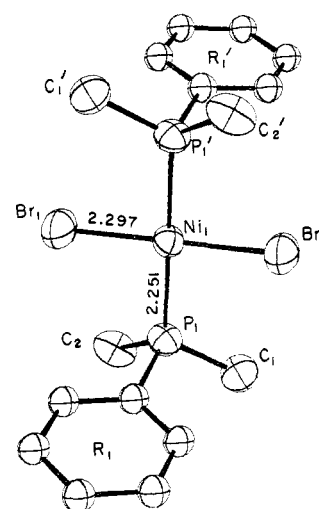


Figure 2.—Perspective drawing of NiBr₂(P(C₆H₅)(CH₃)₂)₂.

The four-coordinate NiBr₂(P(C₆H₅)(CH₃)₂)₂ molecule is crystallographically required to possess a center of symmetry and hence must have a *trans*-planar structure; Figure 2 illustrates this geometry, and also presents selected bond lengths. The Ni₁-P₁ and Ni₁-Br₁ distances of 2.251 (3) and 2.297 (2) Å, respectively, appear to be normal for planar Ni(II) complexes. The corresponding bond lengths in NiBr₂(P(C₂H₅)₂)₂, also with a crystallographic center of symmetry, are 2.26 and 2.30 Å, respectively.²⁵ The P₁-Ni₁-Br₁ angle is 90.45 (8)°, essentially the same as that reported for NiBr₂(P(C₂H₅)₂)₂, 90.5°.

The five-coordinate NiBr₃(P(C₆H₅)(CH₃)₂)₂ molecule possesses a trigonal-bipyramidal geometry, as illustrated in Figure 3; Figure 3 also presents selected bond distances. The three bromine atoms occupy the equatorial plane and the two phosphine groups are axial. The P-Ni-P linkage is nearly linear, with a P₂-Ni₂-P₃ angle of 178.8 (1)°. The Ni atom and three Br atoms are coplanar, with the Ni atom 0.004 (2) Å out of the best-weighted least-squares plane. This plane is essentially perpendicular to the P₂-Ni₂-P₃ vector, as can be seen by examination of the P-Ni-Br bond angles.

TABLE III
OBSERVED AND CALCULATED STRUCTURE AMPLITUDES (X10) (IN ELECTRONS) FOR
NiBr2(P(C6H5)(CH3)2)2 · 0.5NiBr2(P(C6H5)(CH3)2)2 · C6H6

Table with columns for h, k, l, F0, Fc and corresponding observed and calculated values for various reflections. The table is organized into multiple columns and rows, with some cells containing asterisks or other symbols.

TABLE IV
 SELECTED INTERATOMIC DISTANCES AND ANGLES

Atoms	Distance, Å	Atoms	Angle, deg
$\text{NiBr}_2(\text{P}(\text{C}_6\text{H}_5)_2)(\text{CH}_3)_2$			
Ni ₁ -Br ₁	2.297 (2)	Br ₁ -Ni ₁ -P ₁	90.45 (8)
Ni ₁ -P ₁	2.251 (3)	Ni ₁ -P ₁ -C ₁	118.4 (4)
P ₁ -C ₁	1.824 (11)	Ni ₁ -P ₁ -C ₂	117.9 (4)
P ₁ -C ₂	1.826 (11)	Ni ₁ -P ₁ -R ₁ C ₁	107.0 (3)
P ₁ -R ₁ C ₁ ^a	1.822 (9)	C ₁ -P ₁ -C ₂	101.6 (5)
		C ₁ -P ₁ -R ₁ C ₁	105.1 (5)
		C ₂ -P ₁ -R ₁ C ₁	105.7 (4)
$\text{NiBr}_3(\text{P}(\text{C}_6\text{H}_5)_2)(\text{CH}_3)_2$			
Ni ₂ -Br ₂	2.349 (2)	Br ₂ -Ni ₂ -Br ₃	110.63 (9)
Ni ₂ -Br ₃	2.375 (2)	Br ₂ -Ni ₂ -Br ₄	132.71 (9)
Ni ₂ -Br ₄	2.339 (2)	Br ₃ -Ni ₂ -Br ₄	116.66 (9)
Ni ₂ -P ₂	2.263 (3)	P ₂ -Ni ₂ -P ₃	178.8 (1)
Ni ₂ -P ₃	2.273 (3)	P ₂ -Ni ₂ -Br ₂	90.3 (1)
P ₂ -C ₃	1.800 (11)	P ₂ -Ni ₂ -Br ₃	89.3 (1)
P ₂ -C ₄	1.816 (12)	P ₂ -Ni ₂ -Br ₄	90.6 (1)
P ₂ -R ₂ C ₁	1.800 (9)	P ₃ -Ni ₂ -Br ₂	89.7 (1)
P ₃ -C ₅	1.807 (12)	P ₃ -Ni ₂ -Br ₃	89.6 (1)
P ₃ -C ₆	1.820 (12)	P ₃ -Ni ₂ -Br ₄	90.3 (1)
P ₃ -R ₃ C ₁	1.793 (7)	Ni ₂ -P ₂ -C ₃	113.1 (4)
		Ni ₂ -P ₂ -C ₄	113.7 (4)
		Ni ₂ -P ₂ -R ₂ C ₁	112.6 (3)
		Ni ₂ -P ₃ -C ₅	113.1 (4)
		Ni ₂ -P ₃ -C ₆	112.7 (6)
		Ni ₂ -P ₃ -R ₃ C ₁	113.8 (4)
		C ₃ -P ₂ -C ₄	104.5 (6)
		C ₃ -P ₂ -R ₂ C ₁	106.4 (5)
		C ₄ -P ₂ -R ₂ C ₁	105.9 (5)
		C ₅ -P ₃ -C ₆	104.8 (6)
		C ₅ -P ₃ -R ₃ C ₁	105.5 (6)
		C ₆ -P ₃ -R ₃ C ₁	106.2 (5)

^a R_iC₁ refers to carbon atom 1 on phenyl ring *i*.

There is, however, a distortion of the molecule in the equatorial plane; the Ni₂-Br₃ bond length of 2.375 (2) Å is significantly longer than the other two Ni-Br distances of 2.339 (2) and 2.349 (2) Å, and the angle opposite this long bond is expanded to 132.71 (9)° from the expected 120°. This distortion is similar to that found for the two Ni(CN)₂L₃ complexes^{14,15} where L = P(C₆H₅)(OC₂H₅)₂ and P(C₆H₅)(CH₃)₂ and for one of the Ni(CN)₅³⁻ ions in [Cr(NH₂CH₂CH₂NH₂)₃][Ni(CN)₅]·1.5H₂O,²⁶ except that in these previous examples the two axial ligands were also bent away from the long bond. The distortion in the present compound could be ascribed to Jahn-Teller effects as the complex contains Ni(III) in a low-spin d⁷ electronic configuration. Yet Jahn-Teller effects cannot be invoked to explain the similar distortions found in the Ni(II) complexes.

The Ni-P bond lengths of 2.263 (3) and 2.273 (3) Å cannot be considered significantly longer than that of 2.251 (3) Å found for the four-coordinate complex. However, the Ni-P bond lengths are approximately 0.04 Å shorter than the Ni-P distance found in the high-spin tetrahedral Ni(II) complex NiBr₂(P(C₆H₅)₃)₂, 2.33 Å.²⁷ The Ni-P bond lengths appear to be somewhat longer than those observed in many diamagnetic

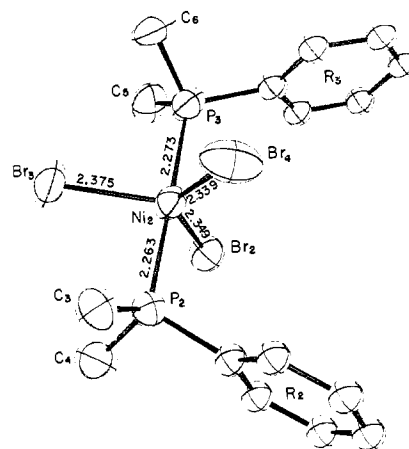


Figure 3.—Perspective drawing of NiBr₃(P(C₆H₅)₂)(CH₃)₂.

five-coordinate Ni(II) complexes, for example, 2.18 (1) Å in Ni(HP(C₆H₅)₂)₃I₂,²⁸ 2.206 (8) Å in [Ni(TAP)CN]ClO₄ (TAP = P(CH₂CH₂CH₂As(CH₃)₂)₃),²⁹ 2.19 Å (equatorial P) and 2.15 Å (axial P) in the trigonal-bipyramidal Ni(P(OCH)₃(CH₂)₃)₅²⁺ ion,³⁰ and 2.223 (3) Å in Ni(CN)₂(P(C₆H₅)₂)(CH₃)₂.¹⁵ The Ni-Br distances of 2.349 (2), 2.375 (2), and 2.339 (2) Å are significantly longer than that of 2.297 (2) Å observed in the four-coordinate complex and are more in keeping with the Ni-Br distance of 2.34 Å found in NiBr₂(P(C₆H₅)₃)₂.²⁷ Apparently, the increase in radius resulting from the presence of an unpaired electron in Ni(III) more than compensates for the contraction that would be expected to result from increased nuclear charge. The apparent increase in the nickel covalent radius on going from Ni(II) to Ni(III) complexes is similar to that observed on going from low-spin square-planar Ni(II) to high-spin tetrahedral Ni(II), for which the observed radius expansion is about 0.05 Å.²⁷

The C-P bond lengths and C-P-C bond angles in both the four- and five-coordinate complexes appear to be normal. If the four P-CH₃ distances in the five-coordinate molecule are assumed to be equivalent, then the standard deviation of a given distance from the mean of 1.811 Å is 0.009 Å, approximately the same as the computed standard deviations. Thus the reported errors are probably reasonable.

This compound is part of a growing series of trigonal-bipyramidal MX₃Y₂ complexes, where M is a transition metal. Insofar as we are aware, such complexes invariably have the X groups in the equatorial plane, despite widely varying steric and electronic properties of the X and Y ligands.

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