Additions and Corrections

1969, Volume 8

D. P. Rillema, W. J. Reagan, and C. H. Brubaker, Jr.: Complexes of Tungsten(V) which Contain Alkoxide as a Ligand.

Page 589. The sentence beginning on column 1, line 3, should read: All reactions were carried out under dry N_2 because the compounds (except for $W_2Cl_4(OC_2H_5)_6$) are air and moisture sensitive.

Dwight A. Fine: On the Spontaneous Reduction of Hexachloroiridate(IV) in Aqueous Solution.

Page 1015. In ref 7 the year should be 1961; in ref 8 the citation should be *Can. J. Chem.*, **44,** 515 (1966).

Werner Giggenbach and C. H. Brubaker, Jr.: The Preparation and Optical and Electron Spin Resonance Spectra of Some Chloro(methanol) Complexes of Titanium(III).

Page 1134. The values in column 1, line 20, should be $g_{\|} = 1.88$ and $g_{\perp} = 1.92$.

Nick Serpone and Robert C. Fay: Dipole Moments and Electric Field Effects on the Proton Chemical Shifts of Halo(acetylacetonato) Complexes of Group IV Metals.

Page 2384. The numerator of the last term in eq 9 should read

$$3z_i(x_i - y_i - \sqrt{2}z_i)$$

1970, Volume 9

N. Dennis Chasteen and R. Linn Belford: Triplet State of a Binuclear Copper *dl*-Tartrate and Electron Paramagnetic Resonance of the Copper Tartrates.

Page 171. In eq 2, B^2 should read $B^2/2$.

Page 173. In Table II, the column 10^4D_{pseudo} should read 10^3D_{pseudo}, 280 should read 360, and 2.10 in footnote *h* should read 2.19.

As a practical matter, however, the magnitudes of the D_{pseudo} values must be much closer to those given in the original paper. Clearly, the dipolar interactions are the principal contributors to D for all the cases cited in Table II, as demonstrated by the reasonable agreement between R_{calcd} and R_{exptl}. Moreover, Bleaney and Bowers [B. Bleaney and K. D. Bowers, *Proc. Roy. Soc. (London)*, **A214,** 451 (1952)] found that D_{pseudo} calculated from eq 6 with the ground-state J value is an order of magnitude too large for the copper acetate dimer. The generally reasonable idea prevails [A. E. Hansen and C. J. Ballhausen, *Trans. Faraday Soc.*, **61,** 631 (1965); G. F. Kokoszka and H. C. Allen, *J. Chem. Phys.*, **46,** 3013 (1967)] that the excited-state J values which ought to be used in the correct version of eq 6 are much smaller than the ground-state value. For dimers of the copper acetate geometrical type (of which the copper adenine and tartrate complexes are examples) in the absence of definitive information about exchange interaction in the excited states, one ought to reduce J by an order of magnitude before applying eq 6. Finally, for the vanadyl tartrates the situation is probably similar; furthermore, recent work [P. G. James and G. R. Luckhurst, *Mol. Phys.*, **18,** 141 (1970)] on a vanadyl tartrate dimer reveals a very small ground-state J value, about 0.1 cm^{-1}.

NOTE ADDED IN PROOF.—Low-temperature magnetic susceptibility measurements on $Na_2Cu(d,l\text{-}C_4O_6H_2)\cdot 5H_2O$ powder (W. E. Hatfield, private communication) now show $|J| \approx 14$ cm^{-1} and thus fully justify our neglect of D_{pseudo}.

E. I. Stiefel, L. E. Bennett, Z. Dori, T. H. Crawford, C. Simo, and Harry B. Gray: Synthesis and Characterization of Tris(maleonitriledithiolato)metalates.

Page 284. In Table III, the X-ray powder pattern for $[(C_6H_5)_4As]_2[Fe(mnt)_3]$ was inadvertently left out. It is as follows: 11.20 w, 10.28 w, 7.70 vvs, 6.71 w, 6.14 m, 6.02 m, 5.72 w, 5.43 m, 5.14 w, 5.05 w, 4.94 m, 4.52 w, 4.39 w, 4.15 m, 4.07 w, 3.85 w, 3.89 w, 3.60 w, 3.50 w.

David L. Cullen and E. C. Lingafelter: The Crystal Structures of Bis[2,2′-iminobis(acetamidoxime)]nickel(II) Chloride Dihydrate and Bis[2,2′-iminobis(acetamidoxime)]copper(II) Chloride.

Page 1865. The authors should have been: David L. Cullen, E. C. Lingafelter, and Lowell P. Eddy.

The sentence beginning at the end of the fourth line up from the bottom of the abstract should read: The other ligand molecule is bidentate with one oxime nitrogen atom filling the fourth position in the plane and the amine nitrogen atom occupying the apical position.