

The Crystal and Molecular Structure of Tetra- μ -chloro-tetra- μ -ethyloxycarbene-di- μ_3 -hydroxo dodecacarbonyl-di-chloro-hexaruthenium(II)

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The crystal and molecular structure of tetra- μ -chloro-tetra- μ -ethyloxycarbene-di- μ_3 -hydroxo-dodecacarbonyl-di-chloro-hexaruthenium(II)-benzene, $[Ru_3(Cl_3(COC_2H_5)_2(OH)(CO)_6]_2 \cdot C_6H_6$, have been determined. The complex crystallizes as monoclinic crystals in space group $P\bar{2}_1/c$ with unit cell dimensions $a = 10.71(2)$ Å, $b = 10.41(2)$ Å, $c = 21.92(3)$ Å, $\beta = 113^\circ 15' (10')$ and $Z = 2$. A three-dimensional X-ray analysis, based on 3101 observed reflections, converged to a R of 0.076. The crystal structure of the complex consists of centrosymmetric hexanuclear molecules. The three ruthenium atoms of the asymmetric unit are octahedrally coordinated. The octahedra share a common vertex occupied by a hydroxo group which bridges three ruthenium atoms. Significant differences in the Ru–Cl bond lengths are accounted by the “trans influence” of the carbene nature of the C_2H_5CO groups.

Introduction

η -Allyl ruthenium complexes of formula $RuX(C_3H_5)(CO)_3$ ($X = Cl, Br$) have been used as catalysts in several reactions involving unsaturated substrates [1, 2]. The reactivity with different olefinic and acetylenic ligands was also widely investigated [3].

In the course of these studies some new ruthenium complexes, obtained in crystalline and amorphous forms, by reaction of $RuCl(C_3H_5)(CO)_3$ (I) with ethylene were isolated [2].

Since the analytical and spectroscopic data (IR and 1H n.m.r.) did not allow to establish exactly the nature of the products and to understand some IR and analytical differences observed for the crystalline and amorphous product, recourse was had to X-ray crystallography.

Experimental

Preparation

By reaction of I (1 g, 3.8 mmol) with ethylene under pressure (50–60 atm) at 70–80 °C for 10 h in benzene solution and by standing the reaction mixture in the autoclave at room temperature for several days, a crystalline compound (II), initially formulated as $[RuCl(COC_2H_5)(CO)_2]_n$ with solvate benzene, was obtained [Found: C, 24.57; H, 2.17; Cl, 14.3; Ru, 40.2; M(benzene, Mechrolab Osmometer), 1600. Calc. for $C_{30}H_{28}Cl_6O_{18}Ru_6$ (on the basis of the formula determined by X-ray structural data): C, 24.09; H, 1.88; Cl, 14.22; Ru, 40.55%; M, 1495.7]. The complex could also be isolated as an amorphous product (III) either by adding n-pentane to the benzene solution or by evaporating to dryness [Found: C, 23.82; H, 2.18; Cl, 14.6; Ru, 39.2. $C_5H_5ClO_3Ru$ requires C, 24.05; H, 2.02; Cl, 14.21; Ru, 40.4%].

IR and 1H n.m.r. Spectra

The IR spectra (Perkin–Elmer 225 spectrometer, in KBr) of the crystalline and amorphous products showed three bands in the carbonyl stretching region at 2130 vw, 2060vs, 2005vs cm^{-1} and the bands assignable to a bridging acyl group (1543s cm^{-1}) [4] and to the terminal (329w cm^{-1}) and bridging chlorine atoms (290w and 260w cm^{-1}) [5]. The IR spectrum of the crystalline product differed from that of the amorphous one for a weak but sharp absorption at 3280 cm^{-1} [$\nu(\text{OH})$], by a more complex pattern of bands in the Ru–C–O deformation and Ru–C stretching region (650 – 400 cm^{-1}) and by the bands due to solvate benzene.

The 1H n.m.r. spectrum of the crystalline product (Varian T 60 spectrometer, in benzene solution) showed a quartet and a triplet at 7.32 and 9.15 τ

TABLE I. Positional and Thermal Parameters for Non-hydrogen Atoms in $[\text{Ru}_3\text{Cl}_3(\text{COC}_2\text{H}_5)_2(\text{OH})(\text{CO})_6]_2 \cdot \text{C}_6\text{H}_6$ ^{a,b}.

	x	y	z	β_{11}	β_{22}	β_{33}	β_{12}	β_{13}	β_{23}
Ru(1)	-0.08384(10)	-0.20176(11)	0.40900(5)	81.7(10)	23.3(15)	20.6(2)	-5.1(8)	17.6(4)	-3.4(4)
Ru(2)	0.06520(10)	0.02238(12)	0.34289(5)	75.0(9)	40.2(15)	19.1(2)	0.8(8)	17.1(4)	0.0(4)
Ru(3)	-0.19243(10)	0.15187(12)	0.37336(5)	70.4(9)	36.0(15)	20.4(2)	1.9(8)	14.1(4)	-1.1(4)
Cl(1)	0.13394(34)	-0.20980(38)	0.50463(16)	88.4(32)	44.7(51)	20.5(8)	9.8(28)	16.0(13)	-1.6(14)
Cl(2)	0.03363(37)	-0.21990(39)	0.33623(18)	108.0(36)	33.5(50)	25.1(9)	3.4(30)	28.4(15)	-5.8(15)
Cl(3)	-0.19267(36)	-0.15407(42)	0.48031(17)	99.0(35)	58.1(53)	23.3(8)	-4.0(32)	25.7(14)	-1.6(15)
C(1)	-0.1038(16)	-0.3769(18)	0.4170(8)	107(16)	75(26)	24(4)	-38(15)	16(6)	1(7)
C(2)	-0.2436(15)	-0.2023(16)	0.3344(8)	93(14)	47(21)	26(4)	-23(12)	27(6)	-10(6)
C(3)	0.1465(15)	0.0427(18)	0.2829(8)	90(15)	92(25)	23(4)	9(14)	14(6)	-3(7)
C(4)	0.2430(14)	0.0003(17)	0.4077(7)	84(14)	83(23)	24(4)	-7(13)	22(6)	-4(7)
C(5)	-0.3455(17)	0.0640(17)	0.3741(8)	122(19)	41(22)	30(4)	11(15)	19(7)	-6(7)
C(6)	-0.3053(15)	0.2876(17)	0.3438(7)	99(15)	93(25)	20(3)	1(14)	21(6)	6(6)
C(7)	-0.2311(14)	0.0923(16)	0.2809(6)	92(13)	56(20)	14(3)	-8(12)	9(5)	1(5)
C(8)	0.0700(11)	0.2065(15)	0.3584(6)	52(11)	67(22)	23(3)	19(10)	14(5)	6(6)
C(9)	-0.3624(15)	0.1058(19)	0.2227(8)	77(14)	108(27)	22(4)	0(14)	-16(6)	7(7)
C(10)	-0.3629(18)	0.2253(19)	0.1834(9)	136(20)	65(25)	28(5)	17(17)	20(8)	15(8)
C(11)	0.1766(15)	0.3009(18)	0.3589(8)	83(14)	101(26)	23(4)	-15(14)	17(6)	-8(7)
C(12)	0.1649(21)	0.4437(20)	0.3825(11)	166(26)	24(24)	51(8)	15(18)	31(12)	-4(9)
C(13)	0.4235(26)	0.0358(32)	0.0363(13)	171(32)	291(55)	36(7)	73(34)	14(12)	1(15)
C(14)	0.4899(34)	0.1302(29)	0.0120(13)	286(48)	165(45)	39(8)	-18(36)	14(16)	-19(14)
C(15)	0.4390(32)	-0.0909(34)	0.0246(14)	254(46)	243(53)	40(9)	-73(39)	4(16)	13(16)
O(1)	-0.1189(14)	-0.4803(14)	0.4243(7)	157(17)	86(19)	38(4)	-32(13)	29(7)	7(7)
O(2)	-0.3461(12)	-0.2025(13)	0.2887(6)	102(13)	127(20)	27(3)	-17(12)	-1(5)	-8(6)
O(3)	0.1997(12)	0.0459(14)	0.2474(6)	135(15)	148(20)	30(3)	18(13)	46(6)	9(6)
O(4)	0.3509(11)	-0.0135(13)	0.4438(6)	94(12)	130(20)	30(3)	25(12)	13(5)	8(6)
O(5)	-0.4450(11)	0.0221(13)	0.3705(7)	87(12)	107(19)	48(5)	-8(11)	35(6)	10(7)
O(6)	-0.3744(13)	0.3776(12)	0.3265(6)	151(16)	57(17)	32(3)	48(12)	17(6)	7(5)
O(7)	-0.1351(9)	0.0436(10)	0.2705(4)	82(9)	54(14)	14(2)	4(8)	8(3)	4(4)
O(8)	-0.0279(9)	0.2556(10)	0.3698(5)	78(9)	12(12)	28(3)	-8(8)	22(4)	1(4)
O(9)	-0.0431(9)	0.0020(10)	0.4069(4)	84(9)	44(13)	17(2)	7(8)	21(4)	4(4)

^a The numbers in parentheses in this and other tables are the estimated deviations in the least significant digits. The anisotropic thermal parameters are of the form $\exp[-10^{-4}(\mathbf{k}^2\beta_{11} + \mathbf{k}^2\beta_{22} + 2\mathbf{k}\mathbf{k}\beta_{12} + 2\mathbf{h}\mathbf{k}\beta_{13} + 2\mathbf{k}\mathbf{l}\beta_{23})]$.

from TMS, in the proper intensity ratio for an ethyl group, and a resonance at 10.9τ assignable to a triple bridging OH group [6]. This last resonance was absent in the spectrum of the amorphous product.

X-ray Crystallography

Unit-cell dimensions and space group for the crystals of the complex II were obtained from precession photographs taken with Mo K α radiation. The crystals belonged to the monoclinic system with $a = 10.71(2)$ Å, $b = 10.41(2)$ Å, $c = 21.92(3)$ Å, $\beta = 113^\circ 15' (10')$. Systematic absences ($h0l$ absent for $l = 2n + 1$ and $0k0$ absent for $k = 2n + 1$) indicated space group $P\bar{2}_1/c$. For $Z = 2$ the calculated density was 2.45 g/cm 3 .

The intensity data were recorded with Ni-filtered Cu K α radiation ($\lambda = 1.5418$ Å) by means of Weissenberg photographs with multiple film technique and integration process. A crystal elongated in the b direction was reduced to nearly cylindrical shape with a radius of 0.013 cm ($\mu R = 2.70$ for Cu K α radiation). Nine layers with b as rotation axis ($k = 0$ through 8) were taken.

A total of 4138 independent reflections (3101 observed) were collected. The intensities measured with a Nonius microdensitometer have been corrected for Lorentz and polarization effects, for the incipient but incomplete $\alpha_1 - \alpha_2$ spot doubling and for the absorption factor using the three-constant formula proposed by Palm [7] for cylindrical crystals.

Solution and Refinement of the Structure

The ruthenium atoms were located from three-dimensional Patterson function. All chlorine, carbon, and oxygen atoms were located by successive Fourier syntheses. Various cycles of block-matrix least squares refinement lowered the value of $R = \Sigma ||F_o| - |F_c||/\Sigma|F_o|$ to 0.092; unit weights were assumed for all observed reflections, isotropic temperature factors were attributed to carbon and oxygen atoms, while anisotropic temperature factors were attributed to ruthenium and chlorine atoms.

The anisotropic refinement was extended to all the atoms and the weighting scheme $\sqrt{w} = 1/(a + |F_o| + b|F_o|^2)$ proposed by Cruickshank [8] was introduced with $a = 40.0$ and $b = 0.005$.

Seventeen strong reflections, probably affected by extinction, were excluded from the last refinement cycles. The final R value was 0.076.

Scattering factors used in the structure factor calculations were taken from International Tables for X-ray Crystallography [9]. All the programs used are incorporated in the X-ray 70 system [10]. Positional and thermal parameters for the atoms in the asymmetric unit are given in Table I. Observed and calculated structure factors are compared in Table II.

Results and Discussion

The resolution of the structure of the crystalline product II allows to formulate this as tetra- μ -chlorotetra- μ -ethoxy carbene-di- μ_3 -hydroxo-dodecacarbonyl-dichloro-hexaruthenium(II)-benzene, $[\text{Ru}_3\text{Cl}_3(\text{COC}_2\text{H}_5)_2(\text{OH})(\text{CO})_6]_2 \cdot \text{C}_6\text{H}_6$.

The crystal structure of II is composed of two discrete hexanuclear molecules located at the inversion centers at $(0, 0, 1/2)$ and $(0, 1/2, 0)$ in the monoclinic unit cell. At another pair of centers at $(1/2, 0, 0)$ and $(1/2, 1/2, 1/2)$ are two molecules of benzene of crystallization.

This structure is closely analogous to that of a toluene-containing form of di- μ -aquo-di- μ_3 -hydroxodeca(1,1,1-trifluoropentane-2,4-dionato)hexanickel(II) [11] which also crystallizes in the space group $P\bar{2}_1/c$ with two hexanuclear molecules at two inversion centers and two disordered molecules of toluene at another pair of centers. The resemblance goes further as both the complex molecules are composed of two trinuclear units, related by a center of symmetry and joined by two bridging water molecules in the nickel complex and by two bridging chlorine atoms in II. In both cases the three metal atoms in the asymmetric unit are each octahedrally coordinated and the three octahedra share a common vertex at the position of the hydroxo group.

Fig. 1 is an ORTEP [12] drawing of the molecule. Bond lengths and angles are given in Tables III and IV, together with their standard deviations.

As shown the molecular structure of II has a number of unusual features that seldom are present in a same molecule.

The first one is the presence of two bridging $\text{C}_2\text{H}_5\text{CO}$ groups between the Ru(2) and Ru(3) atoms; a similar situation was up to now found only in the crystal structure of bis-(μ -phenyloxycarbene-tricarbonyliron) [13]. In that case, however, the C atoms of the two bridging groups are both linked to the same iron atom, whereas the oxygen atoms are linked to the other iron atom. The C–O bond lengths of $\text{C}_2\text{H}_5\text{CO}$ groups [$\text{C}(7)\text{–O}(7)$: 1.246 Å; $\text{C}(8)\text{–O}(8)$: 1.276 Å] longer than that of a keto group (1.23 Å) and analogous to those found in the "carbene compound" $\text{Fe}_2(\text{CO})_6[\text{C}(\text{O})\text{Ph}]_2$ [13] and the Ru–C distances [Ru(2)–C(8): 1.944 Å; Ru(3)–C(7): 1.996 Å], longer than the ruthenium carbonyl lengths (1.89 Å) but shorter than those of a Ru–C (σ) bond (2.07 Å) [14], indicate for the bridging $\text{C}_2\text{H}_5\text{CO}$ groups an arrangement of carbene type [13]. The carbene character of $\text{C}_2\text{H}_5\text{C}(\text{O})$ groups, which display a high σ -donor and a low π -acceptor power, is also evidenced by the lengthening by *trans* effect [15, 16] of the Ru(2)–Cl(2) and Ru(3)–Cl(1)ⁱ bonds.

The C_2H_5 groups in II arise from the insertion of a C_2H_4 molecule into a ruthenium–hydrogen bond formed by β -elimination of pentadienes from the

TABLE II. Observed and Calculated Structure Factors ($\times 10^2$). The reflections marked with a star were excluded from the last refinement cycles.

0,0,L	12	-417	-1469	-12	1082	-1030	2	1516	-1721	-23	496	+421	8,1,L	-1	325	-278	-11	368	27*				
	-2	205	640	-14	151	-116	3	1777	-1930	-24	537	-520		4	215	-206	-12	613	-658				
4	2394*-3681	20	-742	-148	-16	448	-810	4	511	-481	-27	268	-223	0	784	760	-5	537	620				
6	1119	1392	-2	567	-975	-18	822	-741		808	-992	-28	151	-251	1	657	613	-9	580	-752			
8	1737	1878	-4	876	1453	-22	493	-461	4	567	-189		2	288	-222	10	205	-244	-15	277	-240		
12	1081	1066	-6	1794	1180	-24	269	-284	2	748	-775	5,1,I	6	287	306	-16	342	-209					
14	667	614	-8	1794	1180	-26	221	-448	6	1201	1265		7	671	643	-17	243	237					
16	883	-480	-10	500	332	-12	221	-448	9	811	816	0	822	-772	8	336	-305	-19	176	-142			
24	506	440	-14	946	-953		11,0,L		11	608	523	2	966	-975	12	519	553	-21	157	-105			
	-16	445	1529	0	422	391	3	206	-469	3	274	-143	11	153	162		22	399	-351				
1,0,L	-22	765	-1677	2	330	-317	13	714	-638	4	857	-851	14	246	299	13,1,I	-22	173	-153				
	-22	804	-323	4	158	-161	14	360	-331	5	1077	1112	-1	920	885		-25	405	-380				
2	3878*-4376	-22	652	607	-2	705	679	16	953	-851	5	822	-826	2	788	710	-3	115	162				
1	1704*-2208	-25	666	-465	-4	851	-791	17	777	667	7	592	-526	-3	1003	-917	-4	136	166				
4	2930*-3652	-23	119	-137	-6	868	-805	18	819	742	8	990	-1012	-4	796	-778	-5	120	-120				
6	657	680	-8	912	960	19	373	-357	9	971	-1010	-5	431	-371	-6	208	-238	3,2,L					
8	1231	1338	4,0,L	-10	615	568	21	449	-469	11	625	564	-6	1013	-963	-8	297	-327					
10	1276	-1317	-12	743	-691	23	321	360	13	379	306	-7	693	632	-13	357	-369	0	220	186			
12	903	895	3	777	-552	-14	163	-124	28	858	32	14	298	-292	-8	310	270	-14	235	-224			
14	565	512	29	290	-242	-16	310	282	1	620	673	15	599	-569	-9	450	-394	-17	349	360			
16	1332	-1265	5	901	757	-18	453	-424	-12	123	1367	19	161	-79	-10	916	858	-10	123	-103			
18	706	-624	8	581	999	-22	257	259	8	429	461	20	228	-236	-12	345	278		8	258	-27		
20	611	697	16	774	-761	-26	135	-145	-8	1066	1066	-1	655	612	14	248	248	5,2,L					
22	397	347	22	722	-1127	10	555	520	13	918	-913	9	503	399	9,1,L	7	933	970	15	457	-335		
24	1734*-2304	12	722	-1127	12	1203	-10	555	520	-7	235	163	2	194	-205	4	936	959		6	206	165	
	-16	2045	16	693	556		12,0,L		-8	582	523	4	595	-559	-20	284	236	2	942	-1018			
6	513	562	19	217	-237	2	381	-489	-10	1060	1172	5	775	-752	-21	524	-472	3	292	289			
-10	1277	1338	-12	249	-320	-2	637	658	-17	1047	-429	-6	1861	1550	-25	386	362	4	402	327			
12	2362	2396	-18	147	1896	-4	363	-350	-11	1485	-1617	-7	433	377	-27	174	-193	5	818	-221			
14	355	315	-9	273	232	-6	861	-836	-12	1089	-1111	-8	1051	1087	-6	306	-190	14	930	915			
-16	671	-641	-12	207	1203	-10	555	520	-13	918	-913	9	503	399	9,1,L	7	933	970	15	457	-335		
-18	514	694	-12	107	-1024	-12	142	100	-18	461	413	-10	815	-769	8	523	-497	16	259	248			
-22	342	-308	-12	1089	-1098	-18	100	-64	15	1135	-1131	-12	697	-675	0	590	551	9	1225	-1126			
-24	691	647	-16	1917	1923	-16	319	-271	-16	1111	1171	-13	831	789	2	710	-709	12	1489	-1569			
-26	307	330	-20	1750	-1153	-20	331	303	-19	757	-652	-15	287	-253	7	861	846	15	558	602			
	-22	585	-506	-22	585	-506	23	526	-460	-22	748	-694	8	217	-188	16	816	-798	-2	425	329		
2,0,L	-24	585	-506	13,0,L	-22	587	322	13,0,L		21	267	293	-2	19	709	652	9	202	-207	17	812	-779	
	-22	586	-408	22	586	-408	20	549	512	10	247	241	19	435	-376	-5	292	-291		6	206	-165	
2	500	-471	7,0,L	-2	146	208	-24	868	378	-21	681	656	11	191	-216	18	886	-846	-6	651	835		
4	2922*-3549	-6	204	-202	25	877	-380	23	516	-456	20	344	484	29	191	-194	-7	653	665				
6	489	436	8	1955	-1922	-8	265	-259	-26	415	394	-28	504	-471	-1	364	240	21	212	-206			
8	1083	1079	10	173	173	-18	10	149	-139	26	324	-321	-1	655	655	22	562	-560	-10	824	359		
10	704	-644	8	143	1124	-12	411	393	27	343	391	-6	316	-271	23	223	-228	-11	206	-200			
12	625	627	8	358	388	-14	328	319	28	12	179	-2	250	250	24	327	-218	13	616	637			
14	1713	-1522	12	592	-539	-16	549	-589	0	1087	-1118	-9	635	587		-14	670	-600		6	206	-165	
16	471	-445	12	248	-231	-18	253	-276	1	1068	-1061	-10	653	587	1,2,L	-15	774	-749		6	206	-165	
20	1053	873	18	472	433		2,0,L		2	1282	-1320	6,1,L	-10	653	538	1,2,L	17	574	-582		6	206	-165
24	345	-435	16	211	233	0	1,0,L		3	867	-351	0	514	471	-12	1062	1019	0	787	-896			
	-2	1069	1074	-2	275	259	8	1929	-2056	2	598	595	-13	410	366	1	725	-750	-19	331	-306		
4	1766	1893	-8	173	1681	2	153	137	5	1222	1309	3	761	-715	-15	608	-521	3	1182	1212			
6	622	595	-6	635	-556	3	1862	-2670	6	1682	1922	8	353	-320	-16	789	-716	6	595	604			
8	3972*-3799	-10	173	1835	■	982	-1207	5	196	-930	5	272	241	-19	452	-487	5	1567	-1665				
-10	267	250	-12	294	-406	5	565	575	10	882	-922	6	841	-872	-20	385	338	4	328	-275			
-12	2487	2672	-14	174	1115	-15	1565	-1839	11	219	238	8	512	-447	-21	495	-863	7	1017	-1249			
-16	571	-871	-12	1245	1173	7	1212	-1281	12	479	-433	-2	237	218	8	445	-405	-27	177	-240			
-18	894	840	-18	196	1146	8	1180	1304	18	375	375	-11	793	789	-25	459	491	9	1248	-1228			
-22	249	-322	20	315	-299	9	284	-186	12	248	-179	-7	114	-128	10	925	-948	10	1074	1012			
-24	699	666	-24	161	-126	10	1599	1751	19	957	-910	-10	1074	-1102	11	1074	-1102	2	1014	1012			
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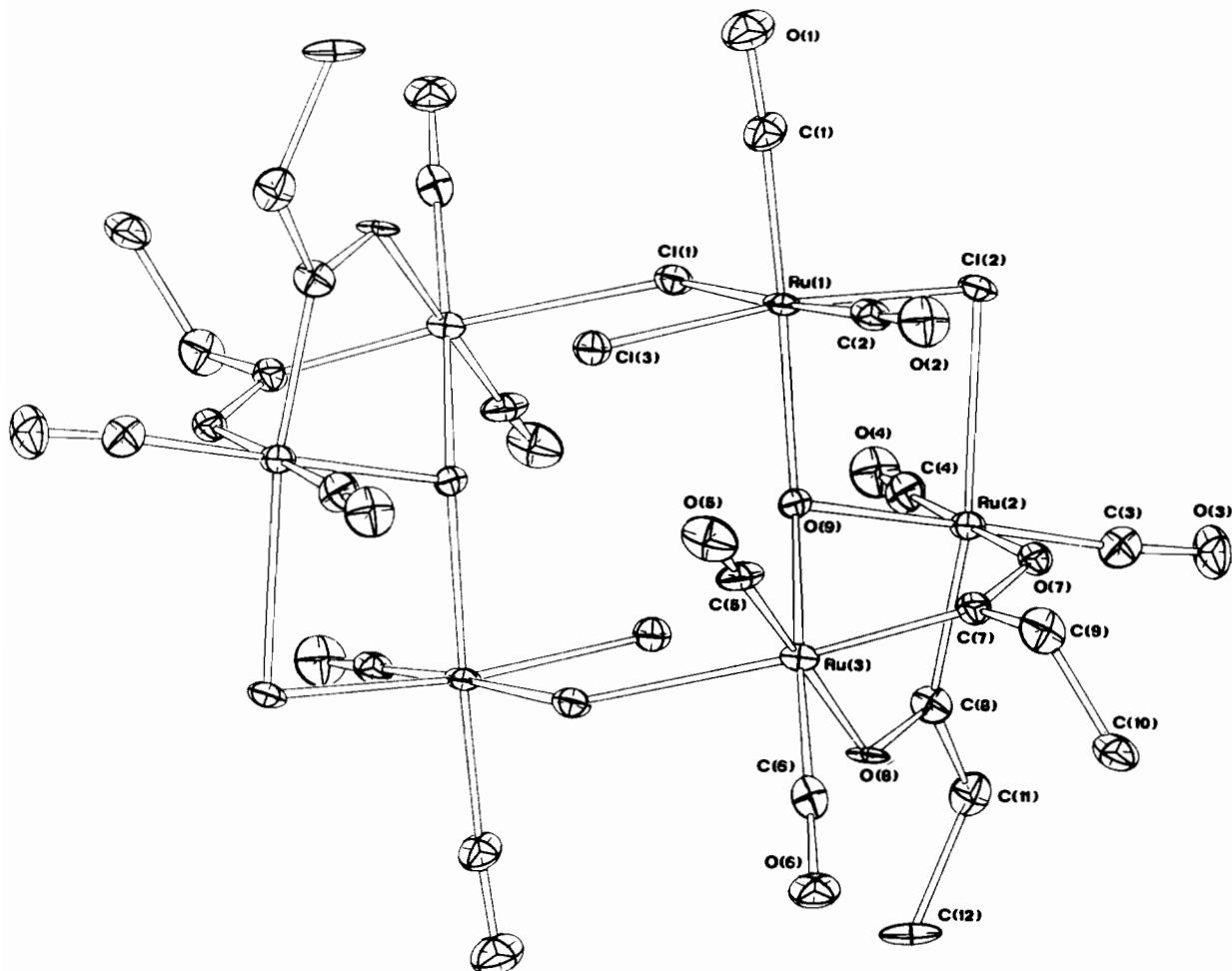
alkyl intermediate derived by a primary insertion of ethylene into a ruthenium–allyl bond of I [2].

The second feature is the presence of a triple bridging hydroxo group: such a group has been previously characterized in the crystal structure of $[\text{Pt}(\text{OH})(\text{CH}_3)_3]_4$ [17, 18], $\text{Ni}_6(\text{CF}_3\text{COCHCOCH}_3)_{10}\cdot(\text{OH})_2(\text{H}_2\text{O})_2$ [11] and $[\text{W}(\text{OH})(\text{CO})_3\text{H}]_4\cdot4\text{OPPh}_2\text{Et}$ [19]. As the hydroxo group is a five-electron donor, in one molecule of II there are a total of sixty ligand-donated electrons, so each ruthenium atom can conform to the closed shell configuration without any metal–metal bond.

The close approach $\text{O}(9)\cdots\text{Cl}(3)^+$ [3.174(9) Å] between the oxygen of the hydroxo group and the terminal chlorine atom is indicative of hydrogen bonding: as appears from Fig. 1, the relative orientation of the two atoms is favourable for hydrogen bonding which assures a further linking between the

two centrosymmetrically related trinuclear units in the molecule. This is also confirmed by the sharp IR absorption at 3280 cm^{-1} , consistent with the O–H stretching frequency for an O...Cl bond distance of 3.19 Å [20]. The origin of the hydroxo group of II, indicated also by a resonance at 10.9 τ in the ^1H n.m.r. spectrum [6], is not clear: it might arise from the moisture of the commercial ethylene used as well as from decomposition of carbon monoxide [21].

The last feature of the molecular structure of II is the different coordination of the three ruthenium atoms in the asymmetric unit: in fact, Ru(2) and Ru(3) are identically coordinated, each being linked to two carbonyl groups, to the “carbene atom” of one ethyloxy carbene group and to the oxygen atom of the other ethyloxy carbene group, to a bridging chlorine atom and to the triple bridging hydroxo group. Ru(1) on the contrary is linked to three chlorine

Figure 1. The molecular structure of $[\text{Ru}_3\text{Cl}_3(\text{COC}_2\text{H}_5)_2(\text{OH})(\text{CO})_6]_2 \cdot \text{C}_6\text{H}_6$.TABLE III. Bond Distances (Å)^{a,b}.

Ru(1)–Cl(1)	2.446(3)	Ru(3)–O(8)	2.094(10)
Ru(1)–Cl(2)	2.398(5)	Ru(3)–O(9)	2.146(9)
Ru(1)–Cl(3)	2.343(5)		
Ru(1)–C(1)	1.851(19)	C(1)–O(1)	1.109(24)
Ru(1)–C(2)	1.842(12)	C(2)–O(2)	1.158(16)
Ru(1)–O(9)	2.169(10)	C(3)–O(3)	1.134(25)
		C(4)–O(4)	1.123(16)
Ru(2)–Cl(2)	2.541(4)	C(5)–O(5)	1.125(23)
Ru(2)–C(3)	1.854(20)	C(6)–O(6)	1.162(21)
Ru(2)–C(4)	1.887(13)	C(7)–O(7)	1.246(19)
Ru(2)–C(8)	1.944(16)	C(8)–O(8)	1.276(19)
Ru(2)–O(7)	2.118(8)	C(7)–C(9)	1.497(17)
Ru(2)–O(9)	2.155(11)	C(9)–C(10)	1.499(28)
		C(8)–C(11)	1.509(23)
Ru(3)–Cl(1) ⁱ	2.571(4)	C(11)–C(12)	1.602(29)
Ru(3)–C(5)	1.885(19)	C(13)–C(14)	1.43(5)
Ru(3)–C(6)	1.805(17)	C(13)–C(15)	1.37(5)
Ru(3)–C(7)	1.996(14)	C(14)–C(15) ⁱⁱ	1.37(5)

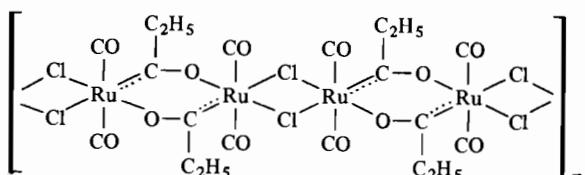
^aThe figures in parentheses are the estimated standard deviations. ^bThe symmetry code used in this and in the subsequent table is the following:

i atom at $-x -y 1-z$

ii atom at $1-x -y -z$

atoms (one terminal and two bridging), to two carbonyl groups and to the triple bridging hydroxo group.

Concerning the structure of the amorphous product, the IR spectrum does not show the presence of the triple bridging OH group and indicates a simpler arrangement of the CO groups as results from the lower Ru–C stretchings between 650 and 400 cm⁻¹. The high value of the molecular weight (~2000) furthermore suggests for this product a “ μ -catena” arrangement of $\text{RuCl}(\text{COC}_2\text{H}_5)(\text{CO})_2$ units of type:



Both the crystalline and amorphous products are stable in the solid state and give by treatment with

TABLE IV. Bond Angles ($^{\circ}$)^a.

Bond Angles at the Ruthenium Atoms			
Cl(1)–Ru(1)–Cl(2)	89.7(1)	O(9)–Ru(2)–C(8)	88.1(6)
Cl(1)–Ru(1)–Cl(3)	89.6(1)	C(4)–Ru(2)–O(7)	178.9(11)
Cl(1)–Ru(1)–C(1)	89.5(4)	C(4)–Ru(2)–C(3)	86.0(7)
Cl(1)–Ru(1)–C(2)	176.6(5)	C(4)–Ru(2)–C(8)	92.2(6)
Cl(1)–Ru(1)–O(9)	85.4(2)	O(7)–Ru(2)–C(3)	94.4(5)
Cl(2)–Ru(1)–Cl(3)	172.3(2)	O(7)–Ru(2)–C(8)	88.7(4)
Cl(2)–Ru(1)–C(1)	95.5(6)	C(3)–Ru(2)–C(8)	91.4(8)
Cl(2)–Ru(1)–C(2)	87.6(6)		
Cl(2)–Ru(1)–O(9)	84.2(3)	O(9)–Ru(3)–Cl(1) ⁱ	88.3(3)
Cl(3)–Ru(1)–C(1)	92.2(6)	O(9)–Ru(3)–C(5)	100.2(6)
Cl(3)–Ru(1)–C(2)	93.4(6)	O(9)–Ru(3)–C(7)	87.2(5)
Cl(3)–Ru(1)–O(9)	88.0(3)	O(9)–Ru(3)–C(6)	174.4(7)
C(1)–Ru(1)–C(2)	88.8(6)	O(9)–Ru(3)–O(8)	82.2(4)
C(1)–Ru(1)–O(9)	174.9(4)	Cl(1) ⁱ –Ru(3)–C(5)	87.6(6)
C(2)–Ru(1)–O(9)	96.2(6)	Cl(1) ⁱ –Ru(3)–C(7)	175.2(5)
		Cl(1) ⁱ –Ru(3)–C(6)	92.4(5)
Cl(2)–Ru(2)–O(9)	81.2(3)	Cl(1) ⁱ –Ru(3)–O(8)	93.1(3)
Cl(2)–Ru(2)–C(4)	89.7(5)	C(5)–Ru(3)–C(7)	91.6(7)
Cl(2)–Ru(2)–O(7)	89.3(3)	C(5)–Ru(3)–C(6)	85.4(8)
Cl(2)–Ru(2)–C(3)	99.3(6)	C(5)–Ru(3)–O(8)	177.6(5)
Cl(2)–Ru(2)–C(8)	169.2(6)	C(7)–Ru(3)–C(6)	92.3(7)
O(9)–Ru(2)–C(4)	98.1(6)	C(7)–Ru(3)–O(8)	88.0(5)
O(9)–Ru(2)–O(7)	81.5(4)	C(6)–Ru(3)–O(8)	92.2(6)
O(9)–Ru(2)–C(3)	175.9(6)		
Bond Angles in the Ligands			
Ru(1)–Cl(1)–Ru(3) ⁱ	130.0(2)	Ru(2)–O(7)–C(7)	124.7(8)
Ru(1)–Cl(2)–Ru(2)	88.7(1)	Ru(3)–C(7)–O(7)	117.2(8)
Ru(1)–C(1)–O(1)	176.1(19)	Ru(3)–C(7)–C(9)	126.7(12)
Ru(1)–C(2)–O(2)	178.1(17)	O(7)–C(7)–C(9)	116.1(13)
Ru(2)–C(3)–O(3)	174.8(16)	C(7)–C(9)–C(10)	110.9(12)
Ru(2)–C(4)–O(4)	176.6(17)	Ru(3)–O(8)–C(8)	124.7(10)
Ru(3)–C(5)–O(5)	172.5(15)	Ru(2)–C(8)–O(8)	117.5(11)
Ru(3)–C(6)–O(6)	177.6(13)	Ru(2)–C(8)–C(11)	127.5(12)
Ru(1)–O(9)–Ru(2)	106.0(4)	O(8)–C(8)–C(11)	115.0(14)
Ru(1)–O(9)–Ru(3)	126.1(4)	C(8)–C(11)–C(12)	116.3(16)
Ru(2)–O(9)–Ru(3)	103.0(4)		
Bond angles in the Benzene Molecule			
C(14)–C(13)–C(15)	118.6(31)		
C(13)–C(14)–C(15) ⁱⁱ	119.1(29)		
C(13)–C(15)–C(14) ⁱⁱ	122.3(33)		

^aThe numbers in parentheses are the corresponding standard deviations.

triphenylphosphine or pyridine, with releasing of carbon monoxide and breaking of the bridging structure, the same mononuclear compounds Ru(COC₂H₅)(CO)L₂Cl·Sv (L = PPh₃, Py; Sv = benzene, methylene chloride, tetrahydrofuran) [2].

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