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The Crystal Structure of a Five-Coordinated Cadmium(II) Complex: Trithiourea-Cadmium Sulphate

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The crystal structure of trithiourea-cadmium sulphate, $\text{Cd}(\text{SC}(\text{NH}_2)_2)_3\text{SO}_4$, has been determined by three dimensional X-ray analysis. Crystals are triclinic ($\bar{P}\bar{1}$) with cell dimensions $a = 9.06(2)$, $b = 9.80(2)$, $c = 8.78(2)$ Å, $\alpha = 110.6(5)$, $\beta = 95.1(5)$, and $\gamma = 92.0(5)$ °. The measured and calculated values for the density (for $Z=2$) are 2.02 and 2.00 g/cm³, respectively. A total of 1568 non-zero reflexions has been obtained by photographic methods. The structure was solved by Patterson methods and refined by the full-matrix least square method to a conventional R factor of 0.097. The crystal structure shows that the molecule is a dimer, and that the coordination around the metal atom is intermediate between square pyramidal and trigonal bipyramidal, although somewhat closer to the former. One sulphur atom from a thiourea molecule coordinates to two adjacent cadmium atoms at distances 2.627(7) and 2.870(7). The other three bond lengths in the coordination polyhedra are: $\text{Cd}-\text{S}(2) = 2.538(6)$, $\text{Cd}-\text{S}(3) = 2.627(6)$ (both S from thiourea molecules) and $\text{Cd}-\text{O}(1) = 2.339(15)$ (the O belonging to the sulphate group).

Introduction

Nardelli and Chierici¹ obtained three complexes of cadmium sulphate with thiourea (hereafter tu); $\text{Cd}(\text{tu})_2\text{SO}_4 \cdot 2\text{H}_2\text{O}$, $\text{Cd}(\text{tu})_4\text{SO}_4$ and $\text{Cd}(\text{tu})_3\text{SO}_4$. The first compound was studied by X-ray diffraction² and the crystal structure showed that it was polymeric with octahedral coordination around the cadmium atom.

In this work, the analysis of the crystal structure of $\text{Cd}(\text{tu})_3\text{SO}_4$ is undertaken, in order to investigate the coordination around the metal atom and the binding of its ligands, since it seems that these cannot be determined from the chemical composition alone.

Experimental Section

The crystals of $\text{Cd}(\text{tu})_3\text{SO}_4$ were kindly supplied by Dr. R. Levitus.

Crystals are triclinic, with cell dimensions $a = 9.06(2)$ Å, $b = 9.80(2)$ Å, $c = 8.78(2)$ Å, $\alpha = 110.6(5)$ °, $\beta = 95.1(5)$ °, $\gamma = 92.0(5)$ °. These

values were obtained from calibrated precession photographs, taken with Mo $\text{K}\alpha$ radiation ($\lambda = 0.7107$ Å) at 20°C, and standard deviations were estimated from repeated film measurements; they are slightly different from those reported previously by Nardelli et al.¹ The space group is $\bar{P}\bar{1}$, as determined by crystal analysis. There are two molecules per cell and the observed and calculated density are 2.02 and 2.00 g/cm³, respectively. The volume of the unit cell is 724 Å³.

The intensities were collected with a single crystal of roughly spherical shape and mean radius of 0.01 cm, which was mounted with the c axis parallel to the spindle axis.

The layers $hk0$ to $hk8$ were recorded in a Weissenberg camera using the equi-inclination technique; the $h01$ and $0kl$ layers with a precession camera (all with Mo $\text{K}\alpha$ radiation).

The intensities were estimated visually on the Weissenberg photographs by comparison with a calibrated scale; for the integrated precession photographs a microdensitometer was used. A total of 1568 non-zero reflexions was obtained within a 2,000:1 ratio of maximum to minimum intensity.

Lorentz and polarization corrections were taken into account and the Phillips spot-shape factor³ was applied to the extended half of the upper level Weissenberg photographs.

The data from the different layers were put on the same scale using the $h01$ and $0kl$ levels.

No absorption correction was applied ($\mu_{\text{MoK}\alpha} = 21.4 \text{ cm}^{-1}$) because if the crystal is assumed to absorb as a sphere of radius 0.01 cm, the largest variation in intensity for a given layer is less than 2 per cent.

Structure Determination and Refinement. The cadmium and sulphur atoms were located in a three dimensional Patterson synthesis. The positions of the other atoms were obtained from successive electron density maps.

The initial reliability index $R = (\sum(|F_o| - |F_c|)) / \sum|F_o|$ with all atoms was 0.29. A full matrix least square refinement was begun at this point, refining the positional parameters, isotropic temperature factor and a scale factor, one for each layer. After four cycles, convergence was achieved with reliability indexes $R = 0.122$ and $R' = (\sum \omega(F_o^2 - F_c^2)) / \sum \omega F_o^2 = 0.142$. The weighting scheme used was $\sigma = 1$ when

(3) D. C. Phillips, *Acta Cryst.*, 7, 746 (1954).

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(1) M. Nardelli and I. Chierici, *Ric. Sci.*, 28, 1016, (1958).
(2) L. Cavalca, P. Domiano, G. Fava and P. Boldrini, *Acta Cryst.* 22, 878 (1967).

$|F_o| < 50.0$ and $\sigma = (1 + 0.1(|F_o| - 50.0))^{1/2}$ when $|F_o| > 50.0$, which gave a fairly constant value of $\omega\Delta^2$ in ranges of $|F_o|$ and $\sin \theta/\lambda$.

The refinement was continued allowing for anisotropic temperature factors of the Cd and S atoms and reducing the number of variable scale factors to one overall scale factor. This new series of refinement cycles converged after three cycles with $R = 0.097$ and $R' = 0.123$. With the coordinates obtained in the last cycle of refinement a difference Fourier synthesis was computed. The resulting electron density map showed no peaks exceeding ± 0.8 e. \AA^{-3} .

The atomic scattering factors for neutral atoms used during the refinement were taken from a paper of Cromer and Waber.⁴

No attempt was made to locate hydrogen atoms. Theoretical positions for H were calculated assuming an sp^2 configuration for the nitrogen atoms and distances N-H of 1.03 \AA . These positions are given in Table I.

Table I. Theoretical Positions for Hydrogen Atoms

Atom	x	y	z
H(1)	0.580	0.340	0.089
H(2)	0.741	0.384	0.009
H(3)	0.141	0.060	0.566
H(4)	0.179	0.224	0.785
H(5)	0.573	0.132	0.626
H(6)	0.568	0.333	0.738
H(7)	0.551	0.125	0.127
H(8)	0.700	-0.012	0.074
H(9)	-0.076	0.452	0.688
H(10)	0.061	0.421	0.834
H(11)	0.426	0.468	0.638
H(12)	0.300	0.385	0.431

Table II. Coordinates and Isotropic Temperature Factors. Standard deviations given in parentheses are in units of the last decimal place.

Atom	x	y	z	B
Cd	0.1125(2)	0.1331(2)	0.2211(2)	*
S(1)	0.0991(6)	-0.1450(6)	0.0367(7)	*
S(2)	-0.0798(7)	0.2003(7)	0.4235(8)	*
S(3)	0.3728(7)	0.0977(6)	0.3546(7)	*
S(4)	0.2732(6)	0.3657(5)	0.1057(7)	*
O(1)	0.1887(16)	0.3711(15)	0.2456(18)	2.8(2)
O(2)	0.4260(20)	0.4310(19)	0.1671(22)	4.3(3)
O(3)	0.1961(19)	0.4512(18)	0.0162(21)	4.0(3)
O(4)	0.2737(18)	0.2085(18)	-0.0056(21)	3.9(3)
N(1)	0.3205(23)	-0.3123(22)	-0.0465(25)	4.0(4)
N(2)	0.1220(23)	0.1795(22)	0.6631(26)	3.6(4)
N(3)	0.5316(22)	0.2425(24)	0.6384(29)	4.5(4)
N(4)	0.3464(22)	-0.0956(20)	-0.0837(26)	3.5(4)
N(5)	0.0018(24)	0.3880(22)	0.7208(26)	4.0(4)
N(6)	0.3843(23)	0.3805(23)	0.5340(27)	4.1(4)
C(1)	0.2695(22)	-0.1826(21)	0.0354(26)	2.3(3)
C(2)	0.0256(23)	0.2592(22)	0.6096(26)	2.6(4)
C(3)	0.4351(29)	0.2513(26)	0.5137(31)	3.7(4)

* were refined anisotropically

Results

Table II gives the positional parameters and the isotropic temperature factor; Table III gives the

(4) D. T. Cromer and J. T. Waber, *Acta Cryst.*, **18**, 104 (1965).

anisotropic thermal parameters for the Cd and S atoms. The observed and calculated structure factors using the numbers given in Tables II and III are given in Table IV. Table V gives the least squares planes passing through the thiourea molecules and the deviations of the atoms from them. Table VI gives the most important bond distances and angles.

Table III. Anisotropic Temperature Factors. Thermal parameters ($\times 10^4$). Standard deviations given in parentheses are in units of the last decimal place.

Atom	β_{11}	β_{22}	β_{33}	β_{12}	β_{13}	β_{23}
Cd	100(2)	84(2)	103(3)	6(1)	31(2)	25(2)
S(1)	63(6)	76(6)	113(10)	3(5)	17(6)	31(6)
S(2)	84(7)	127(8)	122(11)	23(6)	8(6)	13(7)
S(3)	110(8)	78(6)	109(10)	13(6)	-6(7)	15(6)
S(4)	77(6)	72(6)	96(10)	4(5)	13(6)	32(6)

The form of the anisotropic temperature expression is
$$(\sum_i \sum_j \beta_{ij} h_i h_j)$$

Discussion

Figure 1 gives a projection of the structure on the (100) plane of a set of orthogonal coordinates obtained by the transformation given in the footnote of Table V. Figure 2 shows a schematic view of the molecule.

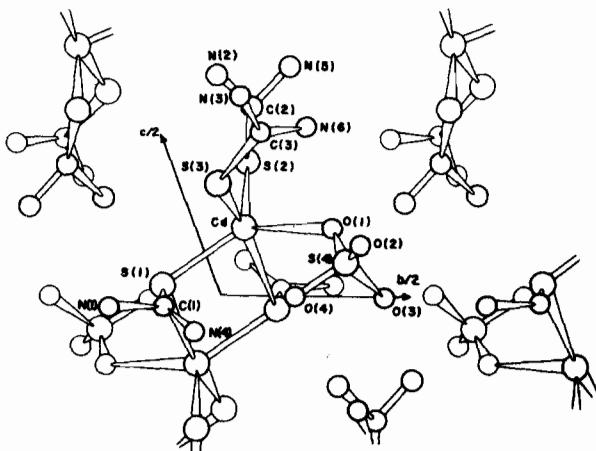


Figure 1. Orthogonal projection of the structure on the plane (100).

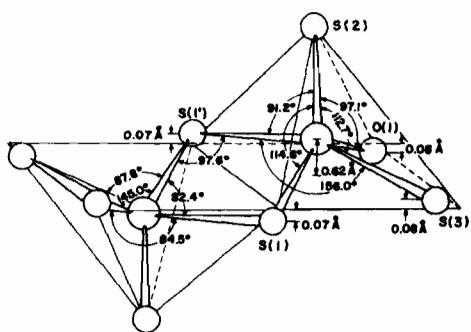


Figure 2. Schematic view of the molecule.

Table IV. Observed and calculated structure factors. The five columns contain values of h , k , l , $10|F_0|$ and $10 F_c$.

0	1	0	606	561	10	0	0	152	221	0	-4	1	406	-462	-11	3	2	254	-201	0	9	2	220	-245	-8	-5	3	263	230	
0	2	0	527	-694	11	-	0	162	174	0	-3	1	668	732	-10	3	2	239	-176	1	-12	2	220	247	-8	-2	3	204	-697	
0	3	0	1176	-940	-12	-3	1	160	-210	0	-2	1	770	605	-9	-4	2	230	210	1	-7	2	237	-236	-8	-1	3	158	-220	
0	4	0	1451	-1144	-11	-4	1	158	-210	0	-1	1	663	197	-8	-3	2	219	265	1	-5	2	315	281	-8	-1	0	150	108	
0	5	0	107	84	-11	-3	1	151	116	0	1	1	169	-653	-9	-2	2	214	194	1	-4	2	930	918	-8	1	3	323	320	
0	6	0	195	204	-11	-1	1	255	201	0	2	1	555	-490	-9	-1	2	210	-161	1	-2	2	250	-236	-8	0	2	3	345	401
0	7	0	382	400	-11	0	1	207	242	0	3	1	811	-411	-9	0	2	304	-329	1	-1	2	426	-556	-8	2	3	345	401	
0	8	0	158	-1139	-10	-4	1	266	-141	0	4	1	305	-365	-9	3	2	275	182	1	0	2	980	-822	-8	0	4	3	152	-187
1	-9	0	206	208	-10	-3	1	284	-230	0	5	1	476	425	-9	4	2	163	161	1	-5	2	325	687	-7	-6	3	200	331	
1	-8	0	106	139	-10	-1	1	308	297	0	6	1	559	583	-8	-5	2	271	220	1	2	2	325	541	-7	-5	3	142	162	
1	-6	0	242	246	-10	0	1	137	138	0	7	1	363	371	-8	-3	2	268	250	1	3	2	259	237	-7	-3	3	353	-343	
1	-5	0	205	-262	-10	3	1	159	-186	0	9	1	177	-200	-8	-2	2	144	-192	1	4	2	378	374	-7	-3	3	450	-486	
1	-4	0	731	-611	-9	-4	1	207	-153	0	10	1	237	-265	-8	-1	2	413	-461	1	5	2	382	-356	-7	-2	3	450	-486	
1	-9	0	c 87	-75	-9	-5	1	159	-142	1	-10	1	156	176	-8	0	2	246	-264	1	6	2	216	-271	-7	1	3	303	386	
1	-2	0	c 358	-351	-9	-4	1	193	-153	1	-9	1	222	227	-8	3	2	557	516	1	0	2	210	-205	-7	1	3	311	292	
1	-1	0	1247	1257	-9	-2	1	260	263	1	-8	1	97	-137	-8	4	2	217	167	1	10	2	146	200	-7	4	3	159	140	
1	0	0	c 586	775	-9	-1	1	248	281	1	-6	1	120	-111	-8	7	2	300	-252	2	-13	2	153	196	-7	4	3	142	142	
1	2	0	225	-193	-9	2	1	259	-256	1	-5	1	234	-275	-7	-9	2	251	-139	2	-10	2	132	-162	-7	3	3	254	290	
1	3	0	1458	-1415	-8	-7	1	280	-202	1	-1	1	200	282	-7	-4	2	325	362	2	-8	2	135	-173	-7	5	3	180	-215	
1	4	0	c 372	-404	-8	-3	1	126	130	1	0	1	894	-712	-7	-2	2	452	-430	2	-5	2	1188	1105	-6	3	3	302	-302	
1	5	0	245	-222	-8	-2	1	232	232	1	1	1	275	-291	-7	0	2	386	-386	2	-4	2	757	675	-6	0	3	320	327	
1	6	0	c 545	-586	-8	2	1	207	-396	1	2	1	208	-194	-7	2	2	531	515	-3	2	245	246	-6	0	1	349	457		
1	8	0	c 245	261	-8	5	1	261	227	4	1	1	726	690	-7	5	2	180	-103	-2	-1	2	1319	-1531	-6	3	3	227	-210	
1	9	0	c 245	-266	-6	6	1	235	220	5	3	1	248	248	-7	6	2	562	-476	4	0	2	102	-112C	-6	4	3	194	-244	
1	10	0	c 156	-170	-7	8	1	203	-183	1	8	1	127	-147	-7	7	2	166	-152	2	1	2	211	362	-5	9	3	222	213	
1	11	0	220	-223	-7	7	1	335	313	1	9	1	280	-305	-6	-9	2	240	-177	2	2	2	820	749	-5	8	3	222	213	
1	12	0	c 151	149	-7	5	1	248	230	1	10	1	127	-130	-6	-8	2	224	-212	2	3	2	658	819	-5	6	3	345	-373	
1	13	0	c 199	243	-7	4	1	267	315	2	-12	1	215	178	-6	-7	2	152	152	2	6	2	83	76	-5	6	3	403	-390	
1	14	0	c 151	196	-7	3	1	230	217	2	-11	1	247	220	-6	-6	2	290	263	2	5	2	535	-486	-5	4	3	436	-416	
1	15	0	c 185	155	-7	2	1	153	-187	2	-10	1	232	267	-6	-5	2	215	250	2	7	2	126	-151	-5	4	3	547	-547	
1	16	0	c 567	-566	-7	1	129	-263	2	-9	1	216	-203	-6	-4	2	426	384	2	9	2	163	160	-5	3	3	366	329		
1	17	0	c 422	-389	-7	0	1	242	-224	2	-8	1	493	-465	-6	-3	2	327	-292	-3	-10	2	168	-242	-6	3	3	377	353	
1	18	0	c 136	-90	-7	3	1	229	260	2	-6	1	356	-419	-6	-1	2	336	-335	3	-8	2	126	142	-5	2	3	350	-323	
1	19	0	c 333	253	-7	4	1	245	245	2	-5	1	238	224	-6	0	2	410	-384	3	-6	2	441	469	-5	3	3	448	-331	
1	20	0	c 468	461	-7	5	1	350	389	2	-4	1	562	506	-6	1	2	122	123	3	-5	2	166	193	-6	4	3	350	-379	
1	21	0	c 6	49	-30	-7	1	358	-310	2	-3	1	431	445	-6	2	2	347	328	3	-4	2	120	97	-2	6	3	352	346	
1	22	0	c 651	-650	-6	2	1	240	-222	2	-2	1	515	907	-6	3	2	466	446	3	-2	2	1039	-1102	-5	7	3	219	210	
1	23	0	c 651	-580	-6	8	1	240	-386	2	-1	1	221	-184	-6	4	2	153	178	3	-1	2	661	-636	-9	3	3	340	329	
1	24	0	c 5	250	-274	-6	6	1	349	-359	2	0	1	305	-1255	-6	5	2	443	-400	3	0	2	132	123	-8	8	3	377	353
1	25	0	c 274	290	-6	6	1	309	637	2	1	1	244	-226	-6	2	2	307	364	3	1	2	713	647	-6	4	3	329	-325	
1	26	0	c 267	267	-6	3	1	301	292	2	3	1	324	224	-6	2	2	267	241	3	0	2	1043	956	-5	5	3	546	-606	
1	27	0	c 165	-210	-6	1	1	324	322	2	4	1	400	-400	-6	-5	2	452	466	3	4	2	454	544	-5	6	3	350	-249	
1	28	0	c 250	300	-6	2	1	266	266	2	8	1	265	185	-6	-3	2	326	-277	4	-10	2	103	313	-6	3	3	313	316	
1	29	0	c 472	-484	-6	3	1	260	252	2	9	1	58	100	-6	-2	2	219	-188	3	6	2	396	-396	-6	1	3	322	301	
1	30	0	c 675	-680	-6	4	1	248	220	3	-11	1	265	209	-6	-1	2	154	-172	3	7	2	235	-193	-7	3	3	321	217	
1	31	0	c 210	-205	-6	5	1	191	-970	3	-1	1	856	-500	-6	4	2	427	-380	4	-3	2	347	-325	-7	3	3	346	-517	
1	32	0	c 118	-180	-6	0	1	551	-539	5	0	1	516	571	-6	3	2	512	-452	4	-2	2	317	-321	-6	3	3	578	-605	
1	33	0	c 497	-516	-6	1	1	182	180	5	1	1	133	-124	-6	1	2	483	955	5	0	2	221	208	-6	3	3	360	345	
1	34	0	c 232	-355	-6	2	1	457	439	5	2	1	322	322	-6	0	2	642	697	4	1	2	362	361	-6	3	3	322	322	
1	35	0	c 814	835	-6	3	1	412	461	5	3	1	602	655	-6	5	2	437	-409	5	-11	2	155	-202	-6	3	3	332	324	
1	36	0	c 8	156	-184	-6	4	1	550	589	5	0	1	193	-167	-6	-4	2	366	-332	3	-3	2	297	-260	-7	10	2	152	150
1	37	0	c 322	330	-6	5	1	434	-414	5	2	1	358	374	-6	-1	2	333	-283	6	-4	2	573	-555	-6	1	3	318	326	
1	38	0	c 277	271	-6	2	1	349	319	5	3	1	128	-114	-6	1	2	118	157	5	7	2	229	251	-2	2	3	512	-496	
1	39	0	c 324	-405	-6	2	-11	1	385	-361																				

Table IV. (Continued)

1	3	258	244	-3	-9	4	250	-237	-6	-7	5	282	-310	-7	4	6	346	356	7	-4	6	176	-170	5	0	7	117	122
1	5	420	-390	-3	-8	4	357	-373	-6	-6	5	270	-296	-6	-5	6	190	210	7	-3	6	206	-203	5	1	7	116	-132
1	6	320	-336	-3	-7	4	210	-221	-6	-4	5	384	-392	-6	-5	6	258	260	7	-1	6	178	207	5	2	7	139	-104
1	8	314	169	-3	-6	4	165	-119	-6	-3	5	624	622	-6	-2	6	262	-245	7	-1	6	190	201	5	3	7	137	-105
1	9	313	292	-3	-5	4	161	-144	-6	-2	5	268	204	-6	-1	6	302	-303	7	2	6	172	-160	6	-9	7	160	173
1	-10	311	-185	-3	-4	4	617	661	-6	-1	5	410	-389	-6	-2	6	234	260	7	3	6	179	-177	6	-7	7	97	92
1	-9	252	-259	-3	-3	4	270	-270	-6	0	5	426	-456	-6	3	6	282	304	8	-4	6	110	-124	6	-5	7	246	-244
1	-8	308	321	-3	-2	4	681	-456	-6	1	5	466	-492	-5	-6	6	178	249	8	0	6	189	180	6	-4	7	97	-130
1	-7	343	303	-3	-1	4	242	-216	-6	3	5	335	324	-5	-4	6	201	167	8	2	6	183	-197	6	-5	7	246	-244
1	-6	337	530	-3	-2	4	401	-335	-6	4	5	320	330	-5	-2	6	311	-316	8	3	6	146	-140	6	-4	7	162	-130
1	-5	333	375	-3	-4	4	355	411	-6	5	5	212	231	-5	2	6	222	260	9	-4	7	217	260	6	-2	7	153	167
1	-4	747	-699	-2	-10	4	151	-129	-6	-8	5	250	-176	-4	-7	6	359	447	-3	-3	7	215	260	6	-1	7	144	166
1	-3	625	-570	-2	-9	4	245	-225	-5	-7	5	263	-264	-4	-6	6	265	-266	-6	-6	7	234	-231	6	-1	7	126	-129
1	0	459	462	-2	-6	4	370	335	-6	5	5	240	-241	-4	-5	6	357	-341	-6	-4	7	242	350	7	-6	7	142	-189
1	1	657	656	-2	-5	4	460	530	-6	5	5	304	307	-6	-2	6	302	-312	7	-1	7	246	-311	7	-5	7	160	-138
1	2	947	480	-2	-4	4	226	179	-5	3	5	527	526	-6	1	6	578	563	-6	0	7	252	-312	7	-3	7	120	124
1	3	312	155	-2	-3	4	411	-355	-5	1	5	359	355	-4	2	6	210	191	-8	2	7	215	187	7	-1	7	123	143
1	4	253	-234	-2	-2	4	251	-289	-5	0	5	425	-449	-4	3	6	225	-201	-8	3	7	223	240	-8	1	7	230	235
1	5	310	-390	-2	-1	4	764	-868	-5	1	5	239	-262	-3	-8	6	367	407	-7	1	7	232	-244	-7	-4	7	145	-210
1	6	318	-257	-2	-0	4	235	-220	-5	2	5	311	320	-3	-7	6	265	257	-7	2	7	255	248	-7	0	7	259	253
1	7	310	244	-2	-1	4	478	422	-5	3	5	390	392	-3	-5	6	147	-134	-7	3	7	213	209	-6	-9	7	230	202
1	8	315	202	-2	-2	4	479	372	-5	4	5	194	189	-3	-4	6	511	-576	-6	-6	7	189	203	-6	-5	7	220	200
1	9	318	119	-2	-3	4	619	622	-4	-9	5	204	-163	-3	-3	6	429	-441	-6	-3	7	175	-171	-6	-5	7	246	-258
1	-6	205	172	-1	-10	4	250	-233	-8	-8	5	270	-259	-3	-1	6	233	261	-6	0	7	180	184	-6	-6	7	135	-203
1	-5	317	-393	-1	-7	4	617	165	-5	5	5	198	196	-3	0	6	565	656	-5	-7	7	349	376	-6	0	7	194	221
1	-4	192	-215	-1	-6	4	324	262	-5	1	5	367	404	-3	1	6	459	484	-5	7	7	210	-213	-5	-9	7	211	212
1	-3	141	-149	-1	-5	4	354	366	-4	-2	5	281	-277	-3	2	6	165	-135	-5	-4	7	333	-231	-5	-6	7	304	-400
1	-2	511	514	-1	-3	4	480	-410	-4	0	5	381	-406	-3	4	6	489	-499	-5	-1	7	231	225	-5	-6	7	236	161
1	0	810	710	-1	-2	4	523	-611	-2	-2	5	284	-276	-5	3	6	109	-159	-5	0	7	168	226	-5	2	7	216	270
1	1	315	153	-1	-1	4	741	-676	-4	-4	5	245	-241	-3	6	6	117	-137	-5	3	7	155	-161	-5	-1	7	255	272
1	2	318	180	-1	-0	4	460	417	-3	-11	5	244	-240	-2	-8	6	324	-312	-8	0	7	242	271	-5	2	7	206	-192
1	3	340	-400	-1	-1	4	565	745	-3	-1	5	244	-240	-4	3	6	369	-408	-2	7	7	217	215	-5	3	7	263	-285
1	4	150	-120	-1	-0	4	460	160	-3	-5	5	295	-325	-2	7	6	266	316	-4	2	7	205	-203	-6	-6	7	336	381
1	5	320	-274	-2	-2	4	503	-553	-5	-2	10	252	-241	-1	-9	6	351	366	-3	8	7	222	230	-3	1	7	153	-167
1	6	318	118	-0	-10	4	374	-251	-3	-5	5	390	379	-2	-1	6	361	322	-6	4	7	349	584	-3	0	7	167	203
1	7	290	-269	0	-9	4	131	-177	-3	-2	5	736	-750	-2	0	6	437	438	-3	4	7	355	-236	-2	2	7	312	-382
1	8	313	-164	0	-7	4	371	400	-3	-1	5	156	-151	-2	3	6	102	-152	-6	2	7	241	330	-3	2	7	196	-120
1	9	445	470	0	-6	4	455	498	-3	1	5	156	-151	-2	3	6	474	-470	-6	1	7	517	498	-3	5	7	246	188
1	0	256	304	0	-5	4	367	287	-3	1	5	483	501	-2	4	6	500	-456	-4	0	7	352	382	-3	7	181	145	
1	1	150	-120	0	-4	4	460	160	-3	3	5	336	322	-2	6	6	255	247	-4	2	7	331	-330	-3	3	7	246	268
1	2	320	-274	0	-2	4	505	-553	-5	-2	10	252	-241	-1	-9	6	351	366	-3	8	7	222	230	-3	1	7	153	-167
1	3	340	164	0	-2	4	416	476	-5	-7	5	281	-277	-2	-3	6	370	217	-3	6	7	241	239	-2	1	7	103	-125
1	4	167	162	0	0	4	435	410	-2	-9	5	283	-230	-2	1	6	516	494	-0	6	7	241	160	-2	0	7	221	-270
1	5	207	227	0	1	4	525	520	-2	-7	5	388	-319	-2	4	6	186	-181	-2	5	7	243	229	-1	2	7	96	128
1	6	348	578	0	2	4	525	456	-2	-5	5	295	-293	-2	0	6	276	292	-2	4	7	136	-93	-1	-9	7	281	-376
1	7	341	-380	2	-6	4	207	-225	-2	-4	5	295	-312	-1	-11	6	181	-181	-2	3	7	343	274	-1	0	7	226	-260
1	8	322	-245	2	-5	4	374	-383	-1	-11	5	298	-312	0	-1	6	195	148	-2	2	7	343	116	-1	-7	7	172	172
1	9	341	404	3	-6	4	150	-159	-1	-1	5	159	159	-1	-5	6	230	-195	-1	7	7	343	116	-0	9	7	243	-276
1	0	347	520	3	-5	4	447	-401	-1	-3	5	358	-426	-1	-5	6	223	220	-1	7	7	347	215	-0	7	7	242	190
1	1	366	-371	3	-4	4	130	121	-1	4	5	280	-301	-2	1	6	316	298	-0	3	7	241	252	-2	1	7	364	324
1	2	333	268	3	-3	4	354	347	-1	7	5	279	-260	-2	12	6	139	136	-0	4	7	251	252	-0	5	7	346	221
1	3	233	235	3	-2	4	346	363	-2	3	5	337	151	-3	0	6	367	318	-2	7	7	343	254	-2	1	7	346	194
1	4	267	-263	5	-7	4	297	-308	-2	-1	5	459	-72	-2	6	6	287	-300	-3	1	7	345	361	-2	-8	7	213	211
1	5	256	-308	5	-6	4	333	-275	2	0	5	540	-509	-2	7	6	217	-233	-1	6	7	188	-171	-2	-6	7	226	201
1	6	267	-199	5	-5	4	232	-229	2	1	5	422	-380	-3	-12	6	141	157	1	7	7	183	-197	-2	-6	7	126	103
1	7	260	194	5	-4	4	161	123	2	2	5	250	-243	-3	-9	6	200	-235	2	-8	7	205	-196	-2</td				

Table V. Equations of least-squares planes and deviations (\AA) from these planes. (In this calculations an unitary weighting scheme was used). Coefficients are direction cosines relative to orthogonal axes. X, Y, Z are also orthogonal coordinates, in \AA^a .

Equation of plane	Atom	Deviation
$(-0.358)X + (-0.096)Y + (-0.929)Z + 2.78 = 0$	S(1) C(1) N(1) N(4) S(2) C(2) N(2) N(5) S(3) C(3) N(3) N(6)	-0.0044(6) 0.01(2) 0.00(2) 0.00(2) -0.012(6) 0.04(3) -0.01(2) -0.01(2) -0.011(6) 0.04(3) -0.01(2) -0.01(2)
$(0.743)X + (0.570)Y + (-0.350)Z - 1.669 = 0$		
$(0.793)X + (0.358)Y + (-0.484)Z - 0.986 = 0$		

^a Orthogonal coordinates in angstroms are obtained by transforming the fractional coordinates with the matrix (a, b, $\cos \gamma$, c, $\cos \beta // 0$, b, $\sin \gamma$, c, $(\cos \alpha - \cos \beta \cos \gamma) / \sin \gamma // 0, 0, V / (a.b.\sin \gamma)$). X is along the axis, Y in the (a, b) plane and Z along the c-axis.

Table VI. Distances (\AA) and Angles ($^\circ$) in Cd(tu)_3SO_4 . Standard deviations, given in parentheses, are in units of the last decimal place. For numbering of atoms see Figure 1.

Coordination Polyhedra					
Cd—S(1)	2.627(7)	S(1)—Cd—S(1')	82.4(2)	S(1')—Cd—S(3)	156.0(2)
Cd—S(1')	2.870(7)	S(1)—Cd—S(2)	114.6(2)	S(1')—Cd—O(1)	84.5(4)
Cd—S(2)	2.538(6)	S(1)—Cd—S(3)	87.8(2)	S(2)—Cd—S(3)	112.7(2)
Cd—S(3)	2.627(6)	S(1)—Cd—O(1)	145.0(4)	S(2)—Cd—O(1)	97.1(14)
Cd—O(1)	2.339(15)	S(1')—Cd—S(2)	91.2(2)	S(3)—Cd—O(1)	91.6(4)
Sulphate Group:					
S(4)—O(1)	1.491(15)	O(1)—S(4)—O(2)	110.0(1.0)	O(2)—S(4)—O(3)	109.4(1.0)
S(4)—O(2)	1.476(19)	O(1)—S(4)—O(3)	108.0(1.0)	O(2)—S(4)—O(4)	111.0(1.0)
S(4)—O(3)	1.487(19)	O(1)—S(4)—O(4)	108.4(1.0)	O(3)—S(4)—O(4)	109.0(1.0)
S(4)—O(4)	1.507(19)				
S—O (mean value)	1.490				
Thiourea Molecules:					
S(1)—C(1)	1.724(21)	S(1)—C(1)—N(1)	117.3(1.6)		
C(1)—N(1)	1.342(31)	S(1)—C(1)—N(4)	123.5(1.7)		
C(1)—N(4)	1.287(30)	N(1)—C(1)—N(4)	119.1(2.0)		
S(2)—C(2)	1.713(22)	S(2)—C(2)—N(2)	125.5(1.7)		
C(2)—N(2)	1.351(31)	S(2)—C(2)—N(5)	118.8(1.7)		
C(2)—N(5)	1.335(31)	N(2)—C(2)—N(5)	115.2(2.0)		
S(3)—C(3)	1.690(27)	S(3)—C(3)—N(3)	119.7(2.0)		
C(3)—N(3)	1.368(34)	S(3)—C(3)—N(6)	123.3(2.0)		
C(3)—N(6)	1.321(36)	N(3)—C(3)—N(6)	116.6(2.3)		
S—C (mean value)	1.709	C—N (mean value)	1.336		

In the following discussion atoms are labelled as follow:

Superscript	Coordinates
none	x, y, z
'	-x, -y, -z
"	1-x, 1-y, 1-z
'''	1-x, -y, 1-z
iv	x, y, 1+z
v	x, 1-y, z
vi	1+x, -y, -z
vii	-x, 1-y, 1-z
viii	x, y, 1+z

It is seen that in Cd(tu)_3SO_4 , the cadmium atom is five-coordinated. Coordination can best be described in terms of a distorted square pyramid. Four sulphur atoms from thiourea molecules and one oxygen atom from the sulphate group surround the metal

atom. The atoms S(1), S(3), S(1') and O(1) determine the base of the distorted pyramid, while S(2) is at the apex. The least squares plane passing through S(1), S(3), S(1') and O(1) is $(0.664)X + (0.208)Y + (-0.718)Z + 0.001 = 0$ where X, Y, Z are orthogonal coordinates (see footnote of Table V). The distances of the atoms to this plane are: S(1) = -0.074(6); S(3) = 0.076(6); S(1') = 0.077(6) and O(1) = -0.078(6). For Cd and S(2) the distances are 0.624(2) and 3.099(6) \AA , respectively. As can be seen in Figure 2, atoms S(1') and S(3) that are 0.077 and 0.076 \AA above the mean plane, respectively, make the angle S(1')—Cd—S(3) greater than S(1)—Cd—O(1). The angles S(2)—Cd—O(1) and S(2)—Cd—S(1') are smaller than S(2)—Cd—S(1) and S(2)—Cd—S(3). This indicates that the S(2)—Cd bond is tilted with respect to the basal plane in the direction of the S(1')—O(1) edge of the base. Two pyramids related by an inver-

sion center are joined through the S(1')—S(1) edge, resulting in a dimer molecule. The angles around S(1): Cd—S(1)—Cd' = 97.6(7), Cd—S(1)—C(1) = 107.8(7) and Cd'—S(1)—C(1) = 109.6(7) $^{\circ}$, indicate a distorted tetrahedral environment. The S(1) atom coordinates to two adjacent Cd atoms with somewhat different distances: 2.627(7) and 2.870(7) Å. The other distances between Cd and S in the coordination polyhedra are: Cd—S(2) = 2.538(6) and Cd—S(3) = 2.627(6) Å.

The ligand thiourea acts as a bridge in other Cd^{II} complexes with distances Cd—S of 2.64 and 2.65 Å in monothiourea cadmium sulphate,² and 2.74 and 2.71 in bisthiourea cadmium formate.⁵ Some of these values are significantly different from those obtained by us, and in all cases they are intermediate between the sum of Pauling's covalent and ionic radii, 2.52 and 2.87 Å, respectively.

The Cd—O(1) distance, 2.339(15) Å, is greater than the sum of Pauling's covalent radii, 2.22 Å, and it differs by less than three e.s.d. from the mean value obtained in Cd(tu)SO₄ · 2H₂O.²

It is convenient at this point to emphasize that Cd(tu)₃SO₄ is the first reported five-coordinated Cd^{II} complex; it is also the first reported structure where thiourea acts as a bridge in a dimer molecule. Distorted square pyramidal coordination has been also found in some Zn^{II} complexes, *i.e.* Zn(acac)₂ · H₂O,⁶ and the angles obtained in the coordination polyhedra are rather similar to those found by us.

Distances and angles in the sulphate group are in good agreement with those observed in other sulphate complexes and are practically equal to those found in the ionic sulphate.⁷ The orientation of the sulphate group is mainly determined by the intra and intermolecular hydrogen bonds involving the oxygen atoms and the NH₂ groups. The angle Cd—O(1)—S(4) is 109.3(8) $^{\circ}$. One of the oxygen atoms, O(4), is located at 2.84 Å from Cd atom. This is a relatively short distance for non bonded atoms; a tentative explanation for this fact is that O(4) participates strongly in intra and intermolecular hydrogen bonding with H(8) and H(4^{IV}).

The three independent thiourea molecules are planar, and their least square planes are given in Table V. The dimensions in the thiourea molecules

are not significantly different from those reported previously for the free ligand.⁸ The S—C bond lengths are between 1.690 and 1.724 Å (mean value 1.709(14) Å), compared with a value of 1.720(9) Å in the free ligand; while the mean value of the C—N bond length, 1.336(15) Å, differs by less than one e.s.d. from 1.340(6) Å, which is the value found in the free thiourea.

The thiourea molecules are tilted with respect to the Cd—S bond. The values of the angles are: Cd—S(1)—C(1) = 107.8(7), Cd—S(1')—C(1) = 109.6(7), Cd—S(2)—C(2) = 103.2(8) and Cd—S(3)—C(3) = 100.0(9). The values are consistent with those obtained in other thiourea complexes, *i.e.* 101.9, 107.7 and 107.0 $^{\circ}$ in trithiourea zinc sulphate;⁷ 100.6 and 101.2 $^{\circ}$ in bisthiourea zinc acetate⁹ and 113.0 $^{\circ}$ in bisthiourea cadmium chloride.¹⁰

Apparently the molecules in the crystal are held together by hydrogen bonds involving oxygen atoms from the SO₄ group, the NH₂ group of the thiourea molecules and probably S atoms of the same ligand. The distances that can be considered as hydrogen bonds are listed below, as well as the N—H—X angle, where X is O or S. In this calculation the theoretical positions for H atoms, as are given in Table I, were used. The distances and angles are:

N(1)—O(2 ^{vii})	2.74(2) Å	N(1)—H(1)—O(2 ^{vii})	165°
N(1)—O(3 ^{viii})	2.79(3)	N(1)—H(2)—O(3 ^{viii})	161
N(2)—O(4 ^{vii})	3.02(3)	N(2)—H(4)—O(4 ^{vii})	153
N(3)—S(3 ^{viii})	3.49(3)	N(3)—H(5)—S(3 ^{viii})	168
N(3)—O(2 ^{vii})	3.04(3)	N(3)—H(6)—O(2 ^{vii})	143
N(4)—O(4)	2.93(3)	N(4)—H(8)—O(4)	161
N(5)—O(1 ^{viii})	2.92(3)	N(5)—H(9)—O(1 ^{viii})	148
N(5)—O(3 ^{viii})	2.86(3)	N(5)—H(10)—O(3 ^{viii})	167
N(6)—O(1)	2.93(3)	N(6)—H(12)—O(1)	170
N(6)—O(2 ^{vii})	2.97(3)	N(6)—H(11)—O(2 ^{vii})	154

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