

The Crystal and Molecular Structure of Bis(thiocarbohydrazide-N,S)cadmium Dichloride

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The crystals of bis(thiocarbohydrazide-N,S)cadmium dichloride, $\text{Cd}(\text{S}=\text{C}(\text{NH}-\text{NH}_2)_2)_2\text{Cl}_2$, belong to monoclinic space group $P2_1/c$. The structure of the compound, determined from three-dimensional data, consists of trans-octahedral complexes where the ligand molecule forms pentatomic chelate rings with N,S as donor atoms. The chlorine atoms lie along the normal to the plane containing metal and organic molecules. The distances $\text{Cd}-\text{N}=2.34(3)$ and $\text{Cd}-\text{S}=2.59(1)$ Å are slightly shorter than those found in non-chelate cadmium complexes; $\text{Cd}-\text{Cl}=2.73(1)$ Å is very close to the sum of ionic radii. The configuration of the ligand, with one $-\text{NH}_2$ group turned towards $\text{S}=\text{C}$, repeats that of crystallized thiocarbohydrazide. Some differences, however, in distances, planarity, and angles between complexed or free ligand can be detected. The complexes are held together in layers parallel to (102) by weak hydrogen bonds $\text{NH}\dots\text{Cl}$.

Introduction

Thiocarbohydrazide, $\text{S}=\text{C}(\text{NH}-\text{NH}_2)_2$, forms complexes which can be assigned structures with pentatomic chelate rings with S,N as donor atoms. Such structures have been assumed to interpret IR spectra¹ and equilibria in solution.² In order to confirm these assumptions, the crystal structure of bis(thiocarbohydrazide-N,S)cadmium dichloride has now been determined.

Experimental Section

Preparation. Crystals of the compound, $\text{Cd}(\text{SC}(\text{NH}-\text{NH}_2)_2)_2\text{Cl}_2$ have been obtained by evaporation of aqueous solutions of thiocarbohydrazide and cadmium chloride in stoichiometric ratio. Small, colorless crystals were obtained.

Crystal Data. Compound: bis(thiocarbohydrazide-N,S)cadmium dichloride, $\text{Cd}(\text{SC}(\text{NH}-\text{NH}_2)_2)_2\text{Cl}_2$, F.W. 395.62.

Crystal class: Monoclinic prismatic.

Unit cell (radiation $\text{CuK}\alpha$, $\lambda=1.5418$ Å, from

rotation and Weissenberg photographs around [100] and [010]):

$$a=8.64(1), \quad b=5.78(1), \quad c=13.78(1) \text{ \AA};$$

$$\beta=119.5(3)^\circ, \quad V=599.3 \text{ \AA}^3, \quad Z=2;$$

$$D_x=2.19, \quad D_m=2.17 \text{ g.cm}^{-3};$$

$$\mu(\text{CuK}\alpha)=224.7 \text{ cm}^{-1};$$

Space group $P2_1/c$ (C_{2h} (5), No. 14).

Intensity Data. Intensities of reflections $0kl, \dots, 7kl$, and $h0l, \dots, h5l$ were recorded on integrating Weissenberg camera and then measured by a microdensitometer (934 independent reflections out of 1369 possible).

Calculations. Usual corrections were applied. Absorption corrections as for cylindrical specimens were introduced ($\mu R_{[100]}=1.01$, $\mu R_{[010]}=0.67$). Atomic form factors according to Cromer and Mann³ were used.

Table I. Fractional atomic coordinates (with e.s.d's $\times 10^4$).

	x	y	z
Cd	1.0000	.0000	.0000
S	.6777(8)	-.1555(5)	-.0947(5)
Cl	.9028(8)	.2939(6)	-.1755(5)
N(1)	.6763(27)	.2649(17)	-.0094(18)
N(2)	.8578(34)	.2493(22)	.0655(19)
N(3)	.4141(24)	.1321(16)	-.1499(17)
N(4)	.3324(34)	.3398(26)	-.1450(26)
C	.5812(23)	.1021(24)	-.0864(17)

Table II. Anisotropic thermal parameters^a (\AA^2).

	B_{11}	B_{22}	B_{33}	B_{12}	B_{13}	B_{23}
Cd	2.825	3.130	3.248	0.131	1.197	0.042
S	2.951	2.898	3.361	-0.004	1.365	-0.139
Cl	3.067	3.076	3.153	-0.023	1.074	0.345
N(1)	2.561	3.378	3.249	0.109	0.657	-0.143
N(2)	3.663	3.046	3.767	-0.398	1.594	-0.788
N(3)	3.604	2.758	3.919	-0.100	1.550	0.283
N(4)	2.445	3.951	4.252	-0.026	0.977	0.128
C	2.800	3.166	2.987	-0.476	1.022	0.126

^a In the last cycle for all the atoms the average shift, $|\Delta B_{ij}|_{\text{av}}$, and the maximum shift, $|\Delta B_{ij}|_{\text{max}}$, were:

$$|\Delta B_{ij}|_{\text{av}} = 0.049 \qquad |\Delta B_{ij}|_{\text{max}} = 0.173$$

(1) G. R. Burns, *Inorg. Chem.* 7, 277 (1968).

(2) A. Braibanti, F. Dallavalle and E. Leporati, *Inorg. Chim. Acta*, 3, 459 (1969).

Table IV. Atomic peak heights (e.Å⁻³), curvatures (e.Å⁻⁵), and e.s.d.'s.

		ρ	A_{hh}	A_{kk}	A_{ll}	A_{hk}	A_{kl}	A_{lk}
Cd	obs.	109.1	1028	1084	1017	12	460	-24
	calc.	110.5	1030	1087	1032	9	470	-25
S	obs.	31.1	304	319	280	-3	134	-15
	calc.	31.3	305	317	286	-2	137	-15
Cl	obs.	31.2	278	297	283	-1	121	11
	calc.	31.6	282	298	286	0	125	7
N(1)	obs.	10.7	79	98	79	-2	32	2
	calc.	11.0	78	100	83	-4	36	4
N(2)	obs.	9.7	71	78	86	-7	37	4
	calc.	9.8	73	79	88	-7	38	6
N(3)	obs.	10.6	99	84	85	-3	42	-7
	calc.	10.7	100	83	85	-4	43	-6
N(4)	obs.	8.2	70	61	57	0	29	-1
	calc.	8.2	69	62	59	1	30	-1
C	obs.	9.9	113	71	85	-10	45	4
	calc.	10.1	115	73	85	-7	48	0
e.s.d.'s		3	4	5	6	3	3	3

Table V. Main interatomic distances and angles (with e.s.d.'s).

Coordination around metal			
Cd-S	2.59(1) Å	S-Cd-N(2)	76.7(6)°
Cd-Cl	2.73(1)	S-Cd-Cl	89.1(2)
Cd-N(2)	2.34(3)	Cl-Cd-N(2)	88.1(5)
Thiocarbohydrazide molecule			
S-C	1.74(2)	Cd-S-C	95.8(7)
C-N(1)	1.35(2)	S-C-N(1)	121.5(17)
C-N(3)	1.28(3)	S-C-N(3)	119.2(13)
N(1)-N(2)	1.39(4)	C-N(1)-N(2)	124.3(17)
N(3)-N(4)	1.41(3)	C-N(3)-N(4)	120.4(18)
		N(1)-N(2)-Cd	111.5(14)
		N(1)-C-N(3)	119.2(17)
Hydrogen bonds			
N(2)-Cl ⁱ	3.23(2)	Cd-N(2)-Cl ⁱ	110.3(11)
N(3)-Cl ⁱⁱ	3.26(2)	N(1)-N(2)-Cl ⁱ	121.4(9)
N(4)-N(1 ⁱⁱⁱ)	3.15(3)	C-N(3)-Cl ⁱⁱⁱ	132.8(9)
		N(4)-N(3)-Cl ⁱⁱⁱ	106.7(16)
		N(3)-N(4)-N(1 ⁱⁱⁱ)	146.0(20)
		N(4)-N(1 ⁱⁱⁱ)-N(2 ⁱⁱⁱ)	86.5(13)
		N(4)-N(1 ⁱⁱⁱ)-Cl ⁱⁱⁱ	145.9(19)
Asymmetric units			
i	-x+2, -y+1, -z	ii	-x+1, y-1/2, -z-1/2
		iii	-x+1, -y+1, -z

The structure was solved by Patterson and Fourier methods and refined by differential syntheses.

Anisotropic temperature factors were introduced. (Final $R=9.7\%$). Hydrogen atoms were not identified in the difference Fourier map. The results of the structure determination are summarized in Tables I→V.

Discussion

The structure (Figure 1) consists of octahedral complexes with chelate molecules lying in the same plane and the chlorine atoms in the apical positions.

The chelating atoms (Figure 2) are N, of $-\text{NH}_2$, and S, thus confirming the assumptions on which were based the interpretations of IR spectra¹ and of

equilibria in solution.² Atoms N, S are in *trans* positions and all of them lie by symmetry conditions in the same plane as the metal atom. The chlorine atoms lie very close to the normal to that plane. The distance $\text{Cd}-\text{Cl}=2.73(1)$ Å is comparable with the sum of ionic radii (2.78 Å) and shows that the bond should be prevalently ionic. The organic ligand forms pentatomic chelate rings as expected; it is therefore in the same configuration, with one $-\text{NH}_2$ group turned toward $\text{S}=\text{C}$, as in the neutral molecule in crystals.⁴ This configuration is different from that found in thiocarbohydrazide dichloride⁵ where both $-\text{NH}_2$ groups are turned toward $\text{S}=\text{C}$.

(3) D. T. Cromer and J. B. Mann, *Acta Cryst.*, A24, 321 (1968).
 (4) A. Braibanti, A. Tiripicchio and M. Tiripicchio Camellini, *Acta Cryst.*, B25, 2286 (1969).
 (5) A. Braibanti, M. A. Pellinghelli, A. Tiripicchio and M. Tiripicchio Camellini, *Inorg. Chim. Acta* (in the press).

The distance Cd–N=2.34(3) Å is only slightly shorter than that found in chelates of hydrazinecarboxylic acid (Cd–N=2.38(3) Å,⁶ 2.40(2) Å⁷ and is equal to Cd–O in octahedral complexes (Cd–O=2.34(3) Å⁸). The distance Cd–S=2.59(1) Å is significantly shorter than that found in complexes of thiourea (Cd–S=2.638(4), 2.647(4) Å⁹); the shortening is probably due to chelation.

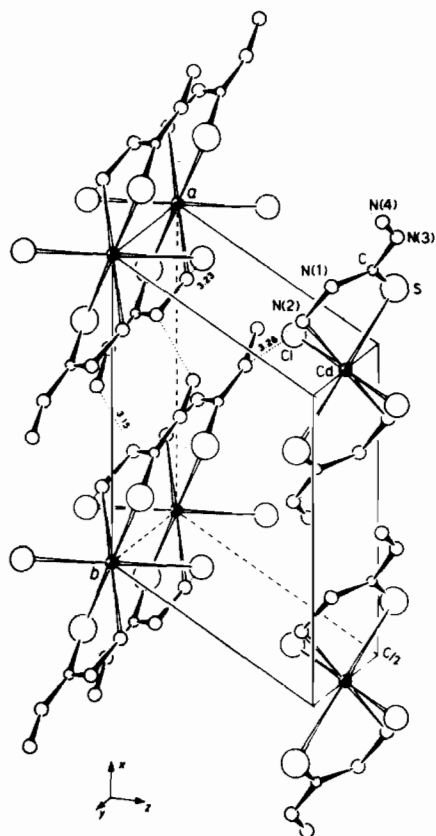


Figure 1. Clinographic projection of the structure of bis-(thiocarbohydrazide-N,S)cadmium dichloride.

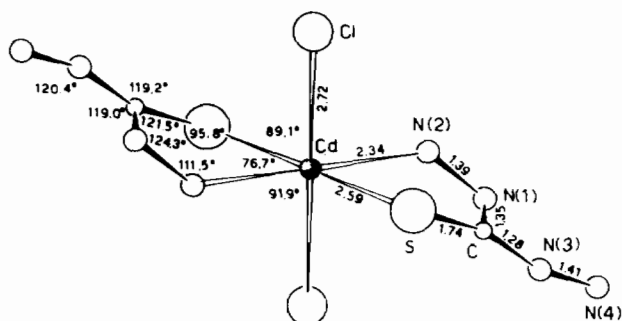


Figure 2. *Trans*-octahedral complex.

(6) A. Braibanti, A. Tiripicchio, A. M. Manotti Lanfredi and F. Bigoli, *Zeit. Kristallogr.*, 126, 307 (1968).

(7) A. Braibanti, A. M. Manotti Lanfredi, A. Tiripicchio and F. Bigoli, *Acta Cryst.*, B25, 100 (1969).

In the chelate ligand some differences with the isolated molecule can be detected. The distance C–S=1.74(2) Å is comparable with that in thiocarbonylhydrazide (C–S=1.724(10) Å⁴), in complexed thiourea (C–S=1.76(1) Å,⁹ 1.73(1) Å¹⁰ and in free thiourea (C–S=1.71(1) Å¹¹). Also the distances N(1)–N(2)=1.39(4) Å and N(3)–N(4)=1.41(3) Å are not different from one another and equal to that found in the free ligand (1.405(8) Å⁴). In any case they are shorter than those found in free (N–N=1.46 Å¹²) and complexed hydrazine (N–N=1.47(2)¹³). The distances C–N(1)=1.35(2) and C–N(3)=1.28(3) Å as compared with single bond C–N=1.47 Å indicate that they possess some double bond character, although in different amounts. They can be compared with values found in thiocarbonylhydrazide (C–N=1.327(7) Å⁴), in thiourea (C–N=1.33(1) Å¹¹), and in thiosemicarbazide (C–N=1.323(2) Å¹⁴). The angle S–C–N(1)=121.5(17)° is smaller than the corresponding angle in the free molecule.

Table VI. Planarity of thiocarbonylhydrazide molecule. Equation of plane $5.2772x + 2.4606y - 12.1516z - 4.3449 = 0$

	Δ	σ	
Cd	0.9323 Å	0.0000 Å	
*S	-0.0004	0.0075	
*C	0.0233	0.0247	
*N(1)	-0.0099	0.0264	
*N(3)	-0.0130	0.0245	
*N(2)	-0.0006	0.0297	
*N(4)	0.0073	0.0369	$\Sigma(\Delta/\sigma)^2 = 1.3542$

* Atoms lying in the plane.

The whole ligand molecule lies in one plane (Table VI). This aspect is again different from that found in the free ligand where both terminal –NH₂ groups are out of the plane of the thioureide group, N–CS–N.

The complexes are held together in the crystal by hydrogen bonds N(3)–H...Cl=3.26(2) Å and N(2)–H...Cl=3.23(2) Å, thus forming layers parallel to (102). Weak contacts N(1)...N(4)=3.15(3) Å join the layers to one another.

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