Table I. Calcination Products

Calcination time, 16 hours

Temp., °C.	Compounds Identified				
	$Cu(NO_3)_2 \cdot 3H_2O$-$2CrO_3$	$Cu(NO_3)_2 \cdot 3H_2O$-Cr_2O_3	$Zn(NO_3)_2 \cdot 6H_2O$-$2CrO_3$	$Zn(NO_3)_2 \cdot 6H_2O$-Cr_2O_3	$1Cu(NO_3)_2 \cdot 3H_2O$ $1Zn(NO_3)_2 \cdot 6H_2O$-$2CrO_3$
260	Cr_3O_8[a] CrO_3 CuO?	Cr_2O_3, hexagonal CuO	?	Cr_2O_3, hexagonal ZnO ?	ZnO CuO ?
480	$CuCr_2O_4$ (50%)[b] Cr_2O_3, hexagonal (25%) CuO (25%)	Cr_2O_3, hexagonal (45%) CuO (45%) $CuCr_2O_4$ (10%)	$ZnCr_2O_4$ (95%) Cr_2O_3, hexagonal ZnO	$ZnCr_2O_4$ (40%) ZnO (30%) Cr_2O_3, hexagonal (30%)	$ZnCr_2O_4$ $CuCr_2O_4$ CuO ZnO
650	$CuCr_2O_4$ (70%) Cr_2O_3, hexagonal (15%) CuO (15%)	Cr_2O_3, hexagonal (35%) CuO (35%) $CuCr_2O_4$ (30%)	$ZnCr_2O_4$ (97%) Cr_2O_3, hexagonal ZnO	$ZnCr_2O_4$ (85%) Cr_2O_3, hexagonal ZnO	$ZnCr_2O_4$ CuO ZnO $CuCr_2O_4$
870	$CuCr_2O_4$ (>90%) Cr_2O_3, hexagonal CuO, weak $Cu_2Cr_2O_4$, weak	$CuCr_2O_4$ (70%) Cr_2O_3, hexagonal CuO $Cu_2Cr_2O_4$	$ZnCr_2O_4$	$ZnCr_2O_4$ (>90%) Cr_2O_3, hexagonal ZnO	$ZnCr_2O_4$ CuO ZnO
1200	$Cu_2Cr_2O_4$ Cr_2O_3, hexagonal $CuCr_2O_4$ (>5%)	$Cu_2Cr_2O_4$ Cr_2O_3, hexagonal $CuCr_2O_4$ (>5%)	$ZnCr_2O_4$	$ZnCr_2O_4$ (>90%) Cr_2O_3, hexagonal ZnO	$ZnCr_2O_4$ CuO

[a] The compounds are listed in the order of their concentration as per x-ray diffraction. [b] Percentage of constituent as per x-ray diffraction.

quently enhanced when one of the reactants is undergoing decomposition.

The conversion of cupric chromite to cuprous chromite is demonstrated at 870° and 1200° C. in the copper series. Zinc chromite formation is complete at 870° C. in the chromium trioxide mixture.

Except for the cupric nitrate trihydrate—chromium(ic) oxide series—there are unexplained peaks in the x-ray diffraction patterns of the samples calcined at 260° C. These unexplained peaks may be associated with either unknown chromium-oxygen compounds or unknown reaction products.

The affinity of zinc as compared with copper for chromite formation is demonstrated in the study of the mixture consisting of 1 mole of cupric nitrate trihydrate, 1 mole of zinc nitrate hexahydrate, and 2 moles of chromium trioxide. This affinity of zinc over copper is most likely related to relative ionization potentials, 9.36 volts for zinc and 7.71 volts for copper (2).

LITERATURE CITED

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CORRECTION

In the article "High Temperature *PVT* Properties of Sodium, Potassium, and Cesium" [J. CHEM. ENG. DATA 11, 309 (1966)] in Table I, Experiment 17, the second temperature should be 2516.1 (not 2416.1); in Table II, Experiment 8, the 12th figure for specific volume should be 5.1371 (not 5.1293); in Table III, Experiment 35, the fourth temperature should be 2492.6 (not 2492.0) and ~~the~~ corresponding pressure should be 311.00 (not 211.00). The subcaption of Figure 6 should read: ■ Tepper; ● this work. In the first reference of Literature Cited the page number should be 315.

In the article "High Temperature Vapor Pressures of Sodium, Potassium, and Cesium" [J. CHEM. ENG. DATA 11, 315 (1966)] on page 320, next to the last line in the second paragraph, first column, delete "with the contact." In reference 10 the volume number should be 11.

In the article "High Temperature Specific Volumes of Liquid Sodium, Potassium, and Cesium" [J. CHEM. ENG. DATA 11, 320 (1966)] reference 9 should read: J. CHEM. ENG. DATA 11, 309 (1966). Reference 10 should read: *Ibid.*, p. 315.

CORRECTION

In the article "Isobaric Vapor-Liquid Equilibrium of *n*-Heptane–*n*-Butanol System" [J. CHEM. ENG. DATA 11, 147 (1966)], the following corrections should be made:

Equation 1 should be $\gamma_i = \frac{y_i P}{x_i p_i^0} \theta_i$

The caption for Figure 3 should read, α_1 *vs.* composition. Page 148, first column, line 13 should read, α_1, are calculated from experimental data with the help of. Page 148, second column, line 3 should read, measurement of temperature, $E(T)$; pressure, $E(P)$; and. Page 149, first column, line 21 should read, $D - J = 4.75 < 10$. Page 149, second column, line 10 should read, α_1 = dimensionless thermodynamic function in Equations 5 and 7.