Reply to Comments by van Noort on *J. Chem. Eng. Data* 2001, *46,* 286−298

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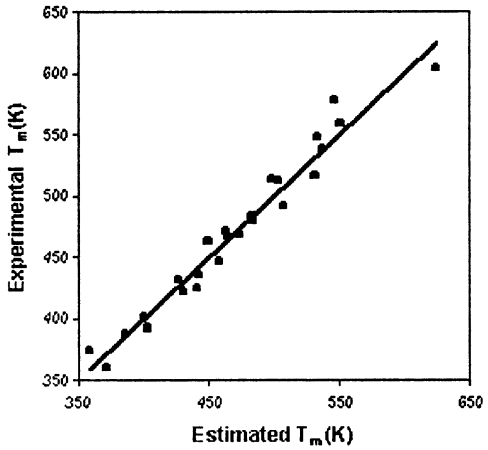

Figure 1. Experimental versus estimated melting temperatures for PCDDs. The line is the $y = x$ line.

The correlation suggested by van Noort[1] in his Comment on our paper[2] is an improvement of the Simamura correlation.[3] Unfortunately, there is not sufficient information in our paper[2] to compare both approaches. The uncertainties cited in ref 2 are not statistical but an expert's estimation of the uncertainties of unknown melting points. The average deviations of experimental and estimated values from our correlation are 10 K for 61 chlorine-substituted dioxins (PCDDs) and 15 K for 59 chlorine-substituted dibenzofurans (PCDFs). The corresponding graphs are presented as Figures 1 and 2. So, our correlation is considerably better for PCDDs and the same for PCDFs.

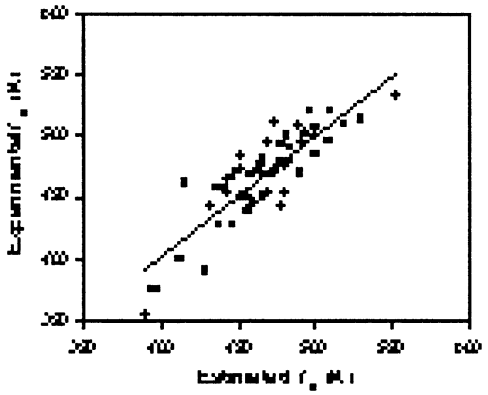

Figure 2. Experimental versus estimated melting temperatures for PCDFs. The line is the $y = x$ line.

Literature Cited

(1) van Noort, P. C. M. Comments on "Thermodynamic Properties of Dibenzo-*p*-dioxin, Dibenzofuran, and Their Polychlorinated Derivatives in the Gaseous and Condensed Phases. 2. Thermodynamic Properties of Condensed Compounds". *J. Chem. Eng. Data* **2002**, *47,* 1558.

(2) Iorish, V. S.; Dorofeeva, O. V.; Moiseeva, N. F. Thermodynamic Properties of Dibenzo-*p*-dioxin, Dibenzofuran, and Their Polychlorinated Derivatives in the Gaseous and Condensed Phases. 2. Thermodynamic Properties of Condensed Compounds. *J. Chem. Eng. Data* **2001**, *46*, 286−298.

(3) Simamura, P.; Yalkowsky, S. H. Group Contribution Method for Predicting the Melting Points and Boiling Points of Aromatic Compounds. *Ind. Eng. Chem. Res.* **1994**, *33*, 1405−1409.

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