

## Corrections

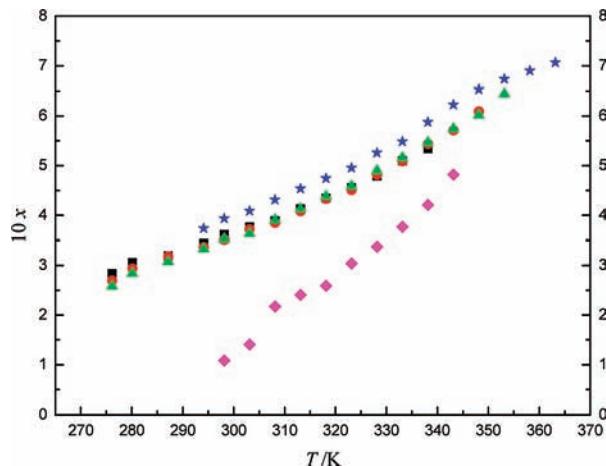
**Solubility of Triethylenediamine in Methanol, Ethanol, Isopropanol, 1-Butanol, and Ethyl Acetate between 276.15 K and 363.15 K.** Dishun Zhao,\* Erhong Duan, Xiaoli Zhang, Juan Zhang, and Jinlong Wang, *J. Chem. Eng. Data* 2007, 52, 1483–1485.

Page 1484. The data in Table 1 and Figure 2 were the solubility of TEDA in  $\text{mol}\cdot\text{L}^{-1}$  and not the mole fraction solubility  $x$ . The mole fraction solubilities  $x^{\text{exptl}}$  and  $x^{\text{calcd}}$  are now given below in Table 1 and Figure 1. Because the wrong values of composition were given, the values in Table 2 were also incorrect. The correct values are given below in Table 2.

Page 1485. The conclusion was based on the solubility as molality in  $\text{mol}\cdot\text{L}^{-1}$  and not mole fraction  $x$ . Hence, some of the sentences in the conclusion have to be revised. The sentences stating “(2) The solubility of TEDA...” should be replaced by “(2) The mole fraction solubility of TEDA in 1-butanol is

**Table 1. Mole Fraction Solubility,  $x$ , of Triethylenediamine in Pure Solvents**

$T/\text{K}$	$10x^{\text{exptl}}$	$10x^{\text{calcd}}$	$T/\text{K}$	$10x^{\text{exptl}}$	$10x^{\text{calcd}}$
Methanol					
276.15	2.84	2.89	280.15	3.05	3.00
287.15	3.19	3.20	294.15	3.44	3.43
298.15	3.61	3.57	303.15	3.77	3.75
308.15	3.89	3.93	313.15	4.13	4.13
318.15	4.34	4.35	323.15	4.56	4.57
328.15	4.79	4.81	333.15	5.10	5.07
338.15	5.34	5.34			
Ethanol					
276.15	2.70	2.80	280.15	2.94	2.91
287.15	3.18	3.12	294.15	3.37	3.35
298.15	3.52	3.49	303.15	3.72	3.68
308.15	3.85	3.88	313.15	4.08	4.09
318.15	4.33	4.32	323.15	4.51	4.57
328.15	4.82	4.83	333.15	5.09	5.11
338.15	5.43	5.41	343.15	5.71	5.73
348.15	6.09	6.07			
Isopropanol					
276.15	2.59	2.66	280.15	2.84	2.80
287.15	3.07	3.05	294.15	3.33	3.31
298.15	3.53	3.47	303.15	3.63	3.68
308.15	3.90	3.90	313.15	4.12	4.13
318.15	4.38	4.37	323.15	4.59	4.62
328.15	4.89	4.88	333.15	5.15	5.16
338.15	5.46	5.45	343.15	5.74	5.75
348.15	6.02	6.06	353.15	6.44	6.40
1-Butanol					
294.15	3.73	3.69	298.15	3.93	3.87
303.15	4.09	4.09	308.15	4.31	4.32
313.15	4.54	4.56	318.15	4.74	4.81
323.15	4.95	5.06	328.15	5.26	5.31
333.15	5.48	5.57	338.15	5.87	5.83
343.15	6.22	6.10	348.15	6.53	6.37
353.15	6.73	6.65	358.15	6.90	6.93
363.15	7.07	7.21			
Ethyl Acetate					
298.15	1.08	1.31	303.15	1.41	1.60
308.15	2.17	1.91	313.15	2.40	2.26
318.15	2.59	2.63	323.15	3.04	3.03
328.15	3.37	3.44	333.15	3.77	3.86
338.15	4.21	4.28	343.15	4.82	4.71



**Figure 1.** Mole fraction solubility of TEDA  $x$  in different solvents: ■, methanol; ●, ethanol; ▲, isopropanol; ★, 1-butanol; ◆, ethyl acetate.

**Table 2. Parameters of Equation 1 for Triethylenediamine in Pure Solvents**

solvent	$A$	$B$	$C$	$10^3$ RMSD
methanol	-48.26	1468	7.829	9.6
ethanol	-58.83	1889	9.434	13
isopropanol	-28.32	430.2	4.936	30
1-butanol	4.017	-1000	0.1209	47
ethyl acetate	243.9	-13948	-34.55	112

higher than in methanol, ethanol, isopropanol, and ethyl acetate. The mole fraction solubility of TEDA in ethyl acetate is the lowest. (3) There were similar mole fraction solubilities of triethylenediamine in methanol, ethanol, and isopropanol at the same temperature”.

**Note Added after ASAP Publication:** This paper was published ASAP on August 9, 2008. Changes were made in Tables 1 and 2 and in Figure 1. The revised paper was reposted April 1, 2009. Additional changes were made to Table 2. The revised paper was reposted June 12, 2009.

JE8005375

10.1021/je8005375

Published on Web 08/09/2008