Apparent Molar Volumes and Standard Partial Molar Volumes of Aqueous Sodium Phosphate Salts at Elevated Temperatures. Genna E. Woolston, Liliana N. Trevani, and Peter R. Tremaine,* *J. Chem. Eng. Data* **2008,** *53,* 1728–1737.

Page 1731. The data entries in Table 2 have been corrected for typographical and minor round-off errors. No changes to the fitted parameters in Tables 4 and 5 are required. The corrected version of Table 2 is presented below.

Table 2. Experimental Relative Densities, $\Delta\rho = (\rho - \rho_w^*)$, and Apparent Molar Volumes, V_ϕ (exp), for Na_2HPO_4 (aq) from $T = 473$ K to 570 K and $p = 15$ MPa[a]

T / K	p / MPa	m / mol·kg^{-1}	$\Delta\rho$ / g·cm^{-3}	ρ[b] / g·cm^{-3}	V_ϕ (exp) / cm^3·mol^{-1}
		$T = (473.11 \pm 0.02)$ K; $p = 14.99$ MPa			
473.12	14.99	0.09825	0.01348	0.88809	-16.81 ± 0.70
473.12	14.99	0.09920	0.01373	0.88833	-18.40 ± 0.81
473.10	14.99	0.20025	0.02684	0.90146	-12.50 ± 0.50
473.11	14.99	0.24952	0.03322	0.90784	-11.32 ± 0.48
473.09	14.99	0.29967	0.03918	0.91382	-8.25 ± 0.43
473.12	14.99	0.34155	0.04444	0.91904	-7.39 ± 0.38
473.12	14.99	0.39991	0.05155	0.92615	-5.84 ± 0.39
473.08	14.99	0.49637	0.06365	0.93830	-4.96 ± 0.35
		$T = (522.85 \pm 0.02)$ K; $p = 15.19$ MPa			
522.86	15.17	0.09825	0.01467	0.82642	-50.86 ± 1.36
522.86	15.16	0.09877	0.01474	0.82649	-50.72 ± 1.36
522.85	15.19	0.09920	0.01468	0.82647	-48.80 ± 1.29
522.84	15.18	0.20025	0.02845	0.84024	-39.32 ± 0.69
522.83	15.19	0.24952	0.03521	0.84702	-37.60 ± 0.70
522.84	15.21	0.29967	0.04159	0.85342	-33.99 ± 0.65
522.84	15.20	0.34155	0.04726	0.85907	-33.16 ± 0.55
522.83	15.21	0.39991	0.05478	0.86661	-30.88 ± 0.52
522.89	15.20	0.49637	0.06669	0.87843	-26.81 ± 0.44
		$T = (548.19 \pm 0.04)$ K; $p = 15.03$ MPa			
548.19	15.01	0.09825	0.01590	0.78813	-85.74 ± 0.85
548.18	15.00	0.09877	0.01599	0.78824	-85.82 ± 1.17
548.15	15.00	0.09920	0.01612	0.78842	-86.85 ± 0.83
548.15	15.01	0.20025	0.03048	0.80279	-68.64 ± 0.58
548.22	15.06	0.24952	0.03713	0.80938	-62.66 ± 0.49
548.16	15.03	0.29967	0.04402	0.81634	-59.10 ± 0.53
548.19	15.02	0.34155	0.04969	0.82194	-56.48 ± 0.44
548.24	15.07	0.39991	0.05719	0.82943	-52.12 ± 0.42
548.24	15.06	0.49637	0.06932	0.84156	-46.20 ± 0.38
		$T = (570.50 \pm 0.05)$ K; $p = 15.05$ MPa			
570.56	15.04	0.09825	0.01706	0.74811	-127.75 ± 1.07
570.46	15.03	0.09877	0.01713	0.74838	-127.28 ± 1.39
570.50	15.04	0.09920	0.01729	0.74846	-128.88 ± 1.22
570.58	15.05	0.20025	0.03261	0.76364	-105.79 ± 0.76
570.47	15.06	0.24952	0.03975	0.77103	-98.44 ± 0.72
570.49	15.06	0.29967	0.04672	0.77796	-91.58 ± 0.64
570.50	15.04	0.34155	0.05269	0.78388	-88.05 ± 0.57
570.45	15.07	0.39991	0.06056	0.79188	-82.21 ± 0.58
570.46	15.07	0.49637	0.07320	0.80451	-74.19 ± 0.52

[a] The number of significant digits reported for m and V_ϕ (exp) exceeds the standard error by one to ensure that the fitted equations can be reproduced. [b] Calculated from $\Delta\rho = \rho - \rho_w^*$ with ρ_w^*, the density of water, obtained from Hill's equation of state.[16]

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