

Preliminary communication

NEW SYNTHETIC PATHWAYS IN DICYCLOPENTADIENYLTANTALUM CHEMISTRY

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(Received September 8th, 1978)

Summary

The dihalides $[\text{Ta}(\eta\text{-C}_5\text{H}_5)_2\text{X}_2]$ ($\text{X} = \text{Cl}, \text{Br}$) are prepared in 80% yield by treatment of TaX_5 with $\text{Sn}(\text{C}_5\text{H}_5)\text{-n-Bu}_3$. The trihydride $[\text{Ta}(\eta\text{-C}_5\text{H}_5)_2\text{H}_3]$ is formed in 42% yield from $[\text{Ta}(\eta\text{-C}_5\text{H}_5)_2\text{Cl}_2]$ and $[\text{NaAlH}_2(\text{OCH}_2\text{CH}_2\text{OCH}_3)_2]$. The trihydride with butyllithium gives a tantalum-lithio derivative which reacts with benzyl chloride giving $[\text{Ta}(\eta\text{-C}_5\text{H}_5)_2(\text{CH}_2\text{Ph})\text{Cl}]$.

Trialkyl(cyclopentadienyl)tin compounds have been found to provide convenient synthetic routes to the new monocyclopentadienyl halides of niobium and tantalum $[\text{M}(\eta\text{-C}_5\text{H}_5)\text{X}_4]$ ($\text{M} = \text{Nb}, \text{Ta}; \text{X} = \text{Cl}, \text{Br}$) [1, 2]. Also, an improved route to dicyclopentadienyltantalum dichloride has been recently described by treating a toluene suspension of $[\text{Ta}(\eta\text{-C}_5\text{H}_5)\text{Cl}_4]$ with $[\text{AlEtCl}_2]_2$ [3].

Here we report that trialkyl(cyclopentadienyl)tin compounds can provide direct access to dicyclopentadienyltantalum complexes in very high yields. When dichloromethane suspensions of the tantalum pentahalides: TaX_5 ($\text{X} = \text{Cl}, \text{Br}$) are treated with 3 equivalents of $\text{Sn}(\text{C}_5\text{H}_5)\text{-n-Bu}_3$ at room temperature deep green solutions are obtained. Addition of toluene and concentration of these solutions give green crystals of $[\text{Ta}(\eta\text{-C}_5\text{H}_5)_2\text{X}_2]$ (I, $\text{X} = \text{Cl}$, yield 87%; II, $\text{X} = \text{Br}$, yield 84%). The yield of compound I is considerably higher than in the original method [4] and quantities of I as high as 35 g may be conveniently prepared in a single experiment.

Treatment of a suspension of $[\text{Ta}(\eta\text{-C}_5\text{H}_5)\text{X}_2]$ in toluene with $[\text{NaAlH}_2(\text{OCH}_2\text{-CH}_2\text{OCH}_3)_2]$ gave in 42% yield the trihydride $[\text{Ta}(\eta\text{-C}_5\text{H}_5)_2\text{H}_3]$ (III). In a typical experiment 7.2 g of $[\text{Ta}(\eta\text{-C}_5\text{H}_5)_2\text{Cl}_2]$ suspended in 500 cm³ of dry toluene at 0°C were treated dropwise by a benzene solution containing 12.2 g of $[\text{NaAlH}_2(\text{OCH}_2\text{CH}_2\text{OCH}_3)_2]$. The reaction mixture was then hydrolysed. Sublimation of the reddish solid obtained, at 105°C under 10⁻³ mmHg gave 2.5 g of white crystals of III. This reaction represents a reliable and reproducible method of synthesising the elusive $[\text{Ta}(\eta\text{-C}_5\text{H}_5)_2\text{H}_3]$ [5].

Treatment of the trihydride III with butadiene in toluene at 80°C for 6 h

gives green crystals of the previously described [4] (1-methylallyl) dicyclopentadienyltantalum (IV) in 56% yield.

When toluene solutions of III are treated with 1 equivalent or an excess of *n*-butyllithium, an orange solution is formed which deposits orange-yellow needles (V) after a few minutes. The crystals of V are extremely reactive towards oxygen and moisture. Hydrolysis of V reforms III quantitatively. Compound V is insoluble or decomposed in most organic solvents. The infrared spectrum of V shows a band characteristic of cyclopentadienyl groups and a broad band at 1590 cm^{-1} which can be assigned to $\nu(\text{Ta}-\text{H})$. We assume V is a tantalum-lithio derivative analogous in type to the compound $[\text{W}(\eta\text{-C}_5\text{H}_5)_2\text{HLi}]_4$ [6].

The presence of at least one tantalum-lithium bond in V can be deduced on the basis of the reaction of V with benzyl chloride. Treatment of a toluene suspension of V at -80°C with an excess of benzyl chloride gives black crystals of the new compound $[\text{Ta}(\eta\text{-C}_5\text{H}_5)_2(\text{CH}_2\text{Ph})\text{Cl}]$ (VI) in 66% yield. Compound VI has been characterized by elemental analysis (Found C, 47.95; H, 4.1; Cl, 7.95. $\text{C}_{17}\text{H}_{17}\text{ClTa}$ calcd.: C, 47.65; H, 3.9; Cl, 8.1%), infrared and the ESR spectrum ($\langle g \rangle$ 1.965, A_{iso} 115 G).

We thank the Royal Society and the C.N.R.S. for an exchange fellowship (to J.J.E.M.).

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