

Luminescence of Vanadium and Rare Earth Ions in Alkaline Earth Sulfates

G. BLASSE AND H. G. PIETERSEN

Solid State Department, Physical Laboratory, State University, Utrecht, The Netherlands

Received December 18, 1973

The luminescence of samples $\text{MeSO}_4\text{-V, RE (Me = Mg, Ca, Ba)}$ depends strongly on the nature of the Me ions. The amount of association of the V^{5+} and RE^{3+} ions can be estimated from the measured quantum efficiencies.

Recently the luminescence of vanadium and rare earth (RE) ions in calcium sulfate has been investigated in this laboratory (1, 2). It appeared that V^{5+} and RE^{3+} ions in CaSO_4 occur as associates in which the energy transfer from the vanadate group to the RE^{3+} ions could be studied. Among others it was found for the first time that efficient energy transfer from the vanadate group to the Tb^{3+} ion occurs.

In this note we report a similar investigation in 2 other host lattices, viz MgSO_4 and BaSO_4 . The Mg^{2+} ion is smaller than the Ca^{2+} ion, the Ba^{2+} ion is larger. The experimental conditions were the same as described before (1, 2). Sample compositions were either $\text{MS}_{0.999}\text{V}_{0.001}\text{O}_4$ or $\text{M}_{0.999}\text{RE}_{0.001}\text{S}_{0.999}\text{V}_{0.001}\text{O}_4$ ($\text{M} = \text{Mg}$ or Ba).

The luminescence of the BaSO_4 samples was weak at room temperature and at liquid nitrogen temperature except for the blue emission of $\text{BaSO}_4\text{-V, Y}$ (with quenching temperature, T_q , about 150°K) and the yellow emission of $\text{BaSO}_4\text{-V}$ (with T_q about 80°K). In comparison with the CaSO_4 samples (1, 2) the intensities and T_q are considerably lower. The MgSO_4 samples show much higher intensities and T_q than CaSO_4 and BaSO_4 samples and were, therefore, investigated in more detail. $\text{MgSO}_4\text{-V}$ as well as $\text{MgSO}_4\text{-V, Y}$ show a yellow emission under shortwave ultraviolet excitation with T_q about 400°K .

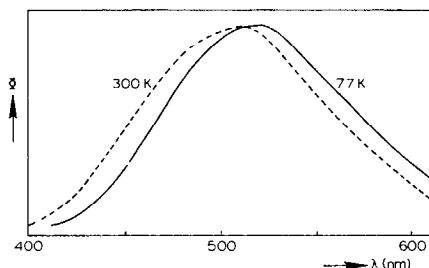


FIG. 1. Spectral energy distribution of the emission of $\text{Mg}_{0.999}\text{Y}_{0.001}\text{S}_{0.999}\text{V}_{0.001}\text{O}_4$ at 77 and 300°K for 240 nm excitation. Φ is the spectral radiant power in arbitrary units.

The intensity of the luminescence increases, if the sample temperature is lowered from room to liquid nitrogen temperature. The spectral energy distribution of this yellow emission is slightly dependent on sample temperature and excitation wavelength. Figure 1 shows the emission spectra of $\text{MgSO}_4\text{-V, Y}$ at 300 and 77°K . Replotted on an energy scale these emission curves are fairly symmetrical with half width about 4.500 cm^{-1} . In the case of CaSO_4 this value is about 5.500 cm^{-1} .

Samples $\text{MgSO}_4\text{-V, RE (RE = Sm, Tb, Dy)}$ show vanadate as well as characteristic rare earth emission upon excitation into the strong absorption band of the vanadate group, which peaks at about 250 nm. This is different from the analogous CaSO_4 samples where only rare earth emission is observed. Table

TABLE I
RATIO OF THE QUANTUM EFFICIENCY OF THE RE^{3+} EMISSION AND THE TOTAL QUANTUM EFFICIENCY OF SAMPLES $Mg_{0.999}RE_{0.001}SO_{0.999}V_{0.001}O_4$ AT $77^\circ K$ FOR EXCITATION INTO THE VANADATE GROUP (THE IONIC RADIUS OF RE^{3+} (ACCORDING TO REF. 4) IS ALSO GIVEN)

	Dy ³⁺	Tb ³⁺	Sm ³⁺
$q_{RE^{3+}}$	0.3	0.2 ⁵	0.1 ⁵
q_{total}			
$r_{RE^{3+}}(\text{\AA})$	0.91	0.92	0.96

I gives the ratio between the quanta emitted by the rare earth ions in $MgSO_4-V, RE$ and the total number of quanta emitted after excitation into the vanadate group at $77^\circ K$.

From these results the following conclusions may be drawn.

a. The quenching temperature of the vanadate emission in the different sulfates behaves as expected from the model given before (1, 2). Figure 2 gives a survey of the values of T_q plotted as a function of the ionic radius of the alkaline earth ion in the host lattice. Note that T_q decreases, if this ionic radius increases. A small ionic radius gives a stiff host lattice and, therefore, a high value of T_q . Note also that the variation of T_q is different for MSO_4-V and MSO_4-V, Y . This indicates that

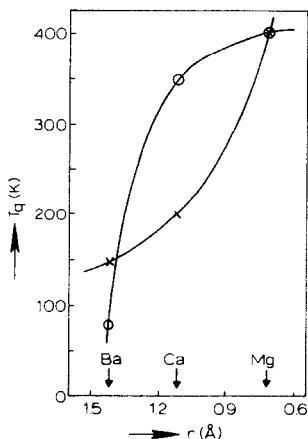


FIG. 2. Quenching temperature T_q of the vanadate emission of $MS_{0.999}V_{0.001}O_4$ (circles) and $M_{0.999}Y_{0.001}S_{0.999}V_{0.001}O_4$ (crosses) as a function of the ionic radius of M^{2+} (according to Ref. 4).

the luminescent centres involved are indeed different. For MSO_4-V, Y only the regular host lattice ions vary. For MSO_4-V the nature of the compensating defect may also vary. If this would be an interstitial alkaline earth ion as supposed in Ref. 2 it is immediately clear that the nature of the compensating defect varies also. These results agree with the observation made by Kotera et al. (3) that vanadium ions luminesce in $MgSO_4$ at room temperature but not so in $BaSO_4$.

b. From the present data we can estimate the amount of association of V^{5+} and RE^{3+} in $MgSO_4$. In $CaSO_4$ this is very near to 100%. Since energy transfer from the vanadate group to RE^{3+} ions is restricted to nearest neighbours, i.e., associated vanadate groups do not luminesce, but transfer their excitation energy to the near RE^{3+} ion, whereas unassociated vanadate groups emit their own luminescence, the amount of association in $MgSO_4$ at a V and RE concentration of 0.1% is about 20% (see Table I). It tends to be lower for larger RE^{3+} ions and higher for smaller RE^{3+} ions. The fact that association is incomplete in $MgSO_4$ in contradistinction to $CaSO_4$ may be due to the fact that in $MgSO_4$ the V^{5+} and RE^{3+} ions are considerably larger than the host lattice ions (ionic radii V^{5+} 0.35 Å, S^{6+} 0.12 Å, RE^{3+} 0.9–1.0 Å, Mg^{2+} 0.72 Å (4)), whereas in $CaSO_4$ the RE^{3+} ions fit into the Ca^{2+} sites (Ca^{2+} 1.12 Å). The association of 2 too large ions is probably unfavourable. The fact that the amount of association tends to be lower for larger RE^{3+} ions agrees with this assumption.

c. Also in $MgSO_4$ energy transfer from the vanadate group to the Tb^{3+} ion occurs.

d. The slight dependence of the vanadate emission spectra upon excitation wavelength may be due to the presence of crystallographically different V^{5+} ions as argued in Ref. 2.

References

1. W. HORDIJK AND G. BLASSE, *J. Luminescence* **6**, 137 (1973).
2. W. T. DRAAI AND G. BLASSE, *Phys. Status Solidi*, (a) **21**, 569 (1974); (b) *Chem. Phys. Lett.* **25**, 167 (1974).
3. Y. KOTERA, M. YONEMURA, AND T. SEKINE, *J. Electrochem. Soc.* **108**, 540 (1961).
4. R. D. SHANNON AND C. T. PREWITT, *Acta Cryst.* **B25**, 925 (1969).