Corrigenda

Gas-phase Rearrangements of Deprotonated Ketoximes, Ketoxime Ethers, and Aldoximes. A Negative-ion Beckmann Rearrangement

Gregory W. Adams, John H. Bowie, and Roger N. Hayes
J. Chem. Soc., Perkin Trans. 2, 1989, 2159

Page 2159: *Add* the following footnote, which refers to the line: 'We suggest that loss of water occurs from (**2**) (Scheme 1) by a negative-ion Beckmann rearrangement, with methyl-anion migration proceeding from N to form ion complex (**3**).§' The Editor would like to apologise for this omission.

§ The alternative mechanism involving direct attack of the methylene anion at nitrogen [a Neber type arrangement (D. J. Cram, 'Fundamentals of Carbanion Chemistry,' Academic Press, New York and London, 1965, p. 249)] seems a less likely possibility and should, in any case, ultimately produce the same product ion.

A Theoretical Study of the Proton Affinities of Water, Alcohols, and Ethers: Absolute *versus* Relative Basicities

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J. Chem. Soc., Perkin Trans. 2, 1990, 565

Page 566, *delete* Table 1 and *replace* with the following:

Table 1. Experimental and calculated proton affinities for water and selected alcohols and ethers.

ROR′			E_{pa}/kcal mol^{-1}			
R	R′	Number	Experimental[a]	STO 3-21G	MNDO	AM1
H	H	(**1**)	167.3	191.6	172.1	164.5
Me	H	(**2**)	182.5	204.8	175.0	171.9
Et	H	(**3**)	187.8	208.7	177.1	178.8
Pr	H	(**4**)	189.6	210.3	178.2	179.0
Bu	H	(**5**)	190.2	211.8	178.8	179.5
Pr^{i}	H	(**6**)	192.0[b]	214.0	180.2	182.8
Bu^{t}	H	(**7**)	193.4[b]	215.5	181.8	186.5
CF_3CH_2	H	(**8**)	169.3	186.8	155.8	157.9
Me	Me	(**9**)	191.1	213.1	177.4	177.4
Et	Me	(**10**)	195.2	216.7	179.4	183.4
Pr	Me	(**11**)	196.1			
Bu	Me	(**12**)	197.3			
Pr^{i}	Me	(**13**)	198.6	220.4	181.1	187.8
Bu^{t}	Me	(**14**)	201.1	223.0	183.6	190.9
1-Adamantyl	Me	(**15**)	206.0			
Et	Et	(**16**)	198.5	219.8	181.0	188.7
Pr^{i}	Et	(**17**)	202.3			
Bu^{t}	Et	(**18**)	205.1			
Pr	Pr	(**19**)	201.1			
Pr^{i}	Pr^{i}	(**20**)	204.8			
Bu^{t}	Bu^{t}	(**21**)	212.0	208.5[c]	188.7	200.9
Cyclic ethers						
$(CH_2)_2O$		(**22**)	186.4	203.8	179.4	177.2
$(CH_2)_3O$		(**23**)	193.2	220.4	183.3	187.0
$(CH_2)_4O$		(**24**)	197.6	223.3	182.0	185.9
$(CH_2)_5O$		(**25**)	198.5	219.1	182.9	186.5

[a] Experimental proton affinities based on $E_{pa}(NH_3)$ = 204.0 kcal mol^{-1} (ref. 2). [b] From ref. 3(*c*). [c] See the text.