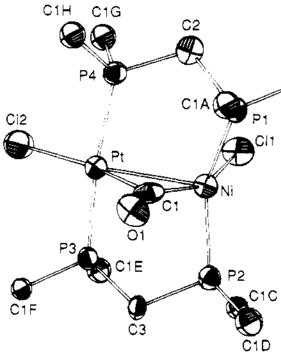

Figure 1. ORTEP of drawing of $[PtNi(\mu\text{-}CO)(\mu\text{-dppm})_2Cl_2]$ showing 50% probability thermal ellopsoids. For clarity, only phenyl carbons bonded to phosphorus atoms have been included. Selected bond lengths (Å): Pt-Ni, 2.689 (2); Pt-Cl2, 2.374 (5); Pt-C1, 2.03 (1); Pt-P3, 2.306 (4); Pt-P4, 2.323 (4); Ni-Cl1, 2.274 (4); Ni-C1, 1.77 (2); Ni-P1, 2.207 (4); Ni-P2, 2.209 (4); C1-O1, 1.22 (2). Bond angles (deg): P3-Pt-P4, 172.8 (1); P3-Pt-Cl2, 91.8 (2); P3-Pt-C1, 87.5 (4); P4-Pt-Cl2, 91.3 (2); P4-Pt-C1, 91.6 (4); Cl2-Pt-C1, 161.6 (5); P1-Ni-P2, 146.0 (2); P1-Ni-Cl1, 92.6(2); P1-Ni-C1, 87.1 (5); P2-Ni-Cl1, 105.8 (2); P2-Ni-C1, 90.5 (5); Cl1-Ni-C1, 149.9 (5); Pt-C1-Ni, 89.8 (6); Pt-C1-O1, 125 (1); Ni-C1-O1, 145 (1).

150.1°, respectively. The Pt-Pt distances in $Pt_2(\mu\text{-dppm})_2(\mu\text{-}HgCl_2)Cl_2$ and $Pt_2(\mu\text{-dppm})_2Cl_2$ are 2.712 and 2.651 Å, respectively.[14] The Pt-Ni distance of 2.689 (2) Å therefore implies the presence of metal-metal bonding which would be consistent with 18- and 16-electron counts on Ni(0) and Pt(II), respectively,[15] a formulation which is supported by the fact that **3** in CH_2Cl_2 reverts to $Ni_2(CO)_2(\mu\text{-}CO)(\mu\text{-dppm})_2$ and $PtCl_2(dppm)$. The structure of **3** is quite different from that of the related compound $Pt_2Cl_2(\mu\text{-}CO)(\mu\text{-dpam})_2$ (dpam = the arsenic analogue of dppm).[16]

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Registry No. 1, 113894-20-5; 2, 32423-72-6; 3, 113894-21-6; $Ni(CO)_2(PPh_3)_2$, 13007-90-4; $Ni_2(CO)_2(\mu\text{-}CO)(\mu\text{-dppm})_2$, 106251-27-8; $PtCl_2(COD)$, 12080-32-9; $Rh_2Cl_2(CO)_4$, 14523-22-9; $Mo(CO)_5(THF)$, 53248-43-4; Ni, 7440-02-0; Pt, 7440-06-4.

Supplementary Material Available: Summary of crystal data, data collection parameters, and refinement details and tables of positional and thermal parameters and interatomic distances and bond angles (13 pages), a listing of observed and calculated structure factors for the X-ray structural analysis of $NiPtCl_2(\mu\text{-}CO)(\mu\text{-dppm})_2$ (58 pages). Ordering information is given on any current masthead page.

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[14] Sharp, P. R. *Inorg. Chem.* **1986**, *25*, 4185. Manojlovic-Muir, Lj.; Muir, K. W.; Solomun, T. *Acta Cryst.*, **1979**, *B35*, 1237.

[15] The Fe-Pt bond distance in $(OC)_3Fe(\mu\text{-dppm})(\mu\text{-}CO)PtBr_2$ is almost identical at 2.647 (4) Å. Jacobsen, G. B.; Shaw, B. L.; Thornton-Pett, M. *J. Chem. Soc., Dalton Trans.* **1987**, 3079.

[16] Brown, M. P.; Keith, A. N.; Manojlovic-Muir, Lj.; Muir, K. W.; Puddephatt, R. J.; Seddon, K. R. *Inorg. Chim. Acta* **1979**, *34*, L 223.

Additions and Corrections

Richard G. Ball, Michael R. Burke, and Josef Takats*: Synthesis and Comparative Study of Iron Triad $M(CO)_4(\eta^2\text{-alkyne})$ Complexes (M = Fe, Ru, Os; Alkyne = Bis(trimethylsilyl)acetylene). **1987**, *6*, 1918–1924.

An error occurred in the reported coalescence temperature, T_c, and the corresponding $\Delta G^{\ddagger}{}_{T_c}$ for carbonyl scrambling in $Ru(CO)_4(\eta^2\text{-BTMSA})$ appearing in Table V. The correct values are T_c = −64 °C and $\Delta G^{\ddagger}{}_{Tc}$ = 9.0 kcal/mol. As a consequence, the trend in $\Delta G^{\ddagger}$ for carbonyl scrambling now should be Fe = Ru < Os. Furthermore, although the inverse relationship between $\Delta G^{\ddagger}$ and $\Delta\delta$ and $\Delta\nu_{CC}$, shown graphically in Figure 4, is still correct, the linear relationship between these properties as the metal is changed is no longer valid. Admittedly the consequences of the error in T_c are not insignificant; nevertheless, they do not adversely effect the major conclusions of the paper. In particular, the importance of both σ/π components of the metal alkyne interaction in this series of compounds is valid. The unexpectedly similar $\Delta G^{\ddagger}{}_{T_c}$ for carbonyl scrambling of the Fe and Ru compounds further emphasizes the special behavior of Ru and other second-row transition metals when compared to their first- and third-row congeners. We regret the error and the attendant changes.

Yan-Lung Shi, Yi-Ci Gao, Qi-Zhen Shi,* David L. Kershner, and Fred Basolo*: Oxygen Atom Transfer Reactions to Metal Carbonyls. Kinetics and Mechanism of CO Substitution Reactions of $M(CO)_6$ (M = Cr, Mo, W) in the Presence of $(CH_3)_3NO$. **1987**, *6*, 1528–1531.

Table V on page 1531 should be corrected as follows.

compd	nucleophile	T, °C	k, M^{-1} s^{-1}	ref
$Cr(CO)_6$	N_3^-	30.0	9.7×10^{-3}	12
	$PhCH_2MgCl$	27.0	0.053	13b
$Mo(CO)_6$	$PhCH_2MgCl$	27.0	0.091	13b
$W(CO)_6$	CN^-	95.0	5.2×10^{-3}	11b
	$PhCH_2MgCl$	27.0	0.108	13b
	Me_3NO	25.0	0.366	*a*

In the abstract and the text, $PhCH_2MgCl$ should replace $PhCH_2MgBr$.