Additions and Corrections

Brett A. Cowans, R. C. Haltiwanger, and M. Rakowski DuBois*: Synthesis and Characterization of Organometallic Molybdenum–Iron–Sulfur and Molybdenum–Cobalt–Sulfur Clusters. **1987**, *6*, 995.

For complex II, the angle γ in the cell dimensions is given incorrectly. The correct cell dimensions are a = 12.542 (2) Å, b = 10.724 (2) Å, c = 10.296 (1) Å, α = 121.35 (9)°, β = 93.90 (1)°, and γ = 107.07 (1)°. Also, on page 998, Table V, the fractional cell coordinates of Mo Fe(1) should read: for Fe(1), x = 0.238 03 (4), y = 0.345 41 (5), and z = −0.021 56 (5).

Giuseppe Bruno, Sandra Lo Schiavo, Enrico Rotondo, Pasquale Piraino, and Felice Faraone*: Synthesis of Triangular Mixed-Metal Clusters by the Addition of Copper Electrophiles to an Electron-Rich Rhodium–Rhodium Bond. X-ray Crystal Structures of $[Rh_2(\eta\text{-}C_5H_5)_2(\mu\text{-}CO)(\mu\text{-}Ph_2PCH_2PPh_2)(\mu\text{-}CuI)]$ and $[Rh_2(\eta\text{-}C_5H_5)_2(\mu\text{-}CO)(\mu\text{-}Ph_2PCH_2PPh_2)(\mu\text{-}AgO_2CCF_3)]$. **1987**, *6*, 2502–2507.

The correct title is published above. The title was incorrectly published as "Synthesis of Triangular Mixed-Metal Clusters by the Addition of Copper Electrophiles to an Electron-Rich Rhodium–Rhodium Bond. X-ray Crystal Structures of $[Rh_2(\eta\text{-}C_5H_5)_2(\mu\text{-}CO)(\mu\text{-}Ph_2PCH_2PPh_2)(\mu\text{-}CuI)]$ and $[Rh_2(\eta\text{-}C_5H_5)_2(\mu\text{-}CO)(\mu\text{-}Ph_2PCH_2PPh_2)(\mu\text{-}AgO_2CCH_3)]$". In the text the crystal structure reported is correctly referred to the $\mu\text{-}AgO_2CCF_3$ adduct.

Greg A. Urove and Mark E. Welker*: Synthesis of a Stable Disulfur Monoxide Precursor and Trapping of Disulfur Monoxide with Transition-Metal Complexes. **1988**, *7*, 1013.

The last line of the first paragraph that reads "The thiosulfinate esters (**3**) have shown interesting biological activities as antibacterials,[4] antifungals,[4] plant growth inhibitors,[5] and platelet aggregation inhibitors.[6]" should read "Certain thiosulfinate esters have shown interesting biological activities as antibacterials,[4] antifungals,[4] plant growth inhibitors,[5] and platelet aggregation inhibitors.[6]"

Book Reviews

Gmelin Handbook of Inorganic Chemistry. Eighth Edition. Be. Organoberyllium Compounds. Part 1. H. Schmidbaur, volume author. U. Krüerke and A. Slawisch, volume chief editors. Gmelin Institut für Anorganische Chemie der Max-Planck-Gesellschaft zur Förderung der Wissenschaften and Springer-Verlag, Berlin. 1987. xii + 247 pages. DM 1189.

Organoberyllium chemistry at present is not a very active area, and, in fact, it never really has been. Yet over the years information on organoberyllium compounds has accumulated, and this is all summarized in the present book. In modern times it was Geoffrey Coates who was the main contributor to organoberyllium chemistry, and, most fittingly, this book is dedicated to him. It is the toxicity of beryllium and its compounds that mainly is responsible for the fact that organoberyllium chemistry has, as the author points out in the preface, come to a virtual standstill. The brief burst of interest in this area, occasioned by the use of organoberyllium compounds as precursors for BeH_2 (a potential high energy fuel) and the fascinating questions of structure and bonding presented by the beryllium cyclopentadienyls and borohydrides, was not sustained, and now only a rare paper on organoberyllium compounds that reports experimental work is encountered.

This book covers all mononuclear organoberyllium compounds: first the simple ones, R_2Be and R_2Be–Lewis base adducts, including salts of $[R_2BeX]^-$ and $[R_4Be_2X]^-$ anions; RBeH and diverse other RBeX compounds in which Be is bonded to halogen, oxygen, sulfur, selenium, nitrogen, and phosphorus. The more exotic beryllocarbaboranes and the various sandwich-bonded beryllium compounds, $(\eta^5\text{-}C_5H_5)_2Be$ and $\eta^5\text{-}C_5H_5BeX$ types, follow. The book concludes with a section on theoretical studies of hypothetical compounds, for beryllium chemistry has been a popular playground for some theoreticians.

Dimethylberyllium is the most thoroughly investigated beryllium alkyl. In fact, it is pointed out by the author that over one-third of the organoberyllium literature is devoted to this single compound. The information on Me_2Be is reported in the usual thorough manner that is a characteristic of the Gmelin Handbook: preparation, association equilibria, molecular parameters and physical and spectroscopic properties, chemical reactions, and Lewis base adducts, all supported by 105 references. Many of the other organoberyllium compounds have not been investigated in such detail, but what is known about them is brought here. The literature is covered completely through the end of 1986 and includes some 1987 references as well. All this information is presented effectively in a well-organized manner with useful formulas and structural figures and tables of data. The book has an empirical formula index, a ligand formula index, and a detailed table of contents. The review literature of organoberyllium chemistry is summarized on p 1.

If ever organoberyllium chemistry becomes an active research field again, this book will be a great help to those who must familiarize themselves with what is known about organoberyllium compounds.

Dietmar Seyferth, *Massachusetts Institute of Technology*