

-0.10 (m, 1H, cyclopropyl CH₂).

Spectroscopic Characterization of 5-Hexenylcyclopropane. A CDCl₃ solution of **2a** (115 mg, 0.25 mmol) was heated for 2 h at 40 °C, and filtration of the resulting red solution through alumina afforded a colorless solution of 5-hexenylcyclopropane (92% yield based on ferrocene; 47 mg, 0.25 mmol). ¹H NMR (CDCl₃, ferrocene as internal standard): δ 5.83 (m, 1H, =CH), 4.97 (m, 2H, =CH₂), 0.65 (m, 1H, cyclopropyl CH), 0.39 (m, 2H,

cyclopropyl HCHHCH), -0.02 (m, 2H, cyclopropyl HCHHCH). ¹³C[¹H] NMR (CDCl₃): δ 139.3 (=CH), 114.2 (=CH₂), 10.9 (cyclopropyl CH), 4.4 (cyclopropyl CH₂).

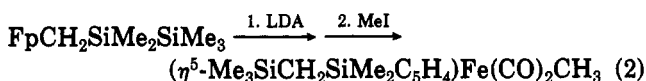
Acknowledgment. Dr. J.-R. Hamon is gratefully acknowledged for helpful discussions.

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Additions and Corrections

Keith H. Pannell,* Toshiaki Kobayashi, and Ramesh N. Kapoor: Thermal and Photochemical Properties of [(1-Heptamethyltrisilyl)methyl]iron and [(2-Heptamethyltrisilyl)methyl]iron Complexes (η^5 -L)Fe(CO)₂CH₂Si₃Me₇ (L = C₅H₅, C₅Me₅, C₉H₇). 1992, 11, 2229–2235.

On page 2229, eq 2 should read



Linda L. Lopez, P. Bernatis, J. Birnbaum, R. C. Haltiwanger, and M. Rakowski DuBois*: Mechanistic Study of Hydrogen Activation by Cationic Dinuclear (μ -Sulfido)molybdenum Complexes. 1992, 11, 2424–2435.

In the Experimental Section the part entitled "Dependence of Reaction Rates on Temperature" on p 2434 should include the following corrections and additions.

In line 10 of the first paragraph in this part, the formula should read ($k_2 = (1/2)k_{\text{obs}}/[\text{H}_2]$). The factor of 1/2 is explained in ref 14. In line 12, the correlation coefficient should read 0.999. The last sentence in the first paragraph should read as follows: From the slope and intercept of this line the ΔH and ΔS values of activation were calculated to be 81 kJ/mol and -23 J/(mol K), respectively.

Paragraph 2 in this part should read as follows: The solubility of H₂ (1 atm) in CD₂Cl₂ at 19 °C was found to be 2.69×10^{-3} M by comparing the NMR integration of the hydrogen resonance to that of the internal standard,

t-BuCl. An average value from multiple NMR experiments was used. Integrations were compared for data collected with a 20-s pulse delay. A similar procedure was used to determine the solubilities of H₂ at 19 °C (3 atm), 40 °C (1 atm), and 60 °C (1 atm). Solubility values for 3 atm of H₂ at the higher temperatures were estimated from the relative solubilities at 19 °C, which generally showed good agreement with Henry's Law. Rate constants were corrected for the differences in hydrogen solubilities determined over this temperature range.

The activation parameters in the abstract, line 13, and on p 2430, lines 3–4 in the first column, should read as follows: $\Delta H^\ddagger = 81$ kJ/mol; $\Delta S^\ddagger = -23$ J/(mol K).

Christopher J. Tokar, Peter B. Kettler, and William B. Tolman*: Enantiomerically Pure Copper(I) and Zinc(II) Complexes of a Novel C₃-Symmetric Polypyrazole Ligand. 1992, 11, 2737–2739.

The incorrect enantiomer was chosen for refinement in the X-ray crystal structure of 6-2CH₂Cl₂ (Figure 1) due to an error made during placement of the rigid group for camphor (ref 10). Refinement of the data using the mirror image of the structure shown in Figure 1 gave improved discrepancy indices of $R = 0.061$ and $R_w = 0.076$. A revised ORTEP drawing (with camphor rings now in agreement with those drawn in Scheme I) and new tables of X-ray data have been deposited as supplementary material.

Supplementary Material Available: An ORTEP drawing and tables of positional and thermal parameters and bond distances and angles for 6-2CH₂Cl₂ (33 pages). Ordering information is given on any current masthead page.