−0.10 (m, 1H, cyclopropyl CH_2).

Spectroscopic Characterization of 5-Hexenylcyclopropane. A $CDCl_3$ solution of **2a** (115 mg, 0.25 mmol) was heated for 2 h at 40 °C, and filtration of the resulting red solution through alumina afforded a colorless solution of 5-hexenylcyclopropane (92% yield based on ferrocene; 47 mg, 0.25 mmol). ^1H NMR ($CDCl_3$, ferrocene as internal standard): δ 5.83 (m, 1H, =CH), 4.97 (m, 2H, $=CH_2$), 0.65 (m, 1H, cyclopropyl CH), 0.39 (m, 2H cyclopropyl *H*CH*H*CH), −0.02 (m, 2H, cyclopropyl HC*H*HC*H*). $^{13}C[^1H]$ NMR ($CDCl_3$): δ 139.3 (=CH), 114.2 ($=CH_2$), 10.9 (cyclopropyl CH), 4.4 (cyclopropyl CH_2).

Acknowledgment. Dr. J.-R. Hamon is gratefully acknowledged for helpful discussions.

OM920269U

Additions and Corrections

Keith H. Pannell,* Toshiaki Kobayashi, and Ramesh N. Kapoor: Thermal and Photochemical Properties of [(1-Heptamethyltrisilyl)methyl]iron and [(2-Heptamethyltrisilyl)methyl]iron Complexes $(\eta^5\text{-L})Fe(CO)_2CH_2Si_3Me_7$ (L = C_5H_5, C_5Me_5, C_9H_7). **1992,** *11,* 2229–2235.

On page 2229, eq 2 should read

$$FpCH_2SiMe_2SiMe_3 \xrightarrow{1.\ LDA} \xrightarrow{2.\ MeI} (\eta^5\text{-}Me_3SiCH_2SiMe_2C_5H_4)Fe(CO)_2CH_3 \quad (2)$$

Linda L. Lopez, P. Bernatis, J. Birnbaum, R. C. Haltiwanger, and M. Rakowski DuBois*: Mechanistic Study of Hydrogen Activation by Cationic Dinuclear (μ-Sulfido)molybdenum Complexes. **1992,** *11,* 2424–2435.

In the Experimental Section the part entitled "Dependence of Reaction Rates on Temperature" on p 2434 should include the following corrections and additions.

In line 10 of the first paragraph in this part, the formula should read ($k_2 = (1/2)k_{obs}/[H_2]$). The factor of 1/2 is explained in ref 14. In line 12, the correlation coefficient should read 0.999. The last sentence in the first paragraph should read as follows: From the slope and intercept of this line the ΔH and ΔS values of activation were calculated to be 81 kJ/mol and −23 J/(mol K), respectively.

Paragraph 2 in this part should read as follows: The solubility of H_2 (1 atm) in CD_2Cl_2 at 19 °C was found to be 2.69×10^{-3} M by comparing the NMR integration of the hydrogen resonance to that of the internal standard, *t*-BuCl. An average value from multiple NMR experiments was used. Integrations were compared for data collected with a 20-s pulse delay. A similar procedure was used to determine the solubilities of H_2 at 19 °C (3 atm), 40 °C (1 atm), and 60 °C (1 atm). Solubility values for 3 atm of H_2 at the higher temperatures were estimated from the relative solubilities at 19 °C, which generally showed good agreement with Henry's Law. Rate constants were corrected for the differences in hydrogen solubilities determined over this temperature range.

The activation parameters in the abstract, line 13, and on p 2430, lines 3–4 in the first column, should read as follows: $\Delta H^{\ddagger} = 81$ kJ/mol; $\Delta S^{\ddagger} = -23$ J/(mol K).

Christopher J. Tokar, Peter B. Kettler, and William B. Tolman*: Enantiomerically Pure Copper(I) and Zinc(II) Complexes of a Novel C_3-Symmetric Polypyrazole Ligand. **1992,** *11,* 2737–2739.

The incorrect enantiomer was chosen for refinement in the X-ray crystal structure of $\mathbf{6}\cdot 2CH_2Cl_2$ (Figure 1) due to an error made during placement of the rigid group for camphor (ref 10). Refinement of the data using the mirror image of the structure shown in Figure 1 gave improved discrepancy indices of $R = 0.061$ and $R_w = 0.076$. A revised ORTEP drawing (with camphor rings now in agreement with those drawn in Scheme I) and new tables of X-ray data have been deposited as supplementary material.

Supplementary Material Available: An ORTEP drawing and tables of positional and thermal parameters and bond distances and angles for $\mathbf{6}\cdot 2CH_2Cl_2$ (33 pages). Ordering information is given on any current masthead page.