

# Synthesis, Characterization, and Photophysical Properties of Iridium Complexes with an 8-Phenylquinoline Framework. The First Six-Membered Chelated Iridium Complexes for Electroluminescence

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A series of six-membered chelated iridium complexes bearing an 8-phenylquinoline framework have been prepared and characterized with an X-ray diffraction study. The photophysical properties of these complexes were examined with appropriate spectroscopic methods. The results, in combination with ab initio approaches, led us to clearly assign various electronic transition states. One salient feature for these red complexes is the appearance of a dual strong absorption band around 425–500 nm, which incorporates a state mixing between <sup>1</sup>MLCT and  $\pi$ - $\pi^*$  manifolds. These complexes show deep red phosphorescent emissions (650–680 nm), with short lifetimes of 1.7–3.0  $\mu$ s and moderate quantum yields of 0.05–0.11 in deaerated CH<sub>3</sub>CN. The electroluminescence performance of these species shows a promising perspective in the OLED display. One representative, Ir-[8-(3,5-difluorophenyl)quinoline]<sub>2</sub>(acac) (**4c**), shows a  $\eta_{\text{ext}}$  value of 2.04% at  $J = 20$  mA/cm<sup>2</sup> and the maximum brightness is 3427 cd/m<sup>2</sup> ( $x = 0.68$ ,  $y = 0.32$ ) with a full width at half-maximum of only 60 nm at 16 V, demonstrating the first six-membered chelated iridium complexes to suit applications in OLED devices.

Cyclometalated iridium complexes have attracted considerable attention in material research because of their outstanding performance in organic light-emitting diodes (OLED).<sup>1–4</sup> The strong spin–orbit coupling caused by the heavy atom (e.g., Ir(III)) effect leads to efficient intersystem crossing from the singlet (S<sub>1</sub>) to the triplet (T<sub>n</sub>) as well as the enhancement of the T<sub>1</sub>–S<sub>0</sub> transition. OLED devices based on these phosphors can significantly improve electroluminescence performance because both singlet and triplet excitons can be harvested for light emissions, theoretically attaining 100% internal quantum efficiencies. Interestingly, all these iridium complexes are prepared in a five-membered chelated

framework (**A**),<sup>1–6</sup> as depicted in Figure 1, and to our best knowledge, no information is known about the

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(1) (a) Baldo, M. A.; O'Brien, D. F.; You, Y.; Shoustikov, A.; Sibley, S.; Thompson, M. E.; Forrest, S. R. *Nature* **1998**, *395*, 151. (b) Kohler, A.; Wilson, J. S.; Friend, R. H. *Adv. Mater.* **2002**, *14*, 701.

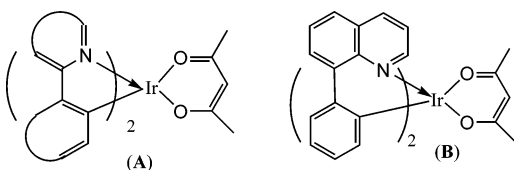
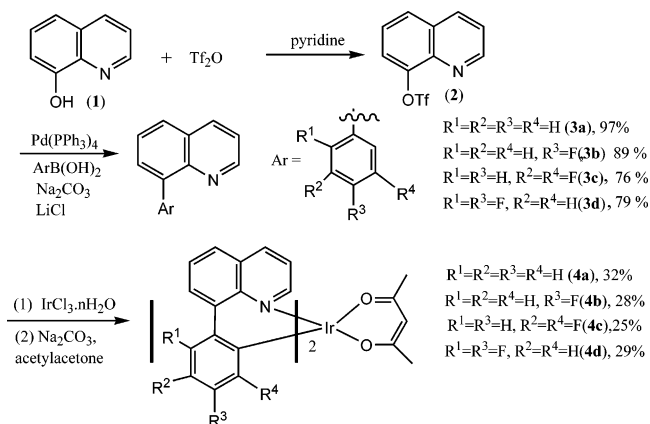
(2) (a) Baldo, M. A.; Thompson, M. E.; Forrest, S. R. *Nature* **2000**, *403*, 750. (b) Pfeiffer, M.; Forrest, S. R.; Leo, K.; Thompson, M. E. *Adv. Mater.* **2002**, *14*, 1633. (c) Holmes, R. J.; Forrest, S. R.; Tung, Y.-J.; Kwong, R. C.; Brown, J. J.; Garon, S.; Thompson, M. E. *Appl. Phys. Lett.* **2003**, *82*, 2422.

(3) (a) Baldo, M. A.; Lamansky, S.; Burrows, P. E.; Thompson, M. E.; Forrest, S. R. *Appl. Phys. Lett.* **1999**, *75*, 4. (b) Adachi, C.; Baldo, M. A.; Thompson, M. E.; Forrest, S. R. *Appl. Phys. Lett.* **2000**, *77*, 904. (c) Tamayo, A. B.; Alleyne, B. D.; Djurovich, P. I.; Lamansky, S.; Tsyba, I.; Ho, N. N.; Bau, R.; Thompson, M. E. *J. Am. Chem. Soc.* **2003**, *125*, 7377. (d) Adachi, C.; Baldo, M. A.; Forrest, S. R.; Lamansky, S.; Thompson, M. E.; Kwong, R. C. *Appl. Phys. Lett.* **2001**, *78*, 1622. (e) Tsuzuki, T.; Shirasawa, N.; Suzuki, T.; Tokito, S. *Adv. Mater.* **2003**, *15*, 1455. (f) Markham, J. P. J.; Samuel, I. D. W.; Lo, S.-C.; Burn, P. L.; Weiter, M.; Bässler, H. *J. Appl. Phys.* **2004**, *95*, 438. (g) Polson, M.; Fracasso, S.; Bertolasi, V.; Ravaglia, M.; Scandola, F. *Inorg. Chem.* **2004**, *43*, 1950. (h) Slinker, J. D.; Gorodetsky, A. A.; Lowry, M. S.; Wang, J.-J.; Parker, S.; Rohl, R.; Bernhard, S.; Malliaras, G. G. *J. Am. Chem. Soc.* **2004**, *126*, 2763.

(4) (a) D'Andrade, B. W.; Brooks, J.; Adamovich, V.; Thompson, M. E.; Forrest, S. R. *Adv. Mater.* **2002**, *14*, 1032. (b) D'Andrade, W.; Thompson, M. E.; Forrest, S. R. *Adv. Mater.* **2002**, *14*, 147. (c) Lamansky, S.; Djurovich, P.; Murphy, D.; Abdel-Razzaq, F.; Lee, H.-E.; Adachi, C.; Burrows, P. E.; Forrest, S. R.; Thompson, M. E. *J. Am. Chem. Soc.* **2001**, *123*, 4304. (d) Kolosov, D.; Adamovich, V.; Djurovich, P.; Thompson, M. E.; Adachi, C. *J. Am. Chem. Soc.* **2002**, *124*, 9945. (e) Adachi, C.; Baldo, M. A.; Thompson, M. E.; Forrest, S. R. *J. Appl. Phys.* **2001**, *90*, 5048. (f) Adamovich, V.; Brooks, J.; Tamayo, A.; Alexander, A. M.; Djurovich, P. I.; D'Andrade, B. W.; Adachi, C.; Forrest, S. R.; Thompson, M. E. *New J. Chem.* **2002**, 1171. (g) Huang, W.-S.; Lin, J.-T.; Chien, C.-H.; Tao, Y.-T.; Sun, S.-S.; Wen, Y.-S. *Chem. Mater.* **2004**, *16*, 2480. (h) Laskar, I. R.; Chen, T.-M. *Chem. Mater.* **2004**, *16*, 111.

(5) (a) Xie, H.-Z.; Liu, M.-W.; Wang, O.-Y.; Zhang, X.-H.; Lee, C.-S.; Hung, L.-S.; Lee, S.-T.; Teng, P.-F.; Kwong, H.-L.; Zheng, H.; Che, C.-M. *Adv. Mater.* **2001**, *13*, 1245. (b) Ikai, M.; Tokito, S.; Sakamoto, Y.; Suzuki, T. *Appl. Phys. Lett.* **2001**, *79*, 156. (c) Chen, F. C.; Yang, Y. *Appl. Phys. Lett.* **2002**, *80*, 2308. (d) Grushin, V. V.; Herron, N.; LeCloux, D. D.; Marshall, W. J.; Petrov, V. A.; Wang, Y. *Chem. Commun.* **2001**, 9, 1494. (e) Tokito, S.; Iijima, T.; Tsuzuki, T.; Sato, F. *Appl. Phys. Lett.* **2003**, *83*, 2459. (f) Holmes, R. J.; D'Andrade, B. W.; Forrest, S. R.; Ren, X.; Li, J.; Thompson, M. E. *Appl. Phys. Lett.* **2003**, *83*, 3818. (g) Tokito, S.; Iijima, T.; Suzuki, Y.; Kita, H.; Tsuzuki, T.; Sato, F. *Appl. Phys. Lett.* **2003**, *83*, 569. (h) Chen, X.; Liao, J.-L.; Liang, Y.; Ahmed, M. O.; Tseng, H.-E.; Chen, S.-A. *J. Am. Chem. Soc.* **2003**, *125*, 636. (i) Niu, Y.-H.; Chen, B.; Liu, S.; Yip, H.; Bardecker, J.; Jen, A. K.-Y.; Kavitha, J.; Chi, Y.; Shu, C.-F.; Tseng, Y.-H.; Chien, C.-H. *Appl. Phys. Lett.* **2004**, *85*, 1619.

(6) (a) Okada, S.; Iwawak, H.; Furugori, M.; Kamatani, J.; Igawa, S.; Moriyama, T.; Miuva, S.; Tsuboyama, A.; Takiguchi, T.; Mizutani, H. *SID Digest* **2002**, 1360. (b) Tsuboyama, A.; Iwawaki, H.; Furugori, M.; Mukaide, T.; Kamatani, J.; Igawa, S.; Moriwama, T.; Miura, S.; Takiguchi, T.; Okada, S.; Hoshino, M.; Ueno, K. *J. Am. Chem. Soc.* **2003**, *125*, 12971. (c) Su, Y.-J.; Huang, H.-L.; Li, C.-L.; Cheng, C.-H.; Tao, Y.-T.; Chou, P.-T.; Datta, S.; Liu, R.-S. *Adv. Mater.* **2003**, *15*, 884. (d) Li, C.-L.; Su, Y.-J.; Tao, Y.-T.; Chou, P.-T.; Chien, C.-H.; Cheng, C.-C.; Liu, R.-S. *Adv. Funct. Mater.*, in press.

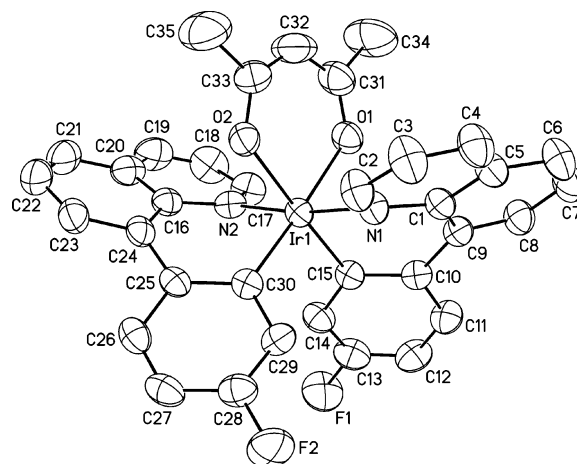
**Figure 1.****Figure 2.**

suitability of six-membered chelated iridium complexes toward application in OLED devices. Herein, we report the preparation, photophysical properties, and the first OLED performances of the iridium complexes chelated with a six-membered 8-phenylquinoline framework (**B**), depicted in Figure 1.

## Results and Discussion

Figure 2 depicts the synthetic pathway of 8-phenylquinoline required for cyclometalated iridium complexes. 8-Hydroxyquinoline was converted to its triflate derivative **2**, which underwent subsequent Suzuki-coupling reactions<sup>7</sup> with a suitable phenyl derivative and gave the desired 8-phenylquinoline derivatives **3a–d** in satisfactory yields (76–97%). These ligands were transformed into cyclometalated iridium complexes **4a–d** via heating with  $\text{IrCl}_3 \cdot n\text{H}_2\text{O}$  in ethoxyethanol, followed by treatment of the resulting diiridium precursors<sup>4c</sup> with a mixture of acetyl acetone/ $\text{Na}_2\text{CO}_3$ ; the overall yields were 25–32%.

The molecular structure of representative compound **3b** has been characterized by an X-ray diffraction study,<sup>8</sup> with the ORTEP drawing and selected bond distances and angles shown in Figure 3. The coordination geometry around iridium approximates an octahedral geometry, with two nitrogen ligating atoms *trans* to each other, whereas the two acetylacetonate oxygen atoms lie *trans* to the ligating carbon atoms of 8-phenylquinoline. The molecular structure of **4b** has near  $C_2$ -symmetry, which accounts for an equivalent pattern in the proton NMR peaks of the two chelated ligands. Notably, the six-membered chelated iridium geometry is not coplanar and is close to a twisted boat form. A dihedral angle of  $28.9^\circ$  is formed between the planes



**Figure 3.** ORTEP diagram of **4b**. Selected bond distances (Å) and angles (deg):  $\text{Ir}(1)\text{--N}(1) = 2.025(4)$ ,  $\text{Ir}(1)\text{--N}(2) = 2.048(5)$ ,  $\text{Ir}(1)\text{--O}(1) = 2.128(4)$ ,  $\text{Ir}(1)\text{--O}(2) = 2.136(4)$ ,  $\text{Ir}(1)\text{--C}(15) = 1.995(5)$ ,  $\text{Ir}(1)\text{--C}(30) = 1.984(5)$ ,  $\text{C}(30)\text{--Ir}(1)\text{--C}(15) = 92.8(2)$ ,  $\text{C}(30)\text{--Ir}(1)\text{--N}(1) = 94.84(19)$ ,  $\text{C}(15)\text{--Ir}(1)\text{--N}(1) = 90.37(19)$ ,  $\text{C}(30)\text{--Ir}(1)\text{--N}(2) = 90.40(19)$ ,  $\text{C}(15)\text{--Ir}(1)\text{--N}(2) = 94.70(19)$ ,  $\text{C}(15)\text{--Ir}(1)\text{--O}(1) = 89.56(18)$ ,  $\text{N}(1)\text{--Ir}(1)\text{--O}(1) = 86.79(17)$ ,  $\text{N}(2)\text{--Ir}(1)\text{--O}(1) = 87.75(17)$ ,  $\text{C}(30)\text{--Ir}(1)\text{--O}(2) = 88.45(18)$ ,  $\text{N}(1)\text{--Ir}(1)\text{--O}(2) = 87.81(16)$ ,  $\text{N}(2)\text{--Ir}(1)\text{--O}(2) = 87.00(17)$ ,  $\text{O}(1)\text{--Ir}(1)\text{--O}(2) = 89.22(16)$ .

**Table 1.** Photophysical Data of Iridium Complexes **4a–d**

	<b>4a</b>	<b>4b</b>	<b>4c</b>	<b>4d</b>
absorption (nm)	448 (12 760) <sup>a</sup> 504 (7818)	436 (13 109) 487 (10 000)	432 (14 286) 487 (7848)	426 (9553) 477 (7480)
PL $\lambda_{\text{max}}$ (nm) <sup>b</sup>	680 (670)	666 (661)	656 (650)	650 (647)
$\Phi_{\text{degassed}}$	0.05	0.07	0.09	0.11
$\Phi_{\text{aerated}}$	0.006	0.008	0.008	0.01
$\tau_{\text{degassed}}$ (ns)	1769	2825	3072	2763
$\tau_{\text{aerated}}$ (ns)	233	292	399	422
$\tau_{\text{solid}}$ (ns)	340	1073	368	340
$E_{1/2}^{\text{ox}}$ (V) <sup>c</sup>	0.85	1.00	1.03	1.02

<sup>a</sup> These values in parentheses represent extinction coefficients in  $\text{M}^{-1} \text{cm}^{-1}$ . <sup>b</sup> Data are obtained in degassed acetonitrile, and data in parenthesis show the property in solid state. <sup>c</sup> The values are calculated relative to  $\text{Cp}_2\text{Fe}/\text{Cp}_2\text{Fe}^+$  vs NHE in  $\text{CH}_2\text{Cl}_2$  solution.

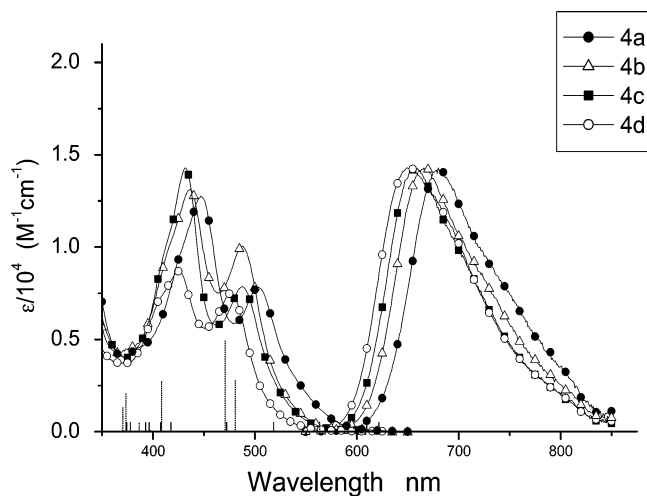
defined by  $\text{C}24\text{--C}16\text{--N}2$  and  $\text{C}25\text{--C}30\text{--Ir}$ . Although this nonplanar structure is expected to weaken the iridium–ligand bonding, it is not reflected in the Ir–C and Ir–N distances of complex **4b**. In comparison to those of the five-membered chelated *fac*- $\text{Ir}(\text{ppy})_3$ ,<sup>9</sup> the Ir1–C30 (1.984 Å) and Ir1–C15 (1.995 Å) bonds of complex **4b** are similar to those (2.003 Å) for *fac*- $\text{Ir}(\text{ppy})_3$ , whereas the Ir1–N bonds (2.025 and 2.048 Å) are significantly shorter than those (2.135 Å) of *fac*- $\text{Ir}(\text{ppy})_3$ .

Table 1 and Figure 4 show the photophysical properties of iridium complexes **4a–d**. By comparison to free 8-phenylquinoline ligands **3a–d**, the absorption spectral features around 330–350 nm in Figure 4 are reasonably assigned to the  $^1\pi\text{--}\pi^*$  transition of the quinoline chromophore. The most salient feature for these red complexes is the appearance of two strong absorption bands

(7) (a) Echavarren, A. M.; Stille, J. K. *J. Am. Chem. Soc.* **1987**, *109*, 5478. (b) Huth, A.; Beetz, I.; Schumann, I. *Tetrahedron* **1989**, *21*, 6679.

(8) Crystallographic data for compound **4b**:  $a = 11.4624(9)$  Å,  $b = 11.8852(9)$  Å,  $c = 12.1307(9)$  Å,  $\alpha = 101.170(2)^\circ$ ,  $\beta = 11.4624(9)^\circ$ ,  $\gamma = 11.4624(9)^\circ$ ,  $V = 1514.7(2)$  Å<sup>3</sup>,  $Z = 2$ . Of the 17 909 unique reflections, 7486 were considered observed having  $I > 2\sigma(I)$ . Final  $R = 0.0331$  and  $R_w = 0.1156$ .

(9) (a) Lamansky, S.; Djurovich, P.; Murphy, D.; Abdel-Razzaq, F.; Kwong, R.; Tsyba, I.; Bortz, M.; Mui, B.; Bau, R.; Thompson, M. E. *Inorg. Chem.* **2001**, *40*, 1704. (b) Colombo, M. G.; Brunold, T. C.; Riedener, T.; Güdel, H. U.; Fürtsch, M.; Zorn, B. *Inorg. Chem.* **1994**, *33*, 545–550.



**Figure 4.** UV-vis absorption and emission spectra of complexes **4a–d**. The line sign shown on the wavelength coordinate indicates the calculated energy level of various transitions in **4b**, in which the dashed and solid line denote singlet and triplet excited states, respectively. See text for a detailed elaboration of the ab initio approach.

in the region of 420–500 nm with peak wavelengths at 425–450 nm ( $\epsilon = 9553\text{--}14\,286\text{ M}^{-1}\text{ cm}^{-1}$ ) and 475–500 nm ( $\epsilon = 7000\text{--}14\,000\text{ M}^{-1}\text{ cm}^{-1}$ ), which are apparently absent in the free neutral 8-phenylquinoline ligands.

To obtain a clear assignment of these absorption features, we have carried out a reliable ab initio approach for **4b** using the B3LYP density functional theory, in which double- $\zeta$  quality basis sets were employed for the Ir atom. In addition, a relativistic effective core potential on Ir replaced the inner core electrons, leaving the outer core [(5s)<sup>2</sup>(5p)<sup>6</sup>] electrons and the (5d)<sup>6</sup> valence electrons of Ir<sup>III</sup>. The ground state geometry was adapted from the truncated X-ray data of **4b** without further optimization. A time-dependent DFT (TDDFT) calculation using the B3LYP functional was then performed to estimate the corresponding energy of transition (for details, see the Experimental Section). The calculated parameters for various lower-lying transitions together with the associated frontier orbitals are listed in Table 2. For a fair comparison, the calculated energy levels are also depicted in Figure 4. On the basis of Figure 4 and Table 2, it is reasonable to assign the two major bands at 420–450 and 475–500 nm to S<sub>2</sub> and S<sub>3</sub> states, respectively, both of which incorporate mainly a spin-allowed metal–ligand (cyclometalated 8-phenylquinoline) charge transfer (<sup>1</sup>MLCT) and, in part, a  $\pi\text{--}\pi^*$  intraligand transition from phenyl to cyclometalated 8-phenylquinoline (see Table 1). Note that the much broader absorption spectral feature in solution, in reality, is due to the vibronic coupling that is omitted in the ab initio approach. The lowest singlet state, i.e., the S<sub>1</sub>, was calculated to be at ~563 nm, which is obscured in the absorption spectra due to the overlap with the much stronger S<sub>2</sub> band. The frontier orbital (HOMO → LUMO) analysis of the S<sub>1</sub> also indicates a state mixing between <sup>1</sup>MLCT and  $\Pi\text{--}\Pi^*$  manifolds. Moreover, the lowest triplet state originates from HOMO–1 → LUMO (+91%) and HOMO → LUMO+1 (9%, see Table 1), which apparently possess a mixed <sup>3</sup>MLCT and intraligand  $\pi(\text{phenyl})\text{--}\pi^*(8\text{-phen-}$

ylquinoline) character. The S<sub>0</sub> → T<sub>1</sub> transition is estimated to be at ~621 nm, which is consistent with the phosphorescence peak wavelength at 666 nm, taking into account the bathochromic shift of the emission due to the solvent stabilization.

In the PL spectra, the parent **4a** shows a single, deep-red emission band with a wavelength up to 680 nm in CH<sub>3</sub>CN. Intriguingly, by comparison, its five-membered chelated 2-phenylquinoline isomer is known to be an orange emitter ( $\lambda = 580\text{ nm}$ ).<sup>4c</sup> As shown in Table 1, the emission is drastically quenched with oxygen. The excited state lifetimes of complexes **4a–d** were measured to be 0.23–0.42  $\mu\text{s}$  in aerated CH<sub>3</sub>CN, but increased to 1.77–3.07  $\mu\text{s}$  in deaerated CH<sub>3</sub>CN. Similarly, the quantum yields in the emissions were only 0.006–0.010 in aerated CH<sub>3</sub>CN, but increased to 0.05–0.11 in deaerated CH<sub>3</sub>CN. The nearly diffusion-controlled oxygen quenching rate in CH<sub>3</sub>CN leads to the unambiguous assignment to a phosphorescent emission.

The presence of a fluoro substituent in the phenyl group leads to a fine-tuning of the emission gap. In comparison the parents **4a** and **4b**, bearing one fluoro substituent, exhibit a nearly 10 nm hypsochromic shift in the emission peak wavelength. Moreover, **4d**, with dual fluoro substituents at resonant ortho- and para-positions relative to 8-phenylquinoline, reveals a shorter peak wavelength (656 nm) than that (650 nm) for **4c**, bearing fluoro substituents at the nonresonant meta-position. The results, in combination with a large Stokes shift of >150 nm,<sup>10,11</sup> render further support for the intraligand  $\pi(\text{phenyl})\text{--}\pi^*$  (8-phenylquinoline) character, in part, predicted from the theoretical approach (vide supra). However, the spectral shift is considered to be small, indicating that <sup>3</sup>MLCT (d– $\pi^*$  (8-phenylquinoline)), as predicted by the theoretical approach, should be incorporated in the T<sub>1</sub> manifold.

In brief, the information provided above concludes that the six-membered chelated complexes **4a–d** are equally active in generating the phosphorescent emission via a strong spin–orbit coupling, with phosphorescence lifetimes comparable to those of five-membered iridium phosphors. The relatively low quantum yields of complexes **4a–d** may be caused by two mechanisms:<sup>12</sup> (1) the low energy gap (650–680 nm), the relaxation of which is often accompanied with a nonradiative process, and (2) the nonrigid nature of the six-membered chelated structures, in which the nonradiative deactivation may be associated with the large-amplitude skeletal motion.

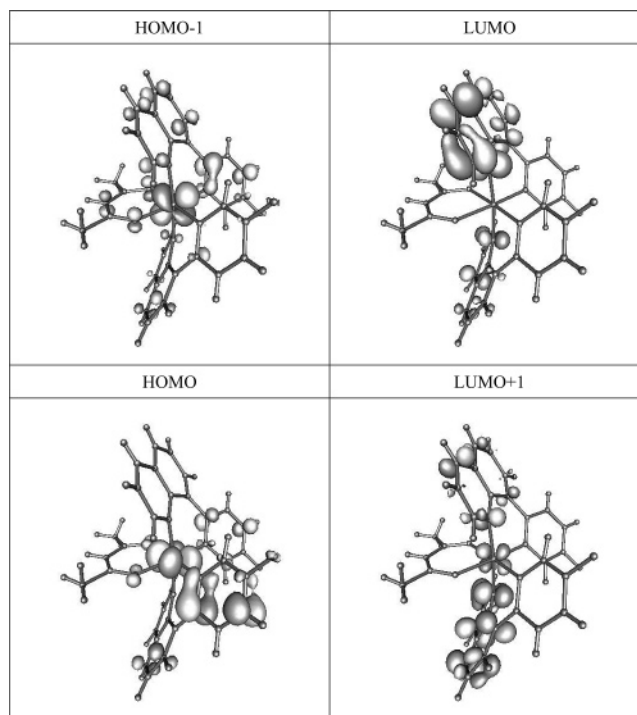
We also examined the electrochemistry of complexes **4a–d** via cyclic voltammetry using ferrocene as the internal standard. These complexes display reversible couples in anodic oxidation within the conditions  $-1.60 < E < +1.80\text{ V}$ . The fluoro substituents **4b–d** shift the potentials [+ (0.99–1.22 V)] to a more positive (anodic) potential when compared to the parent complex **4a** (+0.88V). This phenomenon is consistent with the expectation that the fluoro groups of complexes **4b–d** decrease their HOMO energy levels (vide infra). We

(10) Colombo, M. G.; Hauser, A.; Güdel, H. U. *Inorg. Chem.* **1993**, 32, 3088.

(11) Schmid, B.; Garces, F. O.; Watts, R. J. *Inorg. Chem.* **1994**, 33, 9.

(12) (a) Caspar, J. V.; Meyer, T. J. *J. Phys. Chem.* **1983**, 87, 952. (b) Cummings, S. D.; Eisenberg, R. J. *J. Am. Chem. Soc.* **1996**, 118, 1949.



**Table 2. Calculated Singlet and Triplet Transitions and the Corresponding Frontier Orbitals for Complex 4b**

state	excitation	$E_{\text{cal}}$ (eV)	$\lambda_{\text{cal}}$ (nm)	(f)
Singlet Excited States				
1	HOMO $\rightarrow$ LUMO (+96%)	2.20	563	0.0020
2	HOMO $\rightarrow$ LUMO+1 (+64%), HOMO-1 $\rightarrow$ LUMO (+28%)	2.58	480	0.0121
3	HOMO-1 $\rightarrow$ LUMO (+59%), HOMO $\rightarrow$ LUMO+1 (27%)	2.63	471	0.0227
4	HOMO-1 $\rightarrow$ LUMO+1 (+82%)	3.04	408	0.0126
Triplet Excited States				
1	HOMO-1 $\rightarrow$ LUMO (+91%), HOMO $\rightarrow$ LUMO+1 (9%)	1.99	622	
2	HOMO $\rightarrow$ LUMO (+82%), HOMO-1 $\rightarrow$ LUMO+1 (10%), HOMO $\rightarrow$ LUMO+1 (+8%)	2.03	610	

further measured the HOMO and LUMO energy levels of representative compounds **4a** and **4c** to ascertain the role of fluoro substitution. The HOMO levels for species **4a** and **4c**, according to the photoelectron spectroscopy, were  $-5.25$  and  $-5.42$  eV, respectively. On the basis of their UV absorption spectroscopy, the LUMO levels of **4a** and **4c** are thus calculated to be  $-3.06$  and  $-2.97$  eV, respectively. Compared to parent compound **4a**, the 2,4-difluorophenyl substitution in **4c** leads to a larger decrease in energy for the HOMO than the LUMO orbitals, consistent with the results obtained from both emission spectroscopy and theoretical approaches.

OLEDs using these iridium complexes were fabricated by high-vacuum ( $5 \times 10^{-6}$  Torr) thermal evaporation onto precleaned glass substrates. The devices were prepared<sup>6c</sup> from ITO/NPB(40 nm)/**4b**(6%) in CBP(30 nm)/BCP (10 nm)/Alq(30 nm)/Mg:Ag, in which CBP acts as a hole material, bathocuproine (BCP) as a hole blocker, and 4,4'-bis[*N*-(naphthyl)-*N*-phenylamino]biphenyl (NPB) and tris(8-hydroxyquinoline)aluminum (Alq<sub>3</sub>) as a hole-transport and electron-transport material, respectively. The electroluminescence spectra and corresponding data of complexes **4a–d** are summarized in Table 3. All four devices exhibited deep red emissions in the 634–650 nm region, which are independent of the applied voltage of 6–18 V. No emission bands due to CBP, NPB, or Alq<sub>3</sub> were observed in the electrolu-

**Table 3. Electrophosphorescence Data for Iridium Complexes 4a–d<sup>a</sup>**

complex	brightness (cd/m <sup>2</sup> )	EQE (%)	PE (cd/A)	LE (lm/W)	voltage (V)	CIE coordinate
<b>4a</b>	115 <sup>b</sup>	1.89	0.58	0.26	6.89	X = 0.70
	446 <sup>c</sup>	1.46	0.45	0.15	9.70	Y = 0.30
	1102 <sup>d</sup>	0.88	0.27	0.05	15.82	(668 nm) <sup>e</sup>
<b>4b</b>	213	1.93	1.07	0.47	7.11	X = 0.69
	747	1.38	0.75	0.24	9.70	Y = 0.31
<b>4c</b>	1668	0.76	0.42	0.11	15.22	(650 nm)
	340	2.04	1.70	0.60	8.92	X = 0.68
	1333	1.60	1.33	0.37	11.27	Y = 0.32
<b>4d</b>	3102	0.93	0.77	0.17	14.43	(638 nm)
	448	2.42	2.24	0.83	8.20	X = 0.68
	1604	1.73	1.60	0.45	10.59	Y = 0.32
	3415	0.93	0.86	0.18	14.66	(634 nm)

<sup>a</sup> For each parameter, the data in different rows correspond to those measured at different current densities. <sup>b</sup>  $J = 20$  mA/cm<sup>2</sup>. <sup>c</sup>  $J = 100$  mA/cm<sup>2</sup>. <sup>d</sup>  $J = 400$  mA/cm<sup>2</sup>. <sup>e</sup> The values in parentheses represent the maximum EL emission.

minance (EL) spectra. The corresponding CIE coordinates are  $x = 0.70$ ,  $y = 0.30$  for **4a**,  $x = 0.69$ ,  $y = 0.31$  for **4b**,  $x = 0.68$ ,  $y = 0.32$  for **4c**, and  $x = 0.68$ ,  $y = 0.32$  for **4d**.

The corresponding EL emission peak wavelengths of **4a–d** were observed at 668, 650, 638, and 634 nm, being in good correlation with the tendency of the hypsochromic shift in solution phase. Among these four phosphors, **4c** and **4d** show better EL performance than **4a** and **4b**

in view of maximum brightness and external quantum efficiency. The better EL performance of complexes **4c** and **4d** can be attributed to their hypsochromic shifts of the emission maximum, in which the relatively larger energy gaps are less perturbed by nonradiative processes. Using complex **4c**, a maximum brightness of 3427 cd/m<sup>2</sup> was obtained at 16 V. The external quantum efficiency  $\eta_{\text{ext}}$  of this complex was 2.04 at  $J = 20$  mA/cm<sup>2</sup>, where luminance efficiency and power efficiency were 1.70 cd/A and 0.60 lm/W, respectively. The  $\eta_{\text{ext}}$  value dropped slowly to 0.93 at  $J = 400$  mA/cm<sup>2</sup>, where luminance efficiency and power efficiency were 0.77 cd/A and 0.17 lm/W, respectively. This information reveals that the short phosphorescent lifetime  $\tau = 3.07$   $\mu$ s of complex **4c** reduces the saturation of the emission site (T–T annihilation). These parameters reflect the promising feature of these six-membered iridium complexes in OLED devices.

### Conclusions

In summary, we have prepared the first iridium complexes with a six-membered chelated structure. Similar to the reported five-membered chelated compounds, these species show strong low-lying absorption bands with a mixed <sup>3</sup>MLCT/<sup>3</sup> $\pi$ – $\pi^*$  character, indicating an unusually strong spin–orbit coupling due to the heavy atom (Ir) effect and, perhaps, the short, six-membered coordinating distance. These complexes show deep red phosphorescent emissions (650–680 nm) with short excited lifetimes of 1.7–3.0  $\mu$ s and moderate quantum yields of 0.05–0.11 in deaerated CH<sub>3</sub>CN. The decent electroluminescent performance of these species demonstrates their promising feature in OLED displays.

### Experimental Section

#### Syntheses. 8-Quinolyl trifluoromethanesulfonate (**2**).

To a pyridine solution of 8-hydroquinoline (4.00 g, 27.6 mmol) was slowly added trifluoromethanesulfonic anhydride (9.32 g, 33.1 mmol) at 0 °C. The resulting mixture was stirred at 0 °C for 5 min and warmed to 23 °C with stirring for 25 h. The resulting mixture was poured into water and extracted with ethyl acetate. The ethyl acetate extract was washed sequentially with water, 10% aqueous hydrochloric acid solution, and water, dried over MgSO<sub>4</sub>, and concentrated to yield an oil. Chromatography (hexanes/EtOAc, 10:1,  $R_f = 0.40$ ) afforded 8-quinolyl trifluoromethanesulfonate (**2**) as a colorless oil, which solidified on standing (6.88 g, 90%). IR (Nujol): 2960(m), 2921(m), 2865(m), 2852(m), 1576(s), 1541(s), 1398(s), 1351(s), 1186(s). <sup>1</sup>H NMR (CDCl<sub>3</sub>, 600 MHz):  $\delta$  8.97 (dd,  $J = 4.2, 1.6$  Hz, 1H), 8.13 (dd,  $J = 8.2, 1.6$  Hz, 1H), 7.77 (dd,  $J = 8.2, 1.1$  Hz, 1H), 7.57 (d,  $J = 8.2$  Hz, 1H), 7.49 (t,  $J = 8.1$  Hz, 1H), 7.44 (dd,  $J = 8.3, 4.2$  Hz, 1H). <sup>13</sup>C NMR (CDCl<sub>3</sub>, 150 MHz):  $\delta$  151.4 (d,  $J_{\text{CF}} = 25.8$  Hz), 145.9, 140.8, 135.7, 129.6, 128.2, 125.8, 122.5, 120.9, 118.9 (q,  $J_{\text{CF}} = 318.6$  Hz). HRMS (70 eV): calcd for C<sub>10</sub>H<sub>6</sub>F<sub>3</sub>NO<sub>3</sub>S 277.0020, found 277.0026. Anal. Calcd for C<sub>10</sub>H<sub>6</sub>F<sub>3</sub>NO<sub>3</sub>S: C, 43.33; H, 2.18; N, 5.05. Found: C, 43.25; H, 2.19; N, 5.08.

**8-(Phenyl)quinoline (3a).** To a toluene (30 mL) solution of 8-quinolyl trifluoromethanesulfonate (**2**) (2.26 g, 8.15 mmol) were added benzyl boronic acid (1.04 g, 8.55 mmol), LiCl (1.04 g, 24.4 mmol), Pd(PPh<sub>3</sub>)<sub>4</sub> (0.47 g, 0.41 mmol), and Na<sub>2</sub>CO<sub>3</sub> (2.0 M, 5.3 mL) under N<sub>2</sub>, and the resulting suspension was heated to reflux for 12 h. The mixture was extracted with ethyl acetate, washed with aqueous NaCl solution, dried with MgSO<sub>4</sub>, and concentrated to yield a brownish solid. Chroma-

tography (hexanes/EtOAc, 20: 1,  $R_f = 0.50$ ) afforded 8-phenylquinoline (**3a**) as a colorless oil, which solidified on standing (1.62 g, 97%). IR (Nujol): 2960(m), 2921(m), 2865(m), 2852(m), 1576(s), 1541(s), 1398(s), 1351(s). <sup>1</sup>H NMR (CDCl<sub>3</sub>, 400 MHz):  $\delta$  8.95 (d,  $J = 4.2$  Hz, 1H), 8.19 (d,  $J = 8.2$  Hz, 1H), 7.81 (d,  $J = 8.0$  Hz, 1H), 7.74–7.69 (m, 2H), 7.59 (t,  $J = 8.4$  Hz, 1H), 7.50 (t,  $J = 8.0$  Hz, 1H), 7.43–7.39 (m, 4H). <sup>13</sup>C NMR (CDCl<sub>3</sub>, 100 MHz):  $\delta$  150.2, 146.0, 140.9, 139.5, 136.2, 130.6, 130.3, 128.7, 127.9, 127.5, 127.3, 126.2, 120.9. HRMS (70 eV) calcd for C<sub>15</sub>H<sub>11</sub>N 205.0891, found 205.0887. Anal. Calcd for C<sub>15</sub>H<sub>11</sub>N: C, 87.77; H, 5.40; N, 6.82. Found: C, 87.77; H, 5.39; N, 6.83.

**8-(4-Fluorophenyl)quinoline (3b).** This compound was prepared from 8-quinolyl trifluoromethanesulfonate (**2**) and 4-fluorobenzylboronic acid according to the procedure described for 8-phenylquinoline (**3a**). IR (Nujol): 2960(m), 2921(m), 2865(m), 2852(m), 1576(s), 1541(s), 1398(s), 1351(s), 1176(s). <sup>1</sup>H NMR (CDCl<sub>3</sub>, 600 MHz):  $\delta$  8.95 (dd,  $J = 6.0, 1.8$  Hz, 1H), 8.21 (dd,  $J = 8.4, 1.8$  Hz, 1H), 7.82 (dd,  $J = 8.1, 1.2$  Hz, 1H), 7.69 (d,  $J = 7.2$  Hz, 1H), 7.67–7.64 (m, 2H), 7.59 (t,  $J = 7.8$  Hz, 1H), 7.41 (dd,  $J = 8.4, 4.2$  Hz, 1H). <sup>13</sup>C NMR (CDCl<sub>3</sub>, 150 MHz):  $\delta$  162.4 (d,  $J_{\text{CF}} = 244.8$  Hz), 150.2, 145.7, 139.7, 136.6, 135.2, 132.1 (d,  $J_{\text{CF}} = 7.7$  Hz), 130.3, 128.8, 127.7, 126.3, 121.1, 115.0 (d,  $J_{\text{CF}} = 21.2$  Hz). HRMS (70 eV) calcd for C<sub>15</sub>H<sub>10</sub>FN 223.0797, found 223.0802. Anal. Calcd for C<sub>15</sub>H<sub>10</sub>FN: C, 80.70; H, 4.51; N, 6.27. Found: C, 80.73; H, 4.50; N, 6.22.

**8-(3,5-Difluorophenyl)quinoline (3c).** This compound was similarly prepared from 8-quinolyl trifluoromethanesulfonate (**2**) and 4,5-difluorophenylboronic acid according to the procedure described for 8-phenylquinoline (**3a**). IR (Nujol): 2962(m), 2920(m), 2861(m), 2854(m), 1573(s), 1540(s), 1394(s), 1350(s), 1183(s). <sup>1</sup>H NMR (CDCl<sub>3</sub>, 600 MHz):  $\delta$  8.95 (dd,  $J = 4.1, 1.8$  Hz, 1H), 8.19 (d,  $J = 8.4$  Hz, 1H), 7.85 (d,  $J = 8.2$  Hz, 1H), 7.69 (dd,  $J = 7.1, 1.1$  Hz, 1H), 7.58 (t,  $J = 7.6$  Hz, 1H), 7.42 (dd,  $J = 8.3, 4.2$  Hz, 1H), 7.26–7.22 (m, 2H), 6.85 (tt,  $J = 9.0, 2.4$  Hz, 1H). <sup>13</sup>C NMR (CDCl<sub>3</sub>, 150 MHz):  $\delta$  162.5 (dd,  $J_{\text{CF}} = 245.8, 13.1$  Hz), 150.4 (d,  $J_{\text{CF}} = 16.8$  Hz), 145.5, 142.6, 138.4, 136.4, 130.2, 128.7, 128.5 (d,  $J_{\text{CF}} = 15.9$  Hz), 126.2, 121.3, 113.6 (dd,  $J_{\text{CF}} = 20.8, 4.7$  Hz), 102.7 (t,  $J_{\text{CF}} = 25.2$  Hz). HRMS (70 eV): calcd for C<sub>15</sub>H<sub>9</sub>F<sub>2</sub>N 241.0703, found 241.0696. Anal. Calcd for C<sub>15</sub>H<sub>9</sub>F<sub>2</sub>N: C, 74.68; H, 3.76; F, 15.75; N, 5.81. Found: C, 74.65; H, 3.77; F, 15.75; N, 5.83.

**8-(2,4-Difluorophenyl)quinoline (3d).** This compound was prepared from 8-quinolyl trifluoromethanesulfonate (**2**) and 4,5-difluorophenylboronic acid according to the procedure described for 8-phenylquinoline (**3a**). <sup>1</sup>H NMR (CDCl<sub>3</sub>, 600 MHz):  $\delta$  8.91 (dd,  $J = 4.2, 1.8$  Hz, 1H), 8.17 (dd,  $J = 9.3, 1.8$  Hz, 1H), 7.85 (dd,  $J = 8.1, 1.8$  Hz, 1H), 7.68 (d,  $J = 7.2$  Hz, 1H), 7.59 (t,  $J = 7.5$  Hz, 1H), 7.49–7.45 (m, 1H), 7.39 (dd,  $J = 8.4, 4.0$  Hz, 1H), 7.00–6.95 (m, 2H). <sup>13</sup>C NMR (CDCl<sub>3</sub>, 150 MHz) [ppm]:  $\delta$  162.6 (dd,  $J_{\text{CF}} = 247.1, 11.7$  Hz), 160.2 (dd,  $J_{\text{CF}} = 248.1, 11.9$  Hz), 150.3, 146.1, 136.2, 134.5, 133.1 (dd,  $J_{\text{CF}} = 8.6, 5.0$  Hz), 131.0, 128.4, 128.2 (d,  $J_{\text{CF}} = 29.3$  Hz), 125.9 (d,  $J_{\text{CF}} = 24.3$  Hz), 121.1, 110.9 (d,  $J_{\text{CF}} = 21.0$  Hz), 103.9 (t,  $J_{\text{CF}} = 25.8$  Hz). HRMS (70 eV): calcd for C<sub>15</sub>H<sub>9</sub>F<sub>2</sub>N 241.0703, found 241.0698. Anal. Calcd for C<sub>15</sub>H<sub>9</sub>F<sub>2</sub>N: C, 74.68; H, 3.76; N, 5.81. Found: C, 74.66; H, 3.76; N, 5.80.

**Ir[8-(Phenyl)quinoline]<sub>2</sub>(acac) (4a).** To a mixed solvent (2-ethoxyethanol/H<sub>2</sub>O, 3:1, 60 mL) of 8-phenylquinoline (**3a**) (1.41 g, 6.85 mmol) was added IrCl<sub>3</sub>·3H<sub>2</sub>O (1.18 g, 3.34 mmol), and the reaction mixture was heated to 120 °C for 24 h. The resulting solution was concentrated at 55 °C, the precipitates were filtered, and the solid was collected and washed with 100 mL of water, 50 mL of hexane, and 50 mL of ethyl ether. The mass of this dark brownish solid was 2.02 g (95%). To a 2-ethoxyethanol (15 mL) solution of this solid (2.02 g, 1.59 mmol) were added acetyl acetone (0.39 g, 3.97 mmol) and sodium carbonate (0.37 g, 3.49 mmol), and the reaction mixture was heated to 120 °C for 12 h. After cooling to room temperature, a dark brownish precipitate was filtered off and washed

with 200 mL of water, 100 mL of hexane, and 100 mL of ether. The crude product was chromatographed on a silica column ( $\text{CH}_2\text{Cl}_2/\text{hexane}$ , 1:6,  $R_f = 0.5$ ) to yield the pure ( $C\wedge N$ )<sub>2</sub>Ir(acac) **4a** (0.69 g, 32%). IR (Nujol): 2974 (m), 2923(m), 2865(m), 2853-(m), 1572(s), 1541(s), 1393(s), 1352(s), 1180(s). <sup>1</sup>H NMR ( $\text{CDCl}_3$ , 600 MHz):  $\delta$  8.55 (d,  $J = 6.9$  Hz, 2H), 8.34 (d,  $J = 8.9$  Hz, 2H), 8.08 (d,  $J = 9.7$  Hz, 2H), 7.69 (d,  $J = 9.2$  Hz, 2H), 7.64–7.62 (m, 4H), 7.05 (t,  $J = 8.2$  Hz, 2H), 6.95 (d,  $J = 8.0$  Hz, 2H), 6.70–6.65 (m, 4H) 4.24 (s, 1H), 1.07 (s, 6H). <sup>13</sup>C NMR ( $\text{CDCl}_3$ , 150 MHz):  $\delta$  184.4, 154.3, 145.0, 142.6, 142.3, 141.7, 137.8, 136.5, 128.7, 127.7, 126.6, 126.6, 125.4, 122.7, 120.4, 99.1, 27.3. HRMS (70 eV): calcd for  $\text{C}_{35}\text{H}_{23}\text{F}_4\text{IrN}_2\text{O}_2$  700.1702, found 700.1722. Anal. Calcd for  $\text{C}_{35}\text{H}_{23}\text{F}_4\text{IrN}_2\text{O}_2$ : C, 60.07; H, 3.89; N, 4.00. Found: C, 60.12; H, 3.82; N, 4.03.

**Ir[8-(4-fluorophenyl)quinoline]<sub>2</sub>(acac) (4b)**. This compound was prepared from 8-(4-fluorophenyl)quinoline (**3b**), acetyl acetone, and  $\text{IrCl}_3 \cdot 3\text{H}_2\text{O}$  according to the procedure described for Ir(8-phenylquinoline)<sub>2</sub>(acac) (**4a**). IR (Nujol): 2972 (m), 2922(m), 2865(m), 2851(m), 1579(s), 1543(s), 1392-(s), 1351(s), 1177(s). <sup>1</sup>H NMR ( $\text{CDCl}_3$ , 500 MHz):  $\delta$  8.52 (d,  $J = 5.5$  Hz, 2H), 8.25 (dd,  $J = 6.5$ , 2.0 Hz, 2H), 8.12 (dd,  $J = 8.0$ , 1.5 Hz, 2H), 7.65–7.61 (m, 6H), 7.00 (dd,  $J = 8.0$ , 5.1 Hz, 2H), 6.74 (dt,  $J = 8.5$ , 3.2 Hz, 2H), 6.29 (dd,  $J = 10.5$ , 2.0 Hz, 2H), 4.24 (s, 1H), 1.07 (s, 6H). <sup>13</sup>C NMR ( $\text{CDCl}_3$ , 125 MHz):  $\delta$  184.6, 161.0 (d,  $J_{\text{CF}} = 247.3$  Hz), 154.3, 144.7, 143.7, 141.6, 137.1, 134.2, 128.7, 128.0 (d,  $J_{\text{CF}} = 8.9$  Hz), 127.8 (d,  $J_{\text{CF}} = 4.4$  Hz), 127.1 (d,  $J_{\text{CF}} = 15.5$  Hz), 125.5, 120.5, 110.0 (d,  $J_{\text{CF}} = 22.3$  Hz), 99.3, 27.3. HRMS (70 eV): calcd for  $\text{C}_{35}\text{H}_{25}\text{F}_2\text{IrN}_2\text{O}_2$  736.1513, found 736.1511. Anal. Calcd for  $\text{C}_{35}\text{H}_{25}\text{F}_2\text{IrN}_2\text{O}_2$ : C, 57.13; H, 3.42; N, 3.81. Found: C, 57.03; H, 3.40; N, 3.79.

**Ir[8-(3,5-difluorophenyl)quinoline]<sub>2</sub>(acac) (4c)**. This compound was prepared from 8-(3,5-difluorophenyl)quinoline (**3c**), acetyl acetone, and  $\text{IrCl}_3 \cdot 3\text{H}_2\text{O}$  according to the procedure described for Ir[8-(phenyl)quinoline]<sub>2</sub>(acac) (**4a**). IR (Nujol): 2981 (m), 2933(m), 2861(m), 2859(m), 1570(s), 1544(s), 1390-(s), 1356(s), 1182(s). <sup>1</sup>H NMR ( $\text{CDCl}_3$ , 600 MHz):  $\delta$  8.54 (dd,  $J = 5.4$ , 1.4 Hz, 2H), 8.29 (dd,  $J = 7.4$ , 1.3 Hz, 2H), 8.07 (dd,  $J = 8.1$ , 1.4 Hz, 2H), 7.69–7.62 (m, 4H), 7.37 (dd,  $J = 11.3$ , 2.5 Hz, 2H), 6.98 (dd,  $J = 8.1$ , 5.3 Hz, 2H), 6.32 (dt,  $J = 9.0$ , 2.5 Hz, 2H), 4.23 (s, 1H), 1.06 (s, 6H). <sup>13</sup>C NMR ( $\text{CDCl}_3$ , 150 MHz):  $\delta$  184.1, 169.4 (dd,  $J_{\text{CF}} = 234.7$ , 11.3 Hz), 160.8 (dd,  $J_{\text{CF}} = 234.9$ , 13.5 Hz), 152.8, 144.5, 141.5 (d,  $J_{\text{CF}} = 7.7$  Hz), 141.4, 136.7, 128.7, 128.4, 127.5, 126.6, 120.4, 115.8 (d,  $J_{\text{CF}} = 34.8$  Hz), 110.0 (d,  $J_{\text{CF}} = 22.3$  Hz), 109.9 (d,  $J_{\text{CF}} = 20.1$  Hz), 101.6 (dd,  $J_{\text{CF}} = 33.3$ , 23.7 Hz), 99.4, 27.1. HRMS (70 eV): calcd for  $\text{C}_{35}\text{H}_{23}\text{F}_4\text{IrN}_2\text{O}_2$  772.1325, found 772.1328. Anal. Calcd for  $\text{C}_{35}\text{H}_{23}\text{F}_4\text{IrN}_2\text{O}_2$ : C, 54.47; H, 3.00; N, 3.63. Found: C, 54.40; H, 3.10; N, 3.69.

**Ir[8-(3,5-difluorophenyl)quinoline]<sub>2</sub>(acac) (4d)**. This compound was prepared from 8-(3,5-difluorophenyl)quinoline (**3d**), acetyl acetone, and  $\text{IrCl}_3 \cdot 3\text{H}_2\text{O}$  according to the procedure described for Ir[8-(phenyl)quinoline]<sub>2</sub>(acac) (**4a**). IR (Nujol): 2983 (m), 2932(m), 2860(m), 2855(m), 1572(s), 1545(s), 1393-(s), 1352(s), 1184(s). <sup>1</sup>H NMR ( $\text{CDCl}_3$ , 600 MHz):  $\delta$  8.50 (dd,  $J = 5.3$ , 1.4 Hz, 2H), 8.35 (dt,  $J = 6.7$ , 1.4 Hz, 2H), 8.14 (dd,  $J = 8.1$ , 1.4 Hz, 2H), 7.68–7.62 (m, 4H), 7.04 (dd,  $J = 8.1$ , 5.3 Hz, 2H), 6.51 (dt,  $J = 10.9$ , 2.6 Hz, 2H), 6.12 (dd,  $J = 9.5$ , 2.6 Hz, 2H), 4.18 (s, 1H), 1.02 (s, 6H). <sup>13</sup>C NMR ( $\text{CDCl}_3$ , 150 MHz):  $\delta$  184.4, 161.1 (dd,  $J_{\text{CF}} = 249.7$ , 13.1 Hz), 158.1 (dd,  $J_{\text{CF}} = 253.7$ , 12.0 Hz), 154.2, 146.2, 145.4, 137.3, 135.1, 132.2 (d,  $J_{\text{CF}} = 15.0$  Hz), 128.6, 127.1, 125.7, 123.3 (d,  $J_{\text{CF}} = 15.6$  Hz), 120.4, 99.2, 98.7 (t,  $J_{\text{CF}} = 28.1$  Hz), 27.0. HRMS (70 eV): calcd for  $\text{C}_{35}\text{H}_{23}\text{F}_4\text{IrN}_2\text{O}_2$  772.1325, found 772.1311. Anal. Calcd for  $\text{C}_{35}\text{H}_{23}\text{F}_4\text{IrN}_2\text{O}_2$ : C, 54.47; H, 3.00; N, 3.63. Found: C, 54.39; H, 3.05; N, 3.72.

**Measurement of Lifetime, Quantum Yields, and Photoelectron Spectroscopy.** Steady state absorption and emission spectra were recorded by a Hitachi (U-3310) spectrophotometer and an Edinburgh (FS920) fluorimeter, respectively. Both wavelength-dependent excitation and emission response

of the fluorimeter have been calibrated. A configuration of front-face excitation was used to measure the emission of the solid sample, in which the cell was made by assembling two edge-polished quartz plates with various Teflon spacers. A combination of appropriate filters was used to avoid the interference from the scattering light. To determine the photoluminescence quantum yield in solution, samples were degassed by three freeze–pump–thaw cycles under vigorous stirring conditions. 4-(Dicyanomethylene)-2-methyl-6-(*p*-dimethylaminostyryl)-4*H*-pyran (DCM,  $\lambda_{\text{em}} = 620$  nm, Exciton, Inc.) in methanol was used as a reference, assuming a quantum yield of 0.43 with a 460 nm excitation.<sup>13</sup>

Lifetime studies were performed by an Edinburgh FL 900 photon-counting system with a hydrogen-filled/or a nitrogen lamp as the excitation source. Data were analyzed using the nonlinear least squares procedure in combination with an iterative convolution method. Single-exponential emission decay was obtained for all studied compounds.

**Theoretic Calculation.** Calculations on the electronic ground states of complex **4b** were carried out using B3LYP density functional theory.<sup>14,15</sup> Double- $\zeta$  quality basis sets were employed for the ligands (6-31G\*) and the Ir atom (LANL2DZ). A relativistic effective core potential on Ir<sup>16</sup> replaced the inner core electrons leaving the outer core [(5s)<sup>2</sup>(5p)<sup>6</sup>] electrons and the (5d)<sup>6</sup> valence electrons of Ir<sup>III</sup>. The ground state geometry was adapted from the truncated X-ray data without further optimization. At the geometry, a time-dependent DFT (TD-DFT) calculation<sup>17</sup> using the B3LYP functional was performed. Typically, the lowest 10 triplet and 10 singlet roots of the nonhermitian eigenvalue equations were obtained to determine the vertical excitation energies. Oscillator strengths were deduced from the dipole transition matrix elements (for singlet states only). The ground state B3LYP and excited state TDDFT calculations were carried out using Gaussian98.<sup>18</sup>

**OLED Fabrication and Measurements.** Prepatterned ITO substrates with an effective individual device area of 3.14 mm were cleaned by sonication in a detergent solution for 3 min and then washed with a large amount of doubly distilled water. Further sonication in ethanol for 3 min was done before blowing dry with a stream of nitrogen. The ITO substrates were then treated with O<sub>2</sub> plasma for 1 min before loaded into the vacuum chamber. The organic layers were deposited thermally at a rate of 0.1–0.3 nm/s under a pressure of  $\sim 5 \times 10^{-6}$  Torr in an Ulvac Cryogenic deposition system. Typical devices were constructed with 40 nm of NPB as the hole-transporting layer (HTL), 30 nm CBP doped with iridium complex as the emission layer, and 30 nm of Alq<sub>3</sub> as the ETL. An alloy of magnesium and silver (ca. 10:1, 50 nm) was deposited as the cathode, which was capped with 100 nm of

(13) Drake, J. M.; Lesiecki, M. L.; Camaioni, D. M. *Chem. Phys. Lett.* **1985**, *113*, 530.

(14) Lee, C.; Yang, W.; Parr, R. G. *Phys. Rev. B* **1988**, *37*, 785.

(15) Becke, A. D. *J. Chem. Phys.* **1993**, *98*, 5648.

(16) Hay, P. J.; Wadt, W. R. *J. Chem. Phys.* **1985**, *82*, 299–310.

(17) (a) Jamorski, C.; Casida, M. E.; Salahub, D. R. *J. Chem. Phys.* **1996**, *104*, 5134. (b) Petersilka, M.; Grossmann, U. J.; Gross, E. K. U. *Phys. Rev. Lett.* **1996**, *76*, 1212. (c) Bauernschmitt, R.; Ahlrichs, R.; Henrichs, F. H.; Kappes, M. M. *J. Am. Chem. Soc.* **1998**, *120*, 5052. (d) Casida, M. E. *J. Chem. Phys.* **1998**, *108*, 4439. (e) Stratmann, R. E.; Scuseria, G. E.; Frisch, M. J. *J. Chem. Phys.* **1998**, *109*, 8218.

(18) Frisch, M. J.; Trucks, G. W.; Schlegel, H. B.; Scuseria, G. E.; Robb, M. A.; Cheeseman, J. R.; Zakrzewski, V. G.; Montgomery, J. A., Jr.; Stratmann, R. E.; Burant, J. C.; Dapprich, S.; Millam, J. M.; Daniels, A. D.; Kudin, K. N.; Strain, M. C.; Farkas, O.; Tomasi, J.; Barone, V.; Cossi, M.; Cammi, R.; Mennucci, B.; Pomelli, C.; Adamo, C.; Clifford, S.; Ochterski, J.; Petersson, G. A.; Ayala, P. Y.; Cui, Q.; Morokuma, K.; Malick, D. K.; Rabuck, A. D.; Raghavachari, K.; Foresman, J. B.; Cioslowski, J.; Ortiz, J. V.; Stefanov, B. B.; Liu, G.; Liashenko, A.; Piskorz, P.; Komaromi, I.; Gomperts, R.; Martin, R. L.; Fox, D. J.; Keith, T.; Al-Laham, M. A.; Peng, C. Y.; Nanayakkara, A.; Gonzalez, C.; Challacombe, M.; Gill, P. M. W.; Johnson, B. G.; Chen, W.; Wong, M. W.; Andres, J. L.; Head-Gordon, M.; Replogle, E. S.; Pople, J. A. *Gaussian 98*, revision A.9; Gaussian, Inc.: Pittsburgh, PA, 1998.



silver. The current–voltage–luminance was measured in ambient with a Keithley 2400 source meter and a Newport 1835C optical meter equipped with 818ST silicon photodiode.

The HOMO energy level was measured with a photoelectron spectrometer (model AC-2, Riken Keiki) on an evaporated thin film of the organic compound.

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**Supporting Information Available:** Crystallographic information files for **4b**; table of crystal data and structure refinement for **4b** (Table S1), table of atomic coordinates and equivalent isotropic displacement parameters for **4b** (Table S2), table of bond lengths and angles for **4b** (Table S3), table of anisotropic displacement parameters for **4b** (Table S4), and table of hydrogen coordinates and isotropic displacement parameters for **4b** (Table S5). These materials are available free of charge via the Internet at <http://pubs.acs.org>.

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