

Paramagnetic-Resonance Spectra of the $3d^5$ Configuration of Chromium in ZnSe and ZnTe

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(Received 1 November 1963)

Paramagnetic-resonance measurements have been made for the $3d^5$ configuration of Cr^+ in ZnSe and ZnTe. The g factor, the cubic crystalline field parameter a , the hyperfine interaction with Cr^{53} , $A(\text{Cr}^{53})$, and the hyperfine interaction A_{Zn} with Zn^{67} nuclei occupying any of the twelve equivalent cation sites nearest the chromium have been determined. These resonance parameters are compared with those previously measured for Cr^+ in cubic ZnS and CdTe and the variations from crystal to crystal are found to be consistent with changes in the covalency of the bonds and with lattice distortions caused by incorporating Cr^+ substitutionally for the much smaller divalent zinc. The hyperfine interaction with Zn^{67} (or Cd^{111} and Cd^{113}) at second-neighbor sites is, unlike the interaction of the isoelectronic Mn^{++} with these sites, found to vary from crystal to crystal in a manner consistent with changes in the covalent bonding.

INTRODUCTION

CHROMIUM has previously been shown to act as an acceptor in II-VI compounds.^{1,2} Chromium enters the lattice substitutionally for the divalent cation. In this electrically neutral state, independent of whether the bonding is ionic, or covalent, or some mixture of both, the electronic charge centered on the chromium site can be designated by a $3d^4$ configuration. When chromium accepts an electron, paramagnetic resonance has shown both in the case of¹ ZnS and² CdTe that the electron attaches itself to the chromium as a d electron and the chromium can in this state be designated by a $3d^5$ configuration. In this paper the observation by paramagnetic resonance of the $3d^5$ configuration of chromium in two other II-VI compounds of the zinc-blende structure, ZnSe and ZnTe, is reported. Although the bonding in these compounds is a mixture of covalent and ionic bonding, the $3d^5$ configuration of Cr will, for the sake of brevity, be designated by Cr^+ . For Cr^+ in these compounds the g factor, the cubic field splitting parameter a , the hyperfine interaction $A(\text{Cr}^{53})$ with Cr^{53} , and the hyperfine interaction A_{Zn} with Zn^{67} (4% abundant) occupying second-neighbor positions have been determined. A comparison of the variation of these parameters for Cr^+ in the four zinc-blende lattices for which they have been observed is made. The variations are consistent with changes expected on the basis of differences in the degree of covalent bonding in the various lattices, and also of lattice distortions in incorporating Cr in the zinc compounds. In striking contrast to the isoelectronic Mn^{++} in these compounds, the hyperfine interaction of Cr with second-neighbor sites is found to vary from lattice to lattice in a manner consistent with variations in the degree of covalent bonding.

EXPERIMENTAL

The crystals of ZnSe were grown by L. Suchow and J. Scardefield. These were grown from the melt either

¹ J. Dieleman, R. S. Title, and W. V. Smith, Phys. Letters **1**, 334 (1962); R. S. Title, Phys. Rev. **131**, 623 (1963).

² G. W. Ludwig and M. R. Lorenz, Phys. Rev. **131**, 601 (1963).

in an oven or by rf heating. Chromium was found to be invariably present as an impurity in concentrations of the order of one part in a million. The chromium is believed to have diffused into the ZnSe from the graphite crucibles at the 1500°C temperature of the melt.

The crystals of ZnTe were grown by J. Kucza. They were grown from a melt containing 50% excess tellurium which has the effect of lowering the melting point. Chromium was added in metallic form.

The resonance measurements were carried out at 77°K. At this temperature some of the chromium centers were found to have a $3d^5$ configuration in both ZnSe and ZnTe. The number in this configuration could be increased some threefold by irradiating the crystal with light of energy near the band gap. When the light was removed, the number of Cr^+ would remain undiminished over a period of hours, indicating that the level due to Cr^+ is deep with respect to the conduction band. This effect of light of energy near the band gap is similar to that of Cr in ZnS.¹

RESULTS

The Cr^+ resonance spectra in ZnSe and ZnTe are shown along with that in ZnS in Fig. 1. The spectra were taken with the magnetic field parallel to the [001] direction. The five fine-structure transitions expected for the $^6S_{5/2}$ ground state of the $3d^5$ configuration have been indicated by the labels a to e . The separation of the fine-structure lines depends on the cubic field parameter a and the orientations of the crystal with respect to the field.³ The spectra in the [001] direction of Fig. 1 along with measurements in other directions have been used to determine a .

A hyperfine interaction with the 10% abundant isotope Cr^{53} is observed in each case. It is most clearly resolved in a direction where the five fine-structure lines coincide. One such direction is in the (110) plane approximately 30° from the [001] direction. The spectrum in this direction is shown in Fig. 2 for ZnSe

³ B. Bleaney and D. J. E. Ingram, Proc. Roy. Soc. (London) **A205**, 336 (1951).

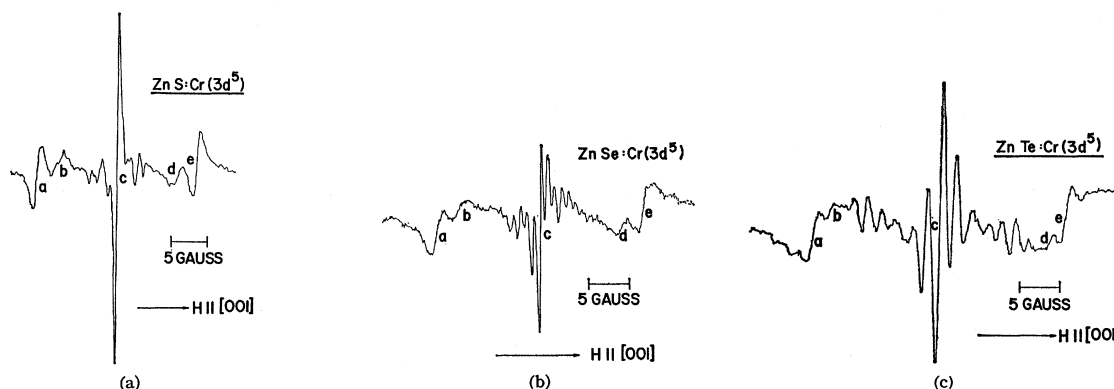


FIG. 1. Paramagnetic-resonance spectrum at 77°K of Cr^{3+} in (a) ZnS, (b) ZnSe, (c) ZnTe—taken with the magnetic field H parallel to the $[001]$ direction.

and the four hyperfine transitions due to $\text{Cr}^{53}(I=3/2)$ are indicated by the arrows.

In Fig. 1 the structure on the $1/2 \rightarrow -1/2$ transition (labeled c in each case) is similar to the structure that has been observed for the $1/2 \rightarrow -1/2$ transition of Mn^{2+} in ZnS,⁴ CdS,^{5,6} and CdTe,⁶ and of Cr in CdTe.² It is attributed to a virtually isotropic hyperfine coupling between the electrons in the $3d^5$ configuration and the magnetic nuclei Zn⁶⁷ (or Cd¹¹³ and Cd¹¹⁵) occupying any of the twelve equivalent cation sites nearest the impurity. The parameter describing this interaction is A_{Zn} (or A_{Cd}) and is equal to twice the spacing between the hyperfine lines.^{5,6} It is seen in Fig. 1 that A_{Zn} increases in going from ZnS to ZnSe to ZnTe.

The resonance parameters that have been measured for Cr^{3+} in ZnSe and ZnTe are given in Table I along with the previous measurements in¹ ZnS and² CdTe.

DISCUSSION OF THE RESULTS

The g Values

Watanabe⁷ has shown that for the ${}^6S_{5/2}$ ground state of the $3d^5$ configuration the g value should be less than

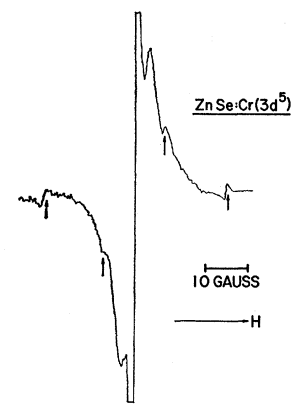


FIG. 2. The paramagnetic-resonance spectrum of ZnSe:Cr³⁺ taken with H in the (110) plane 30° from the $[001]$ direction.

the free-electron value 2.0023. This is seen to be the case for the four crystals in Table I with the g value in ZnTe being equal to 2.0023. The variation in the g values from crystal to crystal is attributed to the effects of covalent bonding. Fidone and Stevens⁸ pointed out that the g value will depend on the degree of covalent bonding. Watanabe⁹ has calculated that an increase in the amount of covalent bonding leads to an increase in the g value. The degree of covalent bonding increases as the anion is changed from sulfur, to selenium, to tellurium, and a corresponding increase is seen in the g value for the three zinc compounds.

The g value for Cr^{3+} in ZnTe is seen to be significantly higher than in the cadmium compound of the same anion, CdTe. This is due to an effect which affects not only the g value but, as will be seen, the other resonance parameters as well. It is a result of a lattice distortion caused by incorporating Cr for the smaller zinc ion. The ionic radius of Cr^{3+} may be extrapolated from the sizes of the isoelectronic Mn^{2+} (0.80 Å) and Fe^{2+} (0.64 Å) to be about 0.96 Å. The Zn^{2+} ion has a radius of 0.74 Å whereas the Cd^{2+} ion has a radius of 0.97 Å. Cr^{3+} should therefore substitute for Cd^{2+} without any lattice distortion. However, its substitution for the much smaller Zn^{2+} ion requires the lattice to undergo a local expansion to accommodate the chromium. The charge on the Cr^{3+} will, therefore, in the zinc-compound lattice overlap more with the charge on the neighboring anions than in the cadmium compound. This implies that the chromium-doped zinc compound is more covalent than the cadmium compound and accounts for the larger g value for Cr^{3+} in ZnTe than in CdTe. Lattice distortions due to size mismatch have been previously reported for Mn^{2+} in II-VI compounds.¹⁰

The a Values

The a value is a measure of the effect of the crystalline field at the Cr^{3+} site produced by the charge on the anions.

⁴ J. Schneider (private communication).

⁵ P. B. Dorain, Phys. Rev. **112**, 1058 (1958).

⁶ J. Lambe and C. Kikuchi, Phys. Rev. **119**, 1256 (1960).

⁷ H. Watanabe, Progr. Theoret. Phys. (Kyoto) **18**, 405 (1957).

⁸ I. Fidone and K. W. H. Stevens, Proc. Phys. Soc. (London) **A73**, 116 (1959).

⁹ H. Watanabe, Bull. Am. Phys. Soc. **8**, 439 (1963).

¹⁰ R. S. Title, Phys. Rev. **131**, 2503 (1963).

TABLE I. The resonance parameters of Cr³⁺ in four zinc-blende lattices.

	Cation-anion separation <i>A</i>	T °K	<i>g</i>	<i>a</i>	<i>A</i> (Cr ⁵³) 10 ⁻⁴ cm ⁻¹	<i>A</i> _{Zn} or <i>A</i> _{Cd}
ZnS	2.35	77	1.9995±0.0005	3.9 ±0.1	13.4 ±0.1	1.23±0.15
ZnSe	2.46	77	2.0016±0.0005	5.35±0.08	13.3 ±0.2	1.81±0.10
ZnTe	2.63	77	2.0023±0.0005	6.60±0.2	12.4 ±0.2	3.59±0.11
CdTe	2.81	4.2	1.9997±0.0003	3.1 ±0.6	12.781±0.005	10.6

The crystalline field should decrease as the separation of the anions and Cr³⁺ is increased. However, as seen in Table I, the *a* value is seen to increase for the zinc compounds in going from smallest lattice ZnS to the largest ZnTe. A similar result is found for Mn²⁺ in these compounds¹¹ and is due to changes in the degree of covalent bonding. As the amount of covalent bonding increases, the charge on the anions and on the cations is no longer concentrated at the lattice points but moves out into the intermediate region. The crystalline field produced by the charge on the anions at the cation site therefore increases because of the greater proximity of the charges. The increase in the *a* value in going from ZnS to ZnTe indicates that the covalency in the bonding is increasing at a rate more than fast enough to offset the increasing separation of the centers of the cationic and anionic charges as given in Table I.

The value of *a*, 3.1×10⁻⁴ cm⁻¹, in CdTe is less than *a* in any of the three zinc compounds, particularly as compared to the value of *a* in ZnTe, 6.6×10⁻⁴ cm⁻¹. This cannot simply be accounted for by the larger cation-anion separation in CdTe as compared to ZnTe. The larger value can, however, arise due to lattice distortion caused by incorporating Cr³⁺ at a Zn site. In the presence of distortion the charge on the Cr³⁺ in the zinc compound will be in more intimate contact with the anionic charge. This results in a higher crystalline field and accounts for the larger *a* value measured in ZnTe than in CdTe.

The *A*(Cr⁵³) Values

The hyperfine splitting characterized by *A*(Cr⁵³) is the result of a magnetic interaction between the nuclear magnetic moment of the 10% abundant Cr⁵³ and the magnetic field produced at the Cr⁵³ nucleus by the 3d⁵ electrons. This magnetic field results from a polarization of the inner paired s² electrons by the aligned spins of the 3d⁵ configuration.¹² The mechanism producing the field is a spin-dependent exchange mechanism and will depend on the degree of overlap of the wave functions of the inner s² electrons and that of the 3d⁵ electrons. The degree of overlap will depend on the covalency in the bonding and would be expected to be less the more covalent the bond. This should result in lower *A*(Cr⁵³)

¹¹ R. S. Title, Phys. Rev. **130**, 17 (1963).

¹² V. Heine, Phys. Rev. **107**, 1002 (1957); R. E. Watson and A. J. Freeman, Phys. Rev. **123**, 2027 (1961).

values for the crystals with the more covalent bonds and is seen to be the case in Table I. The more covalent ZnTe has a lower *A*(Cr⁵³) value than ZnS. In addition, a comparison of ZnTe and CdTe shows the *A*(Cr⁵³) value in ZnTe to be lower. This is an indication of more covalent bonding in ZnTe caused by a lattice distortion in incorporating the larger Cr³⁺ for Zn.

The *A*_{Zn} and *A*_{Cd} Values

The *A*_{Zn} (or *A*_{Cd}) characterize an interaction between the nuclear magnetic moment of Zn⁶⁷ (or Cd¹¹¹ and Cd¹¹³) occupying second-neighbor sites and the magnetic field produced at these sites by the 3d⁵ configuration. Since the spins and moments of the Zn and Cd nuclei are known, the *A*_{Zn,Cd} values may be used to calculate the magnetic field *H*_{Zn,Cd} produced by the Cr³⁺ at the twelve nearest cation sites. The fields are given in Table II. Also included for comparison are the results that have been obtained for the interaction of the isoelectronic Mn²⁺ with second-neighbor sites in ZnS,⁴ CdS,^{5,6} and CdTe.⁶ The magnetic field at the second-neighbor sites is a measure of the contact of the 3d⁵ wave function with these sites.

The results for Mn²⁺ as given in Table II indicate that the magnetic field at the second-neighbor site is independent of lattice size, the covalency in the bonds, and whether the cation is zinc or cadmium. For Cr³⁺, however, the magnetic fields at the second-neighbor sites are much larger than for the isoelectronic Mn²⁺ and there is a variation in *H*_{Zn,Cd} from crystal to crystal. The magnetic field at the second-neighbor sites increases as the covalency in the bonds increase from the sulphide to the tellurides. The smaller value for *H*_{Zn,Cd} in CdTe than in ZnTe may be due to the

TABLE II. The interaction of Cr³⁺ and Mn²⁺(3d⁵) with magnetic nuclei at second-neighbor sites in II-VI compounds.

		<i>A</i> _{Zn} or <i>A</i> _{Cd} 10 ⁻⁴ cm ⁻¹	<i>H</i> _{Zn} or <i>H</i> _{Cd} Oe
Cr ³⁺	in ZnS	1.23	3.46×10 ⁴
	ZnSe	1.81	5.08×10 ⁴
	ZnTe	3.59	10.1 ×10 ⁴
	CdTe	10.6	8.4 ×10 ⁴
Mn ²⁺	in ZnS	0.75	2.1×10 ⁴
	CdS	2.6	2.1×10 ⁴
	CdTe	2.6	2.1×10 ⁴

larger size of CdTe or may possibly reflect lattice distortion in incorporating Cr⁺ for Zn.

The variation in the interaction of Cr⁺ with second-neighbor sites as compared to the lack of variation for Mn⁺⁺ is undoubtedly due to the greater extent of the Cr⁺ wave function as compared to Mn⁺⁺. This is to be expected since Cr in this configuration has accepted an electron. Mn⁺⁺, on the other hand, is electrically neutral.

In summary, the resonance measurements that have been made of Cr⁺ in four zinc-blende lattices show that

there are variations from crystal to crystal of the resonance parameters g , a , A (Cr⁵³), and $A_{Zn,Cd}$. The variations are consistent with changes in the degree of covalency in the bonds and with lattice distortions caused by incorporating the larger Cr⁺ in place of divalent zinc.

ACKNOWLEDGMENTS

I wish to thank Dr. L. Suchow, J. Scardefield, and J. Kucza who grew the crystals and E. E. Tynan for aid in taking the paramagnetic data.

Longitudinal Nuclear Spin-Spin Relaxation*

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(Received 25 October 1963)

A theoretical and experimental study is made of the nuclear magnetization which appears after a sudden unidirectional step magnetic field is applied nonadiabatically to an ordered spin system which has been prepared by the process of adiabatic demagnetization. The magnetization exhibits damped oscillations about a nonvanishing equilibrium value. The oscillation frequencies correspond to the fundamental and harmonic components of Larmor frequency determined by the combined effect of local dipole and external step fields. The oscillations arise as the internally ordered dipole-dipole interaction energy exchanges with the suddenly imposed Zeeman energy reservoir. A density matrix calculation carried to second order in time-dependent perturbation theory, combined with a Gaussian decay model, accounts for the observed oscillations within times comparable to the decay time. Fluorine nuclei in single CaF₂ crystals are studied for various crystal orientations with respect to the applied step field. For long times after step field application the attainment of thermal equilibrium between dipole-dipole and Zeeman reservoirs permits prediction, using the spin temperature concept, of magnetization developed along the step field.

I. INTRODUCTION

FOLLOWING the sudden reorientation of equilibrium nuclear magnetization M_0 by a radio-frequency (rf) pulse, a transient nuclear induction signal appears which is a measure of the Fourier integral of the resonance line shape.^{1,2} The relaxation time of this signal, loosely referred to as T_2 , is a measure of the lifetime of coherent magnetization which precesses in the transverse direction perpendicular to a large polarizing magnetic field H_0 . In this paper the transient response of the spin system is studied after a step direct current (dc) magnetic field H_s is applied to an ordered spin system in zero ($H_0=0$) magnetic field.³ The ordered spin state is prepared by carrying out an adiabatic demagnetization of the initially polarized sample by turn-

ing off H_0 slowly.⁴ The step field H_s is then turned on nonadiabatically in a time much shorter than the Larmor period of the spins in the net field determined by H_s and the mean local dipolar field H_L . A longitudinal nuclear magnetization $\langle M_z(t) \rangle$ develops in time along the direction of H_s after it is applied. A transient longitudinal spin-spin relaxation (LSSR) behavior is displayed by $\langle M_z(t) \rangle$, which is a measure of the oscillatory interchange of energy between the magnetic dipole-dipole energy reservoir and the suddenly imposed Zeeman energy reservoir. In a sense, these measurements yield the Fourier integral of line-shape measurements made by Anderson⁵ on the nuclear magnetic absorption of spins ordered in low magnetic fields comparable to H_L . The transient measurements reveal directly the time evolution toward internal spin-spin equilibrium and toward final thermal equilibrium between Zeeman and dipole-dipole energy reservoirs.

A number of observers^{4,6,7} have investigated various

* Supported by the National Science Foundation and the U. S. Office of Naval Research.

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¹ E. L. Hahn, Phys. Rev. **80**, 580 (1950).

² I. Lowe and R. E. Norberg, Phys. Rev. **107**, 46 (1957); B. Herzog and E. L. Hahn, *ibid.* **103**, 148 (1956).

³ R. L. Strombotne and E. L. Hahn, Bull. Am. Phys. Soc. **6**, 508 (1961).

⁴ A. Abragam and W. G. Proctor, Phys. Rev. **109**, 1441 (1958).

⁵ A. G. Anderson, Phys. Rev. **125**, 1517 (1962); **115**, 863 (1959).

⁶ S. R. Hartmann and A. G. Anderson, Phys. Rev. **128**, 2023 (1962); S. R. Hartmann and E. L. Hahn, *ibid.* **128**, 2042 (1962); F. M. Lurie and C. P. Slichter, Phys. Rev. Letters **10**, 403 (1963).

⁷ R. T. Schumacher, Phys. Rev. **112**, 837 (1958); P. S. Pershan,