



Metal ion induced FRET On–Off in naphthyl-pyrenyl pendent tetrahomodioxacalix[4]arene

Ji Hee Jung^a, Min Hee Lee^a, Hyun Jung Kim^a, Hyo Sung Jung^a, Su Yeon Lee^a,
Na Ri Shin^b, Kwanghyun No^{b,*}, Jong Seung Kim^{a,*}

^a Department of Chemistry, Korea University, Seoul 136-701, Republic of Korea

^b Department of Chemistry, Sookmyung Women's University, Seoul 140-742, Republic of Korea

ARTICLE INFO

Article history:

Received 26 December 2008

Revised 10 February 2009

Accepted 12 February 2009

Available online 15 February 2009

Keywords:

FRET

Fluorescence

Calixarenes

Homooxacalixarene

Copper

ABSTRACT

A novel tetrahomodioxacalix[4]arene (**7**) bearing both naphthyl- and pyrenyl-amide pendants was synthesized as FRET-based fluorometric sensor for Cu²⁺ ion. Intramolecular FRET from the naphthalene emission to the pyrene absorption affords Cu²⁺ ion selectivity over other metal ions. Upon addition of Cu²⁺ ion, the complex solution of **7** gave a significantly decreased pyrene acceptor emission along with an enhancement of naphthalene donor emission via FRET On–Off event.

© 2009 Elsevier Ltd. All rights reserved.

Many heavy and transition metal ions, such as Hg²⁺, Cu²⁺, Pb²⁺, and Cd²⁺, have received considerable attention because of its widespread use in agricultural, chemical and industrial processes, which are becoming threats to living organisms.¹ In particular, Cu(II) is an essential element playing a fundamental role in the biochemistry of the human nervous system, but is toxic with high concentration.² Thus, its accumulation in the human body affects an oxidative stress and disorders associated with neurodegenerative diseases, Menkes disease, Wilson disease, and Alzheimer's disease.³ Therefore, development of the chemosensors specifically for the Cu²⁺ has been greatly interesting in many fields.⁴

Many chemosensors were prepared by attachment of fluorophore units to macrocyclic or chelating molecules. Among the macrocyclic frameworks, a class of calix[4]arene is a good candidate for the chemosensor framework towards specific metal ion because of their characteristic cavity size or geometrically restricted conformation.⁵ However, homooxacalix[4]arenes, having extra oxygen atoms in the macrocyclic ring, have received little attention because they can only be synthesized in relatively low yields.^{6–8} Tetrahomodioxo *p*-*tert*-butylcalix[4]arene was prepared by Gutsche in 44% yield by dehydration of a bishydroxymethylated dimer of *p*-*tert*-butylphenol, which was synthesized by a multi-

step reaction starting from the bromination of *p*-*tert*-butylphenol.⁸ There have only been limited studies on the solution conformations, solid-state structures, and on complexation behavior of homooxacalix[4]arenes.^{9–12} We also previously reported that there are five conformations in tetrahomodioxacalix[4]arene with appropriate nomenclature (cone, partial cone, C-1,2-alternate, COC-1,2 alternate, and 1,3-alternate) as represented in Figure 1.¹³ Recently, we also reported a solid-state structure of C-1,2-alternate *N,N*-diethyl tetrahomodioxacalix[4]arene tetraamide complexing

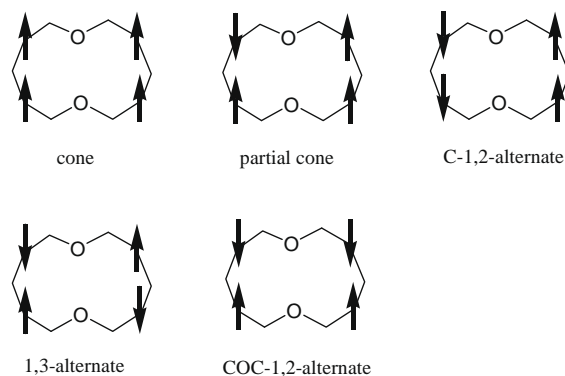


Figure 1. Schematic representation of the five conformations of homooxacalix[4]arenes.

* Corresponding authors. Tel.: +82 2 3290 3143; fax: +82 2 3290 3121 (J.S.K.).

E-mail addresses: hyun@sookmyung.ac.kr (K. No), jongskim@korea.ac.kr (J.S. Kim).

Pb²⁺ proven by X-ray structure.¹⁴ For a corresponding *monoalkyl* amide, the conformation changed to 1,3-alternate due to intramolecular hydrogen bonding, resulting in weak binding of metal ions.¹⁴

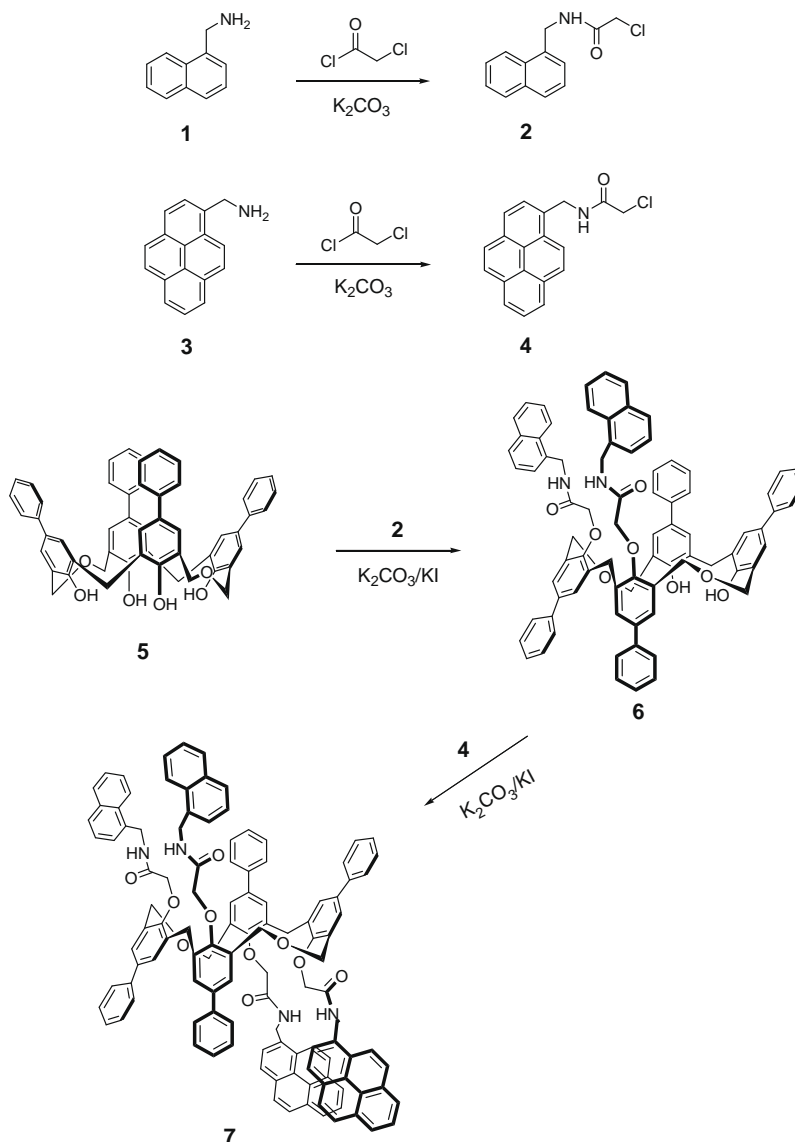
We also previously reported that a series of C-1,2-alternate tetrahomodioxacalix[4]arene pyreneamides were synthesized. Pb²⁺ coordination gave a quenched monomer and excimer fluorescence emission, while upon Ca²⁺ ion binding the receptor provides an enhanced excimer and declined monomer emission with ratiometric response. Upon Ca²⁺ binding, the efficient HOMO–LUMO interaction between Py and Py* induces a formation of strong excimer emission band, whereas there is no such interaction observed upon Pb²⁺ complexation.^{15,16}

Fluorescence chemosensors utilize photophysical changes produced by cation binding: photo-induced electron transfer (PET);¹⁷ photo-induced charge transfer (PCT);¹⁸ excimer/excimer formation and extinction;¹⁹ or fluorescence resonance energy transfer (FRET).²⁰ The FRET is known to be sensitive, selective, and adaptable to a wide variety of systems.²¹ However, reports on FRET-based metal ion sensors are still at a modest number. The FRET arises from a donor–acceptor (D–A) interaction between a pair of fluorophores in their excited state. Excited state of the

fluorescent donor is then non-radiatively transferred to the acceptor, and the donor returns to its electronic ground state. Therefore, the FRET is required to have a certain extent of spectral overlap between emission spectrum of the donor and absorption spectrum of the acceptor.²² From this point of view, we herein report on synthesis and the fluorescence properties of tetrahomodioxacalix[4]arene bearing naphthyl- and pyrenyl-amides (**7**) able to show the FRET On→Off by the Cu²⁺ ion.

The synthetic routes²³ for tetrahomodioxacalix[4]arene derivatives **6** and **7** are described in Scheme 1. Synthesis of **6** was performed by the condensation of *N*-(1-naphthalenemethyl)chloroacetamide (**2**) with tetrahomodioxacalix[4]arene (**5**)²⁴ in the presence of K₂CO₃ as a base and a catalytic amount of KI. Reaction of **6** with *N*-(1-pyrenylmethyl)chloroacetamide (**4**)²⁵ under basic medium gave **7** in 72% yield.

There are two different shapes in 1,2-alternate conformations on tetrahomodioxacalixarene system. 1,2-alternate conformer in which the adjacent *syn* aryl moieties are joined by a CH₂ group is designated as the C-1,2-alternate, while the 1,2-alternate conformer in which the adjacent *syn* aryl moieties are joined by a CH₂–O–CH₂ moiety is designated as the COC-1,2-alternate.¹⁴ The dimethyleneoxy protons of the ArCH₂OCH₂Ar bridge showed AB



Scheme 1. Synthetic pathways to **6** and **7**.

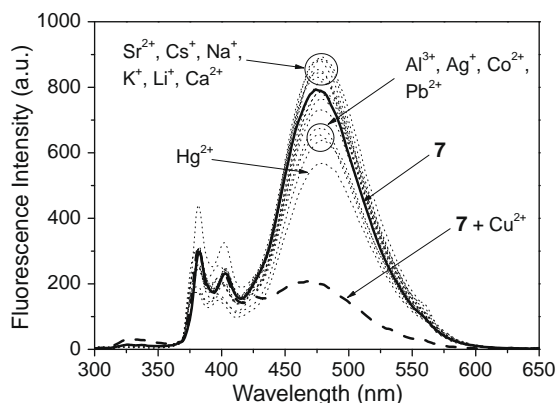


Figure 2. Fluorescence spectra of **7** (6.0 μM) upon addition of various cations (100.0 equiv) in $\text{CH}_3\text{CN}/\text{CHCl}_3$ (40:1, v/v) with an excitation wavelength at 280 nm.

doublets at δ 4.86 and δ 4.07 ($\Delta\nu = 316$ Hz) with a *geminal* coupling constant of 13.0 Hz. A doublet with peaks for the methylene protons of ArCH_2Ar showed at δ 3.75 and δ 3.04 ($\Delta\nu = 284$ Hz) with a *geminal* coupling constant of 11.4 Hz. The ^{13}C NMR spectrum showed one peak at 62.38 ppm for the ArCH_2O bridge methylenoxy carbons and one peak at 30.36 ppm for the ArCH_2Ar bridge carbons also indicating that two adjacent benzene rings are in a *syn* orientation (C-1,2-alternate conformation).

For the detailed study of FRET occurring in **7**, we examined its fluorescence and UV/vis spectral behaviors upon addition of various metal ions such as Li^+ , Na^+ , K^+ , Rb^+ , Cs^+ , Ag^+ , Cu^{2+} , Co^{2+} , Ba^{2+} , Pb^{2+} , Ca^{2+} , Hg^{2+} , Mg^{2+} , Sr^{2+} , Zn^{2+} , and Al^{3+} . As seen in Figure 2, **7** shows a weak naphthalene emission at 331 nm and strong pyrene monomer and excimer emission bands at 382 and 470 nm, respectively, with an excitation at 280 nm. Interestingly, we observed increase in naphthalene emission shown at 331 nm along with decreased pyrene excimer emission at 470 nm upon addition of Cu^{2+} ion. We envisioned that these fluorescence changes may associate with the FRET changes upon Cu^{2+} binding. In the solution of **7**, FRET-On occurs by overlapping of naphthalene emission (donor fluorophore) with pyrene absorption (acceptor fluorophore) as seen in Figure 3a. Thus, **7** exhibits a weak naphthalene emission and strong pyrene emission bands although the excitation wavelength is 280 nm which corresponds to the naphthalene absorption band.

Upon addition of Cu^{2+} ion to a solution of **7**, however, we noticed that the naphthalene emission band at 331 nm increases, but the pyrene emission concomitantly declines. This is attributable to the fact that the pyrene absorption band of **7** decreases upon addition of Cu^{2+} ion (Fig. 3b), then the spectral overlap between naphthalene and pyrene is minimized to give the decline in FRET efficiency. With respect to the extent of FRET changes, **7** indicates a high selectivity toward Cu^{2+} ion over other metal cations.

To gain insight into the FRET efficiency in the energy transfer process from naphthalene to pyrene, the photophysical property of **6** in the absence of the pyrene units as an energy acceptor was also studied. The FRET efficiency can be estimated by the following equation:²⁶

$$E = 1 - (F'_D/F_D)$$

where E denotes FRET efficiency; F'_D and F_D are the donor fluorescence intensity with and without an acceptor, respectively. As seen in Figure 4, **7** without copper ion barely exhibits naphthalene emission band at 331 nm because of FRET occurring. However, addition of Cu^{2+} ion to a solution of **7** increases the naphthalene emission intensity due to the FRET-Off (inset). From the fluorescence changes, E of **7** and **7**· Cu^{2+} was calculated to be 0.58 and 0, respectively.

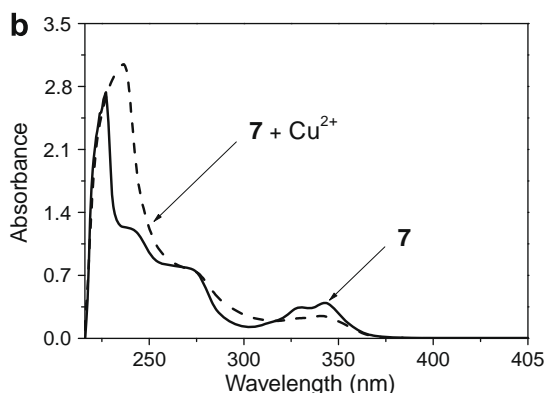
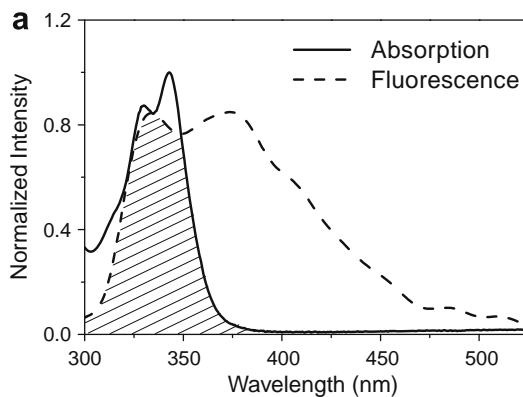


Figure 3. (a) Spectral overlap of naphthalene emission (FRET donor) and pyrene absorption (FRET acceptor). (b) Absorption spectral changes of **7** in the absence and in the presence of Cu^{2+} ion (100.0 equiv).

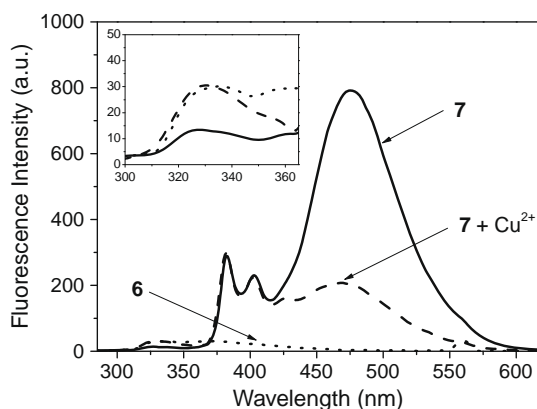


Figure 4. Fluorescence spectra of **6** and **7** (6.0 μM , respectively) in the presence of Cu^{2+} ion (100.0 equiv) in $\text{CH}_3\text{CN}/\text{CHCl}_3$ (40:1, v/v) with an excitation at 280 nm. Inset: enlarged spectra of **6** and **7** between 300 and 365 nm.

Moreover, we observed a selective visual fluorescence change of **7** upon addition of Cu^{2+} ion over other various metal ions, which is obviously due to the FRET-Off to give a decreased pyrene emission of **7** (Fig. 5).



Figure 5. Visual fluorescence changes of **7** (10.0 μM) upon addition of various metal ions in $\text{CH}_3\text{CN}/\text{CHCl}_3$ (40:1, v/v). From left to right: free **7**, Cu^{2+} , K^+ , Li^+ , Na^+ , Cs^+ , Rb^+ , Ag^+ , Co^{2+} , Ba^{2+} , Pb^{2+} , Ca^{2+} , Hg^{2+} , Mg^{2+} , Sr^{2+} , Zn^{2+} , and Al^{3+} (100.0 equiv, respectively).

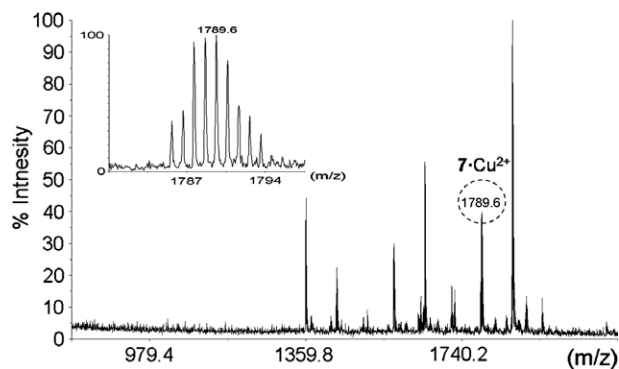


Figure 6. MALDI-TOF Mass spectrum of 7-Cu²⁺.

For the binding mode between **7** and Cu²⁺ ion, MALDI-TOF Mass analysis was carried out. A peak at 1789.6 *m/z* corresponding to 7-Cu²⁺ was observed by addition of excess Cu(ClO₄)₂ to **7** as seen in (Fig. 6). We then noticed that the Cu²⁺ ion is coordinated by **7** with an 1:1 stoichiometry.

In conclusion, FRET-based fluorometric tetrahomodioxacalix[4]arene **7** has been newly synthesized. Derivative **7** exhibits a weak naphthalene emission and a strong pyrene emission to provide a FRET-On due to an energy transfer event from naphthalene to pyrene unit. Complexation with Cu²⁺ increases the naphthalene emission along with decreases in the excimer emission of **7** because of the FRET-Off. With respect to the extent of FRET changes, we could observe the Cu²⁺ selectivity of **7** over other metal ions.

Acknowledgments

The authors wish to acknowledge the financial support of the SRC program (R11-2005-008-02001-0(2008)) and the Sookmyung Women's University Research grant 2008 (1-0803-0137).

References and notes

- Senger, M. R.; Rico, E. P.; Arizi, M. D.; Frazzoni, A. P. G.; Dias, R. D.; Bogo, M. R.; Bonan, C. D. *Toxicology* **2006**, *226*, 229.
- Hultberg, B.; Andersson, A.; Lsaksson, A. *Toxicology* **1997**, *117*, 89.
- Barnham, K. J.; Masters, C. L.; Bush, A. I. *Nat. Rev. Drug Discovery* **2004**, *3*, 205.
- (a) Zeng, L.; Miller, E. W.; Pralle, A.; Isacoff, E. Y.; Chang, C. J. *J. Am. Chem. Soc.* **2006**, *128*, 10; (b) Brunner, J.; Kraemer, R. *J. Am. Chem. Soc.* **2004**, *126*, 13626; (c) Dujols, V.; Ford, F.; Czarnik, A. W. *J. Am. Chem. Soc.* **1997**, *119*, 7386.
- (a) Kim, J. S.; Shon, O. J.; Rim, J. A.; Kim, S. K.; Yoon, J. J. *Org. Chem.* **2002**, *67*, 2348; (b) Kim, S. K.; Lee, S. H.; Lee, J. Y.; Lee, J. Y.; Bartsch, R. A.; Kim, J. S. *J. Am. Chem. Soc.* **2004**, *126*, 16499; (c) Lee, M. H.; Quang, D. T.; Jung, H. S.; Yoon, J.; Lee, C. H.; Kim, J. S. *J. Org. Chem.* **2007**, *72*, 4242.
- (a) Gutsche, C. D.; Dhawan, B.; No, K. H.; Muthukrishnan, R. *J. Am. Chem. Soc.* **1981**, *103*, 3782; (b) Masci, B.; Saccheo, S. *Tetrahedron* **1993**, *49*, 10739.
- Bavoux, C.; Vocanson, F.; Perrin, M.; Lamartine, R. *J. Incl. Phenom. Mol. Recogn. Chem.* **1995**, *22*, 119.
- Dhawan, B.; Gutsche, C. D. *J. Org. Chem.* **1983**, *48*, 1536.
- (a) Thuéry, P.; Nierlich, M.; Vicens, J.; Masci, B. *Acta Crystallogr., Sect. C* **2001**, *57*, 70; (b) Asfari, Z.; Harrowfield, J. M.; Ogden, M. I.; Vicens, J.; White, A. H. *Angew. Chem., Int. Ed. Engl.* **1991**, *30*, 854.
- Harrowfield, J. M.; Ogden, M. I.; White, A. H. *J. Chem. Soc., Dalton Trans.* **1991**, 979.
- Marcos, P. M.; Ascenso, J. R.; Lamartine, R.; Pereira, J. L. C. *Tetrahedron* **1997**, *53*, 11791.
- Felix, S.; Ascenso, J. R.; Lamartine, R.; Pereira, J. L. C. *Synth. Commun.* **1998**, *28*, 1793.
- No, K.; Lee, J. H.; Yang, S. H.; Kim, M. J.; Kim, J. S. *J. Org. Chem.* **2002**, *67*, 3165.
- No, K.; Kim, J. S.; Shon, O. J.; Yang, S. H.; Suh, I. H.; Kim, J. G.; Bartsch, R. A.; Kim, J. Y. *J. Org. Chem.* **2001**, *66*, 5976.
- (a) Iasi, G. D.; Masci, B. *Tetrahedron Lett.* **1993**, *34*, 6635; (b) No, K.; Kim, J. S.; Shon, O. J.; Yang, S. H.; Suh, I. H.; Kim, J. G.; Bartsch, R. A.; Kim, J. Y. *J. Org. Chem.* **2001**, *66*, 5976.
- Choi, J. K.; Lee, A.; Kim, S.; Ham, S.; No, K.; Kim, J. S. *Org. Lett.* **2006**, *8*, 1601.
- (a) Aoki, L.; Sakaki, T.; Shinkai, S. *J. Chem. Soc., Chem. Commun.* **1992**, 730; (b) Jin, T.; Ichikawa, K.; Koyama, T. *J. Chem. Soc., Chem. Commun.* **1992**, 499; (c) Ji, H.-F.; Brown, G. M.; Dabestani, R. *Chem. Commun.* **1999**, 609; (d) Kim, J. S.; Noh, K. H.; Lee, S. H.; Kim, S. K.; Kim, S. K.; Yoon, J. *J. Org. Chem.* **2003**, *68*, 597.
- Leray, I.; Lefevre, J. P.; Delouis, J. F.; Delaire, J.; Valeur, B. *Chem. Eur. J.* **2001**, *7*, 4590.
- (a) Birks, J. B. *Photophysics of Aromatic Molecules*; Wiley-Interscience: London, 1970; (b) Lee, S. H.; Kim, S. H.; Kim, S. K.; Jung, J. H.; Kim, J. S. *J. Org. Chem.* **2005**, *70*, 9288.
- Hecht, S.; Vladimirov, N.; Fréchet, J. M. J. *J. Am. Chem. Soc.* **2001**, *123*, 18.
- (a) Tsien, R. Y.; Miyawaki, A. *Science* **1998**, *280*, 1954–1955; (b) Weiss, S. *Science* **1999**, *283*, 1676.
- Stryer, L.; Haugland, R. P. *Proc. Natl. Acad. Sci. U.S.A.* **1967**, *58*, 719.
- Unless otherwise noted, reagents were obtained from commercial suppliers and used without further purification. Melting points were taken in evacuated and sealed capillary tubes with a Mel-Temp apparatus and were uncorrected. IR spectra were recorded on a Nicolet Impact 400 FT-IR spectrometer. ¹H and ¹³C NMR spectra were recorded with a Bruker AMX 400 spectrometer. Chemical shifts are recorded in parts per million relative to TMS as an internal standard.
- N*-(1-naphthalenemethyl)chloroacetamide (**2**). To a suspension of 1-naphthalenemethylamine (300 mg, 1.91 mmol) and potassium carbonate (1.10 g, 7.64 mmol) in a mixture of water (50 mL) and ethyl acetate (50 mL), a solution of chloroacetyl chloride (0.25 mL, 3.14 mmol) in ethyl acetate (10 mL) was added dropwise. After the reaction mixture was stirred at room temperature for 2 h, organic phase was separated and dried over anhydrous magnesium sulfate. The residue obtained by evaporation of solvent was triturated with hexane to afford 440 mg (98%) as a colorless crystalline solid. mp 120 °C; ¹H NMR (400 MHz, CDCl₃) δ 8.00–7.44 (m, 7, ArH), 6.82 (br. s, 1, NH), 4.95 (d, 2, CH₂, *J* = 5.6 Hz), 4.11 (s, 2, CH₂); ¹³C NMR (125 MHz, CDCl₃) δ 165.79 (C=O), 134.10, 132.73, 131.52, 129.17, 129.11, 127.06, 126.98, 126.34, 125.61, 123.40 (Ar), 42.80, 42.20 (CH₂). Anal. Calcd for C₁₃H₁₂NOCl: C, 66.81; H, 5.18. Found. C, 66.89; H, 5.17.
- 7,13,21,27-Tetraphenyl-29,31-bis[*N*-(1-naphthalenylmethyl)aminocarbonyl]-methoxy]-30,32-dihydroxy-2,4,16,18-tetra-*mo*-3,17-dioxacalix[4]arene (**6**). Tetrahomodioxacalix[4]arene **5** (571 mg, 0.724 mmol), potassium carbonate (100 mg, 0.726 mmol), potassium iodide (50 mg), and **2** (400 mg, 1.81 mmol) in dried acetone (100 mL) was refluxed 80 h. After evaporation of solvent, the residue was extracted with CH₂Cl₂. The organic layer was washed with water, dried over MgSO₄, and evaporated in vacuo. The residue was triturated with MeOH to give the product mixture which was recrystallized from methanol to afford 676 mg (80%) of the desired product as a pale yellow colored crystal. mp 140–141 °C (decomposed); ¹H NMR (400 MHz, CDCl₃) δ 8.15 (d, 1, ArH, *J* = 6.6 Hz), 8.14 (d, 1, ArH, *J* = 6.6 Hz), 8.04 (d, 2, ArH, *J* = 8.8 Hz), 7.65–7.03 (m, 38, ArH, NH and OH), 6.90 (d, 4, ArH, *J* = 9.2 Hz), 5.29 (d, 1, NCH₂Nap, *J* = 14.4 Hz), 5.28 (d, 1, NCH₂Nap, *J* = 14.4 Hz), 4.75 (d, 1, NCH₂Nap, *J* = 14.4 Hz), 4.74 (d, 1, NCH₂Nap, *J* = 14.4 Hz), 4.62 (s, 4, OCH₂CO), 4.61 (d, 2, ArCH₂O, *J* = 14.8 Hz), 4.51 (d, 2, ArCH₂O, *J* = 14.8 Hz), 4.39 (d, 2, ArCH₂O, *J* = 10.4 Hz), 4.26 (d, 2, ArCH₂O, *J* = 10.4 Hz), 3.87 (d, 2, ArCH₂Ar, *J* = 14.0 Hz), 2.96 (d, 2, ArCH₂Ar, *J* = 14.0 Hz); ¹³C NMR (125 MHz, CDCl₃) δ 168.24 (C=O), 153.60, 153.41, 141.14, 140.25, 138.17, 134.36, 133.86, 133.03, 132.57, 131.51, 130.47, 130.07, 129.44, 129.07, 129.01, 128.83, 128.76, 128.21, 127.83, 127.37, 127.18, 127.13, 127.03, 126.93, 126.56, 125.84, 125.49, 125.25, 123.54, 122.56 (Ar), 73.32 (OCH₂CO), 73.04, 71.02 (ArCH₂O), 41.71 (NCH₂Np), 28.88 (ArCH₂Ar). Anal. Calcd for C₈₀H₆₆O₈N₂: C, 81.20; H, 5.62. Found. C, 81.91; H, 5.66.
- 7,13,21,27-Tetraphenyl-29,31-bis[*N*-(1-pyrenylmethyl)aminocarbonyl]-methoxy]-30,32-bis[*N*-(1-pyrenylmethyl)aminocarbonyl]-methoxy]-2,4,16,18-tetra-*mo*-3,17-dioxacalix[4]arene (**7**). A mixture of **6** (421 mg, 0.316 mmol), potassium carbonate (262 mg, 1.90 mmol), **4** (419 mg, 1.79 mmol) and a trace catalytic amount of KI in dried CH₃CN (120 mL) was refluxed for 120 h. The solvent was evaporated and the residue was extracted with CH₂Cl₂. The organic layer was washed two times with water, dried over anhydrous magnesium sulfate, and evaporated in vacuo. The residue was triturated with hexane to give the product mixture which was recrystallized from methylene chloride and methanol to afford the pure product (393 mg, 72%) as a pale yellow colored crystal. mp 113–114 °C (decomposed); ¹H NMR (400 MHz, CDCl₃) δ 7.96–7.74 (m, 20, ArH and NH), 7.62–7.34 (m, 36, ArH), 7.19–6.62 (m, 8, ArH), 5.12 (d, 2, NCH₂Ar, *J* = 12.0 Hz), 4.96 (d, 2, NCH₂Ar, *J* = 12.0 Hz), 4.86 (d, 2, ArCH₂O, *J* = 13.0 Hz), 4.83 (d, 4, OCH₂CO, *J* = 10.8 Hz), 4.61 (d, 2, NCH₂Ar, *J* = 11.2 Hz), 4.07 (d, 2, ArCH₂O, *J* = 13.0 Hz), 3.96 (d, 4, OCH₂CO, *J* = 10.8 Hz), 3.75 (d, 2, ArCH₂Ar, *J* = 12.4 Hz), 3.66 (d, 2, NCH₂Ar, *J* = 11.2 Hz), 3.04 (d, 2, ArCH₂Ar, *J* = 12.4 Hz); ¹³C NMR (125 MHz, CDCl₃) δ 168.18 (C=O), 153.70, 139.44, 134.05, 133.08, 131.45, 131.16, 129.69, 129.0–129.7, 129.00, 128.87, 127.99, 126.96, 126.90, 126.30, 126.21, 125.58, 125.37, 123.59, 123.40 (Ar), 71.26 (OCH₂CO), 62.38 (ArCH₂O), 41.28 (NCH₂Np), 30.36 (ArCH₂Ar). Anal. Calcd for C₁₁₈H₉₂O₁₀N₄: C, 82.11; H, 5.37. Found. C, 81.29; H, 5.36.
- Thuery, P.; Nierlich, M.; Vicens, J.; Masci, B. *Acta Crystallogr.* **2001**, *C57*, 70.
- van der Veen, N. J.; Flink, S.; Deij, M. A.; Egberink, R. J. M.; van Veggel, F. C. J. M.; Reinhoudt, D. N. *J. Am. Chem. Soc.* **2000**, *122*, 6112.
- Lakowicz, J. R. *Principles of Fluorescence Spectroscopy*, 2nd ed.; Plenum: New York, 1999; Chapter 13.