Note

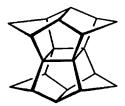
Investigations on the phase behavior of "Pagodan"

R. Krombach^a, J. Ellert^a, G.M. Schneider^a and H. Prinzbach^b

^a Lehrstuhl für Physikalische Chemie II, Fakultät für Chemie, Ruhr-Universität Bochum, D-4630 Bochum (Germany) ^b Lehrstuhl für Organische Chemie und Biochemie, Universität Freiburg, D-7800 Freiburg (Germany)

(Received 17 June 1992)

"Pagodan" (1), the $(CH)_{20}$ polycycle synthesized by the group in Freiburg, was studied to obtain detailed information about its phase behavior. These investigations were of special interest because a phase transition from a crystalline to a plastic crystalline state of pagodan could be presumed by its nearly spherical molecular form and the high melting point of 516 K published earlier [1]. The phase behavior of pagodan was measured with high pressure DTA in the temperature range 293–550 K at normal pressure and at pressures up to 190 MPa. The apparatus and the measuring method have been described elsewhere [2, 4].



Undecacyclo $[9.9.0.0^{1.5}.0^{2,12}.0^{2,18}.0^{3.7}.0^{6,10}.0^{8,12}.0^{11,15}.0^{13,17}.3^{16,20}]$ eicosan (1).

Figure 1 shows two typical thermograms of pagodan at 0.1 and 43 MPa that have been obtained in the present measurements.

Figure 2 shows the pressure-temperature (p-T) phase diagram obtained from the measurements.

The results given in Fig. 2 demonstrate that the predicted phase behavior

0040-6031/93/\$06.00 © 1993 - Elsevier Science Publishers B.V. All rights reserved

Correspondence to: G.M. Schneider, Lehrstuhl für Physikalische Chemie II, Fakultät für Chemie, Ruhr-Universität Bochum, D-4630 Bochum, Germany.

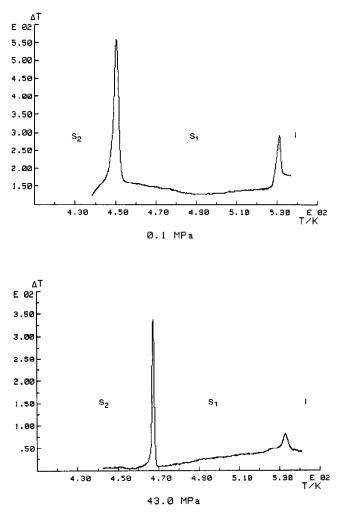


Fig. 1. Original thermograms of pagodan at 0.1 and 43.0 MPa.

could be confirmed. A second solid phase (s_1) has been found, which must be attributed to a plastic crystalline state, because, for example, of the relatively high enthalpy change accompanying the s_2-s_1 transformation compared to that of the melting transition (see Fig. 1). Extrapolation of the s_2-s_1 and s_1-l transition lines in Fig. 2 result in a triple point s_1-s_1-l at approximately 215 MPa and 545 K; at higher pressures the s_1 phase is no longer stable.

The melting point of $T^*(s_1-l) = (529.5 \pm 0.5)$ K found during the present measurements is distinctively higher than the melting point of 516 K reported in an earlier paper [1]. This discrepancy is probably caused by sublimation effects that were avoided by the use of special measuring cells in the present work [3].

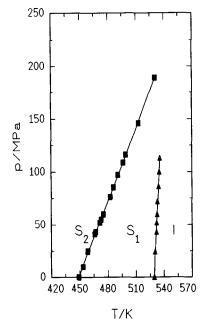


Fig. 2. Phase diagram of pagodan.

ACKNOWLEDGMENT

Financial support of the Fonds der Chemischen Industrie e.V. is gratefully acknowledged.

REFERENCES

- 1 W.D. Fessner, G. Sedelmeier, P.R. Spurr, G. Rihs and H. Prinzbach, J. Am. Chem. Soc., 109 (1987) 4626-4642.
- 2 H.D. Kleinhans, Doctoral Thesis, Ruhr-Universität Bochum, Germany, 1984.
- 3 A. Hamm, Doctoral Thesis, Ruhr-Universität Bochum, Germany, 1989.
- 4 R. Krombach, Doctoral Thesis, Ruhr-Universität Bochum, Germany, 1992.