KINETICS AND REACTION MECHANISM FOR THE THERMAL DECOMPOSITION OF SOME EVEN CHAIN LEAD(II) CARBOXYLATES

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ABSTRACT

The kinetics of the thermai decomposition of lead(I1) dodecanoate, tetradecanoate and octadecanoate have been studied thermogravimetricaiiy in the temperature range 298- 1073 K by the modified method of Freeman and Carroil. The thermogravimetric curves show that while the octadecanoate decomposes in one single step to lead oxide, the dodecanoate and tetradecanoate decompose in two stages to lead oxide. The kinetic results show all the reactions, except the second step breakdown of dodecanoate (which **is nearly unity), to be of zero order. Activation energy values are presented for the desorption of carbon dioxide and the diffusion controlled rate of formation of a degradative intermediate identified as a ketone. Tentative mechanisms are advanced to explain the degradative route of these compounds_**

INTRODUCTION

Many of the early studies on the thermal behaviour of long chain metal carboxylates have tended to concentrate on structural and physical changes up to and beyond the melting temperature but well before the onset of decomposition [l-5]. A survey of the literature reveals that not much information on the kinetics and decomposition mechanism'has been reported so far. However, the thermal decomposition and decomposition kinetics of several metal acetates have been extensively studied and reported $[6-8]$. **These studies have shown that even in these simple acetates, the decomposition mechanisms are very complicated. Thus, Judd et al. [6] have found that the thermal decomposition reaction changed depending on whether the final product of decomposition was a metal or metal oxide. If the metal oxide was the final product, acetone was the major organic intermediate, whilst acetic acid was found when the metal alone was produced. In addition, very early work [9-111 on the thermal decomposition of some metal carboxylates has shown that these compounds decomposed at temperatures in excess of 200°C to yield a ketone as the major organic product. Indeed, the action of heat on calcium laurate has long been used to prepare 12-tricosanone.**

In this work, the thermal decomposition kinetics of dodecanoate, tetradecanoate and octadecanoate of lead(II) have been studied and suggestions **made on the mechanisms of decomposition and the likely side reactions that may occur.**

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EXPERIMENTAL

Materials

The dodecanoate, tetzadecanoate and octadecanoate of lead(II) were prepared by metathesis in alcohol solution $[12-14]$, recrystallised from hot **benzene and characterised as previously described [1,2]. The melting point data and elemental analyses were consistent with those described in the literature.**

Thermogravime try

Thermogravimetric experiments were carried out on a Stanton-Redcroft TG 750 thermobalance on carboxylate samples in the weight range 3-7 mg at a heating rate of 20°C min-' in dry nitrogen at a flow rate between 20 and 25 cm3 min- '. Changes in sample weight were recorded as a permanent trace on a Servoscripe 2s RE 571 X-Y recorder which recorded weight as a function of temperature at a recorder sensitivity of 10 mV and chart speed of 30 mm min-'. The temperature was simultaneously recorded by means of a Pt/Pt-13% Rh thermocouple situated in the microfurnace assembly. The samples, contained in shallow platinum crucibles provided by the manufacturers of the thermobalance, were heated from ambient to 1073 K.

'HNMR spectra

Spectra were measured on a Perk&Elmer RlO spectrometer at *60* **MHz. Solutions in deuterated chloroform were used with tetramethylsilane as internal standard.**

IR spectra

A Pye Unicam SP 3-300 spectrometer calibrated with polystyrene was used to obtain spectra of degraded intermediates in deuterated chloroform.

RESULTS AND DISCUSSION

The continuous % weight loss curves for the thermal decomposition of the lead(II) soaps from 298 to 1073 K are shown in Fig. 1. Decomposition was observed to begin at approximately 513 K and became very rapid between 690 and 730 K. In the cases of the lead(II) dodecanoate (PbC₁₂) and tetradecanoate (PbC₁₄), decomposition occcurred in two stages. The first stage was complete after $\sim 30\%$ weight loss for PbC₁₂ at ~ 682 K and 51-53% weight loss at approximately 729 K for PbC₁₄, and stage two after 62% (PbC_{12}) and 65% (PbC₁₄). The octadecanoate decomposed in one single step **resultirig in a 71% loss in weight. These results indicated that the final product of decomposition was lead oxide where 38%, 35% and 29% weight** remaining for PbC_{12} , PbC_{14} , and PbC_{18} , respectively, were in good agreement

Fig. 1. Thermogravimetric curves for the decomposition of PbC_{12} (A); PbC_{14} (B), and PbC_{18} (C).

with the theoretical weight remaining for the formation of lead oxide of 36.9%, 33.8% and 28.9%.

Kinetic information on the decomposition was obtained from these curves by the modified method of Freeman and Carroll [151, which is baaed on the expression

$$
\Delta \log(\mathrm{d} w/\mathrm{d} t) = n \Delta \log W_{\rm r} - (\Delta E/2.303 \, R) \Delta (1/T) \tag{1}
$$

where dw/dt is the rate of reaction, determined from tangents to the weight loss curves; n is the order of reaction; ΔE is the activation energy; R is the **gas constant; and 2' is the absolute temperature.**

$$
W_{\rm r} = \Delta W_{\rm c} - \Delta W
$$

where, ΔW_c is the total weight loss associated with a given reaction; and ΔW **is the weight loss at the point where dw/dt is taken. The order of reaction** and activation energy can be obtained from eqn. (1) by plotting $\Delta \log(\mathrm{d} w/\mathrm{d} t)$ **vs.** Δ log W_r at constant Δ (1/T). The order of reaction is determined from the slope and the activation energy from the intercept at $\Delta \log W_r = 0$. The rate of reaction and W_r at constant Δ (1/T) were determined from the plot **of the first derivative of the thermogravimetric curve and the corresponding** *W, as* **a function of the reciprocal of absolute temperature. A typical plot is** shown in Fig. 2. The values of dw/dt and W_r were taken at equal intervals of $1/T$ of 2.0×10^{-5} K⁻¹. Typical values obtained for the various parameters are **tabulated in Table 1. Modified Freeman and Carroll data were similarly ob**tained for PbC_{12} and PbC_{14} decompositions.

The kinetic plots for the first stage decomposition of PbC₁₂ and PbC₁₄ and the single step decomposition of PbC_{18} are shown in Fig. 3. The points representing approximately 30% up to \sim 682 K and \sim 52% up to \sim 729 K of the reaction for PbC_{12} and PbC_{14} , respectively, and 71% of reaction in the case

Fig. 2. Plot of dw/dt (\triangle) and W_r (0) as a function of temperature for the decomposition $rfPbC_{18}$.

Of PbC18, do not fall on a straight line. The Arrhenius plots (Fig. 4) indicate zero order kinetics up to this stage of decomposition. In addition, the plots show that in each case there are two activation processes; the first occurring over the first $5-8\%$ of reaction and the other over the next $8-30\%$ (PbC₁₂) and $6-52\%$ (PbC₁₄) of reaction. In the case of PbC₁₈, the second activation **process occurs over the range 6-63s of reaction. The second stage decom**position data were similarily analysed and the kinetic data so obtained for all **the soaps, together with the least squares estimated limits of error, are collected in Table 2.**

The Table shows that with the exception of the second stage decomposi**tion of Pb&, where** *n is* **nearly unity, all the decomposition reactions are of**

TABLE₁

dw/dt	$W_{\mathbf{r}}$ (mg)	log dw/dt	log $W_{\bf r}$	Δ log dw/dt	Δ log $W_{\bf r}$
0.28	4.28	-0.553	0.631	-0.030	0.005
0.30	4.23	-0.523	0.626	-0.114	0.004
0.39	4.19	-0.409	0.622	-0.117	0.005
0.51	4.14	-0.292	0.617	-0.119	0.006
0.67	4.08	-0.174	0.611	-0.098	0.011
0.84	3.98	-0.076	0.600	-0.042	0.016
1.08	3.84	-0.033	0.584	-0.154	0.024
1.32	3.63	0.121	0.560	-0.094	0.034
1.64	3.36	0.215	0.526	-0.095	0.038
2.04	3.08	0.310	0.489	-0.092	0.054
2.52	2.72	. 0.401	0.435	-0.104	0.069
3.20	2.32	0.505	0.366	-0.180	0.140
4.84	1.68	0.685	0.225	0.143	0.243
3.48	0.96	0.542	-0.018	0.360	0.234
1.52	0.56	0.182	-0.252		

TG data for lead(II) octadecanoate

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Fig. 3. Kinetic plots for the decomposition of $PbC_{12} (\triangle)$ **;** $PbC_{14} (\square)$ **; and** $PbC_{18} (\square)$ **.**

zero order. Moreover, the octadecanoate result presented here supports that obtained by Rasheed and Bhobe [lS] for the thermal decomposition of $zinc(II)$ octadecanoate (ZnC_{18}) , where zero order kinetics was also found for **that single step process. However, the activation energy value found by those** workers for ZnC_{18} is different from the value found here for PbC₁₈ decompo**sition. Rasheed and Bhobe [16] presented a value for the overall decomposition which, in view of similar values presented by other workers [7] on like systems, is not unreasonable; the result here indicates that two processes of different energy values are involved.**

It should be borne in mind that these lead(I1) carboxylates all melt 1171 before the onset of decomposition and that the decomposition products are some gas plus a solid (PbO). The reactions are therefore heterogeneous.

Fig. 4. Arrhenius decomposition plots of PbC₁₂ (Δ **), PbC₁₄ (** \Box **); and PbC₁₈ (** \Box **).**

TABLE2 Kinetic data

a first stage decomposition.

b second stage decomposition.

Therefore, zero order kinetics indicate that the reactions are independent of initial carboxylate concentration, and the surface of the molten soap remains covered by gaseous product throughout the reaction. The activation energy values accounting for the first 5-S% weight loss is most probably representative of the desorption of a gaseous product from the surface of the molten soap. Clearly, this term must be dependent on the surface area and surface tension of the molten soap. Since surface area can be assumed to vary only slightly, it is envisaged that surface tension is the most important term here. Unfortunately, this area of molten carboxylate chemistry has not been researched, but, in the case of n-octane where the total surface energy is approximately 12 kJ mole⁻¹ at 20°C [18], the activation energy values reported here of between 11-14 kJ mole⁻¹ are not unreasonable for this **type of process.**

An initial weight loss of 5^{-8%} is indicative of the loss of 1 mole of carbon **dioxide from one mole of carboxylate. That carbon dioxide was indeed the gaseous fraction lost, has been confirmed by standard chemical tests in sepa**rate decompositions carried out in a pyrex heating tube under a stream of **dry nitrogen. The amounts produced over this temperature range were also estimated by titration of the gas dissolved in an excess of standard barium hydroxide solution. The results obtained were, in all cases, in very good agreement with the TG weight loss data. Separate decompositions were also** carried out on the PbC₁₂ and PbC₁₄ in the temperature range 573-663 K in **an attempt to identify the partially decomposed intermediate. Differential thermogravimetry indicated that in this temperature range the most stable intermediates were fcrmed. Thus, each partially decomposed sample was worked up in ether and the purity of the product ascertained by thin layer chromatography (TLC1 in several solvent systems. The TLC results indicated that 'each degraded product was a single compound. Typical IR and NMR** spectra of the degraded intermediate obtained from PbC_{12} are shown in **Figs. 5 and 6, respectively.**

The IR spectrum is that for a typical carbonyl compound where the band at 1745 cm" seems likely to be due to the presence of carbonyl groups.

Fig. 5. IR spectrum of degraded PbC₁₂ intermediate.

O

Absorptions in the region $1410-1440$ cm⁻¹ could be tentatively assigned to $-CH₂-CO-$. The absence of any absorptions in the region 2700-2900 cm⁻¹ **indicates that the compound is not an aldehyde. F'urther confirmation of this tentative assignment as a ketone is provided by the NMR spectrum. The typical aldehydic chemical shifts in the region 9.5-10.1 ppm are absent. The spectrum shows the typical methyl group protons attached to an adjacent carbon atom at 0.82 ppm and methyl group protons attached to a carbonyl** group at 2.35 ppm. A tentative structure (I) for the compound is now pro**posed**

Fig. 6. NMR spectrum of degraded PbC₁₂ intermediate.

This structure should yield, theoretically, protons in the ratio of 4 : **36** : **6, NMR integration gives a proton ratio of 4** : **35** : **6, supporting the correct assignment of this compound as 12-tricosanone. Further evidence for this structure is provided by the melting data where a melting point of 69-70°C compares favourably tith the published melting point for this compound of 6849" C [191.** In **addition, other workers [9] have shown that when metal carboxylates are heated above 473 K, the corresponding ketone is invariably the major product. For the tetradecanoate, the intermediate is proposed to be 14-heptacosanone of structure II.**

$$
\underset{\text{CH}_3(\text{CH}_2)_{11}\text{CH}_2\text{CCH}_2(\text{CH}_2)_{11}\text{CH}_3^{\text{}}}{\overset{\text{II}}{\underset{\text{II}}{\prod}}}
$$

Evidence for this was provided by IR and NMR analyses. Additionally, an analysis for carbon and hydrogen gave values of 82.18% and 13.93%, respectively, compared with 82.2% and 13.7% expected for this compound. The compound melted at 74-75"C, in fair agreement with the literature value of 76.3"C [20].

A general mechanism for the ketone formation is

A side reaction is possible, leading to aldoketene formation thus

 $CH₃CH₂$, $CH = C = 0$ +PbO

The **activation energy values obtained for the second process in the first stage decomposition step must be a measure of the diffusion and chemical attack of the species just described which itself must be a function of the viscosity of the molten soap medium. Ekpe and Sime [2] have shown that the activation energy for viscous flow in a lead soap melt increases with chain** length at any temperature. It then becomes clear that diffusion and attack

would be more difficult in the case of the PbC₁₈ than in PbC₁₂ and PbC₁₄. **Steric factors might also be important. This is reflected in the activation energy data collected (Table 2) though, not unexpectedly, it is not easy to** distinguish between the values for the PbC₁₂ and PbC₁₄.

An explanation of the second stage decomposition data for PbC_{12} and PbC₁₄ is not entirely clear. The kinetic results suggest that the reaction mech**anisms are dissimilar. While the ietradecanoate degradation is concentration** independent and suggests that the surface of the degradative intermediate is **completely covered by some gaseous product, an order of reaction of approximately unity (0.8) in the case of dodecanoate is consistent with a random breakdown mechanism. It is not entirely clear why the two reaction mechanisms should differ. However, it is noticeable that the second stage decomposition process is less pronounced the longer the chain length becomes, until in the octadecanoate it disappears altogether. In the second degradation process, the ketone intermediate can break down to give aldoketene**

/ CHjCHzlnCH

which, being very reactive, might breakdown further in the presence of alkane to csrbene and carbon monoxide

$$
CH_3(CH_2)_nCH=C=O \rightarrow CH_3(CH_2)_n\ddot{C}H + CO
$$

Ketone decomposition of this type at high temperature is not unusual and is indeed the standard method of making ketene from acetone. This type of decomposition mechanism could explain dodecanoate second stage breakdown. As the carboxylate chain increases, it would appear that this type of reaction becomes less important.

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